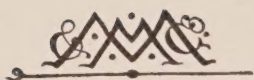


Digitized by the Internet Archive
in 2025

QD
11
.M6

A HISTORY OF CHEMISTRY



QD
11
M6
A

HISTORY OF CHEMISTRY

FROM EARLIEST TIMES TO
THE PRESENT DAY

BEING ALSO

AN INTRODUCTION TO THE STUDY OF THE
SCIENCE

BY

ERNST VON MEYER, PH.D.

PROFESSOR OF CHEMISTRY IN THE UNIVERSITY OF LEIPZIG

TRANSLATED WITH THE AUTHOR'S SANCTION

BY

GEORGE M'GOWAN, PH.D.

DEMONSTRATOR IN CHEMISTRY, UNIVERSITY COLLEGE OF NORTH WALES, BANGOR

WOODSTOCK COLLEGE
CANCELED DUPLICATE
LIBRARY

London

MACMILLAN AND CO.

AND NEW YORK

1891

All rights reserved

540.9E
M575

Prof. Chem
28-14395

QD
11
M6

PREFACE TO THE GERMAN EDITION

NEARLY five decades have passed by since Hermann Kopp's classical *Geschichte der Chemie*¹ began to appear, and it is now fifteen years since this was followed by the same indefatigable author's *Entwicklung der Chemie in der neueren Zeit*.²

The publication of these comprehensive works, in conjunction with which Höfer's *Histoire de la Chimie* must be named, and the further descriptions of the growth of chemistry within particular periods given both by Kopp himself and by other writers, might lead one to suppose that there was no pressing need for further work in the same direction at the present time.

This point can, the author thinks, be best discussed by his making a few remarks here with respect to the aim and plan of the present volume.

In this *History of Chemistry* the attempt has been made to describe within short compass the development of chemical knowledge, and especially of the general doctrines of chemistry which have thus been gradually evolved, from their earliest beginnings up to the present day. After a

¹ "History of Chemistry."

² "The Development of Chemistry in Recent Times."

general account of the main directions followed by chemistry in the various ages, the growth of particular branches of the science has been more or less minutely detailed.

In the general descriptions great emphasis has been laid upon the genesis of particular ideas, and their expansion into important dogmas or comprehensive theories. At the same time, in order that a vivid picture of the various periods and their distinguishing characteristics might be presented to the reader, short accounts have been given of the works, and in some cases of the lives, of the men who originated and developed such views.

In the special sections, on the other hand, the attempt has been made to collect together fundamental facts, which have been sifted and relegated to their proper branch of the science, and thus to offer as clear a description as possible of the state of chemical knowledge at the time in question.

That neither in this nor in the history of the development of theoretical views could completeness be thus achieved, hardly requires to be stated. But the author has at all events endeavoured to give a fair synopsis of the most important theories and facts which constitute the foundation of chemistry as we now know it.

The growth of chemical knowledge during recent times, since Boyle, and especially since Lavoisier, naturally forms the principal subject of the following chapters. The author is fully aware of the many difficulties which have to be met here, difficulties which increase in extent the nearer we approach to the history of our own period. We stand too close to the development of the theoretical views of these latter days to feel certain of always preserving the unbiassed

temperament which is essential to the true historian. But, notwithstanding this, the author has ventured the attempt to carry the record of the history of chemistry up to the present day.

In this task he has done his best to preserve throughout an objective attitude; and he has further been guided by the earnest desire to contribute effectively towards shedding a clear light upon the opposing views held with respect to the development and the importance of the chemical doctrines of to-day. It has also been his duty as an historian to endeavour to apply to the services rendered by eminent investigators of quite recent years a calmer and juster criticism than has hitherto in many cases been meted out to them.

ERNST VON MEYER.

LEIPZIG, *7th October* 1888.

TRANSLATOR'S PREFACE

THE author, in his preface to the original German edition, discusses the question whether there is any necessity for a new history of chemistry in his own language at the present day. That there is full room for one in this country will be admitted upon all hands. It is therefore hoped that the appended *history* will prove not only useful to the student, but also interesting to the general reader who is desirous of gaining some idea of the development of chemical science.

The translator has done his best to reproduce clearly the sense of the German original. And, since Professor von Meyer has been so kind as to read over the first corrected proofs, as well as to answer a great many queries, it is hoped that this has been achieved.

A considerable number of small alterations and additions have been made for this edition, most of them by the author, but some by the translator with the author's concurrence. While these may reasonably be supposed to have improved the book, they have not altered its character in the slightest degree. The translator has further added a number of duplicate references to English journals (to such papers as were published both in German and English), and also a few new ones, for the greater convenience of English readers.

In conclusion, he would express his indebtedness to the various gentlemen who have been kind enough to give him the benefit of their criticism and advice upon different points, with regard to which his own special knowledge was insufficient, and also to those others who have assisted him in the matter of references, etc.

UNIVERSITY COLLEGE OF N. WALES, BANGOR,
March 1891.

AUTHOR'S NOTE TO THE ENGLISH EDITION

It was a great satisfaction to me that the translation of this *history* was undertaken by my former pupil, Dr. McGowan, and I desire to express here my appreciation of the manner in which he has entered into the spirit of the work, and to offer him my hearty thanks for all his trouble in the matter.

May the book find many friends among the English-speaking peoples, and help to stimulate the interest of its readers in the development of our science.

ERNST VON MEYER.

LEIPZIG, *February 1891.*

TABLE OF CONTENTS

LIST OF ABBREVIATIONS	PAGE xxi
INTRODUCTION	1

CHAPTER I

FROM THE EARLIEST TIMES TO THE BIRTH OF ALCHEMY .	7
<i>Theoretical Views upon the Composition of Substances, and especially upon the Elements, 8. Aristotle's Elements, 9.</i>	
<i>The Empirical Chemical Knowledge of the Ancients, 10.</i>	
<i>Metallurgy of the Older Nations—Gold, 13; Silver, Copper, Iron, 14; Lead, Tin, etc., 15; Mercury, 16. The Manufacture of Glass, 16. Pottery, 17. The Manufacture of Soap, 17. Dyeing, 18. The beginnings of Pharmacy, 18.</i>	

CHAPTER II

THE AGE OF ALCHEMY	21
General History of Alchemy	23
<i>Origin and First Signs of Alchemistic Efforts, 23. The Alexandrian Academy, 25. The Alchemy of the Arabians—Geber and his Disciples, 27. Alchemy among the Western Nations, 30. Albertus Magnus, Roger Bacon, 31. Arnaldus Villanovanus, Raymundus Lullus, 32. Basilus Valentinus, 35.</i>	
Special History of Alchemy	37
<i>Theories and Problems of the Alchemistic Period, 37. Geber, 38, 39. Views of Albertus Magnus, Basilus, etc., 40. The Philosopher's Stone, 42.</i>	
<i>Practical-Chemical Knowledge of the Alchemists, 44.</i>	
<i>Technical Chemistry—Gold, 45. Silver, 46. Copper and other metals, 46. Pottery, Glass, Dyeing, 47. Pharmaceutical Chemistry, 47.</i>	

Knowledge of the Alchemists with regard to Chemical Compounds, 49. Alkalies, 49. Acids, 50. Salts, 51. Preparations of Antimony, etc., 53. Organic Compounds, 56.

The Fortunes of Alchemy during the last Four Centuries, 57.

A Short Review of Alchemistic Efforts, 62.

CHAPTER III

	PAGE
HISTORY OF THE IATRO-CHEMICAL PERIOD	63
General History of this Period	65
Paracelsus and his School, 65. The Iatro-chemical Doctrines of Paracelsus, 67. Turquet de Mayerne, 71. Libavius, 71. <i>Van Helmont and his Contemporaries</i> , 72. The Work of van Helmont, 73. Sala and Sennert, 77. Sylvius and Tachenius, 77. <i>Georgius Agricola</i> , 80. Palissy, 82. Glauber, 83.	
Special History of the Iatro-chemical Period	84
Technical Chemistry—Metallurgy, 85. Pottery and Glass Manufacture, 86. Dyeing, etc., 87.	
<i>Development of Pharmacy and of the Knowledge of Chemical Preparations</i> , 87. Inorganic Compounds, 88. Organic Compounds, 93.	

CHAPTER IV

HISTORY OF THE PERIOD OF THE PHLOGISTON THEORY, FROM	
BOYLE TO LAVOISIER	97
Introduction	97
General History of the Phlogistic Period	100
<i>Robert Boyle</i> , 100. Mayow, 104. Lemery and Homberg, 104. Kunkel and Becher, 105. <i>Stahl and the Phlogiston Theory</i> , 107. Fr. Hoffmann and Boerhave, 110.	
<i>The Development of Chemistry, and particularly of the Phlogiston Theory, after Stahl's Time</i> —Neumann, Eller, Pott, Marggraf, 112. Geoffroy, Duhamel de Monceau, Rouelle, Macquer, 113. Black, 116. Cavendish, 117. Priestley, 118. Bergman and Scheele, 119.	
Special History of the Phlogistic Period	121
Pneumatic Chemistry and its Relations to the Doctrine of Phlogiston, 121. The Discovery of Oxygen and the Composition of Air, 123.	
<i>Development of Theoretical Views in the Phlogistic Period</i> , 126.	
Views regarding Elements and Chemical Compounds, 126.	

Views regarding Chemical Affinity and its Causes, 130.
Geoffroy's Tables of Affinity, 131.

Practical Chemical Knowledge in the Phlogistic Age—
The Development of Analytical Chemistry, 133. Boyle, 134.
Fr. Hoffmann, Marggraf, Scheele, 135. Bergman, 136. The
beginnings of Gas Analysis, 137.

Technical Chemistry in the Phlogistic Age—Metallurgy, 138.
The Ceramic Industry, Dyeing, 139.

Technico-chemical Preparations—Acids and Alkalies, 139,
140. The Discovery of Elements, 141. Inorganic and Organic
Compounds, 142-144.

Pharmaceutical Chemistry, 145.

Concluding Remarks upon this Period, 147.

CHAPTER V

HISTORY OF THE MOST RECENT PERIOD (FROM THE TIME OF	PAGE
LAVOISIER UP TO NOW)	149

Introduction	149
------------------------	-----

General History of Chemistry during this Period	151
---	-----

Lavoisier and the Antiphlogistic Chemistry, 151. Lavoisier's
Life and Work, 151 *et seq.* His Combustion Theory, 155.
Triumph of the Antiphlogistic Chemistry, 159. Beginnings of
a Rational Chemical Nomenclature, 161. Guyton de Morveau,
162. Berthollet, 163. Fourcroy, 164. Vauquelin, 165.

*The State of Chemistry in Germany at the End of the Eighteenth
Century*, 167. Klaproth, 167. *The State of Chemistry in
England and Sweden*, 169.

Development of the Doctrine of Chemical Proportions, 170.
Richter, 171. His Law of Neutralisation, 172. The Beginnings
of Stöchiometry, 174. Proust, 174. His contest with Berthollet,
175. Recognition of Constant Combining Proportions, 176.

Dalton's Atomic Theory	177
----------------------------------	-----

Law of Multiple Proportions, 177. Dalton's Attempts to
determine the relative Atomic Weights of the Elements, 179.
His Atomic Weights and Chemical Symbols, 181, 182.

Further Development of the Atomic Theory, 182. Thomas
Thomson, 183. Wollaston, 183. Humphry Davy, his Life and
most important Work 184-187. Gay-Lussac, 187. His Law of
Volumes and Work generally, 188. *Proust's Hypothesis and its
Effects*, 189.

Berzelius—A Survey of his Work	191
--	-----

Biographical Notice, 191. His Influence upon the Develop-

ment of Analytical and Organic Chemistry, 193, 194. His Experimental Researches, 193, 194. Berzelius as a Teacher and Writer, 195. His General Character, 197.

Development of the Atomic Theory by Berzelius, 198. His Determinations of relative Atomic Weights, 199 *et seq.* His Oxygen Law, 200.

Influence of Gay-Lussac's Law of Volumes upon the Atomic Theory, 201. Avogadro's Hypothesis, 202. Application of the Law of Volumes by Berzelius, 203. The Position of the Atomic Theory in 1818, 204. *Dulong and Petit's Law*, 207. *Influence of the Doctrine of Isomorphism upon the Atomic Theory*, 208. Mitscherlich, 209.

The Atomic Weight System of Berzelius after 1821, 210. *Dumas' Attempt to alter the Atomic Weights*, 212. Failure of this Attempt, 213. Faraday, 214. His Law of Definite Electrolytic Action, 215.

The Electro-Chemical Theories of Davy and Berzelius, 216 *et seq.* *The Dualistic System of Berzelius*, 220. His Chemical Nomenclature and Notation, 221-224.

Manifestations against Dualism, 224. Discovery of the Alkali Metals, 224. Recognition of the Elementary Nature of Chlorine, 226. *Theory of the Hydrogen Acids* (Davy and Dulong), 228. *Doctrine of the Polybasic Acids* (Liebig), 230.

Development of the Dualistic Doctrine in the domain of Organic Chemistry, 232. *The Growth of Organic Chemistry previous to 1811*, 232. *The Position of Berzelius with regard to Organic Chemistry*, 234. Development of Views respecting Radicals, 235.

Isomerism and its Influence on the Development of Organic Chemistry, 236. Observations of Liebig, Wöhler, Faraday, and Berzelius, 237. Clearer Definition of the terms *Isomerism*, *Polymerism*, and *Metamerism* by Berzelius, 238.

The older Radical Theory, 239. The Etherin Theory, (Dumas and Boullay), 239, 240. Liebig and Wöhler's Work upon Benzoyl Compounds, 240. The Ethyl Theory of Berzelius and Liebig, 242. Position of the Radical Theory in 1837, 244. Definition of the term *Radical*, 245. Bunsen, 246. His Work upon the Cacodyl Compounds, 246. The significance of the Radical Theory, 247.

Liebig, Wöhler, and Dumas—A Survey of their more important Work, 247. Justus Liebig, his Life and Work, 248. Liebig as a Teacher, 249. His Literary Activity, 250. His experimental Researches, 251. Friedrich Wöhler, 254. Wöhler as a Teacher and Writer, 255. His services to Science, 256. Dumas, his Life and Work, 256-259.

The Development of Unitary Views in Organic Chemistry, 259.

Substitution Theories, 259. *Dumas' Laws of Substitution*, 260. *Laurent's Substitution or Nucleus Theory*, 261. Criticism of the same, 262. *Dumas' Type Theory*, 263. His Unitary System, 265. *The Overthrow of Berzelius' Dualistic Doctrine*, 265. Berzelius' Fight against the Substitution Theory, and his Defeat, 267.

Fusion of the older Theory of Types with the Radical Theory by Laurent and Gerhardt, 269. Laurent and Gerhardt, a Sketch of their Lives, 269. *Gerhardt's Theory of Residues*, 270. His Law of Basicity, 272. *Gerhardt's first Classification of Organic Compounds*, 272. *His Reform of the Atomic Weight System*, 273. *The distinguishing between the terms Molecule, Atom, and Equivalent by Laurent and Gerhardt*, 276. *Work preparatory to the new Type Theory*—Wurtz and A. W. Hofmann, 278. Williamson's Experiments on the Formation of Ethers, 281. His Opinions with regard to the "Typical" View, 282. *Gerhardt's new Theory of Types*, 283. Work preparatory to this, 284. Derivation of Organic Compounds from Types, 285. Gerhardt's Views upon Chemical Constitution, 287. Criticisms upon his Type Theory, 288. *Extension of the Type Theory by Kekulé*, 289. Mixed Types, 290. Marsh Gas as a Type, 291. Position of the Type Theory in 1857, 292.

Development of the Newer Radical Theory by Kolbe—*A Survey of Kolbe's Life and Work*, 292. *The Re-animation of the Radical Theory by him*—*Frankland's Co-operation*, 294. Copulated or Conjugate Compounds, 296. Setting aside of the Notion of Copulation by Frankland, 297. Kolbe's Carbonic Acid Theory, 299. The Derivation of Organic Compounds from Inorganic, 299. *Kolbe's most important Experimental Researches*, 300. *His Attitude towards the older and the newer Chemistry*, 301. Kolbe's real Types, 302.

The Founding of the Doctrine of the Saturation-Capacity of the Elements by Frankland	PAGE 304
--	-------------

Preparatory steps towards this Doctrine, 304. Frankland's services here, 304 *et seq.* Assumption of a varying Saturation-Capacity, 307. Discussions on the Subject by Odling, Williamson, and Wurtz, 308.

The Recognition of the Valency of Carbon, 309. Kekulé's services here, 310. Kolbe and Frankland's share in the Matter, 311.

Development of Chemistry under the Influence of the Doctrine of Valency during the last Thirty Years	313
--	-----

Beginnings of the Structure Theory—*Kekulé and Couper*, 314. Establishment of the true Atomic Weights by Cannizzaro, 317.

Discussions regarding the Nature of "Structure" by Butlerow and Erlenmeyer, 318.

Controversies respecting constant and varying Valency of the Elements, 319. Views upon varying Valency held by Frankland, Kolbe, etc., 319; by Erlenmeyer, Wurtz, and Naquet, 320. Kekulé's Theory of a Constant Valency, 320. Grounds for the Assumption of a varying Valency, 322.

The further Development of the Structure Theory—The chief Directions taken by Organic Chemistry during the last Twenty-five Years, 324. Views upon the linking of Atoms, 325. *Constitution of Organic Compounds according to the Structure Theory*, 326. Saturated and Unsaturated Compounds, 326. *Kekulé's Theory of the Aromatic Compounds*, 328. Modifications in this Theory proposed by Ladenburg, Claus, and Baeyer, 330. Constitution of pyridine, pyrrol, etc., 330, 331. Victor Meyer's more precise Conception of the term *Aromatic Compounds*, 331. *Application of Structural-chemical Conceptions to the Investigation of Isomerism*, 332. Position-isomerism, 333. Tautomerism or Desmotropism, 334. Geometrical isomerism (Wislicenus), 335. Alloisomerism (Michael), 336. The supposed Spatial Arrangement of Atoms, 337. *The Development of Important Methods for investigating the Constitution of Organic Compounds*, 338. *Synthetic Methods* (Wöhler, Kolbe, Frankland, Baeyer, Kekulé, Ladenburg, Fittig, W. H. Perkin sen., and others), 338. Chemical Behaviour of Organic Compounds, 341.

The Main Currents in Inorganic and General Chemistry during the last Thirty Years, 343. Application of the Structure Theory to Inorganic Compounds, 344. Important Researches in Inorganic Chemistry, 345. *The Periodic System of the Elements* (Newlands, L. Meyer, Mendelejeff), 346. Crookes' Hypothesis of a Primary Material, 349. *General Significance of Physico-chemical Investigations*, 350.

CHAPTER VI

SPECIAL HISTORY OF THE VARIOUS BRANCHES OF CHEMISTRY	PAGE
FROM LAVOISIER TO THE PRESENT DAY	353
Introduction	355
History of Analytical Chemistry	358

Qualitative Analysis of Inorganic Substances, 358. Use of the Spectroscope for this purpose, 359. *Quantitative Analysis of Inorganic Substances*, 360. Klaproth, Vauquelin, 360. Lavoisier, Proust, Berzelius, 361. H. Rose, Wöhler, Fresenius, 363. Docimacy, 364. *Volumetric Analysis*, 364. Its Development

by Gay-Lussac, Bunsen, Mohr, etc., 365. *Development of Methods of Gas Analysis*, 366. *The Analysis of Organic Substances* (Lavoisier, Gay-Lussac and Thénard, Berzelius, Liebig), 367. Legal-chemical Analysis, 370. Technico-chemical Methods, 371.

The Progress in Pure Chemistry from Lavoisier to the Present Time	PAGE 373
Special History of Inorganic Chemistry	373

The Discovery of Elements, and the Determination of their Atomic Weights, 374. The Halogens, 375. Selenium, Tellurium, etc., 376. Boron and Carbon, 377. Allotropy, 377. The Metals of the Alkalies and Alkaline Earths, 378, 379. Beryllium, Cadmium, Thallium, Aluminium, Indium, Gallium, 380. Metals of the Cerium Group, 381. Chromium, Titanium, Germanium, etc., 381, 382. Vanadium and allied Elements, 383. Metals of the Platinum Group, 383. Supposed new Elements, 385.

Inorganic Compounds, 386. *Hydrogen Compounds of the Halogens*, 386. *Oxygen Compounds of Hydrogen and of the Halogens*, 387. *Sulphur, Selenium, and Tellurium Compounds*, 388. Compounds of Nitrogen, Phosphorus, etc., 389. Compounds of Boron, Silicon, and Carbon, 392. Compounds of the Alkaline and Alkaline Earth Metals, 393. Compounds of the Metals of the Iron Group, etc., 394. Compounds of Tin, Vanadium, etc., 396. Compounds of Gold, Platinum, etc., 396.

Special History of Organic Chemistry in the Nineteenth Century	398
--	-----

Hydrocarbons and their Derivatives, 399. *The Alcohols and Analogous Compounds*, 402. *Carboxylic Acids*, 405. Acid Chlorides, Anhydrides, and Amides, 407. Oxy- and Amido-Acids, 409. *Aldehydes*, 411. *Ketones and Ketonic Acids*, 412. *Carbohydrates—Glucosides*, 414. *Haloid Derivatives of the Hydrocarbons, etc.*, 416. *Nitro- and Nitroso-Compounds*, 418. *Sulphur Compounds*, 420. *Organic Nitrogen Compounds* (Amines, etc.), 422. Phosphines, Arsines, Stibines, 425. Azo-Compounds, 425. Diazo-Compounds, 426. Hydrazines, Cyanogen Compounds, 427. *Pyridine and Quinoline Bases*, 431. Their Relation to Vegetable Alkaloids, 432. *Pyrrol and Analogous Compounds*, 435. *Organo-metallic Compounds*, 437.

History of Physical Chemistry in Recent Times	439
---	-----

Determination of Vapour Density and the Application of this, 440. *Dissociation*, 443. *The Liquefaction of Gases*, 443. *The Kinetic Theory of Gases*, 444. *Spectrum Analysis*, 444. *Atomic Volumes of Solids and Liquids*, 445. *Laws regulating the Boiling Temperature*, 446. *Raoult's Law of Solidification*, 447.

Specific Heat of Solid Bodies, 449. *Optical Behaviour of Solids and Liquids* (Refraction, Circular Polarisation), 449. *Diffusion, etc.*, 451. *The Electrolysis of liquid or of dissolved Substances*, 452. *Isomorphism, etc.*, 453. *Thermo-Chemistry*, 455. *Photo-Chemistry*, 458.

Development of the Doctrine of Affinity since the Time of Bergman, 460. *Bergman's Doctrine of Affinity*, 460. *Berthollet's Doctrine of Affinity*, 461. *The Supplanting of Berthollet's Opinions by other Doctrines*, 463. *The Revival of Berthollet's Doctrines*, 465. *The latest Development of the Doctrine of Affinity*, 467.

Sketch of the History of Mineralogical Chemistry during the last Hundred Years PAGE 469

Its Earlier History, 469. The Chemical Mineral System of Berzelius, 471. Other Mineral Systems, 472. The more recent Development of Mineral Chemistry, 472. *The Artificial Production of Minerals—Beginnings of Geological Chemistry*, 473.

Development of Agricultural and of Physiological Chemistry 477

Agricultural Chemistry and Vegetable Physiology, 478. The Humus Theory, 478. Reform of Agricultural Chemistry by Liebig, 479. Its further Development by Liebig and his School, 480.

The Development of Phyto-Chemistry, 481. Important Phyto-Chemical Researches, 482 *et seq.*

The Development of Zoo-Chemistry, 484. Researches upon the Constituents of the Animal Body, 484. The Chemistry of the Animal Secretions—Saliva, Gastric Juice, Bile, Blood, 486; Milk, Urine, 487. Metabolism, 489.

Fermentation; Putrefaction, 491. Views regarding Fermentation, 491. Organised and Unorganised Ferments, 492. The Phenomena of Putrefaction, 493. The Ptomaines, 493.

The Relation of Chemistry to Pathology and Therapeutics, 494. Bacteriology, 494. Antiseptics, Anæsthetics, and Antipyretics, 495.

The Relation of Chemistry to Pharmacy, 495.

History of Technical Chemistry during the last Hundred Years 498

Introduction, 498. Development of Technical Instruction, 500. Literature on Technical Chemistry, 500.

The Progress of Metallurgy, 501. Iron and Steel, 501. Nickel, Silver, the Galvano-Plastic Process, Aluminium, 502. Mineral Pigments, 503.

Development of the Great Chemical Industries, 504.

Sulphuric Acid, 504. The Soda Industry, 505. Hydrochloric Acid, Chlorine, and Bleaching Powder, 507, 508. Bromine and Iodine, 508. Nitric Acid, Gunpowder, 509. Other Explosives, Matches, 510.

The Manufacture of Soap, etc., 511. Ultramarine, 511. Glass, Earthenware, and Pottery, 512. Mortar, Paper, 513. Starch, Beet-Sugar, 514.

The Aniline Colours and other similar Dyes, 515. Phthal-eins, Azo-Dyes, 517. Alizarine, the Safranines, Indigo Blue, 518. *Dyeing*, 519. Tanning, 519.

Fermentation Processes, 520. The Manufacture of Spirits, 520. The Quick Vinegar Process, 521.

Various Products from Coal-tar; Illuminating Materials, 521. *Heating Materials*, 522.

The Growth of Chemical Instruction in the Nineteenth Century, more especially in Germany	PAGE 524
--	-------------

The State of Education in Science at the end of the Eighteenth Century, 524. Experimental Lectures, 525. The Development of Practical Instruction (Berzelius, Liebig), 525, 526. The Erection of Laboratories for General Instruction in Germany, 526. The State of Scientific Education in France, Great Britain, etc., 528. Improvements in the Construction of Chemical Laboratories, 530.

Chemical Literature, 531. Text-books, 531. Larger Treatises and Encyclopedias, 532. Periodical Journals, 532. Yearly Reports (*Jahresberichte*), 534. The Necessity for Criticism in Chemical Literature, 534.

INDEX OF AUTHORS' NAMES	537
INDEX OF SUBJECTS	545

ABBREVIATIONS OF THE NAMES OF MOST OF
THE JOURNALS TO WHICH REFERENCE
HAS BEEN MADE

- Ann. Chem.* Liebig's Annalen der Chemie und Pharmacie (begun 1832).
Ann. Chim. Annales de Chimie et de Physique (begun 1816; five series).
Ann. de Chimie The same journal from 1789 to 1815.
Ann. des Mines. Annales des Mines.
Ann. of Philosophy Annals of Philosophy (edited by Thomas Thomson, 1813-26). This journal was subsequently merged in the *Philosophical Magazine*.
Ann. Phys. The new Series (*Neue Folge*) of Poggendorff's Annalen.
Archiv Pharm. Archiv der Pharmacie (begun 1832).
Bayer. Akad. Sitzungsberichte der Bayerischen Akademie der Wissenschaften.
Ber. Berichte der Deutschen chemischen Gesellschaft (begun 1868).
Bull. Soc. Chim. Bulletin de la Société Chimique de Paris (begun 1864).
Chem. Centr. Chemisches Centralblatt (begun 1848).
Chem. News. Chemical News (begun 1860).
Compt. Rend. Comptes Rendus des Séances de l'Académie des Sciences (begun 1835).
Crell's Ann. Chemische Annalen von L. v. Crell (1784-1805).
Dingl. Journ. Dingler's Polytechnisches Journal (begun 1820).
Gilb. Ann. Annalen der Physik von Gilbert und Gren (1798-1824).
Hofmann's Bericht, { Bericht über die Entwicklung der Chemischen Industrie
etc. { während des letzten Jahrzehnts von Hofmann (began 1875, but ceased after the publication of two volumes).
Jahres. Berz. Jahresberichte über die Fortschritte der Chemie und Mineralogie von Berzelius (1821-47).
Jahres. d. Chemie Jahresberichte über die Fortschritte der Chemie von Liebig und anderen (begun 1847).

-
- Journ. Chem. Ind.* Journal of the Society of Chemical Industry (begun 1882).
- Journ. Chem. Soc.* Journal of the Chemical Society (Memoirs and Proceedings, vols. i.-iii., 1841-47; Journal begun 1848).
- Journ. de Phys.* Journal de Physique (1778-94; 1798-1823).
- Journ. pr. Chem.* Journal für praktische Chemie (begun 1834; the new series begun 1870).
- Mon. Scient.* Moniteur Scientifique (edited by Quesneville, begun 1857).
- Phil. Mag.* Philosophical Magazine (begun 1798).
- Phil. Trans.* Philosophical Transactions of the Royal Society (begun 1666).
- Phil. Trans. E.* Philosophical Transactions of the Royal Society of Edinburgh (begun 1788).
- Pogg. Ann.* Annalen der Physik und Chemie von Poggendorff (begun 1824; new series begun 1877).
- Proc. R. S.* Proceedings of the Royal Society [begun 1800. Vols. i.-iv. (1800-1843) are entitled "Abstracts of the Papers printed in the Philosophical Transactions of the Royal Society of London," and vols. v., vi. (1843-1854), "Abstracts of Papers communicated to the Royal Society." The final form of title, "Proceedings of the Royal Society of London," begins with vol. vii., published in 1856].
- Proc. R. S. E.* Proceedings of the Royal Society of Edinburgh (begun 1845).
- Schweigg. Journ.* Journal für Chemie und Physik von Schweigger (1811-33).
- Wagner's Jahresber.* Jahresbericht über die Leistungen der chemischen Technologie von Wagner (begun 1856).
- Ztschr. anal. Chem.* Zeitschrift für analytische Chemie von Fresenius (begun 1862).
- Ztschr. Chem.* Zeitschrift für Chemie (1865-71); this was a continuation of the Kritische Zeitschrift (begun 1858).
- Ztschr. phys. Chem.* Zeitschrift für physikalische Chemie, Stöchiometrie, und Verwandtschaftslehre (edited by Ostwald and van 't Hoff; begun 1887).

INTRODUCTION

CHEMISTRY has of late years been defined as the study of the composition of substances. Its first task, therefore, lies in ascertaining the constituents of which the material world surrounding us is composed, in reducing these constituents to their elements, and in building up new chemical compounds from the latter. Hand in hand with this there goes the further task of determining the laws which regulate the chemical combination of matter.

The problems just indicated occupy, in the widest sense of the word, the attention of chemists to-day. The problems of chemistry in former times were, however, different, and it is precisely these differences in aim which characterise the various epochs into which the history of the science may, therefore, be divided.

The oldest nations with regard to which we possess reliable information—the Egyptians, Phœnicians, Jews, and others—did indeed possess a certain disjointed knowledge of chemical processes acquired accidentally; but these were applied for their practical results alone, and not with the object of deducing any comprehensive scientific explanation from them. We meet with similar conditions among the earliest cultured European nations, the Greeks and Romans, who owed most of their knowledge of chemical facts to the peoples just named. Nowhere do we find in antiquity the endeavour to gain an insight into chemical processes by means of definitely planned experiments. Although the Ancients were wholly without such data, furnished by exact

research, as are nowadays held to be indispensable, this did not prevent them from speculating as to the nature of the universe; indeed, those theoretical views upon the nature of matter, on the "elements" of which the world was composed, have given to the earliest age of chemistry its own particular stamp. Further, a number of these systems—especially Aristotle's system of the elements—continued to hold sway for many centuries, and influenced the whole teaching of the Middle Ages.

From the above-mentioned doctrine of the nature of the elements there was developed the theory of the transmutation of metals, or rather the fixed belief that one element can be transformed into another. Even so far back as the beginning of our own era, at first in Egypt, there began the striving to transmute the base metals into the noble, to "create" gold and silver.

The art by which this was to be achieved was termed *chemia* (χημεία), a name dating, so far as actual proof goes, from the fourth century, but in reality probably from an earlier period.¹

There are many indications that this conception of the aim of chemistry and of the problems which it had to solve predominated for centuries following, *e.g.* it lies at the root of the definition given by Suidas, the author of an encyclopedia, who lived in the eleventh century: "Chemistry, the artificial preparation of silver and gold;" further, "χρυσοποιία" was a very commonly used designation for chemistry during a long period.

This task, the solution of which was the aim of the so-called "Alchemy," characterises the alchemistic period, a period extending from at least the fourth century of our era

¹ This word is of Egyptian origin and is probably founded on the North Egyptian word *Chēmi*, the name for Egypt. It also means, however, "black," and hence there is still some doubt whether the word *χημεία* of that period denotes Egyptian art or, as Hoffmann in the article "Chemie," in the *Dictionary of Chemistry* published by A. Ladenburg, endeavours to prove, the employment of a black-coloured preparation valuable for alchemical purposes. The mode of writing *χημεία*, and the derivation of this word from *χυμός*, may be regarded as incorrect.

² This term with the Arabic prefix "al" very early became naturalised.

to the first half of the sixteenth. It is impossible to state with perfect exactitude the date at which alchemy took its rise, its origin being lost in the mists of the past. The labours of the alchemists, who strove by all imaginable methods to attain to the philosopher's stone (by the aid of which not only were the noble metals to be produced from the base, but also the life of man to be prolonged), had the effect of largely extending the area of the then existing knowledge of chemical facts.

In the first half of the sixteenth century, almost contemporaneously with the Reformation, *i.e.* with the birth of a new epoch in the world's history, chemistry began to develop in a new direction, without, however, losing all at once its alchemistic tendencies. Chemistry, which had already proved itself a valuable helpmeet to medicine in the preparation of active remedies, came to be looked upon as the basis of the whole medical art. Health and illness were reduced to chemical processes in the human body; only by means of chemical preparations could an unhealthy body be restored to its normal condition; in short, the absorption of medicine in chemistry, the fusion of both together, was the watchword which emanated from Paracelsus. Van Helmont, de le Bœ Sylvius, Tachenius, and others were the chief exponents of this doctrine, which characterises the period of Medical or Iatro-Chemistry. The fact that technical chemistry was advanced at the same time, through the labours of individuals such as Georgius Agricola, was without influence on the prevailing tendency of the science of that age.

From the middle of the seventeenth century on, the iatro-chemical current gradually underwent substitution by another. After that date chemistry strove hard to become a self-supporting branch of natural science, quite independent of all others. Indeed, the history of chemistry proper begins with Robert Boyle, who taught, as its main object, the acquisition of a knowledge of the composition of bodies.

The conception of this aim marks the date from which chemistry may be regarded as a science striving towards an ideal goal along the paths of exact research, without

regard to practical results, and solely with the object of arriving at the truth.

The most important problem, whose solution occupied all the eminent chemists of that day, was the question of the chemical reasons of the phenomena of combustion. Since Stahl's attempt to explain the latter, the hypothetical fire stuff Phlogiston, which was supposed to escape during every combustion, was regarded as the universal principle of combustibility. This doctrine held sway over chemists at the end of the seventeenth and during the greater part of the eighteenth centuries to such an extent that we are justified in characterising this period (after the death of iatrochemistry) as the period of the Phlogiston Theory.

The fall of the latter, and its replacement by the anti-phlogistic system of Lavoisier, bring us to the commencement of the chemical era in which we are still living. For, upon the foundation laid by Lavoisier and his co-workers, and firmly fixed by Dalton, Berzelius and others, the structure of the new chemistry rises. The founding and developing of the chemical atomic theory, and its extension to all parts of chemical science, characterise this latest epoch, to which the period of Lavoisier's reform of chemistry was a necessary stepping-stone; it is, therefore, to be designated as the period of the Chemical Atomic Theory. An insight into the conditions which it involved being only possible by careful quantitative researches, the balance has been, since the time of Lavoisier, the most valuable instrument of the chemist. H. Kopp is, therefore, fully justified in naming the epoch which begins with the French *servant* the period of quantitative research. Of late years the first aim of chemistry, *i.e.* the determination of the composition of substances, has been accompanied by the investigation of the relations which exist between their properties and composition. But the light of the atomic theory permeates the whole, so that one is forced to regard it as the guiding star of modern chemistry.

To the characteristics of the more important periods in the history of chemistry may be added here a few general

remarks upon the manner in which it was studied, and on the dissemination of chemical knowledge in the various ages.

In early times actual chemical operations were carried out only by a few of the initiated, whose art was therefore known by the name of "holy" (*ἁγία τέχνη*). In Egypt particularly, the home of chemical knowledge, the latter was jealously guarded by the priesthood, although eminent Greek philosophers, such as Solon, Pythagoras, Herodotus, Democritus, and Plato, succeeded in gaining their confidence and thus penetrating the mysteries. One frequently reads of laboratories attached to the temples, in which chemical operations were conducted; and rooms have actually been discovered, for instance, in Edfu, Dendera, the inscriptions in which show them to have been used for the above purposes.

In the alchemistic ages it was likewise the rule to keep secret those processes by which the transmutation of metals was to be effected. Notwithstanding this, however, numerous operations became known, especially those which bore on applied chemistry, and which were carried out practically in drug- and work-shops, etc. The oldest of such processes were the metallurgic ones followed in smelting works, and which had already in Pliny's time attained to a high degree of excellence. Geber (in the eighth century) used well-constructed apparatus for filtering, subliming, and distilling, and he was also acquainted with water and ash baths.

A systematic, if but slight, instruction in chemistry was first made possible in the seventeenth century by the founding of public laboratories and the establishment of chemical professorships. Even in the century preceding we meet with text-books of chemistry, such as those of Agricola (*De Re Metallica*) and of Libavius (*Alchemia* and *Chemia*), in which the most important of the operations then accurately known are described. With the development of chemistry into an independent science, the means of studying it increased, so that in the eighteenth century theoretical chemistry was one of the subjects taught in every university; notwithstanding this, however, it made but little progress compared with some of

the other branches of natural science. A special section of this book will be devoted to the great development of chemical teaching in our own century, while the following chapters will give a sufficiently detailed idea of the position of chemistry in the various ages.

CHAPTER I

FROM THE EARLIEST TIMES TO THE BIRTH OF ALCHEMY

THE characteristics of the period which have just been mentioned justify one in designating it the period of crude empiricism with regard to chemical facts. In sharp contrast with the disinclination of the Ancients towards experiment, by which the secrets of nature are to be unravelled, stands their great love of speculation, by means of which they did not hesitate to attempt an explanation of the ultimate reasons of all things. Aristotle, to whom the natural sciences owed the direction which they followed for a long time, pointed to deduction as the road which should lead to the goal. Instead of drawing general conclusions from accurately observed facts, the Ancients preferred to advance from the general to the particular. The position of all the natural sciences in far-back times, especially that of chemistry, is sufficient to prove how the most mischievous errors crept in and became firmly established in consequence of the purely deductive method.

The philosophical writings of the Ancients, especially those of the Greeks and Romans, give us a tolerably distinct idea of their theoretical views. Certain writings of Aristotle, and also the “*περὶ λίθων*” of his pupil Theophrastus, are of especial value for the criticism of the empirical chemical knowledge of these times.

The works of Dioscorides on *Materia Medica* and particular chapters of the *Historia Naturalis* of the elder Pliny give us an exceptionally clear glance into the know-

ledge of the Ancients. Dioscorides, who was born about the middle of the first century in Anazarbae, enlarged his acquirements, already eminent, by experiences collected on long journeys. His fame as a physician holds good among the Turkish doctors to this day. The work of Pliny above mentioned contains exceedingly valuable records of the state of scientific knowledge in his time; it also shows, however, that the author was by no means master of the immense amount of material which he had collected¹ from tradition, but which he had not really assimilated.

*Theoretical Views upon the Composition of Substances, and especially upon the Elements.*²

The question of the ultimate constituents of bodies, *i.e.* of the elements which go to build up the world, occupied the minds of the oldest nations. To give an exhaustive description of their speculations on the point does not come within the scope of this work; what is wanted is rather to call special attention to those views which have exercised a permanent influence upon the chemical ideas of later times.

This applies in a particular degree to the doctrine of the elements, which originated with Empedocles, although it usually bears Aristotle's name; also, but to a lesser extent, to the ideas of the older Greek philosophers regarding the original material of which the world, according to them, was built up. Views like that of Thales (in the sixth century B.C.), that water is the ground material, or those of Anaximenes and Heraclitus (in the same century), who ascribed to air and fire respectively the same *rôle*, have had no influence upon the development of chemical knowledge.

Democritus (in the fifth century B.C.) also took a ground material as the basis of his speculations, but subdivided this further in that he imagined it to be made up of the smallest

¹ Pliny the younger characterised the work of his uncle as "*opus diffusum, creditum, ac minus varium, quam ipsa natura*," and similar admiration of it was expressed by other authors of the day.

² Cf. Kopp, *Geschichte der Chemie*, vol. i. p. 29; vol. ii. p. 267; also Höfer, *Histoire de la Chimie*, vol. i. p. 72.

possible particles, of atoms, which differed from one another in form and size, but not in the nature of their substance. All the changes in the world consisted, according to him, in the separation and recombination of these atoms, which were supposed to be in a state of continual motion. This doctrine, which at first sight appears to accord with our present chemical atomic theory, but which in reality has nothing in common with the latter, was further developed by Epicurus, as may be well seen in the didactic poem of Lucretius, *De Rerum Natura*.

The four so-called "elements"—air, water, earth, and fire—were first regarded by Empedocles of Agrigent (about 440 B.C.) as the basis of the world; Aristotle himself did not look upon them as different kinds of matter, but as different properties carried about by one original matter. Their chief qualities (the *primæ qualitates* of the later scholastics) he held to be those apparent to the touch, viz. warm, cold, dry, and moist. Each of the four so-called elements is characterised by the possession of two of these properties, air being warm and moist, water moist and cold, earth cold and dry, and fire dry and warm. The differences in the material world were, therefore, to be ascribed to the properties inherent in matter. From the assumption that these latter can alter, there necessarily follows the immediate conviction that substances can be transformed, one into the other. It is easy to see how, when based upon speculations of this nature, the belief in the transformation of water into air should establish itself, for both have the property of moistness in common, while cold, the individual property of water, can be converted by the addition of heat into the second chief property of air. And it is not surprising that considerations of this kind on the states of aggregation of matter should lead to the conception of transforming one kind of matter into another. It was by the generalisation of such ideas that the belief in the possibility of the transmutation of metals, which formed the chief feature of the alchemistic period, grew to such dimensions.

Aristotle considered that his four elements were insuffi-

cient in themselves to explain the phenomena of nature; he therefore assumed a fifth one, termed *οὐσία*, which he imagined to possess an ethereal or immaterial nature and to permeate the whole world. As the "*quinta essentia*" this played an immense rôle among the followers of the Aristotelian doctrine in the Middle Ages, and gave rise to endless confusion, from the endeavours of many (who, unlike Aristotle, supposed it to be material) to isolate it.

There seems to be a high degree of probability in the assumption that Empedocles and Aristotle did not themselves deduce their theory of the elements, but derived it from other sources; thus the oldest writings of India teach that the world consists of the four elements mentioned above,¹ together with ether.²

It is unnecessary to point out how widely the above views of the Greek philosophers with regard to the elements deviate from the conceptions of modern chemistry.

With respect also to the meaning of the term "chemical combination," one meets, even if only occasionally, with opinions diametrically opposed to those obtaining at the present day; the formation of a substance by the interaction of others was looked upon as the creation of a new matter, and the destruction of the original substances from which it was produced was assumed. Everywhere men were contented with theoretical explanations, without attempting to prove their correctness by actual experiment. This want shows itself very markedly in the manner in which the Ancients regarded the numerous chemical facts which they had learned by empirical methods, and mostly by accident.

*The Empirical Chemical Knowledge of the Ancients.*³

The Egyptians stand out from among the earlier civilised nations as having usefully applied their knowledge of chemical

¹ Instead of air, the element wind is given.

² So teaches Buddha (as Dr. Pfungst has been good enough to inform me); see the *Anguttara Nikâja*, vol. i. fol. c.e. Here consciousness is named as the sixth element.

³ Cf. Kopp, *Gesch. d. Chemie*, vols. iii. and iv.; Höfer, *Hist.*, p. 106, *et seq.*

processes, acquired by chance observations; the needs of everyday life and the desire to make that life a comfortable one were the teachers.

Their country formed a kind of focus in which was concentrated the chemical knowledge of the time, if one may so designate an acquaintance with technical processes. The Egyptians already possessed at a very early date a large experience in the production of metals and alloys, in dyeing, in the manufacture of glass, and also in the making and application of pharmaceutical and antiseptic preparations. There can scarcely be a doubt that the Phœnicians and Jews obtained their knowledge of the manufacture of important technical products from the Egyptians. In like manner, and to an even greater extent, was there a treasury of chemical experiences opened out to the Greeks, and later on to the Romans, by reason of their close relations with the ancient country Chêmi (see p. 2, note 1).

But all the knowledge so gained was and remained empirical; only a very long time afterwards were its items brought together under a general scientific standpoint. In this section of the book only those portions of applied chemistry which were known to the Ancients will be treated of. That a people, so gifted as the Greeks were, should have failed to understand how to group together the numerous observations in those subjects which lay ready to their hand, and to draw conclusions from them, can only be explained by the whole tendency of their thought, and particularly by their undervaluing the inductive method. Aristotle's opinion that "industrial work tends to lower the standard of thought" was certainly of influence here. In accordance with this dictum the educated Greeks held aloof from the observation and practice of technical chemical processes; a theoretical explanation of the reactions involved in these lay outside their circle of interests. To this want of sympathy is certainly to be ascribed the fact that the discovery of even the most important chemical processes is but very seldom to be connected with the names of distinct historical persons; while, on the other hand, the old

historians give detailed records of those men who advanced untenable opinions on the constitution of the world.

An account of the state of practical chemical knowledge in early times will show that much uncertainty often prevailed in consequence of different products being called by one and the same name. Substances were not distinguished according to their chemical behaviour, the investigation of which possessed no interest for the Ancients, but were classified according to their outward appearance and source, a confounding of similar or identification of dissimilar substances thus frequently resulting. Two samples of one and the same compound—soda, for instance—were looked upon as different, if the external appearance seemed to indicate a dissimilarity. Much discrimination has been found to be, and still is, requisite in order to clear up the indistinct points in the records of the old historians.

Metallurgy of the Older Nations.¹

We find in the earliest records of the civilised nations (the Egyptians, Jews, Indians, etc.) an acquaintance with the working of different metals. Mythical persons are held by the younger civilised peoples to have taught this art, *eg.* Prometheus, Cadmus, etc., by the Greeks. If the translations of the Hebrew words in the Old Testament signifying "metals" are correct, then the Jews were acquainted with six, *viz.* gold, silver, copper, iron, lead, and tin; this may be considered certain as regards the first four, which either occur native or are readily reduced from their ores. They are recorded on the oldest monument in the order just given.

The name "metals" is derived, according to Pliny, from the fact of their never occurring separately but in veins together, *μετ' ἄλλα*.² Already at that period glance, ductility, and hardness were held to be characteristics of a

¹ The following works have been used for reference:—R. Andree, *Die Metalle bei den Naturvölkern* (Veit und Compagnie, Leipzig, 1884); Beck, *Geschichte des Eisens* (Vieweg, Braunschweig, 1884); O. Schrader, *Sprachvergleichung und Urgeschichte* (Jena, 1883).

² Herodotus gives *μέταλλον* as signifying a mine.

metal. With regard to the origin of metals and ores in the interior of the earth, the Ancients had formed the most extravagant conceptions; they believed, on the ground of Aristotle's weighty testimony, that they were produced by the penetration of air into the vitals of the earth, and consequently assumed that the amount of metal or ore increased as the mine proceeded inwards.

The Greeks, and especially the Romans, were intimately acquainted with many metallurgical processes; Dioscorides, Pliny, and later historians give fairly exact data for the obtaining and smelting of ores, but not the slightest attempt is made to explain the chemical processes which these involve.

The noble metals gold and silver, whose stability in the fire had not escaped the Ancients, were those earliest known (in prehistoric times), and were highly valued; the fact of their occurring native, and the ease with which they can be worked, afford a sufficient explanation of this.¹ The exceeding malleability of gold excited in a high degree the astonishment of the older nations, and rendered possible the gilding of objects by covering them with thin plates of the metal. The later discovery of affixing a layer of gold by means of the amalgamation process was already known considerably before the time of Pliny.

In the second century B.C. we meet with the first records² of a cupellation process, by which gold was freed from admixtures; in fact, an operation similar to the so-called lead process was then carried out, gold dust being melted with lead and salt for a number of days. The

¹ The gold mines of Nubia (the Egyptian name *nub*, *i.e.* gold, is perhaps connected with the designation of that country) were worked very extensively by the Egyptians. According to the record of Diodorus Siculus, in which he expresses pity for the slaves employed in the work, the finely ground gold ore was washed out and the heavy residue melted. In the time of Rameses II. the mines yielded gold to the value of £125,000,000 sterling per annum. The gold-producing land of Ophir, from which the Phœnicians obtained the precious metal, is supposed to have been in India, Midian (Arabia), or on the east coast of Africa. The same energetic trading nation opened up for the Greeks the first gold mines on the island of Thasos.

² This record, which originated with Agarthides, is to be found in Diodorus.

purification of gold by means of mercury was also well known in Pliny's time.

Silver, which the energetic Phœnicians are supposed to have supplied to the other civilised nations from Armenia and Spain, where rich silver ores occur, was, according to the record of Strabo, *i.e.* at the beginning of our era, purified in a precisely similar manner to gold, *viz.* by fusion with lead. The separation of silver from gold does not appear to have been known before our era, at any rate an extant record¹ states that Archimedes was not possessed of the means to accomplish this. From indications which Pliny gives, however, it appears that in his time a kind of cementation process was practised, which probably consisted in the treatment of silver containing gold with salt and alum shale. Moreover, an amalgam of gold and silver was regarded in ancient times as a particular individual metal, being termed *asen* by the Egyptians, and ἡλεκτρος by the Greeks (amber being distinguished as τὸ ἡλεκτρον). From this also it may be concluded that at that time no method was known of separating the metals.

The data concerning copper (termed χαλκός, *aes*²), which has been known from very primitive times (being first found in the neolithic stone age), frequently refer to its alloys with other metals, especially to bronze; the latter, as is well known, was very early used for making weapons, ornaments, and utensils. Copper, which was universally employed in prehistoric times, was found native in many places (*e.g.* in Egypt), or was readily smelted from malachite or similar copper ores. All the civilised nations, which have been named, were acquainted with bronze before they had learnt to prepare metallic tin, no mention of which is made in old Egyptian records. With regard to the smelting processes by which the "*aes*" of the Ancients was obtained, nothing certain is known.

¹ Archimedes attempted to determine whether the crown of King Hiero contained silver, and, if so, how much; this problem he tried to solve by taking the specific gravity, not by chemical means.

² The Roman *aes* has the same stem as the Sanscrit word *ayas*, signifying ore; the later designation *cuprum* for copper is an abbreviation of *aescyprum* (so called because of its occurrence in Cyprus).

Iron, the obtaining and working of which was not discovered till after that of copper and bronze, but which, nevertheless, goes back to very ancient times,¹ was prepared in smelting furnaces; the old authors do not, however, give any particulars as to the actual process.² The ores used are supposed to have been brown iron ore and magnetite; that meteoric iron was first employed is an improbable and unproven assumption. The tempering of iron was early learnt in Ancient Egypt; already in the time of Pliny the undesirable property of impure iron, which we now term brittleness, was known, and its capability of assuming the peculiarity of the magnet stone when brought into contact with the latter, was also observed.

Lead and tin were likewise known to the Ancients; it is, however, doubtful whether the latter was ever prepared pure before our era. Nothing certain is known as to how either of them was obtained, both being for long regarded as varieties of one metal, and therefore distinguished as *plumbum nigrum* and *plumbum candidum*.³ Lead was extensively used by the Romans in the first century A.D., for example, in the making of water-pipes; the soldering of these with tin, and conversely that of tin with lead, and also the production of tinned vessels, were already known at that date.

What has been said with regard to the application of

¹ According to *Lepsius*, iron has been in use in Egypt for more than 5000 years, having served primarily for the manufacture of hard instruments, while utensils of all kinds were made from bronze.

² Old Roman smelting furnaces with their appurtenances have recently been excavated near Eisenberg in the Pfalz. The form of apparatus used by the Egyptians for the smelting of iron can be arrived at approximately from inscriptions, etc.; it is worthy of note that the ancient Egyptian bellows are in use in the interior of Africa at the present day.

³ The word *stannum*, which now denotes tin, appears in Pliny's time to have signified an alloy of tin and lead. Whether the *κασσίτερος* of the *Hiad* stood for tin is likewise highly problematical. It is equally uncertain from whence the Phœnicians obtained this metal (or an alloy of it); whether from India, with which they had commercial relations, or from Britain and Iberia. The similarity between the Sanscrit word *kastira* and the Greek word *κασσίτερος* has been used as an argument in favour of the former assumption (cf. *Al. v. Humboldt, Kosmos*, ii. § 409).

tin applies also to that of zinc. Both of them, after alloying with copper, were used at the earliest periods for the manufacture of the most various utensils, while the records of the metal obtained from the earth termed *cadmia*, which in all probability was zinc, are both few and unreliable.

Brass, the first description of which is given by Aristotle as the "metal of the Mosynoeci" (from which the German word *Messing*, signifying brass, is undoubtedly derived), was for long regarded as copper which had been coloured yellow by fusing it with an earth (*cadmia*): it was only recognised as an alloy at a much later date. The change in colour produced in copper by certain additions to it played an important part in the alchemistic age.

The first records as to mercury are to be found in Theophrastus (about 300 B.C.), who gives its preparation from cinnabar by means of copper and vinegar, and terms it "liquid silver." Dioscorides describes the production of mercury, which he at first termed *ὕδραργυρος*, from cinnabar and iron, *i.e.* by a process of simple elective affinity, without, however, making the slightest attempt to explain the process. For the carrying out of this operation, an exceedingly imperfect distilling apparatus was used. It did not escape the Ancients that other metals, gold in especial, were altered by mercury (cf. p. 13); indeed Vitruvius gives a minute process for the recovery of gold in worn-out sewn draperies by means of it.

An account will be given later on of several metallic compounds known in ancient times.

The Manufacture of Glass.—The art of making vessels from glass originated in China and Egypt, and had for a long time its chief habitat in Thebes: from there it spread to the Phœnicians and other Eastern nations, the Greeks first acquiring it in the fifth century B.C. Pliny is the first to give a distinct account of the preparation of glass by fusing sand and soda together.¹

¹ The discovery of glass in Egypt was undoubtedly accidental, soda having been added as a flux to sand containing gold for the purpose of obtaining the latter.

The artificial colouring of glass by metallic oxides, especially oxide of copper, was very early discovered. Many of the remains which have been found in Ancient Egypt indicate that the manufacture of glass must at that time have attained to a high degree of perfection, the modes of production of artificial gems, enamels, etc., being then known.

The first preparation of glass presupposes in any case an acquaintance with soda or potash; the former of these was found as a natural product on the shores of certain lakes, *e.g.* in Macedonia and Egypt, while carbonate of potash was obtained from a very early period by lixiviating the ashes of plants, and also, according to Dioscorides, by igniting tartar. These two salts¹ were frequently mistaken for one another on account of their similar action. They were largely used for the preparation of soap, and also directly for washing clothes, cleansing the skin and also the teeth (just as the ash of tobacco, which is rich in carbonate of potash, is often employed as a tooth-powder at the present day), and lastly, as ingredients of medicines.

To the art of pottery must be ascribed an age at least as great as that of the preparation of the nobler metals and of glass. Even the old Egyptians understood how to coat their originally simple earthen vessels with coloured enamel. At a later date the ceramic industry prospered among the Etruscans, and also in many towns of Southern Italy and Asia Minor. Porcelain, which was discovered and employed by the Chinese, remained entirely unknown to the older civilised European nations.

The Manufacture of Soap.—Of no slight interest is the fact that the saponification of fats by means of alkalies, with the object of preparing soap—that is to say, a complicated process of organic chemistry—was already practised in ancient times. Pliny's records on the subject leave no doubt that in Germany and Gaul soap was made from animal fat and the aqueous extract of ashes, the latter

¹ The Hebrew *neter* probably denotes soda, while the Latin *nitrum* is employed by Pliny for both alkaline salts. The designation *alkali* came originally from the Arabs.

being strengthened (rendered caustic) by the addition of lime. Further, there was a distinction drawn between soft and hard soap, according as potash or soda (the latter being obtained from the ashes of shore plants in Gaul) was used in its preparation.

Dyeing likewise belongs to the arts which the Egyptians, Phœnicians, and Jews greatly developed. They knew how to fix certain dyes on cloth by means of mordants, alum¹ playing an important part here; indeed the dyeing of purple had attained to a high state of perfection among the Phœnicians. Indigo blue seems to have been more used at that time for painting than for dyeing, but with this exception mineral substances were employed as paints. The principal of these in Pliny's time were white lead, cinnabar, vermilion, smalt,² verdigris, red oxide of iron, and soot. Native sulphide of antimony served as a cosmetic, the sulphides of arsenic, realgar and orpiment, being also employed. In short, the Ancients had access to a considerable number of colouring chemical compounds, some of these being the earliest chemical preparations to be manufactured on a large scale.

The employment of such artificially prepared products in medicine likewise extends to a period very far back, even although, in referring to this, one can only speak of the first beginnings of a pharmaceutical chemistry. But a connection between the chemical art and pharmacy established itself very early indeed, *e.g.* among the Egyptians, who were doubtless the first to employ actual chemical preparations for medicinal purposes. Thus verdigris, white lead, alum, and soda served for the making of salves and other medicaments, while the preparation of lead plaister from litharge and oil was much practised in the time of Dioscorides. Iron rust was a very old medicine, its use being ascribed to Æsculapius, while sulphur and copper vitriol containing iron

¹ Under *στυπτηρα* or *alumen* of the Ancients must be understood substances of astringent properties generally, although alum itself is what is usually meant; being prepared from alum shale, it contained green vitriol as an impurity.

² Davy found cobalt in certain antique glasses, and assumed from this that smalt had been used in their manufacture.

(*chalcanthum*) were valuable ingredients of the medical treasury before our era; but the important preparations of antimony and mercury can be proved to have first come into notice in the alchemistic period.

Most of the officinal compounds just referred to were also used for other purposes, as has already been mentioned in a few cases. The combustion product of sulphur, for instance, was employed for fumigation (Homer), and for the purification of clothes (Pliny), while copper vitriol and alum were used in dyeing operations.—In closing this short account of the knowledge possessed by the Ancients with regard to chemical compounds, the following substances may be mentioned, substances whose practical application dates from a very early period. In ancient times lime was burnt, and after being slaked, was used for preparing mortar, and also, as already stated, for causticising soda (cf. p. 17). Of the acids, acetic acid¹ in the form of crude wine vinegar was the earliest known, its presence being assumed in all acid plant juices. The mineral acids, which are of such importance in technical chemistry, were only discovered in the succeeding epoch.

Other organic compounds known at the beginning of our era, and doubtless even before then, were sugar (from the sugar-cane), starch² (from wheat), many fatty oils, from seeds and fruits (the oil being extracted either by pressing or by boiling with water), and oil of turpentine, which was obtained by the distillation of pine resin in very imperfect apparatus. Compounds such as spirit of wine, carbonic acid, etc., which are formed in many processes of fermentation, *e.g.* in the making of wine, beer, and bread, remained unknown to the Ancients. It is true that they noticed in

¹ The Ancients had the most extravagant ideas with regard to the solvent power of vinegar upon mineral substances, as may be gathered from the concordant statements of Livy and Plutarch that Hannibal, in his passage across the Alps, cleared the way of rocks by means of it. The story which Pliny tells of Cleopatra may also be recalled here,—how she, in fulfilment of her wager to consume a million sesterces at one meal, dissolved costly pearls in vinegar and drank the solution.

² *ἄμυλον*, so called from its being prepared without millstones, and the production of which is described by Dioscorides.

these cases and also in others—natural emanations of gas, for instance—the presence of a kind of air prejudicial to breathing, but it did not occur to them to recognise in this a gas different from atmospheric air.

This lack of the gift of observation, this disinclination to go to the root of any phenomenon, in fact, a certain indifference with regard to natural events, are characteristics of the attitude of the Ancients with respect to nature. Instead of experimenting with natural products, they infinitely preferred to call speculation to their aid, so that the most superficial observations gave rise to opinions which, when uttered by high authorities, attained to the dignity of dogmas. How otherwise than from an extreme lack of the desire of observation can one explain Aristotle's assertion that a vessel filled with ashes will contain as much water as one which is empty? A further instance of the credulity of that time is given in the conviction expressed by Pliny, and universally held, that air can be transformed into water, and *vice versa*, that earth is produced from water, and that rock-crystal also proceeds from the latter. The assumption that water can be transformed into earth has often come up again at later periods, having exercised the minds of people even in comparatively recent times; as it subsequently assumed the form of an important question of dispute, it will be referred to in detail later on.

CHAPTER II

THE AGE OF ALCHEMY

IN the introduction to this book Egypt is spoken of as the mother-land of Alchemy. The university of Alexandria was especially instrumental in the propagation of the latter during the first centuries of our era: it was the carrier and intermediary for the alchemistic views at the time of the fall of the Western Roman Empire.

The attempts to convert the base metals into the noble ones had their origin in superficial observations, which appeared to give a strong support to the belief in such transmutation. Among such accidental observations was that of the deposition of copper, from the waters which accumulated in copper mines, upon iron utensils left therein. What more natural than to conclude that a transmutation of iron into copper had occurred? For the production of gold or silver from copper, the transformation of the latter into yellow or white alloys by means of earthy substances such as calamine or arsenic appeared to give warrant. Finally, the fact that a residue of gold or silver remained behind when an alloy with lead or an amalgam with mercury was strongly heated, indicated the generation of those noble metals.

To these considerations of a practical nature, which strengthened the conviction as to the transmutation of metals, but which inferred a gross self-deception on the part of the observer himself, there came to be allied, in this epoch for the first time, the tendency to group together chemical facts from common points of view.

It was precisely in the mode in which it was attempted to explain the composition of the metals that there lay a powerful and ever-active charm, leading to the belief in the ennobling of the baser metals and to continually repeated efforts to achieve this. The first beginnings in an experimental direction, which we meet with early in the alchemistic period, although very incomplete, indicate nevertheless a distinct step in advance as compared with the deductive method which had hitherto reigned supreme, and whose fruits consisted, for the most part, in the setting up of mystic cosmogonies. The few observations made remained, however, isolated—that is, were not grouped together in a connected manner.

That the attempts to attain to a knowledge of the processes of nature by the inductive method were but slight at best in the alchemistic period, is explained by the supremacy of the Aristotelian doctrine, which, amalgamated with the Neo-Platonic philosophy, enchained the minds of men almost throughout the whole of the Middle Ages. Even the Christian theology had to compromise with this system, the product of the joint work being scholasticism, which imprinted its stamp upon all the mental efforts of that time and prevented their free development. The relation of the alchemistic tendencies to the Aristotelian philosophy has been already indicated (p. 9).

The limitation of this epoch between the first appearance of alchemistic conceptions (in the fourth century) and the bold attempt of Paracelsus to call in chemistry to the aid of medicine (in the beginning of the sixteenth century) is thus a natural one, since, during the whole of this time, one and the same keynote runs through all the questions bearing upon chemistry, viz. the idea of the ennobling of the metals. People were so convinced of the practicability of this for many centuries, that every one who devoted himself to chemistry, and many others besides, strove towards this long-desired goal. The early mixing up of astrological and cabalistic nonsense with these alchemistic endeavours marks very distinctly the degeneration of the latter.

Alchemy by no means ceased to exist on the appearance of the new iatro-chemical doctrines, but gradually receded as chemistry became more of a science. True, its seductive problems are often seen to light up the chemists' camp uncannily, and to exercise upon even the most eminent of them an undoubted influence; but upon the main lines which chemistry has followed ever since the time of Boyle, the phantasies of alchemy have had no appreciable effect. Notwithstanding, however, that this influence was but slight, a short account of the position of alchemy during the last four centuries cannot properly be omitted, and will therefore be added as an appendix to this section of the book.

GENERAL HISTORY OF ALCHEMY.¹

Origin² and First Signs of Alchemistic Efforts.

The sources from which the belief in the practicability of the transmutation of metals was nourished, and which in the course of centuries gradually expanded into a broad stream of the most mischievous errors, have their origin in the gray mists of antiquity. No actual proof of these must be looked for; we depend, with regard to them, upon mythical and mystical traditions. The first historical sources are, further, small in volume and very muddy. But we find among various nations distinct signs of alchemy having been pursued as a secret science and having been held in honour.

When one recalls to mind that Ancient Egypt was a centre of the higher culture, and, especially, that it was a country where the chemical art was practised, one feels no surprise that the earliest reliable records of alchemy are to be found there. Egyptian sources, partly such as have been preserved to us by the Leyden papyrus, and partly the writings of the Alexandrians from the third to the seventh century A.D., constitute the most valuable aids at our disposal for a historical proof of the origin of alchemy.

¹ Cf. Kopp, *Gesch. d. Chemie*, vol. i. p. 40, *et seq.*; also his work, *Die Alchemie in älterer und neuerer Zeit* (Heidelberg, 1886).

² Cf. particularly M. Berthelot's *Les Origines de l'Alchimie* (Paris, 1885).

The tradition, according to which, among other knowledge, the art of ennobling metals had been brought from heaven to earth by demons, was universally diffused in the first centuries of our era; Zosimos of Panopolis states that the mystical book from which this art was to be learned was termed *χημεν*, and the art itself *χημεία*. This myth doubtless sprang from one exactly similar which is to be found in the apocryphal book of Enoch: indeed indications of it are to be met with even in Genesis. The later alchemists were inclined to refer the origin of alchemy to the time before the flood, thinking that a special sanctity would accrue to their art from this great age. Moreover, they wrote down various biblical characters as alchemists, on the authority of certain passages in Holy Writ, for instance, Moses and his sister Miriam, and the Evangelist John. When legends such as these found credence in the Middle Ages, it is not to be wondered at that the records as to the origin of this art, which remain to us from ancient times, should have retained their authority for a long period.

The first personality with which the origin of alchemy is associated is that of Hermes Trismegistos,¹ "the three times great," who was said to have been the author of books upon the holy art; he was, moreover, generally revered as the discoverer of all the arts and sciences. The expressions "hermetic" and "hermetic art"² recalled this undoubtedly mythical personage even so recently as in the present century. In Romish Egypt pillars were erected in honour of this Hermes, upon which alchemistic inscriptions were cut in hieroglyphics.

Who then was this Hermes? One has to seek in him, as ancient traditions indicate with certainty, the personified idea of strength, viz. the old Egyptian godhead Thot (or Theuth), which, when endowed with the serpent-staff as the symbol of wisdom, was compared by the Greeks with their Hermes, the latter designation being thus transferred to the

¹ This designation is first found in Tertullian (end of the second century of our era).

² The designation "spagiric art" (from *σπάω*, to separate, and *ἀγέλω*, to unite) occurs for the first time in the sixteenth century.

Egyptian god. Alchemy, as a divine art, whose special task consisted in the preparation of the metals, was kept secret and fostered by the priesthood, the sons of kings alone being permitted to learn its mysteries. The estimation in which it was held rose in exact proportion with the belief that Egypt owed to alchemy its riches.

When and in what way the influence of other nations made itself felt upon the alchemy of the Egyptians, it is difficult to determine. The Babylonish astrologers, without doubt, undertook the fusion of astrology and magic; in particular, the mutual relations between the sun and planets and the metals, which were taken for granted for so many centuries, were of old Babylonish origin. According to the account of the Neo-Platonist Olympiodor (in the fifth century A.D.), gold corresponds to the sun, silver to the moon, copper to Venus, iron to Mars, tin to Mercury, and lead to Saturn.¹

Certain passages in the works of Dioscorides, Pliny, and the Gnostics enable us to conclude that the transmutation of copper into silver and gold was regarded as an ascertained fact during the first centuries of our era.² The "duplication of the metals," which is to be found in the writings of first-century authors, and which also plays a part in the Leyden papyrus, appears likewise to refer to the transmutation of metals. The designation of this art as "Chemia" first appears in an astrological treatise of Julius Firmicus (in the fourth century).

The records of the study of alchemy go on increasing from this date, much information regarding it being found in the writings of the Alexandrian *sarants* of that time, especially in those of Zosimos, Synesios, and Olympiodor. In addition to these, various pseudo-authors, in especial pseudo-Democritus, are cited here as witnesses to the spread of alchemy; the philological-historical critic is not yet, however, in a position to fix the dates in which these works

¹ Even in Galen are to be found statements with regard to the influence of the planets upon the metals.

² The Chinese also busied themselves with alchemy at that time, the transformation of tin into silver, and of the latter into gold, being held to have been actually accomplished.

were written. In the Middle Ages people did not hesitate to accept the writings of the false Democritus, and also those of a pseudo-Aristotle, as originating from the ancient philosophers Democritus and Aristotle themselves. The later alchemists also fathered counterfeit writings upon Thales, Heraclitus, and Plato, in order to make use of the great authority of those names for their own ends.

Zosimos of Panopolis, a voluminous author of the fifth century, who was looked upon as one of the greatest authorities among alchemists both of that date and of later periods, is said to have written twenty-eight books treating of alchemy, of which, however, only small fragments remain. His mystical recipes are quite unintelligible, and yet he distinctly speaks of the fixation of mercury, of a tincture which changes silver into gold, and also of a divine water (panacea). Reference is frequently made to the work of the pseudo-Democritus, *φυσικὰ καὶ μυστικά*. The graphic and mysterious language of Zosimos appears to have exercised a permanent influence upon the works of the later Alexandrians, and also, subsequently, upon those of the alchemists of the Middle Ages.

The end of the fourth century and the beginning of the fifth constitute, without doubt, the period in which the study of alchemy reached its zenith among the Alexandrians; but the works upon alchemy and magic of Synesios, and also those of Olympiodor, who bore the surname of "ποιητής," *operator*, do not yield much certain information with regard to definite operations or to the knowledge of chemical facts. How many works which would have been valuable for the history of chemistry were lost through the destruction of the Serapeum, which marked the completion of the overthrow of Hellenic culture in Egypt, cannot at this distance of time be estimated. That the knowledge of chemical operations, and especially of chemical knowledge and skill, was not thereby quite exterminated was due to the relations which were before that developed between the Alexandrians and the Byzantine *savants*; for, from the sixth century on, applied chemistry, which may also be said to

include alchemy, found a foothold at Byzantium. Even in Egypt itself the knowledge of chemistry was not completely extirpated by that catastrophe, but continued to exist by fostering certain branches of industry, which, without it, could never have been developed. Lastly, the conviction that metals could be transmuted had fixed its roots too deeply to allow of this art, by which endless riches were to be attained, dying a natural death.

The Alchemy of the Arabians—Geber and his Disciples.

The germs of chemical knowledge, which had lain hidden in the brains of a few philosophers, attained to a marvellous growth among the Arabians, who overran and conquered Egypt in the seventh century; it might have appeared much more likely that they would crush the arts and sciences rather than be the instruments of their resurrection. It was certainly curious that this people, originally strangers to science, should assume the care of it and cause it to flourish in an undreamt-of degree, at a time when culture remained at its lowest ebb in most European countries, and everything had to give way to the pressure of the conditions produced by the migration of the nations.¹

The first appearance of the Arabians in Egypt, where they destroyed the priceless treasures of the library of Alexandria by fire, did not seem to herald any such change of opinion. They very soon learnt, however, to assimilate the elements of the education of the conquered peoples,² so that we find (especially after the conquest of Spain, in the begin-

¹ Alex. von Humboldt gives expression to this point in the following words: "The Arabians, an original Semitic stock, drove away in part the barbarism which had overwhelmed Europe for two centuries, convulsed as it had been by revolutions. They turned to the everlasting springs of Greek philosophy, and thereby assisted not only in preserving the culture of science, but in widening it and opening out new paths to the investigators of nature."

² Reference may just be made here to the important part played by the Nestorians in the engrafting of the scientific spirit upon the Arabians. Certain it is that the first impressions of Grecian literature and the germs of scientific medical education were conveyed to the Arabians by that sect, and it is not improbable that chemical knowledge reached them through the same channel.

ning of the eighth century) many cities of learning springing up; to these in the following centuries the European nations—especially France, Italy, and Germany—sent crowds of anxious learners, who applied themselves, for the most part, to the study of medicine, mathematics, and optics. From the Arabian universities of Spain, in which also alchemy was ardently studied, it made its way to the other western nations, among which it attained to its full development in the thirteenth century.

The first detailed works which give an exact account of the state of chemical knowledge among the Arabians of the eighth century are those of the celebrated physician Geber, about whose life but little certain is known, although he was held in the highest regard both by his contemporaries and also during the whole of the Middle Ages: his work was accomplished in the second half of the eighth century.¹ The numerous works ascribed to him first became generally known in the sixteenth century, by means of Latin translations; the original Arabic MSS. have not as yet been published, so that it is often doubtful whether, in many cases, the translator has not altered certain passages or added new matter. Further, the authenticity of several memoirs bearing Geber's name has been called in question, but he may be accepted as the author of the most important of them, viz. *Somma Perfectionis Magisterii*; *De Investigatione Perfectionis Metallorum*; *De Investigatione Veritatis*; and probably also the *Testamentum Geberi*.

From these works we recognise in Geber a man of the widest chemical knowledge, such as had never before been pictured in previous writings. In contrast with the utterly unintelligible language of the Alexandrian alchemists, he describes with candour, and often with the most refreshing clearness, definite chemical processes for the making of preparations, several of which are mentioned for the first

¹ We learn from recent researches that the designation Geber was given among western nations to the Greek Dschabir, who was born in Tharsis, a disciple of the Arabian Dschafer, who taught in Medina towards the middle of the eighth century. Formerly both teacher and pupil were regarded as identical.

time by himself (see the section of this book upon practical knowledge). He describes apparatus, such as the water-bath, improved furnaces, etc., which have since come into general use, and also teaches the more important chemical operations which no chemist can dispense with, *e.g.* sublimation, filtration, crystallisation, and distillation with improved apparatus. Thanks to such aid, the chemical art was developed greatly in Geber's time, and the number of material facts appertaining to the domain of chemistry was widely extended.

He was, however, much less successful in his theoretical explanation of these facts. Geber manifestly took his views of the composition of the metals, which will be treated of in detail later on, from his predecessors, and, having extended these, attempted to fit his observations to this hypothesis—a hypothesis not based upon facts. The endeavour to attain to this end by the discovery of the philosopher's stone runs through all his writings. If, as Geber assumed, the metals are different mixtures of sulphur and mercury, the noble metals being particularly rich in mercury and poor in sulphur, and *vice versa*, then the transmutation of lead or copper into silver or gold simply consists in the withdrawal of sulphur from, and the addition of mercury to, them. For the accomplishment of this the so-called medicines of different order and strength were to serve. The alchemists accepted the convincing simplicity of such reasoning as showing the correctness of the hypothesis, and rendering any actual proof of it unnecessary. For following this *ignis fatuus* the whole alchemistic age suffered, since the fixed idea of the metals being composite prevailed during many centuries.

The disciples of Geber, Arabian physicians, appropriated his doctrines and experiences without adding materially to them. Rhazes, Avicenna, Avenzoar, Abukases, and Averrhoes, who lived in different parts of the Arabian dominion, exercised a lasting influence upon the development of medicine and of pharmacy.

Alchemy among the Christian Nations of the West during the Middle Ages.

The doctrines of the Arabian alchemists, especially of Geber, gradually penetrated into France, Italy, and Germany, certain Byzantine *sarants*—Michael Psellus among them—also contributing to the spread of alchemistic ideas. Eastern influence is recognised distinctly for the first time in the earliest appearance—of which there is clear proof—of an alchemist in Germany at the court of Adalbert von Bremen (about 1063), as recorded by Adam von Bremen: a baptized Jew named Paul gave out that he had learnt in Greece the art of transmuting copper into gold, and he appears to have imposed upon the above-named ecclesiastical prince. The next certain records of alchemistic endeavours in Germany date from the thirteenth century, at which period chemistry was studied by men famous for their learning, and was consequently developed in a high degree.

The transformation of the base metals into the noble by means of the philosopher's stone formed at that date the cardinal point towards which all chemical knowledge was directed. Vinzenz of Beauvais¹ (in the first half of the thirteenth century) and, after him, men like Albertus Magnus, Roger Bacon, Arnaldus Villanovanus, and Raymond Lully, whose chief works belong to the same century, regarded the transmutation of metals as an incontrovertible fact. These maintained that the philosopher's stone did exist, and was endowed with the most marvellous powers. In their theoretical views upon the composition of the metals, from which the transmutability of these follows, they are disciples of Geber; and the impress of the scholasticism derived from the Aristotelian philosophy upon their ideas is also easy to recognise. In addition to these, the most distinguished representatives of chemistry, all of whom belonged to the priestly class, must be mentioned the famous Thomas Aquinas, who did not indeed materially

¹ Vincentius Bellovacensis.

advance the knowledge of chemistry, but who stood up at various times for the truth of the doctrine of transmutation of metals.

The influence of the four men above mentioned upon the history of chemistry renders biographical notices of them desirable; their views upon the alchemistic problem, and also their very considerable practical knowledge, will be treated of under special sections. Their writings have to be criticised with some caution, since many of the alchemistic treatises of later times were given out to the world under their names.

Albertus Magnus, or, more properly, Albert von Bollstädt, born at Lauingen on the Danube in 1193, taught publicly as a Dominican in Hildesheim, Regensburg, Cologne, and Paris, and became Bishop of Regensburg in 1260. He retired, however, to the cloister five years later, and died in the Dominican convent of Cologne after having devoted himself for fifteen years to scientific work. Albertus Magnus was held, both by his contemporaries and still more during the later Middle Ages, as a man of the greatest erudition and widest acquirements, the degrees of which are given by Tritheim, an author of the fifteenth century, in the following words: *Magnus in magia naturali, major in philosophia, maximus in theologia*. His noble character also earned for him the highest respect. Of his numerous memoirs, the two—*De Alchymia* and *De Rebus Metallicis et Mineralibus* are of the most value for adjudging his position with regard to alchemy.

Roger Bacon was born in Somersetshire in 1214, and studied science, as well as theology, both at Oxford and Paris. The veneration felt by posterity for his marvellous and many-sided knowledge is shown by the title which it conferred upon him of *Doctor Mirabilis*. Since he did not hesitate to oppose in many points the orthodox beliefs of his day, he was subjected to bitter persecution and penalties. His death probably occurred in the year 1294.

His firm belief in the power of the philosopher's stone, not only to transform a million times its own weight of

base metal into gold, but also to prolong life, seems to us incomprehensible when contrasted with the otherwise enlightened views which he held and propagated. This undisguised recognition of miracle-working, and this bias towards the marvellous, are directly opposed by the fact that Roger Bacon taught the working out of carefully devised experiments as a special kind of research, by which new data for the knowledge of nature should be acquired. He is to be regarded as the intellectual originator of experimental research, if the departure in this direction is to be coupled with any one name—a direction which, followed more and more as time went on, gave to the science its own particular stamp, and ensured its steady development. The most important works of Roger Bacon are the following:—*Opus Majus*; *Speculum Alchemie*; and *Brevi Breviarium de Dono Dei*. He did not apparently do much towards the spread and development of practical chemical knowledge.

In the life and work of the two notable alchemists, Arnaldus Villanovanus and Raymundus Lullus, the alchemistic tendencies of their century are clearly reflected, although much uncertainty exists as to many points, especially in the life of the latter, and also with regard to the works ascribed to Lully. Both of them at all events were held in high esteem, not only during their lives, but also in the centuries following. Arnaldus Villanovanus, whose birth-place is uncertain, practised as a physician in Barcelona in the second half of the thirteenth century. His opinions, however, causing great offence to the priests, he was obliged to flee from there, and after vainly endeavouring to escape persecution in Paris and in various towns of Italy, he at last found an asylum in Sicily with King Frederick II. Summoned to Avignon by Pope Clement V., then seriously ill, he lost his life by shipwreck on the way thither, about the year 1313. In his theoretical views upon the composition of metals he followed the doctrines of Roger Bacon and Albertus Magnus, and therefore also of Geber; his contributions to practical chemistry aided the medical art not

inconsiderably. He had special opinions of his own as to the nature and efficacy of the philosopher's stone, and also with regard to the noble metals obtained through its means. Among his writings may be mentioned: *Rosarius Philosophorum*; *De Vinis*; and *De Venenis*.

A similarly restless life was foreordained for Raymund Lully, a life which comprised in itself the greatest contradictions and eccentricities. Shortly after his death the object of a traditional glorification, Lully possessed among all alchemists the fame of having attained to the highest which it was in the power of their art to achieve. The historical critic has a difficult task in dealing with him; for while, on the one hand, many of the writings ascribed to him are obviously counterfeit, there are, on the other, no sufficient data for deciding as to which are really genuine. Thus there is very great uncertainty whether the alchemist Raymund Lully is identical with the famous grammarian and dialectician of the same name, who was called by his admirers *Doctor Illuminatissimus*; for this view, which has been held by many, is strongly opposed by the fact that criticisms of alchemy are to be found in many of the works of the latter.

Most of the records which we possess of the life of Raymund Lully agree in stating that he was descended from a noble Spanish family, and was born in the year 1235. After living a dissipated life at the court of Aragon, he abjured the pleasures of the world in his thirtieth year and devoted himself to science. It was probably Bacon and Villanovanus who initiated him into the secrets of alchemy. When somewhat aged, he gave himself up to the conversion of the heathen, undertaking several journeys to Africa for this purpose; his reception there, however, was more than once of the worst, and he was at last stoned to death in the year 1315. Tradition has it that he lived for several years after this date in the unresting study of alchemy, but there can be no doubt as to the untenability of this report.

His alchemistic doctrines were very obscure, but

they appear to owe their origin to those of Geber; still more incomprehensible and hidden in deep mystic darkness are his recipes for the ennobling of the metals. Certainly none of the alchemists who preceded him have ascribed to the philosopher's stone such powers as he did; for he was able to cry out presumptuously: "If the sea were of mercury, I would transform it into gold."¹ And not only gold, but also all precious stones, and that highest good—health,—together with long life, were to be obtained through its means. Of the writings which are attributed to him, the *Testamentum*, *Codicillus seu Vademecum*, and *Experimenta* are regarded as genuine.

The history of alchemy in the fourteenth and first half of the fifteenth centuries contains no single name which will compare in eminence with those of the above-mentioned philosophers, as the alchemists themselves preferred to be called.

This must not be taken as meaning that the supposed art of making gold had died out: on the contrary, it bore its strangest fruit during that period. If it be desired to connect specific names with the study of alchemy at that time, then the Frenchman Nicolas Flamel, Isaac Hollandus the elder and the younger, Count Bernardo da Trevigo, and Sir George Ripley may be mentioned as among those who were supposed to be in possession of the wonder-working philosopher's stone. These men did nothing, however, to materially advance the knowledge of chemistry.

Alchemy was at this time fostered and protected at many of the European courts, for nothing appeared to be more simple than to recuperate embarrassed finances by means of artificial gold. Many documents in the history of that century bear record to the frequent disappointments which were certain to come about sooner or later,—decrees against the practice of alchemy, threatenings of those who contravened these with the severest punishments, and accounts of discoveries of the most impudent impositions.

¹ *Mare tingerem, si mercurius esset.*

Alchemy found especial protection at the court of Henry VI. of England, in spite of the fact that the kings preceding him had had to pay heavily for their propensities towards the hermetic art, and that a stringent law against it had been promulgated by Henry IV. The consequence of the favour shown to it by these monarchs was the production of large quantities of counterfeit gold which, in the form of coinage, inundated neighbouring countries. Charles VII. of France, who was then at war with England, was seduced by an alchemist, Le Cor, into a similar experiment, and thereby materially increased the debt of his country; to the alchemistic gold which he set in circulation were added the English "Rose nobles." Counterfeit coining, carried out on such a large scale, was not calculated to raise the reputation in which alchemy was held.

Chemistry, not being enriched during that time by any facts of note, likewise suffered from this depreciation; it first received new life from the work of Basil Valentine, whose acquirements in practical chemistry even now excite our highest astonishment. This remarkable man was the real precursor of the iatro-chemical period, even although he was unable to free himself from the fetters of the alchemistic faith. Of his life practically nothing is known; from his writings we learn his name and the time, approximately, in which he lived, viz. the second half of the fifteenth century, and also that he was a Benedictine monk of Southern Germany. The most important of his works were published in the beginning of the seventeenth century by a city chamberlain Tolde in Frankenhausen, Thuringia; whether foreign matter has become mixed up with them cannot now be determined. So much is certain, that Basil Valentine was regarded as an oracle by alchemists as early as the end of the fifteenth and beginning of the sixteenth centuries, and was held in higher honour than Geber, higher even than Raymund Lully, besides being admired by many who had nothing to do with alchemy. His works were spread abroad by means of copies, and excited the interest of the Emperor Maximilian I. to such a degree that he

caused a searching inquiry to be made in the year 1515 as to which Benedictine convent the famous author had dwelt in; but unfortunately his efforts in this direction were without result, as were also all later ones.

An account will be given later on both of his theoretical views and of his wide acquaintance with practical chemistry. Among the writings which are presumably his, and which are at the same time of the most importance, are the following:—*Triumphwagen des Antimonii* ("Triumphal Car of Antimony"); *Von dem grossen Stein der Uralten Weisen* ("On the Great Stone of the Ancient Philosophers"); *Offenbarung der verborgenen Handgriffe* ("Revelation of the Hidden Key"); *Letztes Testament* ("Last Testament"); *Schlussreden* ("Concluding Words").

In the first-mentioned work we possess what for that time was a marvellous description of an element and its compounds, the knowledge of these being due to Basil Valentine himself. The language which he employs is frequently obscured by mystical pictures and alchemistic conceptions; but, while he thus appears as a visionary on the one hand, he excites on the other our highest admiration from the completeness of his temperate and conscientious observations, as well as from the rational views he takes of subjects which were then, for the most part, judged erroneously. The rich experiences in practical chemistry which he made his own cause him to stand out as the most distinguished chemist of the whole alchemistic period. His boldness, too, in proposing the use of chemical preparations for medicinal purposes led the way in a direction which soon after his time became the prevailing one, viz. the iatro-chemical, which dominated the succeeding age. Notwithstanding all this, Basil Valentine was an out-and-out alchemist, holding, as such, the most exaggerated ideas as to the power of the philosopher's stone, just as the tendency towards alchemy and the firm belief in the possibility of transmuting metals and of prolonging life continued engrained in many of the iatro-chemists.

SPECIAL HISTORY OF ALCHEMY.

Theories and Problems of the Alchemistic Period.

The alchemistic ideas, with the transmutation of metals as their leading principle, have been proved, as already mentioned, to have originated and to have been first systematically fostered in Egypt. The first attempt to explain this assumed transmutation, by a theoretical conception of the nature of metals, was made very early. From a similar endeavour, *i.e.* from regarding transmutation, then looked upon as an incontrovertible fact, as a consequence of the constitution of the metals, there sprang the doctrine advocated by Geber, which in its essentials predominated during the alchemistic period. It was thus always the metals which gave rise to the early chemical theories.

If we penetrate to the kernel of the doctrines of the Alexandrians through the veil of mysticism which envelopes it, we see that these philosophers were permeated with the idea that the metals were alloys of varying composition. From this it necessarily followed that the transformation of one metal into another was possible, either through the addition of new metallic substances or the expulsion of some already present. Such transformations of similar substances into one another appear much less wonderful than those of dissimilar ones like air, water, and earth, which were mutually convertible, according to the teaching of the Platonists and Aristotelians. The means for bringing about these changes in the metals, the substances which it was necessary to add to them, and the operations which had to be gone through, were either kept secret or obscured by indistinct figurative language. The various colours of the metals, and their alteration by melting them with others, played a prominent part in alchemistic processes; in imparting thereby the colour of a noble metal to a base one, much was supposed to have been attained. For the Alexandrians, and also for the alchemists of the Middle Ages, the colouring of metals was synonymous with their transmutation. The chief operations were the so-called *Xanthosis*,

Leukosis, and *Melanosis*, which were compared with the processes followed in the dyeing of cloth. The old designation of *tinctura* for the media by which this transformation was brought about, gives expression to the idea that the latter consisted in a dyeing operation.

As may be imagined, no trace can be found of any distinct chemical conception, or of any knowledge of the actual operations which take place in these transmutations. At the root, however, of these endeavours of the Alexandrian alchemists to produce noble metals from base, lay speculations purely philosophical, which strongly excited and strengthened the belief in the transmutation of metals. These were partly taken from the writings of Plato, especially from his *Timæus*, which was highly esteemed by the Alexandrians, and partly from the philosophy of Aristotle. Both of these Greeks held the opinion that the elements in general were capable of transformation into one another,¹ and an extension of this idea led to the assumption that the same applied to the metals. The observations of the supposed generation of noble metals from base, which have been already discussed, were looked upon as proofs of the correctness of this supposition.

In the works of Geber we meet with a new specific chemical theory of the metals, which did not indeed originate with him, but which he developed, and which was diffused by his authority and found universal recognition. This theory looks upon classes of bodies for the first time from a chemical point of view, and seeks to explain the difference between the substances comprising these by assuming a peculiar composition. The metals, as Geber taught, consist of sulphur and mercury, which are present in them in different proportions and in different degrees of purity.²

The transmutation of metals consists, according to him, in

¹ In one passage of *Timæus* we read: "We believe from observation that water becomes stone and earth by condensation, and wind and air by subdivision; ignited air becomes fire, but this, when condensed and extinguished, takes again the form of air, and the latter is then transformed into mist, which dissolves into water. From this, lastly, are produced rocks and earth."

² Geber sometimes added arsenic to the above-named constituents of the metals as a third possible one, without, however, laying emphasis upon this

an arbitrary alteration of their composition ; the ennobling of them, specially, in a purification and fixation of the mercury. The idea of creating a metal anew, which we find highly developed among Western alchemists, is not to be found in Geber's writings. This, together with the application of his theory, is proved by the following sentences, which comprise in themselves his theoretical and practical chemical programme : " To assert that one substance can be produced from another which does not contain it, is folly. Since, however, all metals consist of sulphur and mercury, we can add to them the constituent in which they are deficient, or abstract the one which is present in excess. In order to achieve this, make use of the art : calcination, sublimation, decantation, solution, distillation, coagulation (crystallisation), and fixation. The active agents are the salts, alums, vitriols, borax, the strongest vinegar, and fire."

While Geber, in many passages of his works, draws no distinction between the supposed two constituents of the metals and natural sulphur and mercury, we frequently find him expressing, in others, the opinion that the former are not identical with the latter. The mercury and sulphur present in the metals were, in this second case, looked upon as being of an abstract nature ; thus mercury conferred glance, malleability, fusibility, and what we consider metallic properties generally, while sulphur, on account of its combustibility, was regarded as being present, because of the alteration of many metals in the fire. The noble metals, those which withstood the fire, therefore consisted of almost pure mercury, which however could not be identical with the ordinary substance of that name, since the latter was volatile ; this property was ascribed by Geber to the fact of ordinary mercury containing sulphur. By means of these and similar assumptions, contradictions between theory and facts were easily set aside, the alchemists of later times especially distinguishing themselves in this way.

extension. Here and there, also, Aristotle's doctrine of the four different states of matter appears to get mixed up with his views upon the composition of the metals, the four elements being regarded to some extent as subsidiary constituents, sulphur and mercury being the principal ones.

For the solution of the possible problem of the transmutation of metals—possible, that is, in the sense of the above theory,—so-called “medicines” are, according to Geber, requisite, these being distinguished as possessing different power and virtue. The medicines of the first order do indeed produce changes in the base metals, but these changes are not permanent. Those of the second order partially alter the properties of such metals into those of the noble ones,¹ but the transmutation proper is only effected by the medicine of the third order, which is variously designated as the *Philosopher's Stone*, the *Great Elixir*, or the *Magisterium* (masterpiece).² The accounts which Geber gives of the preparation of the medicines of higher order are wholly unintelligible; it should, however, be emphasised that there is a wide difference between these and the incredible exaggerations of which later alchemists were guilty, when speaking of the efficacy of such secret preparations.

The endeavour to produce the philosopher's stone and, with its aid, to prepare the noble metals, took possession of Geber's disciples, and also of the Arabian and Western alchemists, all of whom were firmly convinced of the reality of the transmutation of metals. Most of them appear to have had no occasion for combating the theoretical views of their great predecessor. The Arab physicians, mentioned at p. 29, taught the composition of the metals, restricting themselves closely to Geber's conceptions. Albertus Magnus did indeed deviate considerably from these, in so far that he assumed arsenic, sulphur, and water as the constituents of the metals; but in this he was not joined by his contemporaries Arnaldus Villanovanus and Raymund Lully, who adhered to Geber's opinions. Lully, in fact, did not hesitate to say in his *Testamentum* that all substances consisted of mercury and sulphur.

It seems surprising that the alchemists who were equipped with the most complete knowledge of their time should have remained satisfied with such speculations,

¹ The *Particulare* of the later alchemists appear to have corresponded to medicines of the second order.

² At a later period the great elixir was distinguished from the small one, which only transmuted the base metals into silver.

without attempting to prepare the substances whose presence they assumed in the metals and other bodies. Instead of obtaining an insight into the composition of these by actual experiment, they started new hypotheses to controvert obvious objections, such as, that these constituents (*e.g.* sulphur) were different from the substances commonly known under their names.

Geber's theory of the metals underwent an extension by Basil Valentine, who assumed the presence in them of a third constituent, viz. salt.¹ By the term *salt* he did not mean a definite chemical compound, such as common salt, but rather the principle of solidification and power of withstanding fire, just as sulphur determined the combustibility or change in the fire and also the colour, and mercury the metallic character and volatility. Basil Valentine generalised his opinion in this way, that he assumed these three essential principles in all substances, an assumption which Paracelsus appropriated later on, and made the basis of his iatro-chemical doctrine.

Their views upon the composition of the elementary bodies being so very obscure and so entirely wrong, one sees how it was impossible for the alchemists to give a correct explanation of chemical processes, connected as these are with the formation of compounds. Some very incomplete attempts were made to give a theoretical explanation of isolated observations, but these only led to the creeping in of the grossest errors; the calcination of the metals, for instance, was supposed to depend upon the escape of moisture or of some other constituent, an idea which reappeared in another form in the later theory of phlogiston. The above theory of the composition of metals is sufficient witness to the small amount of trouble which was taken to find out the true chemical constituents of bodies.

We may safely say that scientific chemistry only really began with the fruitful endeavours to discover the real composition of substances. It is out of the question to speak of this as applying to a time when it was considered as proved that

¹ Isaac Hollandus had before this spoken of the saline principle of the metals.

the formation of a chemical compound was identical with the annihilation of its original components, a new substance being created. This view was the almost sole predominating one during the later alchemistic period, although Geber himself gives some indications of more correct opinions on the composition of many chemical compounds (recognising mercury and sulphur, for instance, as constituents of cinnabar).

Contemporaneously with the holding of such theories, based upon no facts whatever, the Western alchemists strove in every imaginable way to obtain the philosopher's stone. Those of them who were in happy possession of the means for transmuting metals, attributed to it the most astounding powers. In order to give some idea of the aberration of mind of that time which was caused by the alchemistic problem, a few of the extraordinary assertions of well-known alchemists with regard to the preparation and efficacy of the philosopher's stone may be mentioned here.

For its preparation (we are now speaking more particularly of the thirteenth century onwards) a *materia prima* was requisite, the obtaining of which was the hardest task of all. The most incredible substances, natural products of every kind, were taken as raw materials for the manufacture of this preparation, and worked up in every conceivable way. Those who laid claim to the possession of the philosopher's stone took very good care to keep the secret of their *materia prima*. They described all kinds of operations with it¹ in the most enigmatical recipes, employing at the same time mystical drawings, such as those of the dragon, the red or green lion, the lily, the white swan, etc., and well knew how to keep their imitators, of whom there were formerly shoals (isolated cases being found even in this century), in a continual state of tension. That this was possible is explained by the immovable and almost universal belief in the transmutation of metals, by means of the philosopher's stone, during the Middle Ages.

To the latter the greatest miracles were ascribed; thus,

¹ The process of fixation, a term which indicated the solidification of mercury by the transmutation, was of especial importance.

Roger Bacon does not hesitate to say that it was able to transform a million times its weight of base metal into gold (*millies millia et ultra*). Others, *e.g.* Arnaldus Villanovanus, were more modest in their estimate of its powers, stating that it could convert into gold one hundred times its weight of mercury. Others, again, surpassed even Bacon, as the following passage from the *Testamentum Novissimum*, ascribed to Lully, proves: "Take of this precious medicine a small piece, as large as a bean. Throw it upon a thousand ounces of mercury, and this will be changed into a red powder. Put one ounce of the latter upon one thousand ounces of mercury, which will thereby be transformed into a red powder. Of this, again, one ounce thrown upon a thousand ounces mercury, will convert it entirely into medicine. Throw an ounce of this on a thousand ounces of fresh mercury, and it will likewise turn into medicine. Of this last medicine, throw once more one ounce upon a thousand ounces of mercury, and this will be entirely changed into gold, which is better than gold from the mines." One sees clearly, from these and other fraudulent assertions, that the simple standpoint which Geber assumed with regard to the question of the transmutation of metals, was departed from in the later Middle Ages.

In view of such excesses, which are an insult to the human understanding, it causes no surprise to find attributed to the philosopher's stone other results which are, if possible, even more incredible; health and life were to be preserved and ensured by it, as a universal medicine. Statements as to the power of prolonging life possessed by the elixir were also rife in the later Middle Ages, and it was no unusual assertion that adepts, the fortunate possessors of the panacea, had been able to prolong their lives to 400 years and more. The long lives of the patriarchs were explained by the assumption that they were acquainted with this universal medicine. In Geber's time healing properties were ascribed to gold prepared artificially and brought into the potable form, and from this the belief in the medicinal power of the philosopher's stone appears to have originated.

Alchemistic ideas produced their most absurd results towards the end of the Middle Ages and in still more recent times, the creation of living beings by means of the philosopher's stone being not merely held as possible, but being actually taught; this marks the summit of the mental aberration they induced.

The melancholy picture, which the condition of alchemy presents to us at various periods, becomes still more sombre in colour and involved in deeper shadow from the fact that men did not hesitate to affirm the Divine assistance, in order to explain the marvellous effects of the philosopher's stone. Gross abuse was made in this way of the name of the Deity, and also of prayers and biblical quotations, by the alchemists of the thirteenth century, and still more by their successors. There is no need for going into further details upon this point here, but it is necessary to mention it in order that the methods by which the problems of alchemy were treated at different periods may appear in their proper light.

Upon the development of chemistry as a science, the alchemistic doctrines—especially the theories of the composition of metals—had only a slight and an indirect influence. The excesses to which they gave rise have—as aberrations of mind, enchaining a large portion of the educated—a higher value for the history of civilisation than for that of chemistry. The main significance of alchemy for the latter lies in this,—that the endeavours to solve the problem of the transmutation of metals were the cause of actual work with materials of every kind; and the result of this was a not inconsiderable increase in the knowledge of applied chemistry during the alchemistic age. The following section will be devoted to an account of the latter.

*Practical-Chemical Knowledge in the Alchemistic Period.*¹

When one considers upon what superficial observations the conviction of the transmutability of metals was based,

¹ Cf. Kopp, *Gesch. d. Chemie*, vols. iii. and iv.; Höfer, *Histoire*, etc., vol. i. p. 317, *et seq.*; also Gmelin, *Gesch. d. Chemie*.

and how readily wholly untenable theories upon the composition of bodies were brought forward and accepted, one is not surprised that comparatively little progress was made, during the succeeding epochs, towards explaining the numerous chemical processes already known to the Ancients. The acquirements in chemistry during these centuries themselves likewise remained, for the most part, empirical; it was but seldom that the composition of chemical compounds was even in some degree correctly indicated. The fantastic treatment of chemistry—a treatment wholly foreign to the exact sciences—has been sufficiently detailed in the preceding section. We must not omit to mention, however, that the addition of new facts to those already known, and the gain of experiences in the fields of technical and pharmaceutical chemistry and in the manufacture of chemical preparations, were not inconsiderable.

Technical Chemistry.—Metallurgy, upon which the infant powers of an early developed *technique* were expended, shows, upon the whole, but little progress. Towards the end of the alchemistic period certain other metals were indeed added to those already known, viz. the semi-metal antimony, together with bismuth and zinc; but these can only lay claim to a subordinate position in the circle of metallurgical processes generally. From the eleventh century on, mining increased among the Western nations, in Germany especially in the Harz, Nassau, and Schlesien. So far as our present information goes, only trifling alterations were made in the preparation and purification of the metals.¹

Gold was obtained and purified from other metals and admixtures by the old method of cupellation (working with lead), already accurately described by Geber. The latter knew that the desired result was ensured and its progress hastened by the addition of saltpetre, and, further, that

¹ The work entitled *Schedula Diversarum Artium*, which was written by Theophilus Presbyter, a Benedictine of the eleventh century, gives a true picture of the state of technical industry in his time, particularly of the working-up of metals, something being also said about their production from the ores.

copper and tin, but not silver, could be separated from gold in this way. In the fifteenth century there was added to this the process of purifying gold by fusing it with antimony trisulphide ore (*Spießglanzerz*), a method which is given in detail by Basil Valentine.

The extraction of silver from its ores was accomplished, as in Pliny's time, by fusion with lead, an operation first termed "*Aussaigern*" by Basil Valentine. The only means of separating gold from silver which was known up to a comparatively recent date, was the cementation process of the Ancients. The wet process with nitric acid appears to have been first successful in the time of Albertus Magnus, at least he is the earliest to indicate it; an absolutely certain acquaintance with the process is first to be found in Agricola.

From the weight which was laid upon the successful working-up of gold and silver ores, one understands how at an early period the greatest attention was given to the definite quantitative yield of the noble metals. Accurate balances came into use, their employment in cupellation and cementation processes being made obligatory by law; one thus meets here with the first beginnings of a docimacy.

With regard to the metallurgy of iron, lead, tin, and copper in the alchemistic period, there are no particular improvements to record. Basil Valentine (fifteenth century) states that the last metal was also obtained by the wet process as the so-called cement copper, by precipitating a solution of copper vitriol with iron. The changes undergone by these metals on being heated and on treatment with chemical reagents, especially acids, were ardently studied, and thereby the knowledge of metallic preparations decidedly enlarged.

Mereury, which played such an important part in the theoretical views of the alchemists, was prepared on a large scale for technical purposes by roasting quicksilver ores in improved furnaces, especially after the opening up of the rich Idrian mines in the fifteenth century. The prepara-

tion of the metal by distilling a mixture of sublimate and caustic lime was well known to Basil Valentine. For its purification he gives various processes, some of which had been already described by Geber. Mercury was much used, particularly for the extraction of gold and silver (by the so-called amalgamation process) and for gilding.

Metallic zinc and bismuth, and also cobalt ore, are likewise mentioned by Basil Valentine, but the metals themselves do not seem to have been employed technically; some preparations of zinc, however, were. A special place among chemical preparations is to be assigned to antimony and its compounds, the knowledge of which is due to Basil Valentine himself.

In pottery and glass manufacture, important improvements in single points were made during the alchemistic period; but it is also noticeable here that the interest in the chemical processes remains a purely external one, no attempt being made to give a scientific explanation of the facts empirically arrived at. The general use of glazes containing lead and tin for earthenware vessels is worthy of mention, as is also the burning of colours into glass (the whole mass having formerly been coloured by the addition of metallic oxides during fusion).

Dyeing remained stationary on the whole, so far as the chemical media for fixing the colour on the fibre were concerned; alum was universally employed as a mordant, being manufactured on a large scale in different places. The introduction of the kermes dye (cochineal) into European countries by the Arabians, and that of orchilla (from the East in the thirteenth century), and, lastly, the gradual supplanting of the (blue) dye from woad by indigo, are the most important technico-chemical events in the domain of dyeing.

Condition of Pharmaceutical Chemistry.

Although the Arabians and the later Western *savants* busied themselves with chemical operations, and thereby arrived at preparations of the most various kinds, the pharma-

ceutical chemistry of that period only profited but slightly by this; it was Basil Valentine who inaugurated a new era by his bold attempts to apply chemical preparations to medicinal purposes. The opening up of the intimate connection existing between chemistry and medicine, which led to the high development of pharmacy, was reserved for the period of iatro-chemistry. The Arabians prepared their medicines strictly according to the recipes of Galen, Andromachus, and others, which were transmitted to them, according to Leo Africanus, by the Nestorians.¹ Apothecaries' shops, in which the remedies were almost exclusively prepared from vegetable substances, sprang up at an early date. To the Arabians belongs the credit of having improved and rendered the process of distillation serviceable for this purpose: distilled water, ethereal oils, and other products obtained by distillation (especially spirit of wine, to which the most wonderful results were ascribed), soon came into general use.

These apothecaries' shops, with their fittings, then spread into Spain, Southern Italy (into Salerno in the eleventh century), and, somewhat later, into Germany. The recipes of that time for the preparation of medicines, the imperfect pharmacopeias,² show that the doctrines and axioms of Galen and the Arabian physicians remained the standards up to the end of the fifteenth century. The position of the physician with regard to the apothecary was early fixed by legal statute, it being considered advisable to draw a sharp distinction between the man who had to prescribe the medicines and the man who had to make them.

With respect to chemical preparations proper, a few new ones were added to those already used in medicine, *e.g.* saltpetre, mercury in the form of grey ointment, and—towards the end of the fifteenth century, at the instigation of Basil Valentine—various mercurial and antimonial preparations. Almost all the physicians of that time took up, however, an antagonistic position with regard to the last-

¹ For their influence upon the Arabians see note 2, p. 27.

² The first German pharmacopeia (*Arzneibuch*) was drawn up by Ortholph von Baierland and appeared in 1477.

named, being of opinion that the undoubted poisonous properties of antimony compounds were incompatible with their internal use.

Knowledge of the Alchemists with regard to Chemical Compounds.

It has already been mentioned that the knowledge of the true composition of chemical compounds was but slightly increased during this period; we have therefore to deal here with the state of empirical knowledge as affecting substances prepared artificially, together with a few occurring naturally.

The tendency to group together observed facts under a common standpoint showed itself with respect to salts, of which a large number were known. Geber regarded solubility in water as a general characteristic; later on the generic name *sal* was made to include a variety of substances, e.g. the vitriols, potash, soda, saltpetre, alum, etc. Other chemical compounds of totally different nature, viz. the alkalies and acids, were added to the class of salts by many alchemistic writers, the term *sal* being thus widely extended and distorted; it was reserved for a later century to fix it without any ambiguity. In addition to the common designation *sal* for a number of heterogeneous bodies, we find in the writings of that time the generic name *spiritus* for the volatile acids, e.g. *spiritus salis* for hydrochloric acid; also the name *spiritus urinae* for volatile alkaline salt (carbonate of ammonia). The individual salts are distinguished by the word which follows *sal*, for instance, *sal petrae*, *sal maris*, etc.; for alkalies, such as caustic potash, the expression *nitrum alcalisatum* is frequently used. One but seldom meets in the alchemistic age with a strict distinction between potash and soda, or between their carbonates, while, on the other hand, preparations of carbonate of potash obtained in different ways were regarded as dissimilar products.¹

This acquaintance with the carbonates of soda and

¹ The salt from the ashes of plants was termed *sal vegetabile*, and that from tartar, *sal tartari*.

potash was accompanied by a knowledge of the lyes obtained from them by the addition of lime, the strongly alkaline and solvent power of these lyes being largely made use of, *e.g.* in the preparation of milk of sulphur, according to Geber's recipe. The name "alkali" is first met with in the writings of Geber, while the designation "caustic" had been already employed by Dioscorides for burnt lime, and at a later period for lyes. The question of the occurrence of alkalies in plants was frequently discussed among the alchemists; although it did not escape some of them that different amounts of ash and of alkali were found in different parts of a plant, only a few of them held the opinion that the alkali was really present in the plant itself, most of them believing that it was first produced during the incineration of the latter.

The Arabians possessed a very considerable knowledge of the acids, in comparison with that of the Ancients, who were totally unacquainted with the mineral acids. We already find Geber teaching in his treatise *De Inventione Veritatis* the method of obtaining nitric acid by distilling a mixture of saltpetre, copper vitriol, and alum in certain proportions: it was designated *aqua dissolutiva* or *aqua fortis*. Its preparation from saltpetre and sulphuric acid first became known to alchemists of a later date, but we find Basil Valentine speaking of it as a process which had been in operation for a long time.

Sulphuric acid was certainly obtained by Geber, for he mentions as noteworthy that when alum is strongly heated, a spirit distils over which possesses a high degree of solvent power; he does not, however, appear to have investigated its properties more closely. The writings of Basil Valentine show that the preparation of sulphuric acid by distilling a mixture of iron vitriol and pebbles, and also by setting fire to sulphur after the addition of saltpetre to it, was known not only to himself but also to his predecessors. An aqueous solution of sulphurous acid, the combustion product proper of sulphur, was frequently confounded with sulphuric acid.

Basil Valentine is the first to describe the preparation of aqueous hydrochloric acid, which he terms *spiritus salis*, by heating a mixture of common salt and green vitriol, and also its behaviour towards many of the metals and their oxides. He also knew that a mixture of this acid with *aqua fortis* was the so-called *aqua regis*, now termed *aqua regia*, which Geber had already made use of, obtaining it by the solution of salmiac in nitric acid.

Nitric acid and *aqua regia*¹ (so called because it dissolved gold, the king of metals) were highly prized by the alchemists of the West. The observation that almost nothing was able to withstand this *aqua regia*, even sulphur being "consumed" by it, strengthened the conviction that in it was possessed a liquid which very nearly approximated to the long-sought-for "alkahest," the universal solvent. On the same grounds oil of vitriol was greatly valued, many indeed regarding it as the *sulphur philosophorum*, or, at least, as a substance which would lead to the acquirement of the *materia prima*.

Among the salts which were already known in Pliny's time, and whose properties were carefully investigated by the alchemists, alum and some of the vitriols may be especially mentioned, the former being obtained in various places from alum shale. Geber tells us how to purify it by recrystallisation from water, and terms it *alumen de rocca* (from the name of its chief source, the town Rocca), a term which long remained in vogue in France as *alun de roche*. The fact that alum contained an alkaline salt was overlooked, and its true composition was not recognised. Iron and copper vitriols were largely employed in different chemical operations. Geber describes the preparation of the pure products by crystallisation, and Basil Valentine the production of the first-named by dissolving iron in sulphuric acid, a method which indicated the composition of the salt, although he did not explain the latter correctly.

The important salts, saltpetre, salmiac, and carbonate of ammonia, first became known and used for chemical purposes

¹ Albertus Magnus terms them respectively *aqua prima* and *aqua secunda*.

in the alchemistic period. Geber was well acquainted with potash saltpetre, as it served him for the preparation of nitric acid; and there is every reason to suppose that it was used in even earlier times for the production of fireworks and such like things, after its property of detlagrating with red-hot carbon had been recognised. The oldest designations for it were *sal petræ* and *sal petrosam*. Raymund Lully also termed it *sal nitri*, but distinguished between it and *nitrum*, the fixed alkali of the older writers: in the sixteenth century this latter word was converted into *natron*, while *nitrum* was applied to potash saltpetre. Although Geber had already observed the formation of saltpetre from nitric acid and potash, the composition of this salt was only explained correctly at a much later date.

The same applies to the term salmiac, *sal ammoniacum*, as to that of *nitrum*, in so far that both of them had originally a different meaning from what they now possess: for the *sal ammoniacum* of the Ancients was without doubt rock-salt. In the Latin translation of Geber's treatise, on the other hand, this name, which is also metamorphosed into *sal armeniacum* (Armenian salt), can only mean salmiac. The *sal armeniacum* of Basil Valentine led to the contraction salmiac. In Geber's time this salt appears to have been partly prepared from dung, and partly to have been found as a natural product of volcanic origin.

Carbonate of ammonia, well known to the alchemists of the thirteenth century as volatile alkaline salt (*spiritus urinae*), was obtained by distilling putrefied urine. Basil Valentine taught its preparation from salmiac and fixed (carbonated) alkali, a method which led a long time afterwards to the proper recognition of the composition of the salt. The use of these two ammonia compounds, just named, in pharmacy belongs to a later date.

The knowledge of the metallic salts was very decidedly increased during the alchemistic period. A special interest attached to a solution of gold in *aqua regia*, since from this *aurum potabile* the most wonderful medicinal effects were expected. Geber was the first to become acquainted with

nitrate of silver in the crystalline state, and to observe the precipitation of its solution by one of common salt, a reaction which came to be applied later as a test both for silver and for salt. The alchemists were also acquainted with the beautiful precipitation of metallic silver from a solution of its nitrate by means of mercury or copper.

Compounds of mercury attracted the interest of those who carried out chemical operations so far back as the time of Geber. The latter himself described the preparation of mercuric oxide by calcining the metal, and that of sublimate (mercuric chloride) by heating a mixture of mercury, common salt, alum, and saltpetre; he also taught how to prepare various amalgams.¹ Basil Valentine was acquainted with basic mercuric sulphate, and also with mercuric nitrate. Being an advocate of heroic treatment, he recommended the medicinal use both of the latter and of sublimate.

Preparations of zinc and bismuth (*e.g.* zinc vitriol) were well known towards the end of the fifteenth century, but detailed records are wanting both of their formation and their properties. Antimony and its compounds, on the other hand, were the object of unwearied labours on the part of Basil Valentine, as his treatise *Triumphwagen des Antimonii* ("Triumphal Car of Antimony") sufficiently testifies. He shows how to prepare antimony itself from the native sulphide (which was termed *antimonium* or *stibium* and was known to the Ancients), by fusing it with iron. In his treatise *Wiederholung des grossen Steins der Uralten Weisen* ("Recovery of the Great Stone of the Ancient Philosophers") he writes: "If one adds some iron to the fused *Spiessglas*,² there is produced by a particular manipulation a curious star, which the wise men before me called the signet star of philosophy." Basil Valentine was well aware that antimony did not possess the properties of a metal in full degree, and he regarded it as a variety of one, especially as a variety of

¹ This word is first found in the writings of Thomas Aquinas. The part played by amalgams in the transmutation of metals has been already considered.

² This designation of Basil Valentine's for native sulphide of antimony became altered later on into *Spiessglanz*.

lead; he sometimes talks of it as the *lead of antimony*. Even in his time antimony was employed for alloys, which served for the manufacture of printer's type, mirrors, and bells. It did not escape him, either, that *Spießglas* contained sulphur, and he was also acquainted with amorphous sulphide of antimony and *sulphur auratum* (a mixture of Sb_2S_3 and Sb_2S_5). He gives distinct recipes for the preparation of antimony trichloride (butter of antimony), of powder of algaroth (basic chloride of antimony), of antimony trioxide, and of potassic antimoniate, and there can hardly be a doubt that he recommended and applied those preparations for internal use. With regard to the composition of these, he only appears to have had a tolerably clear idea of that of the sulphide.

Arsenic, which is so closely allied chemically to antimony, and with whose sulphides the Ancients were acquainted, was first prepared by the Western alchemists in the thirteenth century; Basil Valentine regarded it as a "bastard metal" analogous to antimony. Arsenious acid is first distinctly spoken of by Geber, having been obtained by the roasting of realgar. From that time it was known as white arsenic, in contradistinction to the red and yellow varieties (realgar and orpiment). Its occurrence in the smoke from pyrites burners was also noticed by the observant Basil Valentine. Mention has already been made of the important part which was played in alchemistic operations by the property possessed by arsenic of turning copper white;¹ indeed, this contributed materially to the belief in the possibility of the transmutation of copper into silver.

In addition to the metallic oxides which have been already spoken of (those of mercury, antimony, etc.), and the early known oxide of copper and oxides of lead (PbO and Pb_2O_3), oxide of zinc and peroxide of iron may be especially mentioned. The former of these, which separated in woolly flakes when zinc was burnt, and which was therefore termed *lana philosophica*, appears to have been known to

¹ On account of this behaviour, Geber calls arsenic *medicina Venerem dealbans*.

Dioscorides, but it is in the alchemistic period that we first come across an intimate acquaintance with it. The Arabians were familiar with peroxide of iron in the different forms, red and yellow; the designation *colcothar*, for the ignited oxide, is to be found for the first time in Basil Valentine.

The theoretical significance which, from the time of Geber, was ascribed to sulphur as a constituent of the metals, and also of other bodies, leads to the question—How was the actual knowledge of this element and of its compounds acquired? The property of sulphur of dissolving in aqueous alkalies, and of being thrown down from such a solution as sulphur milk upon the addition of acids, is described by Geber in his treatise *De Inventione Veritatis*; the disappearance of sulphur, when acted upon by *aqua regia*, was likewise regarded as solution. Basil Valentine is the first to give definite details regarding flowers of sulphur, and also regarding the taking up of the element by many oils, a property upon which the preparation of sulphur balsam depended.

Mention has already been made of various sulphur compounds, the sulphides of mercury and antimony among others, which were the most valuable materials for the production not only of sulphur itself, but also of other bodies. These had already been grouped together as forming a particular variety of compounds, under the name of *marcasita* (Albertus Magnus), zinc blende, galena, and iron and copper pyrites being included among them. The peculiarity, which these substances had in common, of giving off a product of such characteristic odour as sulphurous acid when roasted, may not unlikely have formed the main reason for thus gathering them into one group. It must not be forgotten, however, that the artificial formation of several metallic sulphides from their components had been observed (*e.g.* that of cinnabar from quicksilver and sulphur by Geber), which may be supposed to have contributed materially to a knowledge of their composition.

In spite of many unequivocal observations to the con-

trary, people still held to the assumption that the metals and almost all other substances contained sulphur. Organic bodies, too, had to conform to this hypothesis; their real constituents remained hidden, no sharp general distinction being drawn between them and inorganic compounds. The meagre attempts made to explain the formation of organic substances, *e.g.* in fermentation processes, only give evidence of confused and untenable views. The organic preparations which were known in the alchemistic age were but few in number. Among them spirit of wine takes a prominent position, its manufacture being gradually simplified and improved after more perfect apparatus had been introduced by the Alexandrians. In accordance with its importance for medicinal and alchemistic purposes, it was usually termed *aqua vitae*, the name alcohol being first met with in Libavius (end of the sixteenth century). The preparation of concentrated spirit of wine—as an excellent solvent—by repeated distillation, and also by dehydration with fused potashes, was already known to Raymund Lully. To test its strength, Basil Valentine recommends that a portion be burnt, in order to see whether any water remains behind or not. The latter alchemist was also acquainted with various chemical transformations of alcohol, although he did not obtain the resulting compounds in a state of purity; among these were the production of common ether by the action of sulphuric acid, and of nitric and hydrochloric ethers by the action of nitric and hydrochloric acids respectively. By the “sweetening” (*Versüssung*) of alcohol is to be understood our term etherification. That alcohol is only formed during the various processes of fermentation, which yield wine, beer, and spirits, was not perceived even by the most distinguished observers of that time; its pre-existence in unfermented materials was thus taken for granted.

Increasing attention was likewise paid to the product of the acetic fermentation. Geber and the later alchemists taught how to concentrate vinegar by distillation, and they also prepared various salts of acetic acid, *e.g.* basic acetate and sugar of lead. Other organic acids, too, were noticed

in different plant juices, but they were mostly mistaken for acetic acid. The addition to the medical treasury of various resins and oils, especially ethereal oils, which were obtained from plants by distillation in improved apparatus, is no evidence of scientific progress; this really begins for organic chemistry with the discovery of methods for arriving at the composition of organic bodies.

The Fortunes of Alchemy during the last Four Centuries.

After the labours of Basil Valentine, and especially after the beginning of the iatro-chemical period, alchemy gradually became separated from chemistry, which was raising itself to the rank of a science. Although, therefore, a record of the alchemistic aims and errors of these last centuries does not properly come within the scope of a short history of chemistry, it cannot be passed over in complete silence; the justification for it lies in the important relations in which the most eminent chemists of the sixteenth and seventeenth centuries stood with regard to alchemy. The support given by such men to the latter undoubtedly accounts to a large extent for the belief in the transmutation of metals as an incontrovertible fact being but seldom affected, and this in spite of the great increase in chemical knowledge. Another effective means by which the life of alchemy was prolonged, consisted in the favour with which it was regarded by many princes; the seductive prospect of easily acquired treasure often rendered the latter a prey to the alchemists.

The actual decay of alchemy, for which the numberless disappointments of honest observers and the exposure of numerous frauds paved the way, may be dated from the first half of the eighteenth century, when the conviction of the practicability of transmuting metals began to die out among most chemists. Even up to the present century, however, we find able and educated men in the thralls of alchemistic chimeras, and directly opposing the simplest rules of reason.

A distinction must be drawn during the iatro-chemical

period between alchemists and chemists, inasmuch as the latter aimed at the solution of a scientific problem, viz. the knowledge of the relations between chemistry and medicine. At the same time this distinction must not be taken as meaning that the most eminent among the iatro-chemists were not firmly convinced that the ennobling of metals was a fact, indeed some of them maintained that they were in possession of the most powerful alchemistic specifics; it was but seldom, however, that chemists were at the same time practical alchemists.

Paracelsus, who inclined greatly to romantic exaggerations, claimed for himself the widest knowledge of alchemy. Van Helmont, whose authority was especially weighty, went so far as to describe in detail the transmutation of mercury into gold and silver, as effected by himself with the aid of a very small quantity of a gold- and silver-producing stone. The opinion held by the highly esteemed Libavius respecting alchemy and what it could effect is equally significant of the judgment of that period upon the subject: he regarded the transmutation of metals as an accomplished fact. Other influential physicians of the sixteenth century, such as Agricola—famed as an observant and accomplished metallurgist,—Sennert, and Angelus Sala, were more cautious in their assertions with respect to alchemy, but they never seriously contended against the possibility of transmutation. Tachenius alone, the last iatro-chemist of eminence, took up a sceptical position with regard to the alchemistic problem: he considered the evidence adduced in favour of the ennobling of metals as insufficient, notwithstanding that his famous teacher Sylvius had given himself up unreservedly to the belief in their transmutation.

The power of this belief was still so great at that time, when the phlogistic period was just beginning and chemistry was striving to develop itself independently, that it took firm root in the minds of even the most discerning men, with Boyle at their head. The latter was firmly convinced of the possibility of transmuting individual metals into one another, as were also many of his contemporaries and

successors, *e.g.* Glauber, Homberg, Kunkel, Stahl, and Boerhave, of whose earnest desire to arrive at the truth there can be no doubt whatever. That the wished-for goal was not attained in spite of the most untiring efforts, did not shake their belief in the correctness of the assumptions of alchemy; Stahl alone began to doubt these towards the end of his life, and warned his brethren against alchemistic frauds. The vitality of the belief in transmutation depended chiefly on the theoretical opinions which these men held regarding the composition of metals; the primal error of Geber and his disciples was thus propagated far into the alchemistic age.

Boerhave was the last distinguished chemist to support with his great authority some of the alchemistic ideas, while he did not criticise others of the fraudulent assertions with sufficient sharpness. After his time no notable exponent of chemistry—which had now attained to the rank of a science—spoke in their favour; but all the greater was the number of swindlers and cheats who cultivated the lucrative field of gold-making even during the eighteenth century. The conviction of the impossibility of transmutation, which was at that time establishing itself among scientific chemists, made its way but slowly into outer circles. Credulity, and the hope of obtaining riches for nothing, were the means of leading many into very doubtful paths even so late as the end of last century and the beginning of this one.¹ The final echoes of the alchemistic problem, which had for such a long period of time held the cultured of all nations in a state of tension, and had even blinded eminent scientific men, only appear to die away during the last decades of our own century.

Seeing the marvellous results which alchemy produced, it is natural to inquire more nearly into the supposed evidence in favour of the ennobling of metals, and to ask what kind of observations led to its being regarded as a

¹ For details on these points, especially for an account of the interesting relations of the Rosicrucians to alchemy, and of secret alchemistic associations, etc., see H. Kopp's *Die Alchemie in alterer und neuerer Zeit*, a book which gives us a clear insight into the workings of the alchemists.

matter of fact. If the greatest weight is to be laid upon the statements of men who had established their claim as practised observers, then the first place must be given to the records of the eminent physician and chemist van Helmont (towards the middle of the seventeenth century), respecting transmutation as carried out by himself; these records afford the most remarkable testimony to the power of alchemistic illusions. Van Helmont had received from an unknown source a small specimen of the philosopher's stone, and with this he states that he transformed several portions of mercury into pure gold, giving the exact proportions by weight; one part of this preparation sufficed to transmute 2000 parts of mercury.

Soon after the death of van Helmont, Helvetius, body-physician to the Prince of Orange, published a detailed account of the transmutation of lead into gold, by means of a trifling quantity of a preparation which had come to him from the hand of a stranger. It appeared impossible to doubt the testimony of such men, who were held in high esteem by all the scientific investigators of that time.

More palpable proof of the actual transmutation of metals was held to be furnished by the coins or ornaments prepared from alchemistic gold up to and in the eighteenth century.¹ The evidence, which came for the most part too late, that these consisted of worthless alloys (*e.g.* bronze gilt over), was all too soon forgotten. The findings of courts of justice, too, in favour of alchemistic operations, were looked upon as proofs of transmutation having been actually accomplished.

As has been already mentioned, a large number of German princes gave unremitting support to the efforts of the alchemists, being led thereto by the hope of large gains. Many of them worked zealously at transmutation themselves, among others John, Burggrave of Nurnberg, who received the surname of "the Alchemist"; the Emperor Rudolf II., the most powerful protector of the makers of gold: the Elector Augustus of Saxony, the Elector John George of Brandenburg, etc. etc. The courts of these princes were the field-

¹ Cf. H. Kopp's *Alchemie*, vol. i. p. 90, *et seq.*

grounds of adepts, who for long succeeded, by means of clever experiments, in maintaining a belief in their art among these Mæcenases, until, as usually happened, they were unmasked as cheats, and generally severely punished, after having been the cause of excessive expenditure on the part of their patrons.

It is impossible to enter here into details of the romantic lives of alchemists like Leonhard Thurneysser, physician at the court of John George of Brandenburg, Sendivogius, Caëtano (on whom the title of Count was bestowed), St. Germain, Cagliostro, etc. The two last named lived at a time when chemistry was strong enough as a science to protect itself against the frauds of alchemy. The opposition to the latter which was raised in the course of the preceding century by chemists of repute, *e.g.* Geoffroy the elder (the earlier warnings of Erasmus of Rotterdam, Athanasius Kircher, and Palissy having had no effect), led to its ultimate fall, which the amalgamation of alchemistic aims with those of the secret societies (Rosicrucians, Illuminates, etc.) could not retard. The belief in the possibility of the transmutation of metals received its actual deathblow from the new chemistry which began with Lavoisier.¹

The melancholy errors which arose from the introduction of the mystical religious element into alchemy can only be indicated here; the assertion, frequently made by adepts, that the secret of making gold was revealed to them through the grace of God, only excites feelings of repugnance.² Other frauds, which were likewise the products of alchemistic

¹ Schmieder, who published a history of alchemy in 1832 (in Halle), did not hesitate to accept the transmutation of metals as having been actually accomplished by various adepts. He expresses himself with more caution regarding the assumed efficacy of the philosopher's stone as a medicine and a means of prolonging life.

² Had such misuse of the name of God and of the Bible been made in the time of Luther, as was later the case, or had he been aware of it, his opinion of alchemy would have been a much lower one; as a matter of fact he valued it because of its bearing upon religious feeling. In contradistinction to this stands Melancthon's criticism of alchemy, a criticism which testifies to the sobriety of his judgment (he called it *imposturam quendam sophisticam*).

aims during the preceding centuries, only provoke satire; among these may be mentioned the endeavours to prepare from the air the so-called "substance of shooting stars" (the plant *Nostoc commune*, which is found in wet ground, was regarded as this), and the *materia prima* from "air-salt."

The benefits which have accrued to chemistry during the last four centuries from the mania for producing gold from the base metals, can only be estimated as very slight. It was but very seldom that a discovery of technical importance, like that of the making of porcelain by Botticher, sprang from alchemistic work. On the other hand, it did a great amount of harm during that period, for it crippled the usefulness of many eminent men who would undoubtedly have advanced science had they not been influenced by chimeras of an exciting nature: as it was, they were led away into the most tortuous paths.

We are thus forced to the above unfavourable criticism of the work of the alchemists on their problem of the transmutation of metals, in spite of the striking and seemingly incontestable evidence in favour of the latter: in spite, also, of a strong inclination at the present time to a belief in the mutual convertibility of elements chemically similar—a belief grounded upon speculations which do not seem to be without foundation. But in no single case, as yet, has there been any positive evidence brought forward in favour of this.

If we review the work of the alchemists during the last fifteen centuries, we arrive at the conclusion that it was based upon a series of falsely interpreted chemical problems. The expectations of the easy acquirement of boundless riches, the *auri sacra fames* to which it led, formed the powerful stimulus to the useless and yet continually renewed efforts of an unsatisfied mankind.

CHAPTER III

HISTORY OF THE IATRO-CHEMICAL PERIOD

INTRODUCTION.—Traditional belief, which dominated every branch of science during the Middle Ages, exercised its power not least in the domain of alchemy, for almost all those engaged in chemical pursuits were deluded by the idea that gold and other bodies could be artificially prepared. In the course of the fifteenth century, however, this yoke, which had hindered the development of free inquiry, was cast off in many quarters. The sciences, hitherto studied almost alone in the cloister, now found a foothold in the universities of France, England, Germany, and other countries, which were then both increasing in number and expanding rapidly; the free exchange of ideas among these rendered a development of the sciences possible, as it had never been before. That the discovery and spread of the art of printing contributed materially to this, hardly requires to be stated: for new ideas, which were opposed to those prevalent up till then, and which had hitherto been restricted to a narrow circle, became quickly disseminated by its aid. Any one could inform himself as to the range of any particular science by means of the encyclopedias and special memoirs which were being printed in increasing numbers. As a consequence of this, the capacity for independent criticism spread, one of the most effectual remedies against the domination of the scholastics being thereby created. A further aid to the controverting of scholastic principles was found in the inductive method, then gradually forcing itself forward,

by means of which experimental science was called into life.

In addition to these impulses of a freer spirit, chemistry received a powerful impetus from the increase in scientific knowledge which resulted from the discovery of the New World and of the ocean route to the East Indies. All these events testified to the birth of a new era, which found its most powerful expression in the works of the Reformation.

At that time chemistry strove to free itself from the exclusive domination of the alchemistic idea. Even although the latter was not completely supplanted, still another aim came into prominence, an aim to which a scientific character could not be denied; the chemical knowledge of that day was, however, so very imperfect that a solution of this new problem was not to be expected. Chemistry was, in fact, to be conjoined in the most intimate manner with medicine; each (so many opined) was to help the other. The chemist was to discover the medicines, prepare them carefully, and investigate them chemically, while the physician was to examine and explain their action. The mutual interaction of chemistry and medicine is the main idea which runs through the iatro-chemical age, and which gives to the latter its own particular stamp. What benefit, then, accrued to both of them from this? The answer is, a mutual enriching, which did almost more for chemistry than for medicine; for the former was thereby transferred from the hands of laboratory workers, who were mostly uneducated, to those of men belonging to a learned profession and possessing a high degree of scientific culture. The iatro-chemical age thus formed an important period of preparation for chemistry, a period during which the latter donned a portion of the armour, equipped with which she was enabled in the middle of the seventeenth century to stand forth as a young science by the side of her elder sister physics. That period was for chemistry an apprenticeship in the fullest sense of the word, during which she laboriously acquired the capacity to recognise the untenability of the iatro-chemical doctrines, and to apply herself to her true vocation.

GENERAL HISTORY OF THE IATRO-CHEMICAL PERIOD, IN
PARTICULAR OF ITS THEORETICAL VIEWS.

The main currents of the iatro-chemical age emanated from Paracelsus, van Helmont, and de le Boë Sylvius, with whose name must be coupled that of his most distinguished pupil, Tachenius, their doctrines being spread abroad by schools of greater or lesser importance. Besides these there were some men who worked independently, or who at least did not entirely subordinate themselves to their authority, of whom Libavius, Glauber, and Sala may be mentioned. Men like Agricola, Palissy, etc., employed their energies in a totally different direction, giving all their attention to technical chemistry.

Paracelsus and his School.—Paracelsus was the man who, in the first half of the sixteenth century, opened out new paths for chemistry and medicine by uniting them both. To him is undoubtedly due the merit of freeing chemistry from the restraining fetters of alchemy, by a clear definition of scientific aims. He taught that “the object of chemistry is not to make gold, but to prepare medicines.” True, chemical remedies had been employed now and again before his time, Basil Valentine in particular having suggested their use; but Paracelsus differed from his predecessor in the theoretical motives which led him to employ them. He regarded the healthy human body as a combination of certain chemical matters; when these underwent change in any way, illnesses resulted, and the latter could therefore only be cured by means of chemical medicines. The foregoing sentence contains the quintessence of Paracelsus’ doctrine; the principles of the old school of Galen were quite incompatible with it, these having—indeed—had nothing to do with chemistry.

Paracelsus entered the lists with great boldness and with a marvellous vigour, to combat the old doctrines long accepted by all physicians. Although his exaggerations here are distinctly to be condemned, still he effect-

ally obviated by his action the growing stagnation of medicine, and partly carried through valuable innovations, partly incited others to do so.

His career was not calculated to raise him in the esteem of his opponents, that is, of nearly all the physicians of the time. Paracelsus (his full name was Philippus Aureolus Paracelsus Theophrastus Bombastus) was born at Einsiedeln¹ in Switzerland in 1493, and returned to his native country about 1525 as a physician celebrated for his wonderful cures, after an exceedingly unsettled life and the most romantic wanderings in almost every country in Europe. The chair of Medical Science (therapeutics) at Basle was conferred upon him, and this position, together with his fame as a doctor, he made use of to spread abroad the iatro-chemical doctrine, and to fight against the old medical school with every possible dialectic weapon. He discredited the hitherto undisputed authority of Galen and Avicenna, and succeeded by means of popular lectures given in German, as well as by his rude originality, in gaining a large number of adherents. A quarrel with the Basle Municipal Council soon compelled him, however, to leave that town (in 1527), and after moving about restlessly in Alsace, Bavaria, Austria, and Switzerland, he came at last to Salzburg, where he died in 1541 in wretched circumstances.

There has at all times been much difference of opinion in criticising this gifted man, whose life offered such a rude contrast to his mental capacity. Rated too high and even extolled by his disciples,² and also by many who disapproved of his doctrines, he was, on the other hand, disparaged by his opponents and by chemists who criticised him as historians. The good to which he incited by his reforming labours seldom found the recognition it deserved, from its being so much mixed up with charlatanism and coarseness, while the over-

¹ The name *Eremita* (Hermit) which was given to him by many, recalled that of his native town (*einsiedeln* means "to live like a hermit").

² Cf. A. N. Scherer's memoir *Theophrastus Paracelsus* (St. Petersburg, 1821). Francis Bacon criticised him more reasonably, praising his endeavours to get at the truth through the light of experience.

weening estimation in which he held himself helped to make him ridiculous in the eyes of thoughtful physicians.

At the root of his iatro-chemical doctrines, which he imagined were grounded upon ample experience, lay the conception already mentioned,—that the operations which go on in the human body are chemical ones, and that the state of health depends upon the composition of the organs and the juices. With respect to the constituents of organic bodies, Paracelsus adhered to Basil Valentine's assumption that the latter were composed of mercury, sulphur, and salt. Indeed, in spite of many contradictions in the details of his theoretical views, this hypothesis forms the foundation of his whole system.¹ When one of these elements predominates or when it falls below its normal amount, illnesses ensue. This idea is expressed in the most fantastic manner in the writings of this strange man, as the following sentences show:—

An increase of the sulphur gives rise to fever and the plague, an increase of mercury to paralysis and depression, and an increase of salt to diarrhœa and dropsy. By the elimination of the sulphur, gout results, and by distilling it from one organ into another, delirium, and so on.—However unfounded such opinions are, it is possible to find a certain sense in them; on the other hand, his utterances upon the relations of the individual organs and the secretions of the human body to the metals and planets, to both of which he ascribes a mystical influence, are quite unintelligible. Not less incomprehensible is his assumption of a connection between the plague and shooting stars. He designates *tartarus* as the cause of various illnesses, meaning by this expression precipitates from juices which in the healthy state contain no solid particles. The deposition of concretionary matter, which he may have observed in the affected organs during many diseases (such as gout, stone in the kidneys, and gall-stones), no doubt led him to this partially sound conclusion. The comparison of

¹ Medicine rests, according to the confused statement of Paracelsus, upon four pillars, of which chemistry forms one; the three others are philosophy, astronomy, and virtue.

such secretions with known sediments, particularly with tartar, led to the general designation *tartarus*; the word had possibly also a double meaning, recalling the severe pains which people afflicted with these ailments had to endure.

While Paracelsus endeavoured in this semi-rational manner to reduce pathological processes to chemical causes, he assumed nevertheless for his iatro-chemical doctrine the action of particular forces in certain cases, which forces he, in his drastic manner, pictured to himself as personified. Digestion, in especial, was regulated by the action of *Archæus*, who—as a good genius—rendered the nutriment consumed digestible, effected the separation of indigestible matters, and provided generally for the preservation of a proper equilibrium. Diseases in the stomach were produced by *Archæus* becoming ill. In this interpretation of such a specific chemical process as digestion, Paracelsus was disloyal to his own principles. It fell to the later iatro-chemists to clear their doctrinal system from this incongruity.

Diseases were to be cured by medicines (*arcana*), the preparation of which, as we have already seen, was—according to Paracelsus—the aim of chemistry. Due recognition must be given here to the fact that this axiom infused new life into the effete medical doctrines: Paracelsus enriched medicine with a large number of valuable preparations. The manner in which he applied most of these must remain unknown to us; but it is certain that he effected numerous brilliant cures in cases of serious illness. With regard to the preparations which he used, we know that he was the first to stamp copper vitriol, corrosive sublimate, sugar of lead, and various antimony compounds as medicines, these metallic compounds having hitherto been looked upon with dread, on account of their poisonous properties. Further, he brought into use dilute sulphuric acid, “sweetened oil of vitriol” (sweetened by spirit of wine, and which was known at a later date as Haller’s acid), tinctures of iron, and iron saffron; and he also taught better methods for preparing and utilising various essences and

extracts. He appears to have attained great success by the judicious prescription of laudanum.

That Paracelsus gave a tremendous impetus to the higher development of the apothecary's calling by such large additions to the medical treasury goes without saying; for before his time apothecaries' shops were nothing more than stores for roots, herbs, syrups, and confections of every kind, the preparation of the latter being carried out exclusively in them. The making of new medicines presupposed an acquaintance with chemical facts and processes; pharmacists had therefore to be continually striving to attain to this knowledge, pharmacy, in the proper sense of the word, taking its beginning here. The service which Paracelsus rendered in instigating physicians and apothecaries to busy themselves with chemistry was a very great one, but Scherer goes too far when he says that "pharmacy owes everything to Paracelsus."

The trenchant innovations which Paracelsus strove to introduce gave rise to violent agitations among his contemporaries, agitations which were continually receiving new food from his numerous memoirs, circulated in various languages, and dating for the most part from the time after his departure from Basle. These gave frequent opportunities for vehement contradictions on the part of the old medical school. So far as their composition goes, his writings stand upon a very low level, faithfully reflecting as they do the unsettled life and the rude attitude of their author. Every one of them shows an illimitable self-conceit, many indeed being written in a manner quite unworthy of an educated man. His chemical knowledge and his views with regard to the origin of diseases are best seen in the following works:—*Archidora*; *De Tinctura Physicarum*; *De Morbis ex Tartaro Oriundis*; *Paramirum*; *Grosse Wundarznei*.

The consequences of the labours of Paracelsus were not long in manifesting themselves. His pupils, inspired by the new doctrines, glorified him as the reformer of medicine; while the adherents of the old school, on the other hand, resisted desperately the innovations and attacks

which undermined their views. A violent fight ensued and continued for a long time, until it was decided, if not in favour of Paracelsus, at least in that of the more moderate iatro-chemists. It does not lie within the scope of this work to enter minutely into these controversies, sufficing as it does to indicate here the signification of the new medico-chemical views for the development of chemistry. But we may mention that the Swiss physician Erastus (whose German name was Lieber), who remained faithful to the doctrines of Galen, was Paracelsus' chief opponent, and was especially instrumental in exposing the mischievous contradictions which were accumulated in his later writings. The medical world was agitated during the sixteenth century by polemical writings on both sides. Of the disciples of Paracelsus, who, less gifted than their master, paraphrased his ideas and imitated his less amiable peculiarities, especially his charlatanism, but who did not come up to him as scientists, Leonhard Thurneysser (called *zum Thurm*) was the best known. The latter achieved nothing of any note for chemistry, but his unsuccessful appearance as an adept ensures for him a place in the history of alchemy (cf. p. 61).

The acts of men of this calibre, who wrought immense mischief by the reckless use of poisonous preparations, render intelligible the attempts which were made to put a stop to their excesses by legal statute. This is seen, for instance, by the parliament of Paris prohibiting the prescribing of antimonial preparations, and by the sentence of condemnation which the medical faculty of Paris hurled against every attempted innovation in the healing art.

But there belonged also to the school of Paracelsus men of scientific eminence who did not concur in all his doctrines, but rather regarded them from a critical point of view, and who endeavoured in a rational manner to extract the good which they contained. The most prominent of these at the end of the sixteenth and beginning of the seventeenth century were Turquet de Mayerne and Libavius, Oswald Croll and Adrian van Mynsicht. These were for some time contemporaries

of van Helmont, and formed the connecting link between Paracelsus and that remarkable man. They greatly enriched not only medicine but also chemistry.

Turquet de Mayerne was born at Geneva in 1573, and practised as a physician of note in Paris. Holding, however, as he did, that the antimonial preparations now in ill-repute were necessary, and therefore prescribing them, he found it impossible to get on with his professional brethren in that city, and preferred to become body-physician to the King of England, in which country he died in 1655. His knowledge of chemistry was very highly developed for that age, as a consequence of which he laboured earnestly for the rational application of chemical remedies, without falling into the exaggerations of Paracelsus on the one hand, or rejecting all the medicines of the school of Galen on the other.

The physicians Croll and van Mynsicht busied themselves in a similar manner and at about the same time. Having a good knowledge of chemistry, they brought into vogue many of the medicaments of Paracelsus, together with other new preparations; among the latter, Croll was the first to recommend the use of sulphate of potash (K_2SO_4) and of volatile salt of amber (succinic acid), and van Mynsicht that of tartar emetic.

Andreas Libavius (Libau) attracts our attention in a high degree by the critical position which he took up with regard to the gross errors of the school of Paracelsus, and especially also by many new observations which he contributed to chemistry. He was the first chemist of note in Germany who stood up energetically against the excesses of Paracelsus, and who vigorously combated the defects in his doctrines, the obscurities in his writings, his phantasies and sophisms, and the employment of "secret remedies." Originally a physician, Libavius attained to a wide knowledge of chemistry, which he helped to extend, although latterly he devoted himself chiefly to historical and philological studies. He died in 1616 as director of the gymnasium at Coburg. Thanks to his medical knowledge and to his thorough general

education, Libavius was able to appreciate better than his contemporaries the influence which chemistry ought to exercise upon medicine; he took up a position midway between those of Paracelsus and his opponents, the latter of whom wished nothing less than to banish chemistry from medical science. Notwithstanding his sound judgment, however, of which he gave many proofs, he could not quite free himself from the predilection of his time towards alchemy.

Libavius did chemistry a real service in writing his text-book, which under the title *Alchymia* was published in 1595, and which contained all the most important facts and theories germane to the subject at that date. His other writings, in which he combated the weak points of the Paracelsian school (as indicated above), and also described new chemical observations, appeared in three volumes shortly before his death, under the title *Opera Omnia Medico-chymica*. We shall still have frequent occasion to refer to his practical chemical knowledge, which was attested by the discovery of important facts.

Van Helmont and his Contemporaries.

A distinguished place and a detailed description in the history of the iatro-chemical period is due to van Helmont, one of the most eminent and independent chemists of his time. Endowed with rich acquirements and experiences in medicine and chemistry, he surpassed those of his contemporaries who worked in the same field. His life was for the most part that of a scholar working in quiet, although his brilliant outward circumstances (he belonged to a noble Brabantine family) were not in accordance with this. Born in Brussels in the year 1577, he applied himself at an unusually early age to the study of philosophy and theology; but not finding satisfaction in these, he renounced them to devote himself to medicine. He was at first an adherent of the old school of the Galenites, but soon recognised its deficiencies and turned to the doctrines of

Paracelsus, accepting them, however, only in part. With a growing enthusiasm for his physician's calling, he fought against the old medical system, and materially contributed by his brilliant services in bringing about its fall. Without van Helmont, iatro-chemistry would never have attained to the height to which it was raised later on by Sylvius and Tachenius. In addition, he enriched pure chemistry by a very great number of valuable observations. So attached did he become to his scientific pursuits that he declined the tempting offers of princes, preferring to investigate the secrets of nature in his laboratory at Brussels, in which city he died in 1644.

In van Helmont wonderful contradictions were united in harmonious action. In contrast with his gift of sharp and temperate observation was an intense inclination towards the supernatural,—possibly the result of his mystical and magical studies, to which, as well as to theology, he had applied himself. Thus this same man, who laid the foundation of the first knowledge of gases, and showed thereby a keenness of perception unapproached before his time by any observer, defended the transmutation of the base metals into gold with his utmost vigour (cf. p. 58); his belief in this was grounded so firmly that illusions arose from it which are to us incomprehensible.

After this it is easy to understand that van Helmont was not free from fantastic conceptions of a less questionable nature. His theoretical views upon the elements and his iatro-chemical doctrines yield many proofs of this; but, on the other hand, much of his knowledge was so sound, and he was able to expound it so much better than any of his predecessors, that the good service which he rendered far outweighed the bad influence of any of his mistakes.

Van Helmont had his own opinion with regard to the primary substances of which matter was composed; he neither accepted all the four Aristotelian elements¹ nor

¹ With respect to air, it is not certain whether van Helmont looked upon it as an element or not. He denied altogether that fire could be of a material nature, which is evidence of his extraordinary clearness of perception.

those which were assumed by Basil Valentine, but looked upon water as the chief constituent of all matter. That it was present in organic bodies he concluded from the fact of invariably finding it as a product of their combustion. He imagined that he contributed a strong proof of this by an experiment which showed that plants could be made to grow luxuriantly in pure water alone, which, he believed, was their only nutriment under the circumstances. That he was thereby convinced of the transformation of water into earthy matter is therefore quite intelligible.

Whilst van Helmont thus subscribed to the same error that held possession of many minds both before and after his time, he nevertheless recognised in numerous instances the unchangeableness of matter much more clearly than his contemporaries; thus he contributed more than any one else to do away with the belief that the copper thrown down from a solution of copper vitriol by means of iron was newly created. He further showed that the same substance continued to exist in many of its compounds, *e.g.* silver in its salts, and silica in water glass, the latter yielding, on decomposition with acids (according to his memorable observations), the same amount of silicic acid as was originally used to prepare it. These were views and observations of the greatest importance; for, in place of the former obscure conceptions as to the formation of chemical compounds, he substituted the doctrine that the original substance, even after undergoing chemical changes, remains present in the new products. He had therefore clearly grasped the fundamental law of the conservation of matter in particular cases.

Van Helmont thus stands out as unique in those ideas, which pointed out new paths to chemistry. The relations between chemistry, and medicine too, the latter of which he also ardently fostered, led him to views which likewise possess a partial originality, since he endeavoured to decide theoretical questions by means of experiments with juices and other secretions of the animal body. The reactions which go on in the liquids of the body were in his opinion of especial importance, for, according as the latter were acid or neutral,

they regulated its most important functions. Besides the chemical nature of the juices, fermentation was, according to him, the principal cause of the organic processes; but he expresses himself less clearly upon this point than upon the significance of the chemical reactions. Indeed, he could not quite free himself from the idea of *Archeus* governing digestion and the processes connected with it. On the other hand, he stood on solid ground in his explanation of vital processes, when he took into account the chemical nature of the juices. He held that the acid of the gastric juice brought about digestion, but this, if present in excess, gave rise to discomfort and illnesses, which were the more serious the more acid there was; and the latter could not then, as under normal conditions, be neutralised by the alkali of the bile, which mixes with the gastric juice in the duodenum. To cure any of the ailments produced in this way, van Helmont declared that medicines of an alkaline nature (alkaline salts) must be used: while those of an opposite kind, which arose from a deficiency of acid, were to be treated by medicines of an acid nature. He also recommended the latter in cases of gout, stone, and similar diseases, which likewise originated (in his opinion) from an insufficient admixture of the juices. These views show a distinct advance upon those of Paracelsus. For, while the latter assumed the presence of arbitrary constituents—incapable of preparation—in organic matter, van Helmont searched for the actual substances themselves, and compared the interactions of the various juices which mingle with one another with similar reactions of solutions outside the organs; a procedure which laid the first foundation, however insecure, of chemical physiology.

Van Helmont proved himself an original investigator of the first rank, who opened out new paths for chemical science by his researches on gases—researches which constitute him the real founder of pneumatic chemistry; this indeed only attained to a considerable development a century after his time, when the discoveries connected with it brought about the great reform of the science. If we consider

that before van Helmont's labours the most various gases, such as hydrogen, carbonic acid, and sulphurous acid, were not looked upon as differing materially from ordinary air, and that he was the first to characterise gaseous substances as different, by investigating their properties, we gain some idea of the immense services which he rendered. He it was who gave to them the generic name of "gas,"¹ and he further distinguished them from vapours, in so far that the latter were condensed to liquids upon cooling, while the former were not.

Van Helmont specially examined carbonic acid, and showed how it can be produced from limestone or potashes with acids, from burning coal, and in the fermentation of wine and beer; he also pointed out its presence in the stomach, and its occurrence in mineral waters and in many natural cavities in the earth. He usually termed it *gas sylvestre*.² To the want of suitable apparatus for collecting gases are to be ascribed the imperfections in many of his observations, and also the confounding of carbonic acid with other gases which were non-supporters of combustion like itself; nevertheless he described the two combustible gases—hydrogen and marsh gas—as peculiar varieties of air. His collected works were published in 1648 by his son under the title, *Ortus Medicinarum vel Opera et Opuscula Omnia*.

Van Helmont's influence upon his contemporaries and upon the development of the iatro-chemical doctrines must be rated very high. By his introduction of chemical ideas into medical science, the latter was advanced, because the use of chemical medicines seemed natural from thenceforth; moreover, in his *Pharmacopolum ac Dispensatorium Modernum*, he published suitable prescriptions for the preparation of medicines. The scientific spirit which he

¹ In choosing this designation, van Helmont had Chaos in his mind,—possibly also the process of fermentation (the Dutch word for the verb "to ferment" is *gisten*).

² By the designation *gas sylvestre*, he doubtless meant to indicate the impossibility of condensing the gas; at least he says in one passage: *Gas sylvestre, sive incöercibile, quod in corpus cogi non potest visibile*.

endeavoured to introduce into the healing art tended to its more healthy development, in contrast with the crude empiricism of the Paracelsian school.

In a similar manner, if in lesser degree, various other physicians of that time were also active. Well equipped with chemical knowledge, they pursued the practice of their calling, and were enabled by their clearness of vision to recognise and combat many evils, *e.g.* those which arose from the use of secret remedies: among them we must mention Angelus Sala and Daniel Sennert. Sala, who practised as body-physician at the Mecklenburg Court in the first half of the seventeenth century, surprises us by his able criticisms both of the Paracelsian and of the old medical schools, and also by his (for that time) wide knowledge of chemistry. This knowledge, conjoined with his solid medical experience, was of the utmost value not only to pharmacy but also to pure chemistry; for he formed correct ideas with regard to the composition and reactions of many chemical compounds, such as had never been advanced before his time. Thus he tells us that salmiac consists of hydrochloric acid and carbonate of ammonia (*flüchtiges Laugensalz*), and he also knew that sulphuric acid was able to drive out nitric acid from its salts,—and so on.

Sennert, who taught as professor at Wittenberg in the first quarter of the seventeenth century, chiefly devoted his energies to proving to the medical world the wonderful efficacy of chemical remedies, when these were properly applied. It is true that he could never disentangle himself from many of the erroneous conceptions of Paracelsus, for instance, from the doctrine of the three primary elements; but he worked effectively against the serious abuses which had crept into medicine through the influence of the last named, especially against the so-called universal remedies.

Sylvius and Tachenius.—F. de le Boë (Dubois) Sylvius was born at Hanau in 1614, and, after a thorough grounding in scientific and medical studies, practised at first

as a physician of note, and later on, until his death in 1672, was famous as professor of medical science in Leyden. In his knowledge of medicine he far surpassed most of his contemporaries. He was aware of the difference between arterial and venous blood, and ascribed the red colour of the former to the air absorbed in breathing. Combustion and respiration were in his view precisely similar processes. He directed all his efforts, as instanced in this latter case, to proving that the processes which go on in the human body—whether they be normal or pathological—were purely chemical ones. The spiritualistic element which was mingled with the doctrines of Paracelsus and van Helmont was to be entirely set aside. Digestion, for instance, which only appeared possible to the two latter by the intervention of a spirit (*Archæus*), was regarded by Sylvius as a chemical process in which the saliva primarily, but also the gastric and pancreatic juices and the bile, were the most important acting agents. To the acid, alkaline, or neutral reactions of the juices of the body he ascribed an equal, if not a higher, significance than van Helmont himself, following the latter in this as in similar questions. Sylvius had a predilection for comparing chemical with physiological and pathological processes, which frequently led him into error. Medicine as a whole, he considered, ought simply to be applied chemistry. That these one-sided endeavours were bound to miscarry completely, considering the state of chemical knowledge at that time, requires no demonstration. And it is equally easy to comprehend why his chemical doctrines brought less benefit to medicine than to chemistry, seeing that educated physicians, if they wished to understand them, were compelled to go minutely into the study of chemical questions. This applied in a very especial degree to the new remedies, the preparation and rational application of which presupposed a knowledge of chemistry. Sylvius, addicted as he was to the use of heroic medicines, did not hesitate to prescribe *lapis infernalis* (nitrate of silver), sublimate, and zinc vitriol for internal use; and he was particularly enthusiastic about antimonial and mercurial preparations.

While there are but few discoveries in pure chemistry by Sylvius himself to chronicle, his pupil Otto Tachenius proved himself an independent investigator, to whom the science is indebted both for extremely valuable observations and for speculations deduced from these. Of his life we only know that he was born at Herford in Westphalia, and that, after moving about from place to place as an apothecary's assistant, he applied himself to the study of medicine in Italy towards the middle of the seventeenth century, and practised in Venice as a physician. Although he attached the greatest weight to clear relations between chemistry and medicine, he had no hesitation in working mischief with secret remedies. Tachenius was the last iatro-chemist of note who followed the doctrines of Sylvius with enthusiasm. In addition to him may be mentioned here the famous English physician Willis (*ob.* 1675), who likewise advocated similar opinions.

Tachenius, among his other valuable observations, contributed materially to elucidating that problem which Boyle considered the most important of all, viz. a knowledge of the composition of bodies. It was with him that the first serviceable definition of the term "salt," as a compound of an acid and an alkali, originated. His statements on the composition of various compounds show great acuteness, which is also seen in the value which he laid upon certain reactions as tests for different substances. While Tachenius thus laid the foundations of qualitative analysis in a more systematic manner than his predecessors, his attention was also directed to the proportions by weight in which substances react chemically—a point to which hardly any attention had hitherto been paid; and this he exemplified with tolerable accuracy by noting the increase in weight which took place when lead was transformed into minium. His writings, and also those of his master Sylvius, treat for the most part of subjects chiefly of medical interest, but, as we have just seen, facts and opinions of importance to chemistry are also recorded in them.

If we wish to arrive at the main result which the iatro-

chemical doctrines produced upon the development of chemistry, we must bear in mind particularly the point already touched upon, viz. that the study of chemistry by physicians who had had a thorough education helped materially to shape its course on scientific lines. Notwithstanding the numerous errors and fantastic conceptions in which the iatro-chemists were involved, we come across many very striking views,—views which exercised a marked influence upon the whole tendency of the succeeding epoch. Of these we would mention here: (1) the recognition of the more intimate components of salts, and the clearer comprehension of what was meant by the terms “chemical compound” and “chemical affinity,” by a knowledge of which the chief aim of chemistry, *i.e.* the investigation of the true composition of bodies, was effectively advanced; and (2) the recognition of the analogy between the processes of combustion and the calcination of the metals on the one hand, and respiration on the other. These were doctrines of very great weight indeed. The phlogistic hypothesis, too, which predominated during the greater portion of the eighteenth century, was indicated by many of the iatro-chemists. Lastly, van Helmont’s work upon gases exercised the greatest influence on the development of pneumatic chemistry, from which the impulse to the great reform of our science at the end of last century sprang.

It is thus evident that many of the aims of the phlogistonists were intimately connected with the observations and opinions proper of the iatro-chemists. And while the medico-chemical opinions of the latter were rudely upset after the middle of the seventeenth century, their facts and theories appertaining to chemistry were the means of guiding it along scientific paths.

*Georgius Agricola and the other Promoters of Applied
Chemistry during the Iatro-chemical Age.*¹

Independently of the main iatro-chemical current, chem-

¹ Cf. Kopp, *Gesch. d. Chem.*, vol. i. pp. 104, 128; and Höfer, *Histoire de la Chimie*, vol. ii. pp. 38, 68 *et seq.*

istry in its applications to industries was fostered by men who possessed, for their time, sound chemical knowledge. The chief of these were Georgius Agricola, who directed his attention specially to metallurgy; Bernard Palissy, who developed the ceramic art; and Johann Rudolf Glauber, who, without ceasing to be an iatro-chemist, devoted his powers for the most part to technical chemistry. The following paragraph gives a few details explanatory of the services rendered to the science by the knowledge and experiences of those men; but what we are chiefly concerned with here is their significance from a more general point of view.

Georgius Agricola was born at Glauchau in 1494, and died, while a physician of note and Mayor of Chemnitz, in 1555; he was thus a contemporary of Paracelsus. Although, like the latter, a physician, he followed totally different paths. Without troubling himself about the storms which raged round medicine in his time, he devoted himself by choice to the study of mineralogy and metallurgy, being impelled thereto by the flourishing mining and smelting industry of Saxony, while at the same time he continued to practise as a physician. His chemical knowledge and wide experiences are detailed by him in his principal work: *De Re Metallica, libri XII*, which remained for a long time the most important text-book of mineralogy. Through this, as well as through his other writings—of which *De Natura Fossilium, libri X*, and *De Ortu et Causis Subterraneorum* were also of especial mineralogical value,—there runs quite a different tone from what we find in Paracelsus. They are characterised by a clearness of expression, a temperate conception of the operations described, and a distinct description both of the apparatus employed and the processes followed,—qualities which stamp Agricola as a true investigator. It was through his writings, especially through the first of those named above, that the more important operations in the working up of ores for their metals first became generally known; and he was likewise the first to explain intelligibly the manufacture of other products obtained by smelting, and of various preparations of technical importance.

His quiet objective modes of thought and investigation did not, however, prevent him in his more mature age from attributing a certain degree of likelihood to the alchemistic problem, to which he had devoted himself warmly in his youth; at the same time he had no sympathy with the wild exaggerations which even then prevailed.

Working on lines similar to those of Agricola, and at about the same period, the Italian Biringuccio of Sienna busied himself with the processes of metallurgy, as detailed in his work *Pirotechnia*, which appeared in 1540. This too is marked by the clearness and exactitude with which various technical procedures are described. Biringuccio held aloof from the iatro-chemical questions and the alchemistic doctrines of his day.

Bernard Palissy became distinguished as a prominent investigator, and one who allowed himself to be guided solely by the results of experiment, at a time before the inductive method was commonly recognised as the means of attaining to the truth. It was in the domain of ceramic art that his principal work lay; and, although frequently disappointed in the results he obtained, his untiring efforts at improvement in it were ultimately followed by success. The simple and clearly written works of Palissy enable us to appreciate the labours and struggles of this remarkable and steadfast man, who, beginning as a common potter destitute of the higher education, became an eminent authority on his subject.¹ He took his first lessons from the book of nature, as he himself tells us; ² putting observation and experiment in the foreground, he combated every speculation which was not based upon these, especially such doctrines as had merely the stamp of authority to back them. There could hardly have been any man of his time more free from prejudice; his clear understanding and circumspect criticism enabled him to cast aside the doctrines of Paracelsus, and

¹ Hüfer, who was the first to recognise the services of Palissy as they deserved, speaks of him as "*un des plus grands hommes, dont la France puisse s'enorgueillir*" (*Histoire de la Chimie*, vol. ii. p. 92).

² "*Je n'ai point eu d'autre livre, que le ciel et la terre, lequel est connu de tous, et est donné à tous de connoître et lire ce beau livre.*"

to make use of the weapons of ridicule against the mistaken beliefs of alchemy. His life extended over nearly the whole of the sixteenth century, and might be said to consist of a series of vicissitudes. Along with Agricola he may be looked upon as the chief exponent of experimental chemistry in his time. His acuteness was further evidenced in the domains of mineralogy and agricultural chemistry, to the founding of which branches of science he largely contributed.

Johann Rudolf Glauber, who was born at Karlstadt in Franken (Bavaria) in 1604, and who died in 1668 at Amsterdam, fostered applied chemistry ardently, and enriched it by valuable observations. It was in this direction that he chiefly worked, his iatro-chemical labours holding but a secondary place. His life was an extremely restless one, which may not improbably account for the unsettled and almost discontented tone which runs through many of his writings. Without a classical education, and imbued with the prejudices of his age, he has been well designated the Paracelsus of the seventeenth century. He was, in fact, addicted to fantastic and superstitious ideas, and therefore also to the extravagances of alchemy; on the other hand, he possessed exceptional talents for observation and invention, regarding which some details will be given in the next section of this book. In theoretical points of chemistry, too, he gave proof of his clear-sightedness, explaining, for example, many of the effects of chemical affinity in the decomposition of salts by acids or bases,—and so on. He was the first to explain a case of what we call double decomposition,—the mutual action of mercuric chloride and antimony trisulphide upon one another. Mention must also be made here of his perspicacity in questions of national economy, his writings upon which are to be found mixed up with his chemical papers, especially in the six-volume work *Teutschland's Wohlfarth* ("The Weal of Germany"). Time after time Glauber sought to demonstrate that his country should work up and improve its own products, and not leave this for other nations to do; instead of buying at a dear rate manufactured articles whose raw material was

obtained from Germany, that country ought to make and export them herself.

With Glauber and Tachenius the iatro-chemical period closes. Both of these belonged in many of their chemical ideas, and also in point of time (during the last years of their lives), to the succeeding era, between which and the previous one it is impossible to draw an absolutely sharp line. Both aided chemistry by observations of extreme value, and materially advanced the experimental method, which became from thenceforth the sure guiding star of chemical research.

EXTENSION OF PRACTICAL CHEMICAL KNOWLEDGE IN THE IATRO-CHEMICAL AGE.¹

As was to be expected from the whole tendency of this period, during which chemistry became so intimately united with medicine, the gain of knowledge lay chiefly in the domain of chemical preparations, which it was hoped to apply as medicines. The endeavours to discover new remedies had the effect of causing chemical compounds, whether new or already known, to be investigated more carefully and scientifically than had ever been done before. The products of the animal body were zealously studied, and a small beginning was made in physiological chemistry by the examination of milk, blood, saliva, etc. And this in its turn increased the interest felt in organic compounds, and the knowledge of these thereby acquired. In technical chemistry less progress was made than in chemistry which was related to medicine. An advance in the knowledge of the composition of substances and in the observation of reactions, *i.e.* in qualitative analysis, first became noticeable towards the end of the iatro-chemical period.

Technical Chemistry.—The most eminent exponents in this direction, chief among whom were Agricola and Palissy, have already been referred to. In their works, as also in

¹ Cf. Kopp, *Gesch. d. Chemie*, vol. ii. pp. 111, 126; vols. iii. and iv. *passim*.

the writings of Birringuccio, Cæsalpin, and others, which are devoted to technical chemistry, special importance is laid on the particular operations by which technical products are obtained, these operations being minutely described.

In Metallurgy Agricola and Libavius were the first to point out a method by means of which it was possible to estimate approximately the amount of metal in an ore; the science of testing then gradually developed itself from such beginnings. The more scientific treatment of applied chemistry is further shown by the fact that by-products began to be used which had previously been neglected, *e.g.* the sulphur which escaped during the partial roasting of pyrites was condensed, the tutty from zinc ores was worked up for brass, and so on.

A knowledge of the individual metals, and how they could be obtained and worked up, became extended in the sixteenth century by Agricola and other authors making into common property what had hitherto been only known to the few; *e.g.* the separation of gold from silver by means of nitric acid, which was first carried out on a large scale in Venice towards the end of the fifteenth century, and the amalgamation process, first applied in Mexico about the middle of the sixteenth century for extracting silver from its ores, but only introduced into Europe towards the end of the eighteenth. It is in the sixteenth century that we find the first reliable observations on the production of ruby glass by means of gold. Salts of the latter metal and also of silver were more carefully investigated, with reference particularly to their medical application; and some of their characteristic reactions—by which it became possible to distinguish them from other substances—were also noticed.

With respect to copper and its precipitation from a solution of copper vitriol by means of iron, we find even chemists of discernment like Libavius holding fast to the old idea that a transmutation had occurred; but others, *e.g.* van Helmont and Sala, recognised the pre-existence of the copper. The metallurgical operations necessary for obtaining iron became generally known through Agricola's writings,

thus the production of steel by the puddling process was first described by him. Steel was at that time regarded as a very pure iron. Of the other metals, a knowledge of zinc and bismuth was gradually acquired, although there was often uncertainty about them, and they were frequently confounded with antimony. Tin, lastly, was much used in the sixteenth century for tinning iron. But the iatro-chemical age interested itself less in the metals themselves than in the salts prepared from them, since there was always the chance of these proving useful in medicine. (See under *Preparations*.)

Pottery and Glass Manufacture.—The ceramic industry in particular made considerable progress through the untiring efforts of Palissy; his only teacher was the experience gained from innumerable trials, but he succeeded in affixing beautiful and durable enamels on earthenware vessels, especially on those of Fayence pottery. His observations on this point, and also on the application of different clays for ceramic purposes, and the burning-in of colours, are given in his work *L'Art de Terre*, which at the same time aims at showing the value of the experimental method as opposed to theory alone. Porta was also busy about the middle of the sixteenth century with work similar to Palissy's.

The manufacture of glass did not lag behind that of pottery. From the Venetian factories, whose sixteenth-century productions still astonish and delight the connoisseur, the art of making glass of the most various colours and of different degrees of refrangibility spread to other countries. The work of the Florentine Antonio Neri, entitled *De Arte Vitraria*, which appeared in 1640, not improbably contributed materially to spreading a knowledge of special operations, his rich experiences on the subject being detailed in this book. Great skill was also attained to even at that date in the imitation of precious stones, as Porta's recipes show. One of the most important discoveries of the time was that of cobalt blue by Christoph Schurer, a Saxon glass-blower, who obtained it on fusing the cobalteous

residue from the manufacture of bismuth with glass; it soon became a much-prized article of commerce, being known under the names *zaffre* (from sapphire), and, later on, *smalt*.

Dyeing.—One of the results of the discovery of America and of the ocean route to the East Indies was seen in the increased importation of indigo and cochineal, which gave a fresh impetus to the dyeing industry. Many improved methods of fixing these and other colours upon cloth—*e.g.* the use of a solution of tin, the judicious mordanting of the stuffs with alum, iron solution, etc.—were found out in the sixteenth century. The dyer of that time might consult the first text-book on this subject, written by the Venetian Rosetti, which appeared in 1540. Glauber, too, made numerous observations on dyeing processes, and aided not a little in advancing a knowledge of these.

A new industry sprang up towards the end of the fifteenth century in the rapidly extending distillation of brandy; up to that time spirit of wine was looked upon as a medicine only, but now it began to be more and more used, sufficiently diluted, as a drink. The development of this branch of trade resulted in great improvements in distilling apparatus, which also came to be of service in laboratories. The interest which this industry excited is seen from the numerous works upon the art of the distiller, which appeared at that time.

Chemistry extended its applications in the most various directions, among others to agriculture, if only in a modest degree: thus we find the gifted Palissy calling attention to the importance of soluble salts in manures, and recommending the addition of mineral substances, *e.g.* marl, to these.

Development of Pharmacy and of the Knowledge of Chemical Preparations.

Pharmaceutical chemistry is most distinctly a creation of the iatro-chemical age, during which it was taught that

the chief aim of chemistry lay in the discovery of medicines which could be prepared artificially. In accordance with this dictum, not only were the preparations which were already known, but also those others which had been newly discovered by assiduous application, tested for their action upon the organism. The circle of chemical facts was thus greatly enlarged by these iatro-chemical labours. The influence of the latter upon chemistry was made further apparent by the fact of the drug-shops, in which artificial preparations were made, becoming the nurseries of energetic chemists, who, especially in the succeeding generation, played an important part in the building up of the scientific system.

Inorganic Compounds.—The preparation of mineral acids showed improvements, and their investigation was marked by advances which only became of practical value later on, when the acids began to be employed technically. Glauber taught how to prepare hydrochloric acid from rock-salt and oil of vitriol, and also fuming nitric acid from saltpetre and white arsenic.

To Libavius belongs the merit of simplifying the mode of preparing sulphuric acid, and of proving that the acid obtained in various ways—from alum, vitriol (sulphate of iron), or sulphur—was one and the same substance. The behaviour of the acids just named to metals, salts, and organic compounds led to a knowledge of a great number of bodies which had been either unknown hitherto, or at least had never been produced in a like manner; and thus, from their modes of preparation, deductions as to their composition often became possible. Among such substances were the chlorides formed by the action of hydrochloric acid upon many of the metals, which before then had been prepared by heating the latter with sublimate, and hence the presence of mercury in the resulting products was assumed. Glauber, to whom we owe a knowledge of many of them,—*e.g.* zinc, stannic, arsenious, and cuprous chlorides—refuted this erroneous assumption; he and his contemporaries regarded these salts as compounds of the metal and hydrochloric acid.

Salts were called upon to play a very great part in medicine. Especial interest was taken in the alkaline salts, both from a theoretical point of view—their composition being a frequent theme of discussion—and also from a practical, on account of their technical and officinal applications.

Potash saltpetre, which was prepared on a large scale on account of its increasing use in the manufacture of gunpowder,¹ was also prized as a medicine in the fused state. The observation made by Geber—so important for a knowledge of its composition—that saltpetre results on saturating potashes with nitric acid, was first made use of technically in the iatro-chemical age. Sulphate and chloride of potash, which were prepared by many different methods and known under various names, were employed as medicines,—the former by Paracelsus, and the latter by Sylvius and Tachenius (as *sal febrifugum Sylvi*). Carbonate of potash, too, prepared from tartar and the ashes of plants, was another medicament. Even iatro-chemists of eminence like Tachenius believed in a chemical difference between various potashes, according to their modes of preparation,—an error which Boyle was the first to correct; still more frequently do we meet with a confounding of potash salts with those of soda, *e.g.* their carbonates and chlorides. Glauber's sulphate of soda, obtained from the residue left in the manufacture of hydrochloric acid, and known under the name of *sal mirabile*, was highly prized by physicians. Whether borax, which was used in soldering during the iatro-chemical period, was also employed as a medicine is doubtful.

Salts of ammonia were largely used, both officinally and technically, especially sal ammoniac, whose manufacture was attempted in Europe so early as the seventeenth century; its artificial formation from volatile alkaline salt and hydrochloric acid was known to the iatro-chemists of that time (Sylvius, Tachenius, Glauber), but it was only at a much later date that its true composition was indicated.

¹ Agricola describes the preparation of saltpetre in his work *De Re Metallica*.

The near relation thus found to exist between carbonate of ammonia and salmiac led conversely to the preparation of the former from the latter by means of carbonate of potash: from the apparently different action of samples of volatile alkaline salt of different origin (from blood, urea, salmiac), it was supposed that they were different compounds, but this error was recognised by Tachenius. Of other salts of ammonia may be mentioned the sulphate, discovered by Libavius, the nitrate, by Glauber, and the acetate: the last, known as *spiritus Mindereri* (from its discoverer, the physician Raymund Minderer), was much valued as a medicine.

But few of the salts of the earths were known, and there was uncertainty as to their composition. Lime and alum earth (alumina), for instance, were supposed to be pretty much the same. Of their salts, alum—prepared by adding putrefied urine to the crude alum lye (the aqueous extract from roasted aluminous shale)—was much prized for its technical value, and was manufactured in large quantity; the alum of that day was thus essentially ammonia alum. Agricola himself characterised gypsum as a compound of lime, while chloride and nitrate of calcium were known in the seventeenth century, and possibly even before then. Agricola and his contemporaries were also aware that silica (*i.e.* pure sand)—which was for long reckoned as one of the earths—fused with potashes to a glass which was soluble in water, and the clear-sighted Tachenius saw in this behaviour an indication of the acid nature of that substance.

The salts of the heavy and of the noble metals and various preparations of the semi-metals (arsenic, antimony, bismuth) were of much significance for iatro-chemistry, and therefore also for the development of the chemical knowledge of that time. Basil Valentine, as we have already seen, had worked up a large number of antimonial preparations and had recommended them for internal use prior to the appearance of Paracelsus and his school. And although, in consequence of the abuses resulting from secret medicines containing antimony, sharp edicts were issued

prohibiting their employment, still antimonial preparations came more and more into favour, this being greatly due to the efforts of Sylvius. Metallic antimony itself was prescribed in pills, which were called "the everlasting," since it was believed that they acted merely by contact, and that therefore, after passing through the body, they could be used again and again.¹

It was during this period that "Kermes" mineral, *sulphur auratum*, and powder of algaroth² were added to the medical treasury; antimoniate of potash—prepared by detonating antimony trisulphide with saltpetre—was also much used as a medicine. To Glauber more than any one else is due a clearer chemical knowledge of this and of other antimony compounds.

There was still great obscurity with regard to white arsenic and the metal prepared from it, and also with respect to other arsenic compounds; among the latter we may mention arseniate of potash, which was prepared by fusing the trioxide with saltpetre, and for which Paracelsus appears to have had a great predilection as a medicine (*arsenicum fixum*). Glauber was the first to prepare chloride of arsenic (AsCl_3). Preparations of bismuth were less used for medicinal purposes, although the similarity between bismuth and antimony, which often led to confounding the one with the other, did not escape the iatro-chemists. Basic nitrate of bismuth was much prized as a cosmetic, while the oxide, according to Agricola, was used as a paint.

Of the compounds of zinc, the oxide, zinc vitriol (which Agricola terms *chalcanthum candidum*), and the chloride became better known; the last of these was prepared by Glauber by strongly heating calamine with hydrochloric acid, and it therefore contained basic salt. From tin Libavius obtained its tetrachloride, by distilling it with

¹ Lemery in his *Cours de Chimie* (1675) remarks upon the use of these pills as follows: "*Lorsqu'on avale la pillule perpetuelle, elle est entrainée par sa pesanteur, et elle purge par bas; on la lave et on la redonne comme devant, et ainsi perpétuellement.*"

² So called after the Veronese physician Victor Algarotus, who praised it as *pulvis angelicus*.

sublimate; assuming in this the presence of mercury, he termed it *spiritus argenti vivi sublimati*, but later on it was commonly known as *spiritus fumans Libanii*. The solution of this compound, obtained by treating tin with aqua regia, began to be applied by Drebbel in many dyeing operations about the year 1630.

The discovery and investigation of ferric and lead chlorides, the latter of which was used instead of white lead as a paint, is likewise due to Glauber. The modes of preparing many metallic salts already known were also much improved, as is seen, for instance, in the description given by Agricola of the preparation of iron and copper vitriols.

The iatro-chemists devoted much attention to the production and medical application of quicksilver compounds. It was given to Paracelsus to overcome by his exertions the prejudices of many against mercurial medicines, although most of the physicians of the old school were not at all convinced of their admissibility. Paracelsus and his disciples did not hesitate to make use of metallic mercury—finely divided in pills,—sublimate, and the so-called turpeth mineral (*i.e.* basic mercuric carbonate or sulphate, both of which went under this name). In this way a much better knowledge of various mercury compounds was gradually arrived at, some of these compounds being already known and some newly discovered. Among the latter were calomel and white precipitate (from sublimate and ammonia), both of which were prized as medicines. It was during this period that chemists gradually learnt that cinnabar consisted of mercury and sulphur, and that mercury itself belonged to the true, and not to the half-metals.

Of the compounds of silver, *lapis infernalis* (the nitrate) was found useful in medicine, principally through Sala's recommendation, and the sulphate and chloride of silver were also known. The production of the latter on precipitating a solution of silver with hydrochloric acid or common salt solution, was made use of analytically as a test both for silver and for chlorides.

Indeed the beginnings of qualitative analysis in the wet

way are to be sought for in the iatro-chemical age, in so far that conclusions regarding the presence of one or another constituent were drawn from the appearance and behaviour of precipitates, and of salts which crystallised out from solution. Tachenius laid especial weight on the distinguishing of such precipitates by their colours, and he was himself able to detect several metals in solution together by means of certain reagents, such as tincture of galls, the carbonates of potash and ammonia, caustic potash, etc.

Organic compounds became known in rapidly augmenting numbers, in consequence of the increasing attention paid to the products of vegetable and animal assimilation; the actual knowledge of such bodies continued, however, very superficial and incomplete, since their composition remained quite obscure. Of the acids, acetic acid became better known, and several of its salts were used in medicine with good effect. It did not escape Glauber that the distillate from wood contained an acid which strikingly resembled that of vinegar. The iatro-chemists taught how to prepare concentrated acetic acid by the distillation of verdigris, whence it was known as copper spirit or radical vinegar; and this latter Tachenius was inclined to regard as van Helmont's alkahest. The two acetates, sugar of lead and the basic acetate, were examined more accurately by Libavius, and employed as medicines.¹

Salts of tartaric acid, of which tartar had been known for a long time, became valued as medicines in the sixteenth century; the discovery of the free acid itself belongs to a much later date. The designation *tartarus*, which was applied to tartar, was likewise the generic name in the iatro-chemical age for other very different salts, *e.g.* for the salts of potash, in so far as they were prepared from tartar, and also for sediments from solutions, especially those from animal secretions. The part which *tartarus* played in the

¹ The liquid which distils over on heating sugar of lead and which we now know to contain acetone, was investigated repeatedly: from its designation of *quintessence*, a specially high value seems to have been put upon it.

theoretical considerations of iatro-chemistry has already been spoken of. The salts of other vegetable acids were also frequently termed *tartarus*, e.g. salt of sorrel, which appears to have been often confused with tartar. Neutral tartrate of potash, known as *tartarus tartarizatus*, from its preparation from tartar and salt of tartar (K_2CO_3), and the double tartrate of potash and soda, called Seignette salt after the man who accidentally discovered it, likewise became known to chemists.

A compound of even greater importance to the medical treasury than the tartrates just mentioned was tartar emetic, the preparation of which from oxide of antimony and tartar was described by the Dutch physician Mynsicht, and afterwards more accurately by Glauber.¹ A tartar containing iron (*tartarus chalybeatus*) became known through Sala's *Tartarologia*. Paracelsus further made use of the distillate from tartar—which is now known to contain pyro-tartaric acid besides other substances—as a medicine (*spiritus tartari*).

Succinic acid, the near relation of which to tartaric has only become clear in our own time, is described by Libavius and Croll under the name of *Bernsteinsalz* (*flos succini*), what they referred to being the distillation product of amber; Lemery was the first to recognise its acid nature, about 1675. The acid juice of the apple and other fruits was employed for preparing various medicines (e.g. the *unctura martis pomata*), before any attempt was made to isolate the acid itself. Free benzoic acid, however, obtained by subliming gum benzoin, was discovered and minutely described by the French physician Blaise de Vigenère (1522-96) towards the end of the sixteenth century, while Turquet de Mayerne described the improved method of preparing it in the dry way, which is still practised at the present time. The juice of gall apples, which contains tannic acid, and the extract of oak bark were used by many iatro-chemists from the time of

¹ It may just be mentioned here that the taking of small quantities of tartar emetic, prepared by allowing wine to stand in goblets made of antimony, had been a common practice before this.

Paracelsus to test for iron in solutions, especially in mineral waters; but no one succeeded in isolating either tannic or gallic acid itself.

The old observation—that the fats were altered chemically by the alkalies and metallic oxides—did not indeed lead the iatro-chemists to a knowledge of the fatty acids, but it guided many of them, in especial the acute Tachenius, to the correct assumption that “oil or fat contains a hidden acid.” It was only one hundred and sixty years later that Chevreul’s work upon fats laid the firm foundation for the present views upon their chemical constitution.

Spirit of wine—the *aqua vite* of the alchemists—continued to grow in importance during the iatro-chemical age, as it had done in the alchemistic. This applied to it not merely from a theoretical point of view, as being a product of various fermentation processes to which much attention was paid, but also from a practical, since Paracelsus and his disciples used it largely in the preparation of essences and tinctures.¹

To the German physician Valerius Cordus is due the first exact knowledge of the ether produced from alcohol by acting upon it with sulphuric acid, although his instructions for preparing it were only published after his death, and the ether then accepted in the Pharmacopeias as *oleum vitrioli dulce verum* (about 1560). His work, however, became forgotten so soon, that we find such an accomplished chemist as Stahl unaware of it. A mixture of alcohol and ether, which later on enjoyed a wide popularity under the name of Hoffmann’s drops, had probably been employed by Paracelsus as a medicament. The knowledge of compound ethers remained very slight indeed, hardly any addition having been made to it since the observations of Basil Valentine.

The work done with other organic substances led to their practical application in medicine and in daily life, and also

¹ The name alcohol (*alcohol*) for spirit of wine, which has been in common use since the time of Libavius, had formerly quite another meaning, having been applied indifferently to antimony sulphide, vinegar, and various other compounds.

to improvements in their modes of preparation, *e.g.* in the extraction of sugar from the sugar-cane, the juice being clarified by white of egg and lime: but scientific knowledge with regard to such bodies remained at the lowest level.

CHAPTER IV

HISTORY OF THE PERIOD OF THE PHLOGISTON THEORY, FROM BOYLE TO LAVOISIER

THE reasons for naming this period of about one hundred and twenty years the period of the phlogiston theory, or of phlogistic chemistry, have been already stated shortly (p. 4). For the first portion of this era the designation is in truth not absolutely fitting, since Robert Boyle—the man who above all others gave a new direction to chemistry at the time—did not concur in the phlogistic views. The development proper of the phlogiston theory really took place after his death. Nevertheless the period from Boyle to Lavoisier may be so named, because the most important part of chemical research during that time had to do with the phenomena of combustion and with the analogous calcination of the metals. All the most eminent chemists of that day directed their attention to this problem, both theoretically and experimentally. It formed, especially towards the end of this period, the centre around which the whole of chemistry circled; it became a stumbling-block to the adherents of the old doctrines, and led to a reform of the science so fundamental and far-reaching that the chemistry of to-day still lives under it.

The iatro-chemical theories strove after the impossible, and therefore quickly succumbed; the marked one-sidedness apparent in them, the gratuitous explanations of life-processes, and the total neglect of the anatomy and

morphology of the organs, made their decline inevitable. An opportunity was thus given to chemistry to loosen and finally break the bands which medicine had wound around her, and to take up an independent position of her own. She still remained for a time under the protection of the healing art, to which she was indeed an indispensable aid; but even from the time of Boyle on, the aim of chemistry was recognised as being the discovery of new chemical facts, purely for the sake of arriving at the truth alone.

The spirit of true investigation which penetrated the natural sciences at the end of the sixteenth and beginning of the seventeenth centuries, also began to impart itself to chemistry, the development of physics exerting an especially powerful influence upon the younger sister-science. The inductive method also acquired, as a guide, a continually growing and a lasting influence, the nature of which was indicated by Francis Bacon¹ substantially as follows:—

“The true kind of experience is not the mere groping of a man in the dark, who feels at random to find his way, instead of waiting for the dawn or striking a light. . . . It begins with an ordered—not chaotic—knowledge of facts, deduces axioms from these, and from the axioms again designs new experiments.” Equipped with such axioms, chemistry might enrol itself among the exact sciences.

The learned societies which came into existence in the second half of the seventeenth and beginning of the eighteenth centuries, and whose periodicals spread abroad the results of chemical investigations, aided materially towards the healthy development of the science. The incitement given by them to researches, which could then be submitted to verification by other workers, was also not to be undervalued. Finally, they promoted the reciprocal action of

¹ *Novum Organon*, Aphorism 82, paragraph 3. Bacon in the above paragraph gave expression to no new idea, but merely called special attention to the value of experience, a point already recognised by his predecessors Palissy, Leonardo da Vinci, Paracelsus, and others. Liebig in a series of essays has proved conclusively how unjustifiable it is to designate Bacon as the originator of the inductive method, and how little he was permeated by the spirit of true research (see his *Reden und Abhandlungen*, 1874).

chemistry and allied branches of science upon each other, an action so fruitful in its results, by bringing their exponents into closer connection.

The Royal Society, which was formed about the middle of the seventeenth century by the amalgamation of the two smaller scientific societies of Oxford and London, and which began to publish the *Philosophical Transactions* in 1665, furnishes a good instance of what has just been said. The Italian academies, especially the *Accademia del Cimento* of Florence (1657), devoted themselves mostly to physical and mathematical studies. In Vienna the *Accademia Naturæ Curiosorum* was started in 1652, taking the name of *Cæsarea Leopoldina* in honour of its patron Leopold I. The *Académie Royale* originated in Paris in 1666 out of friendly meetings which were held at the house of the physicist Mersenne; the *Mémoires de l'Académie des Sciences* began to appear in 1699. The Berlin Academy was founded in 1700 by Frederick I., Leibniz being its first president; and during the first half of the eighteenth century the northern countries followed suit with similar learned societies, that of St. Petersburg being started in 1725, that of Stockholm in 1739, and that of Copenhagen in 1743.

That an extraordinary interest was felt at this time in scientific questions is readily seen from the literature of the day, which reflects the excitement—sometimes feverish in its intensity—raised by isolated discoveries, like that of phosphorus, or by disputed problems, such as the question of the cause of combustion.

The modes in which chemical questions were treated did indeed approximate to the methods followed in recent times, but in one respect there was a wide distinction between them. The chemical investigation of the phlogistic period took very little note of the proportions by weight in which substances entered into reaction; it turned its attention almost alone to the qualitative side of the phenomena. The setting up and the subsequent development of the phlogistic doctrines were only possible because of the utter neglect of quantitative relations. Even sagacious thinkers who, by

observing that metals increased in weight upon calcination, came thus into direct conflict with the phlogistic view, evaded the only correct explanation of this,—and, with it, of the phenomena of combustion—by far-fetched conceptions. This blinding of the mind by an erroneous theory, consequent upon the refusal to look into all the conditions which might have contributed to clearing up the question, is peculiar to the period of phlogistic chemistry.

In spite, however, of the fundamental error which ran through it, this period was a highly fruitful one for chemistry: it forms the indispensable introduction to the most recent phase of development of the science. And although it was itself fettered by many erroneous ideas, still the phlogistic age contributed very largely to the refutation of mischievous errors, *e.g.* those belonging to the iatro-chemical doctrines and the false beliefs of alchemy.

GENERAL HISTORY OF THE PHLOGISTIC PERIOD.¹

Robert Boyle and his Contemporaries.

Boyle has been rightly spoken of as the investigator who, by his creative genius, pointed out the new path to the period just then beginning: it would be even better to say that with him this new path originated. The spirit of pure investigation, free from the fetters of alchemistic and iatro-chemical conceptions, animated this remarkable man, whom chemistry has to thank for teaching her the real aims which she should pursue. The leading ideas of his scientific programme, which are laid down in the *Preliminary Discourse* (in Shaw's edition of Boyle's works, three vols., 1725), deserve to be quoted here:—

P. xxvi. “. . . I saw that several chymists had, by a laudable diligence, obtain'd various productions, and hit upon many more phenomena, considerable in their kind, than could well be expected from their narrow principles; but finding the generality of those addicted to chymistry,

¹ Cf. H. Kopp, *Gesch. d. Chemie*, vol. i. p. 146 *et seq.*; Höfer, *Hist. de la Chimie*, vol. ii. p. 146 *et seq.*

to have had scarce any view, but to the preparation of medicines, or to the improving of metals, I was tempted to consider the art, not as a physician or an alchymist, but a philosopher. And, with this view, I once drew up a scheme for a chymical philosophy : which I shou'd be glad that any experiments or observations of mine might any way contribute to complete."

P. xviii. ". . . And, truly, if men were willing to regard the advancement of philosophy, more than their own reputations, it were easy to make them sensible, that one of the most considerable services they could do the world is, to set themselves diligently to make experiments, and collect observations, without attempting to establish theories upon them, before they have taken notice of all the phenomena that are to be solved."

Experimental methods,¹ taken in conjunction with the careful observation of actual phenomena, form therefore, according to Boyle, the only sure foundation for speculations. To have made this the common property of chemistry, which from thenceforth strove to work out its fundamental principles by means of experiment, is the undying service rendered by Boyle.

His life was devoted to fostering the natural sciences, especially chemistry. Born in the year 1626, he became an earnest student at Geneva while still young, and continued his studies in the quiet of his estate of Stolbridge until 1654, when he settled at Oxford, carrying on there a constant intercourse with other eminent men of learning. From 1668 on, he lived in London, where he continued to work actively, as he had done at Oxford, for the Royal Society, then just starting into life ; he became its president in 1680, and held that office until his death in 1691. His noble and unpretentious character, with its accompanying modesty, and his simple religious tone, called forth astonishment and admiration both from his contemporaries and his successors. What a contrast between this modesty

¹ Thus he says that from these alone can we look for progress in all useful knowledge.

and the rude assumption of Paracelsus, or the self-appreciation of van Helmont and many other *servants* of the iatro-chemical age!

The services which Boyle rendered in the development of chemistry stretch over the most various provinces of the latter. Isolated observations of importance, by which he enriched—indeed, fundamentally extended—applied chemistry, the knowledge of chemical compounds and their analysis, the chemistry of gases, and pharmacy, will be treated of in the special part of this book. We have at present only to do with the general significance of his work and of his theoretical views for chemistry.

The term “element,” which before Boyle’s time was a very fluctuating and therefore uncertain one, received through him a more positive meaning. In his work, *Chemista Scepticus* (1661), he criticises the Aristotelian and the alchemistic elements, which were still accepted by many in the iatro-chemical age. He enunciated the axiom that only what can be demonstrated to be the undecomposable constituents of bodies are to be regarded as elements; and he considered it hazardous to advance opinions as to the properties of the elements in general, without having first obtained a firm foundation in their actual properties individually. With a far-seeing glance he looked forward to the discovery of a much greater number of elements than was at that time assumed, at the same time contending that many of the substances then held to be elementary were not really so.

Hand in hand with this wholesome clearing up of views upon the elements, there went fruitful ideas upon the union of the elements to compounds, and also upon affinity as the cause of chemical combination. Boyle was the first to state with perfect clearness that a chemical compound results from the combination of two constituents, and that it possesses properties totally different from those of either of its constituents alone. This clear opinion enabled him to draw a sharp distinction between mixtures and chemical compounds.

In order to explain the formation or decomposition of

compounds, Boyle advanced a corpuscular theory which gave evidence of his acuteness and showed how far he was ahead of his contemporaries. In his opinion all substances consisted of minute particles, and chemical combination took place when particles of different matter which mutually attracted each other came together. If another substance interacted with this new body, whose particles possessed a greater affinity for those of one of the components of the latter than these components had for each other, then decomposition ensued. In such simple manner did Boyle endeavour to explain the formation and decomposition of chemical compounds.

No one before him had grasped so clearly and treated so successfully the main problem of chemistry,—the investigation of the composition of substances. In doing this he had the solid ground of experience and experiment under his feet, and could always bring forward evidence for the probability of his views. His endeavours to get at the bottom of the composition of bodies gave a refreshing impetus to analytical chemistry, which before his time could hardly be said to exist.

Boyle likewise devoted much attention to the question of the cause of combustion and other similar phenomena, and although his attempts at explaining these were not very successful, his remarkable experiments on the part played by air in combustion helped materially to the later solution of the problem. His work on air and gases led him in 1660 to the memorable discovery of the now well-known law that "the volume of a gas varies inversely with the pressure" (Mariotte found this out independently in 1676).

Boyle's writings, which were already widely read in his own lifetime, are characterised by simplicity of style and clearness of expression; they offer an agreeable contrast to the works of many of the other chemists of his time, who sought to hide their deficiencies in clear thought and accurate knowledge by metaphorical and mysterious language. In addition to other papers published in the *Philosophical Transactions*, the following works of his, which were brought

out both in English and Latin, are to be especially mentioned:—*Chemista Scepticus* (1661); *Tentamina quadam Physiologica* (1661); and *Experimenta et Considerationes de Coloribus* (1663).

Among the men who, together with Boyle, advanced the natural sciences, especially chemistry, and of whom Willis, Hooke, Wren, and Hawksbee must be mentioned here, there was one in particular who, although a practising physician himself, rendered good service to chemistry by his observations on combustion and calcination, viz. John Mayow (born 1645). His assumption—that atmospheric air contained a substance¹ (also present in saltpetre) which combined with metals when they were calcined, and which sustained respiration and converted the venous blood into arterial—was bound to result in the right interpretation of the phenomena of combustion, when the observations which had led to it were sufficiently extended. Mayow's early death in 1679 was perhaps the reason why this did not come about, the development of the new chemistry being greatly retarded in consequence.

Lemery and Homberg.—The *Académie Royale des Sciences* formed in France the centre of union for chemists in that country, the chief exponents of the science in Boyle's time, particularly during the last quarter of the seventeenth century, being Wilhelm Homberg and Nicolas Lemery. Both of them being good observers, their work tended chiefly to the development of practical chemistry, which was especially indebted to Homberg for many valuable contributions. In the scientific explanation of technical processes they come a long way after Boyle; Homberg, in particular, was still enchained by alchemistic errors, and held fast to the idea that substances consisted of sulphur, mercury, and salt.

Lemery, born in 1645, hardly did any independent work on the treatment of theoretical questions, but he well knew how to sift and put together the facts already known.

¹ Mayow termed this substance *spiritus igno-aëreus* or *nitro-aëreus*.

This is shown in his *Cours de Chymie*¹ brought out in 1675, which was for long held to be the best text-book of chemistry, and was so widely used that the author himself lived to see thirteen editions of it published.

In addition to this literary work Lemery was exceedingly active as a teacher, the last thirty years of his life being taken up in that way; in his earlier years he was much involved in religious polemics, and could not therefore turn his chemical knowledge to the best account during that period.

Lemery designated chemistry a "demonstrative science," and therefore sought to elucidate chemical operations by suitable experiments. In theoretical questions, *e.g.* in his views upon combustion and upon the composition of substances, he was for the most part an adherent of Boyle.

While Lemery was chiefly exercised, then, about the effective propagation of his science, Homberg—permanently settled in Paris after a restless life and many-sided studies—found particularly good opportunity, as body-physician and alchemist to the Duke of Orleans, of making numerous and sometimes important observations in practical chemistry. Some of his researches, *e.g.* that upon the saturation of acids by bases, contained fruitful germs which became developed later on in the hands of other workers. Most of the writings of these two men, both of whom died in the same year (1715), were published in the *Memoirs of the French Academy*.

Kunkel and Becher.—The most eminent German chemist in Boyle's time was Kunkel, to whose name that of Becher must also be added. Closely connected with the latter was Stahl, the originator of the phlogiston theory, the beginnings of which are to be seen in the views of both of the men first mentioned.

Johann Kunkel, born at Rendsburg in 1630, rendered

¹ Shortly before the publication of the *Cours de Chymie*, two other text-books appeared in Paris, both entitled *Traité de Chymie*, by Lefèvre (1660) and Chr. Glaser (1663), under the latter of whom Lemery had begun his studies. Glaser's book treats chiefly of pharmaceutical, and Lefèvre's of theoretical chemistry, which latter, however, was not much advanced by it.

excellent service to practical chemistry as an able experimenter and acute observer. Originally a pharmacist, he early showed the leaning towards alchemy which was decisive and fateful as regarded the whole course of his life; he was too honest not to see through many of the frauds of adepts, but at the same time was so firmly convinced of the possibility of the transmutation of metals that he gave his life-work to solving the problem. Employed as an alchemist by various princes (among whom were the Dukes of Lauenburg, the Elector John George of Saxony, and the great Elector), whose desires he was unable to gratify, he led a restless life which came to a close at Stockholm in 1702, where, by the favour of Charles XI., he had found a more honourable position than any previously allotted to him. Kunkel's preconceived opinions caused his writings to be permeated by mischievous errors, and to contain work bearing upon alchemy. What a contrast between him and Boyle! While the latter was seeking to ascertain the real composition of substances, and to get at their demonstrable constituents, the former still held to the tenet that all metals contained mercury. Nevertheless, as a promoter of experimental chemistry, and therefore of practical chemical knowledge, Kunkel deservedly holds a high place.

Johann Joachim Becher, who was born at Speyer in 1635 and died in London in 1682, worked almost contemporaneously with Kunkel, but more for the theoretical explanation of already observed facts than for the practical side of the subject; in his unsettled life and his propensity towards new projects, he resembled the latter. He worked as an alchemist at various courts (in Mainz, Munich, and Vienna), but he was too honourable to deceive his patrons, and too candid to allow of his remaining long in any one place. His bold technical projects almost always came to nothing; they show only too clearly their author's deficiency in practical chemical knowledge. In theoretical questions as to the composition of substances Becher attempted to revive the old ideas of Basil Valentine and Paracelsus in another form. In place of mercury, sulphur, and salt, he set up three "earths," of

which all inorganic ("sub-terrestrial") bodies should consist, viz. the mercurial, the vitreous, and the combustible (*terra pinguis*). The nature of any material depended upon the proportions in which these three fundamental earths were contained in it. Of especial importance was Becher's assumption that, when substances were burnt or metals calcined, the *terra pinguis* escaped, and that in this escape lay the explanation of combustion; it was from this conception that Stahl's phlogiston theory originated. The opinions of Becher upon the production of salts and acids from these earths were also received with approbation by his disciples.

These theoretical views are to be found in Becher's first work, *Physica Subterranea* (1669), and in his last, *Theses Chymicæ* (1682). His doctrines acquired great celebrity through Stahl, whose work belongs for the most part to the eighteenth century, upon which he conferred a character of its own by his development of the phlogiston theory.

Stahl and the Phlogiston Theory.

The theory of the phenomena of combustion and other analogous processes, which were to be explained by the assumption of the hypothetical *phlogiston*, was the point round which chemists in general gravitated during the eighteenth century; until the appearance of Lavoisier the phlogiston theory received the unqualified assent of most investigators.

Georg Ernst Stahl, born at Anspach in 1660, devoted himself to the study of medicine, and acquired, first at Jena and later on at Halle—to whose university he had been called as professor of medicine and chemistry in 1693—the reputation of a distinguished physician and academic teacher. Appointed physician to the king in 1716, he removed to Berlin, where he laboured with success for the extension of chemical knowledge until his death in 1734. He worked at chemistry in the true scientific spirit; himself guided by the ardent desire to discover the truth, he was able to draw

around him pupils animated by a similar aim. The most eminent among the Berlin chemists of the succeeding generation were trained by him.

Already in his own lifetime his doctrines, together with a number of valuable detached observations, were widely distributed by means of his writings, and especially by means of his lectures, the latter of which were published by several of his pupils.¹ Stahl, however, exercised his greatest influence both upon his contemporaries and upon the succeeding generation by his phlogiston theory, which eclipsed all his other chemical work.

Stahl himself freely recognised the close connection between his views upon combustion and calcination and the original ones of Becher: he went to work, however, quite differently from the latter, although his doctrine was grounded upon Becher's idea regarding the combustible constituent. This assumption of a constituent common to combustible bodies (a "fire material," a "sulphur," and so on) was indeed of older date than that of Becher's *terra pinguis*, which Stahl at once utilised, in order to build up his phlogiston theory thereon. This rests upon the hypothesis that combustible substances—among which the metals capable of calcination were reckoned—contain phlogiston as a common constituent, which escapes on combustion or calcination. Since, as was then held, every phenomenon bearing upon this could be readily explained by the aid of such an assumption, it was considered unnecessary to prove the actual existence of phlogiston itself directly. Stahl was able by means of it to group uniformly together and to explain a large number of chemical reactions. The more violently the combustion of any substance went on—so he taught,—the richer it was in phlogiston; coal, which can be almost entirely consumed, was therefore to be regarded as nearly pure phlogiston. In order to reproduce the original substance, its combustion-

¹ Among Stahl's writings we may name the *Zymotechnia Fundamentalis*, etc. (1697); *Specimen Becherianum*, etc. (1702); and *Zufällige Gedanken über den Streit von dem sogenannten Sulphure* ("Occasional Thoughts on the Dispute regarding the so-called Sulphurs"), (1718). Of his pupils, Juncker was especially active in propagating the views of his master.

products had to be added to it again; in this manner the metals were "revived" from their calces, which, according to Stahl's notion, had resulted from the former through the escape of the phlogiston. When a metallic calx was heated along with coal, the phlogiston so abundantly contained in the latter combined with it, the metal being thus reproduced; consequently a metallic calx was a constituent of a metal. Upon a like sophism rested Stahl's assumption that sulphur consisted of sulphuric acid and phlogiston. He saw in the production of sulphur, on heating sulphuric acid or a sulphate with coal (phlogiston), a synthesis of the former, and therefore a proof that sulphur was a compound body. Upon the further logical conclusion,—that the products of combustion of any substance must be lighter than that substance itself, seeing that they are constituents of it, no weight was placed. And no attention was paid to the numerous observations which showed that this was not the case,—that, indeed, a calcination of the metals was accompanied by an increase in weight. It was facts like those just named which, after a prolonged struggle, brought about the overthrow of the phlogiston theory.

To Stahl belongs the merit of grouping together the phenomena of oxidation and reduction, as we now term these, albeit by the aid of a false hypothesis. The addition of phlogiston is equivalent to reduction, and its withdrawal or escape to oxidation. The analogy between respiration and the decomposition of animal matters on the one hand, and combustion on the other, did not escape Stahl, who likewise conferred the chief *role* in those processes upon phlogiston.

The value of his theory lay therefore in the interpretation which it afforded of a variety of processes from one common point of view. The simplicity of this explanation blinded both himself and the generation which followed him to such a degree that they did not notice all the glaring contradictions between actual facts and the phlogistic doctrine. Notwithstanding this, however, the latter was in no way detrimental to the development of chemistry, seeing that

chemists like Black, Cavendish, Marggraf, Scheele, Bergman, and Priestley, who so greatly enriched their science by wide-reaching discoveries, were phlogistonists in the full sense of the word.

Hoffmann and Boerhave.—Before speaking of the further destinies of the phlogiston theory, and, in connection with this, of the state of chemistry at that date, the work of two of Stahl's contemporaries who contributed materially to the advancement of the science must be considered, viz. Friedrich Hoffmann and Hermann Boerhave. Both of these men were eminent physicians and accomplished chemists, but they were not exactly adherents of Stahl's phlogiston doctrine, although they held similar views with regard to combustion.

Hoffmann, born at Halle in 1660, *i.e.* in the same year as Stahl, after acquiring a thorough knowledge of medicine, mathematics, and the natural sciences, practised first as a physician, and then became professor of the science of medicine in Halle, where he ultimately died in 1742, after an interregnum spent in Berlin. His most important work was done in medicine and in pharmaceutical and analytical chemistry. He combated with success the iatro-chemical doctrines of Sylvius and Tachenius, which still held their ground with many physicians, exposing their absurdities and showing to what nonsensical deductions such exaggerations led. Many of his investigations and discoveries in pharmaceutical and analytical chemistry will be touched upon in the special history of this time. Hoffmann's views on combustion were very similar to those of Stahl. With respect to the calcination of the metals and the reduction of their oxides, however, he expressed opinions which approximate to those held at the present day, believing, as he did, that metallic calces contained a *sal acidum* in addition to a metal, the former of which escaped when the calces were reduced. This assumption did away with the similarity between combustion and calcination: these phenomena became indeed rather opposed to one another thereby, and

with this the special use of the phlogiston theory vanished. Hoffmann was a very voluminous author, and his collected works, entitled *Opera Omnia Physico-medica*, show clearness of style and precision of expression.

Hermann Boerhave, born in 1668 at Voorhout, near Leyden, was originally destined for the study of theology, but devoted himself to medicine, gaining at the same time an excellent knowledge of the natural sciences, and especially of chemistry. From the year 1709 on, he was able to utilise his catholic education to advantage as professor of medicine, botany, and chemistry in Leyden, and attained to the highest distinction; he died there in 1738.

Boerhave's place in the history of chemistry is due not to any wonderful experimental researches, but to the exceptional acuteness which he showed in noting and collating chemical phenomena from one common point of view. His large text-book *Elementa Chemicæ* (1732) was intended to contain all the most important work done in chemistry, and for a long time it remained by far the best guide to the study of the science. His conception of the latter as an absolutely independent science, subordinate to no other, and whose aim should be the investigation and perception of chemical facts, was at once a beneficial and an elevated one. In accordance with this view we find him condemning the abuses which the iatro-chemists had introduced into chemistry. The work of the alchemists he did not criticise sharply enough; in his endeavours to test the assertions which they made, he found here and there some corroboration of them, and was thus probably not disinclined to decide in favour of the adepts in cases where experience had not as yet spoken her last word. On the other hand, he refuted many statements, such as those which told of the fixation of mercury and of the production of the latter from lead salts, and thus contributed to clear up and set right alchemistic opinions and assertions.

Boerhave appears to have concurred in the phlogiston theory in many points, at least he expressed no opinions contrary to Stahl's fundamental views, although he did not

agree in regarding the calces of the metals as the earthy elements of these latter. Like many other investigators, Boerhave studied the processes involved in calcination, and to him is due the valuable experimental contradiction of the view put forth by Boyle and others,—that, during calcination, a ponderable fire-stuff is taken up, and thus the increase in weight of the metals explained.

*The Development of Chemistry, and particularly of the
Phlogiston Theory, after Stahl's time.*

The influence of Stahl's doctrine manifested itself more immediately in Germany, where it received the almost unqualified support of chemists, Berlin remaining the centre point of this theory. Among the men who upheld and sought to propagate it, Marggraf was the most eminent. Kaspar Neumann (born 1683) and Johann Theodor Eller (born 1689), contemporaries of Stahl, were also active adherents of the doctrine in the capital town of Prussia. Both of them, as professors at the medico-surgical institute, were in a high degree active in maintaining and spreading a knowledge of chemistry. Their own observations were, however, of little importance: those of Eller were chiefly upon medico-physiological subjects, and are full of untenable speculations. Stahl's disciple and pupil Johann Heinrich Pott (born 1692) enriched chemistry by many valuable observations, but he was unfortunate in his explanation of these, regarding boracic acid, for instance,—a substance which he had himself investigated carefully—as consisting of copper vitriol and borax. The results which he achieved were not at all commensurate with his untiring perseverance, which he showed, among other ways, in his endeavours to prepare porcelain. Although an adherent of the phlogistic doctrine, Pott did not bring forward anything new in its favour; with regard to the nature of phlogiston itself, he expressed the opinion that it was “a variety of sulphur.”

Andreas Sigismund Marggraf (1709-1782) was the

last and most eminent adherent of phlogistic views in Germany. Destined originally for an apothecary, he acquired a knowledge and practical experience of chemistry as assistant to Neumann at Berlin, and by sedulous study at the high schools of Frankfurt on the Oder, Strasburg, and Halle, and finally at the Freiberg School of Mines; this knowledge, accompanied as it was by exceptional gifts of observation, put him in a position to carry out researches of the greatest merit. One has only to think of the observations made by him in his work on phosphoric acid,—observations which, considering the highly defective state of chemical analysis at that date, fill us with admiration; of the proof which he furnished of the difference between alum and the so-called bitter earth (magnesia), substances which had hitherto been generally confounded; and, above all, on his investigation of the juice of the red beet, in which he discovered cane sugar (see special section of this book). It was during this research that Marggraf introduced the microscope into chemistry, as a valuable aid in distinguishing different bodies.

With this great talent of observation he united the gift of drawing what were generally sound conclusions from his work. In one point, however, Marggraf, like all phlogistonists, was not in a position to do this; although he had himself proved that phosphorus increases in weight by conversion into phosphoric acid, he could not free himself from the idea that phlogiston escaped during this process of combustion. And he could never be brought to see that this conception was an erroneous one, although the anti-phlogistic doctrine was established several years before his death. Marggraf's papers are almost all contained in the *Memoirs of the Berlin Academy*; most of them were published during his lifetime in two volumes, under the title *Chemische Schriften*.

The French Phlogistonists.—The chief exponents of chemistry in France during the eighteenth century until the downfall of the phlogistic system were Geoffroy, Duhamel, Rouelle, and Macquer, who concurred essentially

in Stahl's views. They enriched the science not only by important facts, but also by useful working theories.

Stephan François Geoffroy (the elder, to distinguish him from his less celebrated younger brother, Claude Joseph, whose work was chiefly pharmaceutico-chemical) was born in Paris in 1672, and helped for some time in his father's drug shop; he gave himself up, however, to chemical and medical studies, and laboured with great success as professor of medicine in the *Jardin des Plantes* from the year 1712 until his death in 1731. Geoffroy made himself a name especially by his researches upon chemical affinity; his *Tables des Rapports* (tables of affinity), in which the results of his most important observations are collected, exercised a prolonged influence upon the doctrine of affinity. His theoretical views were less happy, *e.g.* he looked upon the iron found in the ashes of plants as having been produced artificially during the process of ignition. In the questions of combustion and calcination he approximated to Hemberg's view, already detailed; the metals, for example, he regarded as composed of earths and a species of sulphur.

Geoffroy's treatises were published partly in the *Philosophical Transactions*, and partly in the *Memoirs* of the French Academy. His long-celebrated work, *Traetatus de Materia Medica*, shows what a high value he placed upon chemistry as a sister science and aid to medicine.

Duhamel de Monceau (born 1700, died 1781), of the school of Lemery and Geoffroy, spent his entire life in Paris, where his many-sidedness gained for him a high reputation. His sterling work was not by any means in pure chemistry alone, but also in physics, meteorology, physiology, botany, and—particularly—in chemistry as applied to agriculture. We must make especial mention here of the fact that he furnished definite proof of the difference between potash and soda, by preparing the latter pure; he also showed that it was the base of rock-salt, borax, and Glauber's salt. The first proposals to prepare soda artificially from rock-salt came from him, a fact which shows his far-seeing glance.

Whilst Duhamel worked purely as an academician,

Guillaume François Rouelle (born 1703, died 1770) was mainly occupied in teaching,¹ in which he greatly excelled; some of his pupils, particularly Lavoisier and Proust, arrived at the highest eminence. At the same time he was also busy as an investigator, as many admirable observations and conclusions drawn from these show. Rouelle fixed the meaning of the term "salt" (in the Memoirs of the Academy for 1745) from a far more general point of view than van Helmont or Tachenius had done. The composition of a substance alone was sufficient to tell him whether it belonged to the class of salts or not. Salts were produced by the combination of acids of every kind with the most various bases; and, in addition to neutral salts, he drew a distinction between acid and basic ones. With views so clear as these, Rouelle was far ahead of his contemporaries.

Among the latter was Pierre Josèphe Macquer (born 1718, died 1784), who was likewise an active and successful teacher at the *Jardin des Plantes*, and who also aided effectively in the spread of chemical knowledge by means of his text-books.² His own individual work lay less in theoretical than in applied chemistry, to which he made valuable contributions (especially in the manufacture of pottery and in dyeing).

From the beginning of his career to its end Macquer was a phlogistonist, and did all that he could to reconcile the continually augmenting discrepancies between theory and facts; he paid no heed to proportions by weight, for it

¹ The numerous records of Rouelle's activity as a teacher, which have come down to us, enable us to form a clear picture of the conditions of chemical teaching in those days, and at the same time to appreciate the remarkable personality of the man. The lectures on chemistry were delivered by two professors, one of them treating the theory of chemical processes, whilst the other, in conjunction with him, showed and explained how they were carried out practically. While the former (Bourdelin) fatigued his audience by his abstract reasonings, Rouelle inspired the students of practical chemistry by the vivacity of his discourse, during which he frequently became so excited as to throw off his periwig and some of his articles of clothing (cf. Hüfer, *Hist. de la Chimie*, vol. ii. p. 378).

² The principal of these were: *Éléments de Chimie Théorique* (1749); *Éléments de Chimie Pratique* (1751); and his *Dictionnaire de Chimie* (1778).

was only in this way that he could maintain the phlogistic hypothesis. And even although it was proved to be erroneous and untenable several years before his death, he was still unable to renounce it.

The English and Swedish Phlogistonists.—Joseph Black, professor of chemistry in the Universities of Glasgow and Edinburgh, who was born in 1728 and died in 1799, advanced chemistry in an exceptional degree by his splendid experimental researches, which were published in the *Philosophical Transactions*; in especial by his, for that time, masterly investigations on carbonic acid and its compounds with the alkalies and alkaline earths, which were planned and carried out with the utmost ingenuity. His observations led to a clear knowledge of processes which had formerly been explained quite wrongly, and they drew the attention of investigators in a special manner towards gases; the work done with these had the effect of causing chemistry to proceed on new lines, and was, in fact, the necessary forerunner of the latest epoch of the science. In addition to this Black threw open a new field to physics by his discovery of latent heat in 1762, in which his wonderful gift of experimenting came to his aid.

In order to appreciate his labours at their true value, and to compare them with those of other chemists who busied themselves with similar questions, we have only to fix our attention on his researches upon the alkaline earths and the alkalies. The carbonates of these were before Black's time regarded as simple substances; and it was further assumed that when limestone was burnt fire-stuff was taken up, and that this went over into potash or soda when these were causticised by means of lime. Black, on the contrary, proved by his researches that when limestone or *Magnesia alba* was calcined, something escaped which led to a loss of weight and which was identical with van Helmont's *gas sylvestre*. This gas—which he termed *fixed air*, because of its being held bound by caustic alkalies, lime, etc.—he proved to be also present in the mild

alkalies; and these latter became caustic when deprived of their carbonic acid by lime or magnesia. That Black devoted great attention to the proportions by weight of the compounds which entered into the reaction is seen in all his investigations; and it is thus easy to understand how he gave up the phlogiston theory and concurred in the doctrine of Lavoisier as soon as the correct explanation of combustion and similar processes became possible through the discovery of oxygen.

Black, by his fundamental labours, did away with numerous errors, and thereby prepared the way for the definite knowledge of the true composition of important chemical compounds. Notwithstanding this, the evident conclusions which followed from his researches on causticity were unfavourably criticised by many of the chemists of his time, and indeed their correctness disputed; it is strange to find that even Lavoisier could not bring himself candidly to recognise Black's services in this respect, and that he rather ranged himself on the side of the latter's antagonists, who were not in reality able to weaken one of his arguments.

In his countryman, Henry Cavendish, Black had a most distinguished co-worker, who, while investigating quite independently of him, did so upon similar lines, and greatly enriched the science also. Cavendish, born at Nice in 1731, devoted himself very quietly, but not the less efficiently, to the natural sciences, which he had studied thoroughly, especially to physics and chemistry; he died in London in 1810.

His masterly researches—so important both from a physical and from a chemical standpoint—upon hydrogen (*inflammable air*), which he was the first to distinguish as a peculiar gas differing from all others, and also those upon carbonic acid, constitute him the founder of pneumatic chemistry, and the originator of the new era. To him we owe the proof, of what value need not be said, that water consists of hydrogen and oxygen; further, the proofs that atmospheric air is a mixture of nitrogen and oxygen in constant proportions, and that nitric acid can be produced by the chemical

combination of these two latter gases in presence of water. All these were discoveries of the greatest moment. In them Cavendish forged the most powerful weapon for the overthrow of the phlogiston theory, notwithstanding which we find him faithful to the latter, and defending it with all his might. His opposition to the antiphlogistic doctrine, which he himself helped to found by his own investigations, can only be explained by the fact that he did not pay enough attention to the proportions by weight in the processes of combustion, but explained the latter in a way which appeared to him sufficiently convincing, viz. by regarding hydrogen as identical with phlogiston.

The most zealous champion for the phlogistic idea at that time was Joseph Priestley, to whom the chemistry of gases owes an extraordinarily large number of new observations and important discoveries. Eccentric, and of a restless, fiery nature, Priestley combated the antiphlogistic doctrines until his death (in 1804) as no other man did, although his own researches often went to strengthen, even to lay the foundations of, the latter. In contrast with the quiet existence of Black and Cavendish, wholly devoted to science, a wandering life full of vicissitudes and persecutions was destined for Priestley, doubtless for the most part because of his relations to the English Church and his own intolerance. Theology was his own special subject, and he was already a minister when he first came more closely into contact with scientific questions. Endowed with an unusual gift for experimenting, he was able to treat the most difficult problems of pneumatic chemistry, although lacking a thorough scientific education. He prepared and investigated a large number of gases which, with the exception of air, carbonic acid, and hydrogen, were practically unknown before his time. Of all his discoveries, that of oxygen (in 1774) was the most important; it will be treated of later on. His beautiful researches on this gas did not, however, lead him to the correct explanation of combustion; he remained, on the contrary, as already seen, true to the doctrine of phlogiston. But his erroneous ideas respecting this and similar processes

did not prevent him drawing from his own observations sagacious conclusions with regard to the series of recurrent changes which oxygen undergoes in animal and vegetable metabolism,—a far more complicated process than that of combustion, which he was unable to explain.

Contemporaneously with the three last-named English chemists, two most distinguished investigators, Torbern Olof Bergman and Karl Wilhelm Scheele, were labouring in Sweden as upholders of the phlogistic theory, which their brilliant discoveries and observations only served deeply to undermine. Bergman had acquired such a wide knowledge of the natural sciences that he taught with eminent success as professor of physics, mineralogy, and chemistry at Upsala. Born in the year 1735, he died at the early age of forty-nine, doubtless from the effect of overwork upon a weak constitution. His chief services to chemistry, to which from 1767 he principally devoted himself, were in the domain of analysis, which he treated systematically and enriched by valuable methods. He knew well how to make his chemical experiences useful for the definition and classification of minerals, and thereby laid the foundation of mineralogical chemistry and chemical geology. The current views upon chemical affinity thus gained through him precision and clearness; the scientific character of chemistry was materially raised by such observations, and a general survey of chemical processes rendered much easier. His papers appeared originally in the *Memoirs of the Academies of Stockholm and Upsala*; later on they were collected together, and published under the title *Opuscula Physica et Chemica*.

Karl Wilhelm Scheele will remain for all time one of the most distinguished of chemists; and his fame is not lessened by the fact that he continued all his life through a zealous supporter of the phlogistic doctrine. In spite of the unfavourable conditions under which he lived, and of the short span of his life, he contributed to chemistry a wealth of new observations—many of them most valuable discoveries—which furnished a rich mine for the experimental work and theoretical discussions of future generations. Scheele, born

in 1742 at Stralsund, the capital of Pomerania, which at that time belonged to Sweden, began at fourteen years of age his apprenticeship in various apothecaries' shops (in Gothenburg, Malmoe, and Stockholm), and gained by private study and by experimenting a sound knowledge of chemistry, so that we find him able to account for the production of nitrite of potassium on heating saltpetre, a point that even Bergman could not explain. Becoming known to the latter by this chance, Scheele continued after that time in relation both with Bergman and with the other eminent chemists of his country. The discoveries which he made in the quiet of his apothecary's shop soon rendered him famous in European scientific circles, which showed their appreciation by electing him to their academies. He died in 1786, when barely forty-four years old.

Endowed with a most wonderful gift of observation, Scheele was able to bring to a successful conclusion researches carried on with but very limited means at command. A brilliant proof of this is given in his investigations upon black oxide of manganese (*Oxide of Manganese*), which many competent workers before him had studied without, however, succeeding in making its nature clear. Scheele discovered during this research four new substances—chlorine, oxygen, manganese, and baryta, of which the two first especially were of the utmost importance for the proper understanding of chemical processes.

In a manner nothing short of marvellous Scheele brought his inventive genius to bear upon organic chemistry, which had till then been left almost untouched; working out in every direction new methods for isolating the products of vegetable and animal metabolism, he prepared a large number of acids and other organic compounds hitherto unknown. Scheele was a pioneer in nearly every branch of chemistry, being unique in power of observation and in the quick comprehension of facts, although, it is true, not always happy in his interpretation of these, fettered as he was by the phlogiston theory.

In order to properly appreciate the condition of the

latter in the seventh and eighth decades of last century—that is, shortly before its downfall,—the development up to that date of a special section of chemistry, viz. pneumatic, must be considered. The work done with gases, and the knowledge acquired of their properties, had finally led to the correct interpretation of combustion. The special history of the phlogistic period thus falls to be treated of now.

DEVELOPMENT OF PARTICULAR BRANCHES OF THEORETICAL AND PRACTICAL CHEMISTRY IN THE PHLOGISTIC PERIOD.

Pneumatic Chemistry and its Relations to the Doctrine of Phlogiston.—The influence which the investigation of gases, especially of oxygen, exercised in shaping chemistry is sufficiently well known. Oxygen forms to some extent the centre-point of chemical research during the last quarter of the eighteenth century, since the knowledge of the part played by it in combustion and similar processes led to the setting aside of a doctrine which had dominated all theoretical opinions for a hundred years; and, further, because the most important results were conjoined with its study, inasmuch as this contributed materially to the development of the atomic theory.

The services of the men whose observations aided most in building up the chemistry of gases have already been mentioned generally; it will suffice here to treat in more detail certain of these observations and a few others. Boyle's researches show a marked advance over those of van Helmont in the mode in which he collected gases and worked with them; at the same time neither he nor his contemporaries felt quite sure whether carbonic acid and hydrogen, whose characteristic properties he knew, differed materially from atmospheric air. This uncertainty is also seen in the work of later investigators, *e.g.* Hales: the erroneous idea that gases were ordinary air with various admixtures, had fixed itself firmly in the minds of chemists. To Black is due the merit of proving the distinct difference between

carbonic acid and air, by showing the "fixation" of the former by caustic alkalis. Cavendish, who recognised in hydrogen a peculiar gas, likewise aided in doing away with the misconception. Finally, we would mention here the remarkable supplemental researches of Bergman on carbonic acid (1774).

The methods of collecting gases had improved considerably since Hales—and, before him, the little-known Moitrel d'Elément—had effected a separation of the generating vessel from the receiver. Air was found to be a fluid capable of measurement which possessed weight, and which, like other fluids, could be transferred from one vessel to another. The apparatus which Black, Priestley, and others used, and those which we employ at the present day, gradually developed themselves from that of Hales. Priestley was the first to collect gases over mercury, and succeeded by this device in discovering gaseous ammonia, hydrochloric acid, silicon fluoride, and sulphurous acid,—all of which had been overlooked so long as water only was used for this purpose.

The discovery of so many gaseous substances of such different characters greatly excited the chemical world. The properties of each gas were carefully examined; and, after Mayow's researches, and especially after the more exact determinations by Cavendish, the density was taken as the criterion of one gas differing from another and from atmospheric air. Due regard was also paid to the greater or lesser absorption of gases by water, as a distinct test for some of them; Bergman, for instance, determined with fair accuracy the solubility of carbonic acid in water. But the true composition of gaseous bodies remained unknown during this epoch, great uncertainty prevailing even about the simplest of them until Lavoisier had made clear the elementary nature of oxygen and hydrogen. How could this indeed be otherwise, so long as the presence of phlogiston was assumed in most gases? Hydrogen was considered identical with phlogiston by many chemists soon after the middle of the eighteenth century, Cavendish and Kirwan setting the

precedent for this; others looked upon coal as being rich in phlogiston, if not as the latter itself. The most confused opinions were expressed regarding the composition of carbonic acid, carbonic oxide, nitric oxide, sulphurous acid, and sulphuretted hydrogen (the last of which was accurately investigated for the first time by Scheele in 1777), these opinions being made to fit in with the views of the phlogistic doctrine prevalent at that time.

Of greater moment than these varying opinions upon the constitution of the gases just named were the long unsettled questions: "Is atmospheric air a simple or a compound body, and what are its constituents or ingredients?" These questions were solved experimentally by chemists belonging to the phlogistic era, in particular by Scheele and Priestley; but it was left to Lavoisier to interpret their observations correctly. We must now speak of the most important of the facts then brought to light, which bear upon the composition of the air.

The first observation which aided in overthrowing the old assumption of air being a simple substance, was the behaviour of an enclosed volume to a body burning and to metals heated in it. Boyle was forced by his researches in this direction to the supposition that one ingredient of the air was necessary to respiration and combustion, and to the calcination of the metals; but he was unable to isolate this ingredient, as was also Mayow, who, with his assumption of a *spiritus ignis-aëreus*, which brought about combustion (cf. p. 104), came pretty near to the right interpretation. It was, however, only a hundred years later, after oxygen and nitrogen had been prepared successfully, that the question approached its solution. Nitrogen, which various investigators had already worked at, was isolated by Rutherford in 1772, by the absorption of the carbonic acid produced by combustion or respiration in an enclosed volume of air; it followed from his observations that this gas, which was incapable of sustaining either of these, must be one of the ingredients of the atmosphere. The other was isolated and examined by Priestley and Scheele in 1774

quite independently of one another,¹ being obtained by both of them by heating red oxide of mercury, and by Scheele from other oxides also (including black oxide of manganese), and from saltpetre. Both observed that this gas was capable of supporting combustion and respiration in an intensified degree. Priestley named it "dephlogisticated air," and Scheele "fire air," in accordance with the views forced upon them by the phlogistic doctrine. The momentous discovery of oxygen enabled both of them to recognise air as being a mixture of two kinds of gas;² Priestley calls nitrogen "phlogisticated air," and Scheele terms it "spent air." They both found substances which absorbed the one constituent of the air (oxygen), Priestley using for this purpose saltpetre gas (nitric oxide), and Scheele hydrate of protoxide of iron, phosphorus, etc. They made the further important observation that, upon burning a candle in an enclosed volume of air, exactly as much "fixed air" was generated as oxygen had vanished.

Notwithstanding all this they did not get at the right explanation of combustion, respiration, and calcination, whose analogy to one another they clearly saw; so prejudiced were they by the idea that phlogiston escaped during these processes, that the path distinctly marked out by their own observations was left for another to tread. Lavoisier was destined to do this, as he easily threw aside the trivial phlogistic prepossessions which he cherished at the beginning of his scientific career. The others, indeed, upheld a contradictory explanation of combustion and analogous processes, in order to remain true to the phlogistic doctrine. But that it was Priestley and Scheele who, by their exhaustive researches on oxygen, and the part which it played in the processes just mentioned, furnished the experimental material for the correct understanding of these, and not Lavoisier, is beyond all question.

¹ Hales, Bayen, and Priestley himself had observed oxygen previous to this, without, however, recognising its peculiar nature.

² Scheele, in his treatise *Von Luft und Feuer* ("On Air and Fire"), says: "The air must be made up of elastic fluids of two kinds."

After the discovery of oxygen and of its chief properties the days of the phlogistic theory were numbered, although many eminent chemists still held to it in spite of accumulating contrary evidence. The greatest difficulty in the way of the old doctrine was the fact, already known for a long time, that, in those cases where phlogiston was supposed to escape, the products became heavier instead of decreasing in weight. The exact researches on the calcination of the metals,¹ had their results been studied without any preconceived opinions, ought to have led to the correct explanation, viz. that one ingredient of the air combines with the metals to form calces; for not only was the increase in weight observed, but also the disappearance of a portion of the air. But instead of drawing from this the conclusion that the phlogistic hypothesis was untenable, chemists tried to make the observed facts fit in with the latter by putting a strained interpretation on them. Even Boyle, acute as he was, tried to help himself by the false assumption that the increase in weight was due to a ponderable fire-stuff.² It was sought to show by pure philosophy alone, without the faintest shadow of proof, that air was essential to calcination and similar processes, by assuming that it must be present in order to take up the escaping phlogiston. This expedient, first brought forward by Becher and Stahl, was made use of again and again by later phlogistonists.

While these latter imagined that they had thus correctly interpreted the part played by the air, they followed Stahl's example in paying no heed to the observed alteration in weight, either regarding this as accidental or making the

¹ The earliest of such investigations, which yielded extremely valuable observations on the increase in weight of the metals and the part played by the air in their calcination, were undertaken by Jean Rey, Hooke, Mayow, and Boyle in the seventeenth century. Rey and Mayow came very near to explaining the results of their experiments correctly.

² Boerhave showed the weakness of such an assumption by proving that the weight of certain metals, *e.g.* silver, remained the same, whether they were at the ordinary temperature or at a red heat. He, therefore, expressed the opinion that an increase in weight on calcination depended upon the addition of a "saline ingredient" (*salziges Theilchen*) from the air.

most unhappy attempts at explaining it. Thus we find Juncker, a pupil of Stahl's, pointing out that the metallic calces were denser than the metals, and therefore heavier,—an utter confounding of the absolute weight with the specific gravity, and also a wrong assertion, since Boyle had already shown that the calces were specifically lighter than their corresponding metals. Equally unscientific was the assumption that the phlogiston which escaped in these processes possessed a negative weight, and that, therefore, the residual product must be the heavier: even Guyton de Morveau and Macquer fell into this gross error. True, the most able chemists of the phlogistic period did not concur in such absurd opinions, but maintained that it was the business of physicists to investigate such matters.¹ As a matter of fact, it remained to the physicist Lavoisier to give the right explanation of this, and, with it, that of combustion and similar processes.

*Development of some particular Theoretical Views in the
Phlogistic Period.*

It is necessary to make oneself acquainted with the growth of the more important chemical ideas of this time, in order to properly appreciate the advances which they show upon those of the preceding periods, and also in order to comprehend the connection existing between the theoretical views of the phlogistic era and of that new one which begins with Lavoisier. We have here to deal with the meanings attached to the terms "element" and "chemical compound," and also with the ideas of the phlogistonists upon chemical affinity.

Views regarding Elements and Chemical Compounds.—The position which Boyle took up with respect to

¹ Some chemists there were who did not regard the above observations on the increase in weight of metals when calcined as meaningless; Tillet, for example, who made a communication to the French Academy in 1762 upon the increase in weight of lead, calling special attention at the same time to the fact that a fit explanation of this had still to be given.

the question of the elements has already been spoken of; he it was who established the scientific term "element," in that he regarded as elements those actual constituents of compound bodies which were capable of isolation, and which could not be broken up into simpler substances. With the increase of means for deciding the question whether any substance is in this sense an element or not, the boundary line between elements and chemical compounds became more and more altered in position, but at the same time sharper. Boyle further cherished the idea that the elements attainable by chemists were not the ultimate constituents of matter.

Notwithstanding the clearness with which Boyle set forth the conditions which an element, according to his view, must fulfil, we find amongst his contemporaries' and their successors a tendency to go back to the alchemistic elements, and even to the Aristotelian. Willis, Lefèvre, and Lemery associated earth and water with the three elements of Basil Valentine and Paracelsus; Becher also adhered to those three under other names, adding water to them; and even Stahl could not free himself from ideas of this kind.

The erroneous assumption of the phlogiston theory—that the products of combustion and calcination, *i.e.* acids and metallic oxides, were simple, and the original substances compound—had the most serious consequences in keeping back a knowledge of the true elements. While Boyle appeared inclined to reckon the metals among the latter, their compound nature was never questioned from the time of Stahl until the fall of the phlogistic doctrine; and, conversely, the metallic calces and compounds produced in an analogous manner (*e.g.* sulphuric acid, phosphoric acid, and water) were regarded as elements. Sulphur and phosphorus belonged of course to the compounds. Phlogiston itself, the supposed existence of which was due to this inversion of actual relations, was regarded, on the other hand, as an element. Only after this hypothetical state of matters had been set aside by the proof that instead of the escape of phlogiston the absorption of oxygen must be substituted,

and instead of the assimilation of phlogiston the withdrawal of oxygen, did Lavoisier bring light into the prevailing confusion—a confusion which was being continually augmented by the addition of contradictory facts.

With respect to the term “chemical compound,” and the formation of such, ideas were developed during this period which contained much that was sound, and which indicated an advance over previous ones: this is, of course, apart from the erroneous assumption that those bodies which were afterwards recognised as being simple (many metals and some metalloids) were compounds of their oxides with phlogiston. Boyle contributed materially, by the clearness of his views, to an insight into the nature of chemical compounds, and to a recognition of their dissimilarity to simple substances. Boyle, Mayow, and especially Boerhave, gave utterance to the weighty tenet that the characteristic properties of substances which combine together chemically do indeed disappear after such combination, but that nevertheless they are not lost, but are still present in the compound. At that time it was necessary to defend this truth, which became more distinctly formulated later on in the law of the Conservation of Matter, against the old delusion that the formation of a compound was synonymous with the creation of a new substance. How clearly the investigators just named had grasped the meaning of the term “chemical compound,” is shown by the sharp distinction which they drew between such a one and a mixture of its components.

Analytical chemistry, which was gradually developing, helped towards the understanding of the composition of substances, for by its means certain constituents of salts and of other compounds could be distinguished. So long, however, as analysis remained merely qualitative, and no account was taken of the proportions by weight in which substances combined, any considerable development of the meaning of the term “chemical compound” was not possible; this was reserved for the succeeding age.

The defective knowledge of the quantitative composition of substances forced chemists back upon conclusions drawn

from analogy, when they wished to obtain a survey of the compounds known. It was to the endeavour to explain similar phenomena by the assumption of a common principle that the phlogistic theory owed its origin. Acids, salts, and metallic calces were looked upon as being of analogous composition, both on account of their behaviour and their modes of formation. The distinct recognition of the fact that salts were produced by the combination of acids with bases was one of the greatest achievements of the phlogistic period. Before the term "salt" assumed such a definite form, indistinct ideas on the subject were very prevalent; we have only to recall that even such a man as Stahl used the word for acids and alkalies as well as for salts proper. After Boerhave, Geoffroy, and Duhamel had succeeded in giving greater precision to the conceptions regarding these classes of compounds, Rouelle was able (in 1745) to define salts once for all as the products of the union of acids with bases,—and he further drew a sharp distinction between neutral salts (*sels neutres parfaits*) on the one hand, and basic and acid salts on the other.

The characteristics of salts which formerly obtained—their solubility in water and their taste,—therefore fell to the ground, seeing that Rouelle included the insoluble silver and mercurous chlorides among them.

While Rouelle's views regarding the alkaline salts were perfectly sound, he could not throw off the old idea that the vitriols and other metallic salts consisted of metal and acid; it fell to Bergman to show that this was erroneous by the proof that it is the metallic calces and not the metals themselves which combine with acids to salts.¹ What an advance is shown by those definite conceptions on the composition of salts, as compared with the vague ideas that even Stahl not long before had given utterance to, viz. that salts were made up of an earth and water!

¹ If the following passage from Geber's *Testamentum* is genuine, then he was already on the way towards the true explanation of this. The passage is: *Ex metallis fiunt sales post ipsorum calcinationem.*

Views regarding Chemical Affinity and its Causes. —The old assumption that those bodies have an affinity for one another which have something in common, that affinity, in fact, is conditioned by this, according to the axiom *similia similibus*, held its ground in speculative minds even into the eighteenth century. The word *affinitas*, which expresses this idea, and which was already employed by Albertus Magnus, presupposes therefore the similarity of substances which interact with one another. Boerhave, on the contrary, stoutly maintained that it is unlike substances which show the greatest tendency to combine with one another; and, notwithstanding that the reason given for the combination of bodies is exactly the opposite of what was originally taught as such, viz. their dissimilarity, still the name “chemical affinity” or “affinity” for this force has been generally retained.¹

After the time of Glauber, and especially after that of Boyle, much attention was paid to the processes in which the forces of affinity manifest themselves. Cases of so-called simple elective affinity (*attractio electiva simplex*, a term which originated with Bergman) were interpreted correctly by both the chemists just named, and also by Mayow; for instance, the expulsion of ammonia from salmiac by fixed alkali, by the assumption that the attraction of the latter for hydrochloric acid was greater than that of this acid for the ammonia (*fluchtiges Leugensalz*). Observations of this kind on the expulsion or precipitation of bases or acids from salts, by substances endowed with stronger powers of affinity, soon induced chemists to work out the order in which analogous bodies were separated from their compounds by others. The observations on the precipitation of metals and on the expulsion of various acids from salts by means of sulphuric and nitric acids, among others, may have tended in an especial degree to make clear the different strengths of affinity in analogous bodies. The collation of numerous investigations on the behaviour of acids and bases to salts, and of metals to metallic salts, yielded tables of

¹ These terms were temporarily replaced by others, e.g. *rapport* (Geoffroy), *attractio* (Bergman).

affinity, *Tables des rapports* (first published by Geoffroy in 1718 in the *Memoirs of the Paris Academy*), in which similar substances were so arranged that their affinity to the dissimilar ones placed outside the table gradually decreased.

The following table will serve to elucidate Geoffroy's principle :—

SULPHURIC ACID.	FIXED ALKALI.
Fixed alkali	Sulphuric acid
Volatile alkali	Nitric acid
Absorptive earth	Hydrochloric acid
Iron	Vinegar
Copper	Sulphur.
Silver.	

These tables of affinity remained in use for a considerable period, although it was apparent that they stood in need of amendment, and were frequently modified and enlarged. Their deficiencies became especially obvious when chemists began to recognise more fully the influence of heat upon the progress of chemical reactions, and observed that some, whose course under ordinary conditions was perfectly well known, proceeded in an exactly opposite direction at a higher temperature; Stahl, for instance, had noted this correctly in the interaction of calomel and silver at a lower, and of chloride of silver and mercury at a higher, temperature. Such reciprocal reactions led to the proposal to prepare tables of affinity for medium and high temperatures, both for wet and dry (*i.e.* fusion) reactions. Bergman made the attempt in 1775 to work out this proposal of Baumé's by investigating the mutual behaviour of a very large number of compounds, with the result that the doctrine of chemical affinity was materially advanced, in so far as this was possible by such empirical work.

The results of his extended researches were utilised by Bergman for the setting up of a theory of affinity, which will be most conveniently considered in conjunction with Berthollet's doctrine of affinity (see the history of the doctrine of affinity in recent times). But even previously to the efforts of both of these men, the cause of this affinity was a

subject of frequent reflection and of far-reaching speculation. Boyle's lucid conception—that the small particles (of which, in his view, different bodies were made up) attract each other—has been already mentioned. The greater or lesser degree of this mutual attraction of heterogeneous substances depended upon the form and position of each small particle. He did not, however, specially work out this idea, which lay at the root of his corpuscular theory, doubtless because he was so sagacious as to see that he could not arrive at any knowledge with regard to the shape of atoms. Lemery, on the other hand, gave a loose rein to his fancy upon this question. According to him, the combination of two substances—*e.g.* of an acid with a base—depended upon one of the small particles being sharp and the other porous; by the fitting of the points into the cavities, combination was effected. He further attempted to explain the throwing down of precipitates, the solution of metals in acids, etc., in a similar manner.

The force which the mutual attraction of the particles calls forth was regarded by many, *e.g.* by Buffon (who occasionally took part in the discussion of theoretical chemical questions), as identical with that of gravitation. But Bergman, who was also inclined to this assumption, justly pointed out that, since these particles act upon one another at the smallest possible distances, this force must be exerted differently from that of gravity; and Newton, who also turned his attention to the point, likewise assumed a difference between affinity and gravitation.

It was impossible that this subject which dealt with the phenomena of affinity could develop greatly in the phlogistic period, since the proportions by weight in chemical processes were hardly thought of at all. But the purely qualitative investigation of a large number of reactions, from whose outcome conclusions were to be drawn regarding the interaction of individual components, had the effect of maturing much good fruit, so that the restless endeavours of chemists to enlighten themselves upon such questions turned out by no means useless.

This indeed applies generally to the attempts of that age in questions of theoretical chemistry—attempts which were on the whole unhappy. The chief gain was on the practical side, in the rich material accumulated by observation, the complete application of which was reserved for the new era.

The most important achievements in practical chemistry during this period will be touched upon briefly in the following section, in so far as they have not already been gone into in the general part.

History of Practical Chemical Knowledge in the Phlogistic Age.

The question of the composition of substances—that problem which had been recognised as fundamental from the time of Boyle—could only be solved by the experimental method; it was analytical chemistry, which had developed since that time, which was to lead to this knowledge. This indispensable branch of the science proved itself especially useful to applied chemistry, whose growth also falls to be recorded here. The products of technical importance lead us, lastly, to those chemical compounds, a knowledge of which was of moment at that time, and therefore also to the pharmaceutical preparations and to a description of the state of pharmacy during the phlogistic period.

Development of Analytical Chemistry.—Although the question of the composition of chemical compounds was still in a rudimentary stage, and a solution of it in such a sense as we understand that word to-day was not to be expected, yet great attention was paid during the phlogistic period to those reactions by which it was possible to detect substances with certainty. Qualitative analysis, of which we had only the small beginnings to record in the iatro-chemical age, was developed by the labours of Boyle, Hoffmann, Marggraf, Scheele, and especially Bergman, in such a way that the observations of antiphlogistic chemistry which bore upon it could be accepted as valuable contribu-

tions. When we take into account the then prevailing neglect of the proportions by weight of reacting substances, it causes us no surprise that methods of quantitative analysis were hardly at all applied: and yet, in spite of this, we meet with several notable advances both as regards solid and gaseous bodies.

The analytical investigation of substances in the wet way was greatly advanced by Boyle, and this in a systematic manner as compared with the more scattered, although valuable, observations of Tachenius. Boyle it was who introduced the word *analysis* for those chemical reactions by which individual substances could be recognised in presence of one another. For the carrying out of such reactions he employed certain reagents, of which he possessed, for his time, an extensive knowledge. It was with him that the systematic employment of plant juices originated, either in solution or fixed upon paper, for the recognition of acids, bases, and neutral substances, and for this purpose he studied and made particular use of the colouring matters in the juices of litmus, violets, and corn-flowers. Besides these general reagents, which served to distinguish important classes of compounds, Boyle introduced many other characteristic ones which allowed of the recognition of individual substances in the form of precipitates. For the detection of sulphuric and hydrochloric acids, respectively, he used solutions of calcium and silver salts, and *vice versa*. Ammonia he recognised by the production of a cloud when it came in contact with hydrochloric or nitric acid; copper salts by the blue solution which they gave with excess of volatile alkaline salt; solutions containing iron by the black coloration they yielded with infusions of tanning stuffs¹ (from gall apples, oak leaves, etc.) He was also sometimes happy in the way in which he applied careful observations on the precipitation of certain metals by others, as tests for these.

The salt solutions found in nature, mineral springs in

¹ The prescription for preparing black iron ink from gall apples and iron vitriol is due to Boyle.

particular, had before this time stimulated chemists to find out the substances which they contained. Some advances in the analysis of mineral waters became noticeable at the end of the seventeenth and in the eighteenth centuries, and we find at the same time the chemists engaged on the subject inspired with the wish to prepare those natural products artificially; but the knowledge requisite for doing this, *i.e.* a knowledge of the true, and especially of the quantitative, composition of these, was wanting even at the end of last century. Hoffmann investigated a large number of mineral waters, and proved the presence in them of carbonic acid, iron, common salt, and salts of magnesia and lime, showing at the same time how to test for these; he showed further what were the characteristics of alkaline and sulphur waters. On the other hand, he demonstrated the incorrectness of previous statements as to the presence of gold, silver, and arsenic in such waters, and explained the connection between the occurrence of such exceptional salts as alum and copper vitriol and the nature of the soil at those places. He frequently made use of the crystalline form to distinguish different salts.

The observations of Marggraf materially enlarged the acquaintance with reagents suitable for the detection of substances, and also the knowledge of the composition of many compounds. He used, for instance, a solution of prussiate of potash to test for iron, and applied the different colorations which potash and soda salts impart to a flame for their detection. The behaviour of many salts to caustic potash enabled him to arrive at their composition; thus he proved that gypsum consisted of sulphuric acid and lime, and that this acid was also present in heavy spar. As already mentioned, he made use of the microscope for getting at the crystalline forms of different substances.

That Scheele owed his mastery in the discovery of new substances to the gift of deducing their presence from certain reactions, and that he, therefore, greatly enriched analytical chemistry by a multitude of observations, hardly requires to be stated. But, although in his knowledge of

the chemical behaviour of bodies he was equalled by hardly one of his contemporaries, he did not apply this knowledge systematically, as Bergman did, thereby laying the firm foundation for the methodical use of reagents, and, with it, of qualitative analysis. The reactions which the latter made use of as tests for baryta, lime, copper, sulphuretted hydrogen, and sulphuric, oxalic, arsenious, and carbonic acids, etc., are those in vogue at the present day. Bergman also drew attention to the general application of the fixed alkalis for precipitating solutions of metals and earths: to many other reagents, such as sublimate, sugar of lead, and liver of sulphur; and also to modes for estimating precipitates and separating salts. The first methods, by which it was possible to test minerals, especially ores, completely, were due to him, viz. their digestion with hydrochloric or nitric acid, or their fusion with carbonate of potash.

Qualitative analysis in the dry way made considerable advances in the eighteenth century by the increasing use of the blowpipe, the value of which in the examination of ores was recognised especially in Sweden. Bergman and Gahn, together with the mineralogist Cronstedt, were chiefly instrumental in introducing it into chemistry: they employed at the same time borax, soda, cobalt solution, and other reagents, and also made use of the difference between the inner and outer flames. But it was through Berzelius that the blowpipe became universally employed as a most important aid in analysis.

Attempts not merely to test for substances qualitatively, but also to determine their quantity, were few in number up to the time of Lavoisier, and yet it is evident from many statements made by Boyle, Homberg, Marggraf, Bergman, and others, that they sometimes endeavoured to take the proportions by weight into account. How otherwise is it possible to explain Marggraf's accurate determination of the weight of the precipitate obtained by dissolving a given quantity of silver and precipitating the solution with common salt; or Black's estimation of the weight of the precipitate obtained by adding carbonate of soda to a solution of sulphate of mag-

nesia which corresponded to a definite amount of *magnesia alba*, in order to prove the constant proportion of fixed air in the latter? Mention must also be made here of the determination of the weights in cases of metallic precipitates by Bergman and others. Bergman was the first to proceed on the principle that an element should not be itself isolated and estimated according to its own weight, but separated in the most convenient form as an insoluble precipitate, *e.g.* lime earth as oxalate of lime, and sulphuric acid as sulphate of baryta.

In pneumatic chemistry, too, the necessity became strongly felt of being able to detect different gases in presence of one another by means of reagents, and to estimate their relative volumes, *i.e.* estimate them quantitatively. For this purpose special absorptives were used, by the action of which the differences in the gases had first been noticed. Thus caustic potash was found to be suitable for the absorption of carbonic acid, and saltpetre gas (nitric oxide), hydrate of protoxide of iron, moist sulphuret of iron, or phosphorus, for the absorption of oxygen. Of course the results of such quantitative analysis were very inexact.¹ Cavendish succeeded in making a very accurate determination of the oxygen in air by the method suggested by Volta, *viz.* by exploding with hydrogen. Unlike previous experimenters, he found the composition of the air constant, the oxygen amounting on the average to 20.85 per cent; the mean, as determined at the present day, is 20.9 per cent.

As the foregoing short account shows, a great deal of preparatory work, which chiefly required perfecting in the quantitative direction, stood ready to hand at the period which began with Lavoisier. The most important features and principles of chemical analysis were contained in those preparatory researches, and only waited for development.

¹ Priestley and Scheele found that the proportion of oxygen in air varied between 18 and 25 per cent. The term "eudiometry" [*eûdos*, fine (applied to weather), and *μέτρον*, a measure] came into use then because it was supposed that the purity of the air was arrived at by the determination of its oxygen; and it has continued to be employed in gas analysis in spite of its inaptness.

The State of Technical Chemistry in the Phlogistic Period.

The efforts of many eminent chemists, among whom we may mention Boyle, Kunkel, Marggraf, and Macquer, were frequently directed to applying their scientific experiences of chemical processes to the advancement of particular branches of industry. Technical chemistry thus made good progress during this period. We come across the beginnings of great chemical industries, and are able to perceive the development of a knowledge of technically important chemical preparations, the manufacture of which has increased during this century in an unlooked-for degree.

The distinction between applied and pure chemistry was universally recognised towards the middle of the eighteenth century. Serviceable text-books, treating of particular branches of technical chemistry, were not wanting, the conjunction of theory and practice, so necessary for the welfare of the latter, being thus cared for. Analysis was also successfully brought into the service of chemistry, especially in the working-up of ores. Even so early as 1686 Charles XI. of Sweden had recognised the value of such investigations, and had caused a technical laboratory to be built. Here, under Harné's superintendence, all sorts of natural products (such as ores and other minerals, soils, etc.) were examined, and researches were instituted, with the object of rendering chemical products of practical use, and of applying in daily life the various results obtained.

In metallurgy the several modes of procedure underwent only slight changes, but light was thrown upon many processes, which had hitherto been wrongly explained, in consequence of the clearer comprehension of chemical reactions. The results of the researches of Bergman, Gahn, and Rinman came to be of use in the manufacture of iron and steel, the difference between these being traced to its true reason only at the end of the phlogistic period. Marggraf taught an easier mode of preparing zinc from calamine in closed chambers, with exclusion of air as far as possible, and thus made this useful metal more available. The

manufacture of brass was materially improved by Duhamel de Monceau, and that of cast-iron and steel by the many-sided Réaumur. The production and working-up of particular metals, *e.g.* the engraving, tinning, and gilding of iron, the silvering of copper, etc., were developed in many ways by Boyle and Kunkel.

A highly productive field was opened up for the ceramic industry by the accidental discovery of porcelain, the manufacture of which, although carried out on a large scale at Meissen, remained a secret until it was successfully solved at Sèvres in 1769 by the carefully planned experiments of Réaumur and other later chemists, notably Macquer. Improvements and novelties in the manufacture of glass were introduced by Kunkel and Boyle, *e.g.* in the preparation of ruby glass and in glass painting. Dyeing was likewise enriched by the experiences of energetic chemists. New colours, chief among which was Prussian blue (discovered quite accidentally in 1710), together with paints, such as mosaic gold and Scheele's green, were made available for industrial purposes. And chemists, among whom Stahl, Hellot, and Macquer must be particularly mentioned, endeavoured not only to prepare and apply colours by practical recipes, but also to aid the manufacturer by speculations upon the modes in which dyeing processes are brought about. Dyes were divided into two classes, according as they were capable of being fixed upon cloth with or without mordants.

Those technically important preparations, of which an intimate knowledge was first gained in the phlogistic age, constituted a valuable introduction to the chemical industries of to-day. At that time the tendency of chemists was to inquire whether this or that substance was technically useful, just as in the preceding period they had tested chemical compounds for their application to medicine. The manufacture of acids and alkalis, the chemical industry which constitutes the basis of nearly all others, was last century only in its infancy, although even then some of these products began to be made in considerable quantities.

Thus Boyle tells us that nitric acid was manufactured from saltpetre in special "distilleries" (*Brennerien*) to more advantage than was the case before, by improved methods worked out by Stahl and others. Rouelle was the first to show how it could be concentrated by distilling it with oil of vitriol. Sulphuric acid was first manufactured on the large scale in England (by Ward of Richmond) about the middle of the eighteenth century, by burning sulphur with the addition of saltpetre. The perishable and at the same time costly glass balloons in which the process was carried out were soon replaced (at first in Birmingham) by leaden chambers, which are still indispensable for this manufacture; the continuous working of these chambers is an achievement of our own century. The preparation of fuming sulphuric acid from "weathered" iron vitriol had been known long before that of oil of vitriol itself, which last, moreover, received its name because of its production from this salt. The manufacture of the fuming acid, based upon the old observations of Geber and Basil Valentine, was first carried on at Nordhausen in the Harz (whence its name of Nordhausen sulphuric acid, still in vogue), being removed later on to Bohemia. The time for the technical application of hydrochloric acid and the chlorine generated from it was not yet come; the corresponding hydrofluoric acid, however, was used for etching glass by Schwanhardt of Nurnberg so far back as the seventeenth century.

The alkalies and their carbonates were, as in ancient times, obtained from the ashes of plants, carbonised tartar, and incrustations on the soil, to be used for the production of soap, glass, etc. The discovery of the practical preparation of soda from common salt, which revolutionised industrial chemistry, was reserved for the beginning of the present epoch; but even so early as the first half of the eighteenth century some remarkable observations were made which showed that it was possible to convert salt first into sodic sulphate, and then the latter into soda,—reactions which, as he himself tells us, were turned to use by Leblanc, the gifted originator of the soda industry.

Duhamel de Monceau, one of those who showed how to transform common salt into soda, deserves praise for introducing suitable processes for the preparation of various products of technical importance,—salmiac, starch, soap, etc. We find the clearer knowledge of chemical reactions resulting in improvements in old processes generally, and many new manufactures created or at least prepared for, *e.g.* the beet sugar industry by Marggraf's discovery, already mentioned.

*Knowledge of other important Compounds during the
Phlogistic Period.*

The increase in the knowledge of the elements and of chemical compounds—which, although of no special technical value then, were partly destined to become so—was quite remarkable in the phlogistic period, so that it is worth while to take a short survey of these here. To the elements known at that time (although they were not regarded as such) various new ones were added, of which we may mention phosphorus, chlorine, manganese (isolated by Gahn in 1774), cobalt (Brandt, 1742), nickel (Cronstedt, 1750), and platinum (Watson, 1750). The discovery of these was usually preceded by a thorough investigation of their compounds, although chance sometimes came into play, *e.g.* in the isolation of phosphorus. This last discovery produced an extraordinary sensation among the educated circles of Germany, England, and France, on account of the marvellous properties of the new body, and excited chemists in an unwonted degree. Brand, a Hamburg alchemist, succeeded in 1669 in obtaining phosphorus by distilling the residue from evaporated urine, and gave it the same name as the Bologna stone or *phosphor* (which was sulphide of barium, prepared by heating the sulphate with carbon), already known. The two leading chemists of the day, Boyle and Kunkel, endeavoured for years to discover the secret of its preparation, and ultimately succeeded, contributing thereby at the same time to a better knowledge of the element.

Of the chemical compounds prepared artificially, it was the combustion- and calcination-products of the elements, *i.e.* acids and metallic oxides, which awakened the most interest, in accordance with the tendency of the age: and accompanying this, the salts formed from these bodies were carefully studied. A good deal has already been said with regard to the knowledge of these substances. Although the views as to their composition were quite erroneous, the correct interpretation which came later was materially aided by the accurate investigation of their behaviour.

Of acids as combustion products, phosphoric acid deserves the first mention. It was discovered by Boyle, and its nature elucidated by an admirable research of Marggraf's, who showed how it was produced by burning phosphorus, and also by treating the latter with nitric acid: he likewise explained its production from urine. Further, that the amount of phosphorus present in the latter depended upon the nutriment taken, was distinctly stated by him. Scheele and Gahn were the first to prove the presence of phosphoric acid in bones. That the earliest accurate knowledge of the combustion-products of sulphur, coal, and of gases containing oxygen generally, belongs to the second half of the eighteenth century, has been already mentioned. Cavendish proved the composition of nitric acid by its synthesis from nitrogen and oxygen (in presence of water), but he obscured the clear result of his researches by phlogistic accessories.

The many investigations of the products of calcination of the metals and semi-metals greatly advanced the knowledge of these. We may mention here the recognition of white arsenic as the calx of the metallic arsenic, the discovery of molybdic and tungstic acids, and the investigation of the behaviour of quicksilver calx upon heating—so pregnant in its results, etc.

The knowledge that a salt consisted of an acid and a base facilitated the survey of many compounds widely apart from one another. Marggraf, for instance, showed that sulphate of potash had an analogous composition to gypsum and heavy-spar, although these were so unlike it.

The definite distinguishing of alum earth from lime earth, that of the latter from magnesia¹ (Hoffmann and Black), and that of potash from soda (Duhamel) belonged, with many other discoveries, to the phlogiston theory in its prime, and were of great service to the succeeding period. A large number of new salts became known, among others salts of manganese and bismuth (including the basic nitrate of bismuth, so much valued as a cosmetic), compounds of cobalt, nickel, platinum, etc. And the qualitative composition of many salts, whose nature had hitherto been quite misunderstood, was correctly explained, *e.g.* that of alum, borax, calamine, and other compounds.

Organic Preparations.—The knowledge of organic compounds was likewise much advanced, especially by Scheele, who devised methods for discovering and isolating organic acids. While new fields were thus opened up at the close of the phlogistic period, those organic substances which were already known were also further investigated. It is true that the real composition (even qualitative) of all these carbon compounds remained unrecognised, and this complete ignorance hid itself behind meaningless expressions and periphrases; thus oil and water, or a combustible and a mercurial principle, were assumed as the constituents of alcohol. It was Lavoisier again who pointed out the right path here, by proving that carbon, hydrogen, and oxygen were the constituents of this as of most other organic substances, and by devising modes for determining the proportions by weight of the elements just named.

Spirit of wine and the ethers which could be prepared from it, together with common ether itself, were the subjects of frequent investigation, so that they came to be prepared fairly pure. Spirit of wine in especial was employed in analysis for the separation of different salts, and endeavours were made to deduce the amount of alcohol in aqueous solutions of it from its specific gravity: the

¹ Silicic acid was still reckoned for a long time among the earths as "vitrifiable earth," although Tachenius had recognised its acid nature (cf. p. 90).

beginnings of alcoholometry are to be found with Réaumur in 1733 and Brisson in 1768. With respect to its formation in spirituous fermentations opinions were very confused; many, indeed, disputed this formation, assuming its pre-existence in the wine must, etc.

Ether, which was termed *spiritus vini vitriolatus* or *æthereus*, became known through the labours of Frobenius (about 1730), Hoffmann, Pott, Baumé, and others, and was used medicinally admixed with spirit of wine (Hoffmann's drops). The erroneous idea that it contained sulphur prevailed for a long time, until it was finally done away with by the investigation of Valentin Rose the younger (in 1800).¹ The name "sulphur ether" arose from this. At that time any pungent volatile liquid was termed an ether.

Nitrous² ether, muriatic ether, and acetic³ ether, so named because of their origins, were likewise carefully investigated, and were valued as officinal preparations. Scheele's acuteness of observation is well shown by the fact that he recognised the necessity of having a mineral acid present in the formation of ethers of weak acids, such as acetic and benzoic, a point which had been overlooked before his time.

The knowledge of the organic acids was materially extended during the phlogistic period, especially towards its close. Acetic acid, which had been longest known of any, was now prepared in the concentrated pure state as the glacial acid, and its combustibility was observed by Lauraguais. Kunkel, Boyle, and others believed in the identity of the acetic acids prepared by fermentation and by the distillation of wood, without, however, being able to adduce a definite proof of this; the latter was furnished by Thénard in 1802. The similarity between formic acid, discovered by Wray in 1760, and acetic acid was early noticed, and led to confounding the one with the other. Scheele showed how to prepare a large number of acids from

¹ Prior to this date, Hoffmann and Macquer correctly assumed that ether was formed from alcohol by the elimination of water.

² Our present nitrous ether, admixed with a little nitric ether, aldehyde, etc.

³ Ethyl acetate.

plant juices, by first forming their lime or lead salts, and then decomposing these with suitable mineral acids, usually sulphuric. In this way he discovered tartaric acid, which had hitherto been overlooked in spite of the fact that tartar had been known for a long time; also citric, malic, and oxalic acids, the last of which he prepared by acting upon sugar with nitric acid, and which he recognised as being identical with the compound obtainable from wood-sorrel. By treating milk sugar with nitric acid he was led to the discovery of mucic acid, and by investigating sour milk to that of lactic acid, while he found uric acid in (bladder) stones. For other acids, already known, he devised improved methods of preparation, *e.g.* for gallic and benzoic. Lastly, his discovery of prussic acid in 1782, by decomposing yellow prussiate of potash with sulphuric acid, is worthy of note; the masterly investigation of it which he made enabled him to give its qualitative composition with tolerable accuracy.

The fatty oils and animal fats were frequent subjects of investigation, without their composition and chemical behaviour, especially towards the alkalies, becoming any clearer; and this in spite of an important observation made by Scheele in the discovery of glycerine, or *Oelsüss*, as he termed it, by acting upon olive oil with litharge. The importance of this observation was only recognised at a much later date. Only the small beginnings of preparatory researches are to be seen in the chemistry of the sugar varieties and of other products of animal and vegetable metabolism, such as the ethereal oils, albumens, etc.

Condition of Pharmaceutical Chemistry.

The interests which chemistry and pharmacy had in common resulted in their mutual beneficial action upon one another. A large number of famous investigators owed to the practice of pharmacy their incitement to the study of purely chemical phenomena; of these we may mention Kunkel, the Lemerys (father and son), Geoffroy, Rouelle, Neumann, Marggraf, and Scheele. While they and others

contributed a wealth of the most valuable observations, indeed of fundamental discoveries, to chemistry, pharmacy was at the same time materially advanced, not only by those discoveries, but also by special pharmaceutical researches. The chief gain for pharmacy lay in its intimate fusion with pure chemistry. On the other hand, the work required in apothecaries' shops proved itself the best preparatory training for future chemists. The scientific taste was nourished by excellent text-books on pharmaceutical chemistry, *e.g.* Baumé's *Éléments de Pharmacie Théorique et Pratique* (1762), Hagen's *Lehrbuch der Apothekerkunst* (Hagen's Text-Book of Pharmacy), and firmly fixed by the founding of pharmaceutical laboratories: the growth of the latter belongs, however, more to the present epoch.

Many additions were made during this period to the medical treasury by pharmaceutical chemistry. Of the new medicines which then came into vogue, and whose nature was sometimes involved in mystery until they ceased to be secret remedies, the following important ones may be mentioned:—Carbonate of ammonia, which was contained in the famous "English drops"; sulphate of potash, valued under Glaser's designation of *sal polichrestum*, which was obtained by detonating sulphur with saltpetre; sulphate of magnesia, first prepared from the Epsom (spring) water by Grew in 1695, and termed *sal anglicum*, and, later on, bitter salt; and *magnesia alba*, obtained from the mother liquors in the preparation of saltpetre by means of carbonate of potash. Among the preparations of antimony, the *Kermes mineralis*, whose composition was only arrived at correctly during the present century, came into repute. Ferric chloride in alcoholic solution was a favourite remedy in the first half of the eighteenth century under the name of "gold drops" or nerve tincture: the nature of the latter soon became recognised. Hoffmann's drops and the compound ethers were likewise used officinally. Goulard introduced basic acetate of lead after the middle of the last century as a remedy for external use, and it is still called by his name to this day.

Many observations were made with regard to substances of special antiseptic action, Kunkel pointing to the mineral acids for this. The antiseptic properties of iron vitriol and alum were made use of in the impregnation of wood with these salts, according to the proposal of the Swede Faggot.

Concluding Remarks.—The period of phlogistic chemistry must be looked upon as the indispensable forerunner of the new era which began with Lavoisier. The erroneous conception which attended the important phenomena of combustion and calcination, and which spread itself over many other processes, did not prevent the young science of chemistry from developing in a healthy manner. Undoubtedly it was the experimental method which contributed most to this. Hand in hand with this development we find an increasing improvement in the means for observing chemical processes and for establishing the properties of bodies. These advances were due partly to improved apparatus (for instance, the apparatus required for collecting gases), and partly to the use of physical methods of research; and here we may note the more frequent determinations of the specific gravity of bodies in different states of aggregation, and the use of the microscope. The time had not yet arrived when the balance was to be employed with so great advantage for the exact determination of proportions by weight in chemical reactions, although a number of noteworthy beginnings of quantitative analysis are to be found.

It is especially to be noted as characteristic of this period that chemistry now became fully awake to her own proper task, which was to investigate the composition of substances, and to find out the constituents from which they could be prepared. Analytical chemistry was to aid in solving this problem, but useful and important results were achieved by the synthetic method also.

The independent scientific character of chemistry showed itself in the forms which its relations to other sciences assumed. The previous dependence upon medicine and pharmacy ceased; instead of being their servant, chemistry

became their helper and adviser. It also came into close connection with physics, mineralogy, and botany, which resulted in mutual advantage to all of them, and made chemistry the indispensable helpmeet of these. We have only to think of the services rendered to those sciences by eminent chemists, *e.g.* to physics by Boyle, and to physics and mineralogy by Bergman. This coalition with the various other sciences had the effect of opening up new common ground both for these individually and for chemistry. We find the first scientific treatment of mineralogical and physical chemistry during the phlogistic period, and the advances made in organic prepared the ground for physiological chemistry.

Nothing is less justifiable, therefore, than to assert that chemistry was at that time no science, and that it was Lavoisier who created one out of what was, before his time, a science only in name. The record of the services of Boyle, Stahl, Black, Bergman, Scheele, Cavendish, Priestley, and others, is sufficient to prove the erroneousness of such an assumption.¹ In spite of the false hypothesis which lay at the root of the phlogiston theory, it was the latter itself, together with the work which resulted from it, which formed the necessary foundation for the correct standpoints of the succeeding period.

¹ Cf. Dumas's *Leçons sur la Philosophie Chimique* (1837), p. 137; and also the sentence with which Wurtz began his *Histoire des Doctrines Chimiques* (1868): "*La chimie est une science française; elle fut constituée par Lavoisier,*" etc. Volhard investigated this statement and completely overthrew it (*Journ. pr. Chem.*, N. F., vol. ii. p. 1 *et seq.*) The most eminent among the antiphlogistonists, moreover, never thought of calling in question the scientific tendency of the chemical views which they themselves combated.

CHAPTER V

HISTORY OF THE MOST RECENT PERIOD (FROM THE TIME OF LAVOISIER UP TO NOW)

THE beginning of the latest period of chemistry, to which the present generation of investigators still belongs, is rightly associated with Lavoisier's reforms, which turned the chemical science of his day into new paths; he demonstrated the importance of the proportions by weight in chemical reactions, which were wrongly interpreted when these were disregarded. This applied in an especial degree to the processes of combustion and similar phenomena, which Lavoisier was the first to explain correctly. Of course this explanation only became possible after Priestley's and Scheele's discovery of oxygen. If we desire, therefore, to associate the commencement of the new era with any particular event, it must be with the latter important discovery, which has been already described in the history of the period preceding.

Lavoisier's combustion theory, with oxygen as its centre-point, now stepped into the place of the phlogistic doctrine, which had attained to the dignity of a dogma; the chemistry dominated by the latter was thus changed into the so-called antiphlogistic system. A complete transformation of all the ideas respecting combustion and calcination, and therefore respecting the composition of the most important substances, took place, — truly a reform in the fullest sense of the word. For, all the reactions in which the escape of

phlogiston had hitherto been assumed depended, as Lavoisier taught, upon the taking up of oxygen; and, conversely, those processes which had been explained by assuming the absorption of phlogiston, depended upon the separation of oxygen.

Lavoisier showed that substances like sulphuric and phosphoric acids and the metallic calces, which according to the phlogistic doctrine were looked upon as elements, really were compounds; while those regarded as compounds, *e.g.* the metals, sulphur, and phosphorus, he assumed to be elementary.

It will be appropriate here to enter shortly again into the chief points of dispute in which the phlogistic doctrine became involved at the time of the discovery of oxygen (1775), and by which its fall was accelerated. The facts to which the phlogiston theory could not accommodate itself were many in number. To chemists who regarded hydrogen as phlogiston—a frequent assumption—the great difficulty arose of proving whence the phlogiston came which escaped during the calcination of the metals, and the combustion of sulphur, phosphorus, and coal in closed vessels. The reduction of the metallic oxides by hydrogen did indeed appear to allow of a perfect explanation from the phlogistic standpoint, if one paid no heed to the simultaneous formation of water and the diminution in weight of the oxides. But, then, how could a reduction of the metallic calces take place without the presence of phlogiston (hydrogen)? This occurred in the case of those calces which were converted into metal when heated alone in closed vessels. For the production of quicksilver from red oxide of mercury by heat, the phlogistic doctrine was able to offer no explanation. It was, indeed, this reaction, which led to the discovery of oxygen, that brought about the collapse of the theory, and rendered possible the setting up of the antiphlogistic system. And a few years later the keystone was added to the latter by the proof that water, which had hitherto been looked upon as an element, was a compound of oxygen and hydrogen.

GENERAL HISTORY OF CHEMISTRY DURING THIS PERIOD.

Lavoisier and the Antiphlogistic Chemistry (from 1775 to the end of the Eighteenth Century).

Lavoisier's great achievement consisted in doing away with old prejudices, and in the masterly application of scientific principles to the explanation of chemical processes. A rich material of important facts was handed down to him by the phlogistonists; he himself did not add much to this in the way of new chemical observations, but he sifted and collated, from a point of view hitherto unattained, that already at hand, giving at the same time the correct explanation of many processes. We shall not be wrong if we place such services to the credit of his highly-trained physical and mathematical mind, which early freed itself from the bonds of the phlogistic hypothesis. As a physicist Lavoisier was bound to take into account alterations in weight, *e.g.* in the calcination of metals; the properties of the products obtained interested him in a lesser degree. This explains why he himself made no independent chemical discoveries; but the unique service which he rendered in being the first to give a comprehensive and correct explanation of the observations of others remains incontestable.

Lavoisier lived to see his work appreciated in the highest degree; he saw the fruit of his labours, the antiphlogistic system, come out victorious in the fight with the phlogistic, and propagate itself beyond France. — Anton Laurent Lavoisier was born in 1743, a year after Scheele, but how different were the outward circumstances of the two! While the latter was early thrown upon his own resources, Lavoisier had a splendid education given him, and enjoyed special opportunities for acquiring a thorough knowledge of mathematics and physics, which exercised a permanent influence upon the whole tendency of his thoughts and methods of investigation. It was Rouelle who initiated him into chemistry. Even whilst still very young, Lavoisier gained great repute by his scientific investigations, so that

we find him received (as Associate) into the French Academy in 1768, the immediate cause of this being a prize essay upon the most suitable method of street-lighting for large towns.

His earliest chemical work¹—particularly the research upon the supposed transformation of water into earth, the results of which he published in 1770—afford clear evidence of his physical methods. In this he proved that the total weight of the closed glass vessel and of the water which had been boiling in it for a long time remained unaltered, but that the weight of the earth produced was exactly equivalent to the loss in weight of the vessel: the logical conclusion to be drawn from this was that the earth came from the glass and not from the water. What this earth was he did not investigate; on the other hand, Scheele was led to the same conclusions as Lavoisier by examining it qualitatively.

The latter here recognised and laid stress on the use of the balance as a reliable guide in chemical work. Soon after this he busied himself with investigating the reactions involved in the combustion of substances and in the calcination of the metals, making use here of some previous observations by others on the increase in weight during such calcination. With the aid of an exceedingly delicate balance he sought, in the first instance, to estimate exactly the alterations in weight which occurred during these processes, and to get at the reason for this. The results of these labours, amplified by the addition of Priestley's and Scheele's observations on oxygen and its chemical behaviour, formed the foundation of Lavoisier's theory of combustion.

His position had, in the meantime, become a brilliant one; as Farmer-general and, shortly after, as chief director of the saltpetre industry (of which the Government had a monopoly), he had plenty of leisure to devote himself to his own investigations, and to aid the State both by his advice

¹ With regard to Lavoisier's writings, the reader is referred to the *Œuvres de Lavoisier* (publiées par les soins du Ministre de l'Instruction Publique), which were published in Paris in 1862; and to the analyses of his most important papers, given by H. Kopp in his *Chemie in der neuen Zeit* (1874), and by Höfer in his *Histoire de la Chimie*, vol. ii. p. 490 et seq.

and by the introduction of valuable improvements (*e.g.* in the manufacture of potash saltpetre, gunpowder, etc.) Closely related to his work upon combustion were the important researches which he carried out in conjunction with Laplace upon the latent heat of ice and the specific heats of various bodies. It was his clear physical conception of the nature of heat, as opposed to that of many phlogistonists (who could not get rid of the assumption of a ponderable caloric), which enabled Lavoisier to interpret correctly those chemical reactions in which heat was evolved, in particular, the phenomena of combustion.

Notwithstanding the extraordinary services which Lavoisier rendered to science, and, through the latter, to his country, by applying his knowledge and experience for her benefit, he did not escape the fate which befell so many of his fellow-citizens. Impeached under the Reign of Terror, he was condemned to death, and executed on the 8th of May 1794.¹ Amongst all his numerous friends and admirers, only one chemist, Loysel, had the courage to protest against this, but without effect. His more influential colleagues, like Guyton de Morveau, Monge, and especially Fourcroy,² who took part in politics, did not dare to offer any opposition to this crime.

Lavoisier published most of his work in the *Memoirs of the French Academy*, over sixty papers by him being contained in its volumes for the years 1768-87; some others

¹ Much light has been thrown upon this sad event by documents published by Ed. Grimaux, which relate to the death of Lavoisier. It has been conjectured that Marat hastened the proceedings against him from a feeling of petty revenge, because of Lavoisier having unfavourably criticised a treatise of his, entitled *Recherches Physiques sur le Feu*, which appeared in 1780. In the sentence, which was passed after an imprisonment and inquiry extending over five months, it was stated that he was condemned to death "as convicted of originating or participating in a plot against the French nation, the aim of which was to aid the enemies of France; especially in that he had practised every kind of extortion upon the people, and had caused tobacco to be admixed with water and pernicious substances, to the detriment of the health of the citizens who used it."—Cf. Grimaux's work, *Lavoisier, 1743-94, d'après sa Correspondance, ses Manuscrits, etc.* (Paris, 1888).

² Grimaux's publication, just cited, reflects seriously upon Fourcroy's indifference to Lavoisier's fate.

are to be found in the *Journal de Physique* and in the *Annales de Chimie*.¹ His projected plan of publishing an edition of his collected works was only carried out long after his death (in 1862). His *Opuscules Physiques et Chimiques*, which appeared in 1774, contained his ideas upon the nature of gases and his views upon the processes of combustion. In his *Traité Élémentaire de Chimie (présenté dans un ordre nouveau et d'après les découvertes modernes)*, published in 1789, he gave a summary of the most important facts of chemistry, and explained them in accordance with the antiphlogistic theory, which thus received its first textbook: by means of translations of this book the new doctrine was materially propagated.

The researches of Lavoisier which were of greatest moment for the development of chemistry were those which contributed to the founding of the antiphlogistic system, and which led to the overthrow of the phlogistic; these, namely, which treated of the phenomena of combustion, calcination, and respiration. The chief work of his life consisted in his recognising and explaining the part played by oxygen in these processes, and in this lies his abiding service.

The previous observations of Rey, Mayow, and others, who had attributed the increase in weight of the metals during their calcination to an absorption of air, contained only the first germs of a correct explanation of these processes. From the year 1772 Lavoisier busied himself with investigations bearing upon this subject, the results of which he delivered in a sealed note to the French Academy on November 1st of that year. This note stated that by the combustion of sulphur and phosphorus, and by the calcination of the metals, the weight of these substances increased

¹ The dates upon which Lavoisier's papers appeared are of importance for their criticism; we have especially to remember here that the yearly volumes of the *Mémoires de l'Académie* did not correspond with the dates of their publication, but that they were usually brought out several years afterwards (*e.g.* the *Mémoires* for 1772 in 1776, and those for 1782 in 1785). The effect of this disarrangement has been great confusion with regard to the actual time at which this and the other treatise was written by Lavoisier. But, so far as it has been found possible to verify them, those dates are given above.

from the absorption of a large amount of air; and that, by the reduction of litharge with coal in an enclosed space, a considerable quantity of air—a thousandfold the volume of the litharge—was generated. Lavoisier was at this time still quite uncertain as to which portion of the air caused this increase in weight, as to the air itself being a mixture of gases, and especially as to the nature of the process which went on in the reduction of the litharge; he was inclined to regard the generated gas, *i.e.* carbonic acid, as the fluid originally combined with the lead. This uncertainty was brought about by his paying too little heed to the qualitative side of the chemical reactions.

By repeating these and similar researches, however, Lavoisier soon arrived at a clearer perception of the matter, and especially recognised his error with regard to the reduction of the oxide of lead. In 1774 he gave further details of these observations, in particular of the calcination of tin;¹ the investigation was in its main points a repetition of Boyle's, but Lavoisier was able to draw more correct conclusions from it than Boyle had done. A sealed retort, in which some tin had previously been placed, was weighed both before and after being heated, and found equally heavy each time, whence the conclusion was drawn that no fire-stuff had been absorbed; on the retort being opened after cooling, air rushed in, and the whole apparatus showed an increase in weight exactly equal to that which the tin had undergone by calcination. Lavoisier concluded from this that calcination depends upon the absorption of air, *i.e.* that the latter is the cause of the increase in weight.

But although we find in these results the beginnings of his combustion theory, there was still wanting the definite knowledge as to which portion of the air combined with the metals and the combustible substances. Oxygen was discovered independently by Priestley and Scheele in the same year, and they recognised in it the constituent of the air which was necessary for combustion; but Lavoisier held the key to the explanation of his researches in his hand as soon as he

¹ *Œuvres de Lavoisier*, vol. ii. p. 105.

received news of this discovery. How he turned this to advantage is shown in a paper written in 1775,¹ in which the *role* of oxygen for the general explanation of the reactions in question is duly appreciated; it was this gas which combined with the metals, sulphur, phosphorus, coal, and so on. The production of carbonic acid from saltpetre and coal led him to the conclusion that oxygen must likewise be present in this salt,—a point that Mayow indeed recognised a hundred years before this, only that the latter terms it *spiritus nitro-aëreus* instead of oxygen. Strangely enough, no reference is made by Lavoisier to the influence which Priestley's discovery of oxygen (communicated to him by Priestley himself) exercised upon his researches with oxide of mercury and upon his explanation of previous experiments.² Lavoisier in due course arrived at perfect clearness in his explanations, for instance, with regard to the composition of atmospheric air; it was in 1776 that he observed that the combustion-product of the diamond consisted of carbonic acid alone, and in the following year he showed that, by burning phosphorus in a closed vessel, one-fifth of the volume of air in the latter was used up, and non-respirable air remained behind. The results of these researches, together with the observations made by Scheele

¹ Cf. *Œuvres*, vol. ii. p. 125.

² The attitude which Lavoisier sometimes took up with regard to the observations and discoveries of others frequently awakens painful feelings; it is melancholy to see an investigator of so much acuteness and such splendid gifts so unjust with respect to the services of others. Thus Lavoisier makes no mention in his first chemical paper, on the composition of gypsum, of Marggraf's important researches, although these were among the best known of any, while more than their due recognition was awarded to the other chemists who had worked at the same subject. In a similar manner he ignored, in the account of his researches on the composition of water, those of Cavendish which proved the same point (*i.e.* its composition), and of whose results he had positive knowledge. Black's splendid investigations upon fixed air, from which Lavoisier without doubt received the greatest assistance towards his conception of the fixation of gases, he treated in a cold and depreciatory manner, whilst the most trivial objections raised against Black were examined with the utmost minuteness and care. These are unfortunately blots upon Lavoisier's reputation, notwithstanding the lustre with which it has become surrounded through the idealistic historical writings of Dumas, Wurtz, and others.

and Priestley, of which he had, in the meantime, obtained fuller knowledge, and the investigations which he made in 1777 on the combustion of organic substances, the products of which he proved to be carbonic acid and water, enabled Lavoisier to establish the main points of his Combustion or Oxidation Theory as follows¹—

(1) *Substances burn only in pure air (air éminemment pur).*

(2) *This air is consumed in the combustion, and the increase in weight of the substance burnt is equivalent to the decrease in weight of the air.*

(3) *The combustible body is, as a rule, converted into an acid by its combination with the pure air, but the metals, on the other hand, into metallic calces.*

The last sentence contains an idea of great moment, which Lavoisier developed later into his theory of the composition of acids, according to which these latter contain oxygen as the oxygenating or acidifying principle (*principe oxygène ou acidifiant*). To establish this assumption, he both made investigations himself and referred to and made use of those of others; in this way he states that sulphuric acid consists of sulphur and oxygen, phosphoric acid of phosphorus and oxygen, and nitric acid of saltpetre gas (nitric oxide) and oxygen. The true composition of the last acid was first determined by Cavendish, through its synthesis from nitrogen and oxygen in presence of water. Hydrochloric acid being a powerful acid, likewise contained oxygen, according to Lavoisier's assumption, and this applied in still stronger degree to the chlorine produced by its oxidation. Lavoisier further occupied himself with the question—What kind of oxygen-compound does hydrogen yield? without, however, arriving at the correct explanation of this independently; for he expected to find an acid as the product of its combustion, and therefore looked for one. It is the undisputed merit of the phlogistonist Cavendish to have

¹ *Œuvres*, vol. ii. p. 226.

proved that water alone is produced by the combustion of hydrogen.¹

This fundamental observation first proved itself fruitful, however, in the hands of Lavoisier, who was thus enabled to give at once the real composition of water (out of hydrogen and oxygen), while at the same time estimating the relative proportions of these approximately. He also correctly interpreted the decomposition of water by red-hot iron, and its formation from the reduction of metallic oxides by means of hydrogen. The generation of the latter on dissolving metals in acids was likewise satisfactorily and clearly explained. It was precisely this reaction which had strengthened the phlogistonists in their opinion that the metals contained phlogiston, which, being identical with hydrogen, escaped on dissolving these in acids. The composition of water having been arrived at, Lavoisier now saw that the hydrogen came from the water, and that the oxygen of the latter united with the metal to oxide, which then in its turn combined with the acid.²

With the knowledge of this, which came in the year 1783, the last obstacles with which the antiphlogistic system had to contend were overcome; the phlogistic theory could maintain itself no longer, but collapsed. Up to this date Lavoisier was almost alone in the fight against it, having only received material aid from eminent physicists and mathematicians, like Laplace, Monge, Cousin, etc. But now chemists of position began to apply his ideas, at first in France, and very soon in other countries also. Lavoisier's critical treatises, which were directed to showing the untenability of the phlogistic theory, conjoined with his *Traité de Chimie*, gave the final blow to that doctrine.

The main features of Lavoisier's work, which was the means of leading chemistry into new paths, have now been described; but some of his observations and speculations, *e.g.* his researches on the composition of organic compounds, and

¹ With regard to this point and also to Watt's share in recognising the composition of water, cf. H. Kopp's detailed memoir: *Ueber die Entdeckung der Zusammensetzung des Wassers* (Braunschweig, 1875).

² Laplace and Meusnier took an active share in these investigations.

his ideas regarding metabolism in the organic world, will be treated of in the special history of this time. The systematic application of quantitative methods of research, and the unbiassed recognition of chemical processes from a rather physical point of view, led him to interpret correctly the most important phenomena of chemistry, the explanation of which had been sought for in vain by several generations of investigators, enchained as they were by the phlogiston theory. The material which these latter had collected together, especially the observations of Black, Priestley, Scheele, and Cavendish, were indispensable to Lavoisier; we have only to recollect that the discoveries of most importance for his system—of oxygen, and of the true composition of water—were not made by himself. But his genius, far transcending that of any of his contemporaries, enabled him to get at the root of phenomena which they did not comprehend. After recognising that phlogiston had no existence, and that oxygen was the gas necessary for combustion, calcination, and respiration, he translated the obscure and wholly erroneous reactions in which phlogiston was assumed into simple antiphlogistic language.

Although the quantitative method of research, followed by particular chemists both before and during the time of Lavoisier, *e.g.* by Boyle, Black, Marggraf, Cavendish, and especially Bergman, was duly valued, still none of these investigators made use of the balance as an aid to chemical work with such a definite aim and perfect conviction of its significance as he. Lavoisier was penetrated by the truth that no matter is lost during chemical reactions; what many others accepted as being correct, without emphasising it particularly, was for him a law upon which he based his speculations and researches. The weight of a compound body was equal to the aggregate weights of its constituents. Although this last sentence now sounds so simple and self-evident, it had to be proved to those who regarded heat as material; for the evolution of heat which took place during chemical combination was bound to be accompanied by a decrease in weight, if a caloric

was assumed. Lavoisier was kept from falling into this grievous error by his conception of the nature of heat. His *matière de chaleur* had no weight; this he concluded from experiments in which he burnt substances in closed vessels, proving thereby that no diminution in weight occurred. Many expressions of his show that his views upon its nature approximate to the Mechanical Theory of Heat. The phlogistonists, on the other hand, who saw in heat a ponderable substance, were bound to suffer shipwreck with such a false basis to start from.

The antiphlogistic system, the outcome of the proper interpretation of those processes which were designated combustion, calcination, reduction, etc., meant, in fact, a complete reform of chemistry. The more important of the changes which the latter underwent have been already detailed, but it will be convenient here to speak shortly of the most striking alterations thus effected in the views regarding elements and chemical compounds. Hand in hand with the definite shaping of these opinions went the attempts to introduce a scientific nomenclature, which likewise fall to be treated of now.

Boyle's view with respect to the term "element" was retained by Lavoisier; the latter, therefore, regarded as elements those substances which could not be decomposed into simpler ones. But then what immense alterations he made in details here! The metals and the most important metalloids were placed among the elements; compound bodies like the alkalis, ammonia, and the earths were indeed numbered among these also, but not without doubt being expressed as to their elementary nature. Oxygen, also recognised as an element, became, on account of its part in combustion and its capacity for combining with so many other elements, the centre-point of the antiphlogistic system, which, indeed, owed its origin to the knowledge of the behaviour of other elements towards oxygen. The significance which Lavoisier attached to this gas is clearly shown in his theory of acids, just mentioned, and in the tenet that the bases which combine with acids likewise contain

oxygen. The composition of a large number of compounds—oxides, acids, and salts—was thus now rightly interpreted, the phlogistic hypothesis having regarded the substances belonging to the first two of these classes as simple.

The extent of Lavoisier's knowledge and that of his disciples, in particular their views with respect to elements and compounds, is to be seen in the work entitled *Méthode de Nomenclature Chimique*, which was published by the former in 1787 in conjunction with Guyton de Morveau, Berthollet, and Fourcroy. The three last were the first French chemists of note to give up the phlogiston theory and to follow the "new chemistry." To Guyton de Morveau belongs the credit of making the first attempt towards a convenient chemical nomenclature, and thereby of inciting to the publication of the above book.

In this work all substances are divided into elements and compounds. To the former belonged—in addition to light and heat—oxygen, hydrogen, and nitrogen; these formed the first class. The second group contained the acid-forming elements,—sulphur, phosphorus, and coal, to which were added the hypothetical radicals of hydrochloric, hydrofluoric, and boracic acids. The third class comprised the metals, the fourth the earths, and the fifth the alkalies; but Lavoisier considered the elementary nature of the last of these as so improbable that in his *Traité de Chimie* (1789) he no longer included them among the elements. For the nomenclature of the latter, the old names of the metals and of some of the metalloids (*e.g.* *soufre*, *phosphore*, etc.) were retained, while Lavoisier's new names for others of the metalloids (*e.g.* *oxygène*, *hydrogène*, *azote*) were introduced.

Compounds were classified as binary and ternary, and these designations were in great part retained later on, although it was found necessary to extend their meaning as chemistry developed. To the binary compounds belonged the acids, whose names were composed of two words, one of which (*acide*) was common to all, and the other special to each acid, *e.g.* *acide carbonique*, *sulphurique*, *azotique*. In the

case of two acids of one and the same element, the name of that one which contained the less oxygen ended in *eux*, e.g. *acide sulphureux*. The second group of binary compounds embraced the oxygen compounds of the metals, which, as bases, were placed opposite the acids; they were given the generic name of *oxydes*, that of the particular metal in question being added (e.g. *oxyde de plomb*, etc.) The *sulphures* (e.g. sulphuretted hydrogen and the metallic sulphides), *phosphures*, and *carbures* likewise belonged to the class of compounds of two elements, and also the compounds of the metals with one another.

The principal ternary compounds were the salts, produced by the combination of bases with acids: their generic name was got from the latter, with the addition in each case of that of the metal, alkali, or earth in question (e.g. *nitrate de plomb*, *sulfate de baryte*, etc.)

The advance which is shown by this classification of chemical compounds is very great. In place of false assumptions and designations devoid of any system, we find a correct idea of the qualitative composition of substances, and a rational nomenclature according with this. The development of the latter, and the international form which was given to it by Berzelius, will be treated of below.

Guyton de Morveau, Berthollet, Fourcroy.

These three investigators, who, along with Lavoisier, laid the foundation of a scientific chemical nomenclature, exercised a further influence on the development of chemical doctrines by their other labours, the most important of which fall to be considered here. Guyton de Morveau, born at Dijon in 1737, began life as a lawyer (*avocat*), but gave up this career in order to devote himself wholly to chemistry. His first attempt at a chemical nomenclature brought him into close connection with the French Academy, and in particular with Lavoisier, the outcome of which was the book cited above. Elected a deputy in 1791, Guyton de Morveau did his best to render his chemical knowledge and its practical application of use to his country; we have only to recall here his efforts to

employ the air-balloon for strategic purposes in the battle of Fleurus, his activity in helping to found the *École Polytechnique*, in which he subsequently became a professor, and his services as Director of the Mint, etc. The part which he played in politics was less beneficial—it was, in fact, pernicious; for, although an influential member of the National Assembly and of the Convention, he did nothing which could tend to lessen the excesses of the Revolution. He died in Paris in 1816.

To the main service which he rendered, viz. that of having been efficacious in introducing a rational system of nomenclature for chemical compounds, in place of the unmeaning names and confusing synonyms¹ hitherto in use, he added the further one of developing this system by experimental researches in analytical and technical chemistry. He also aided in spreading abroad a knowledge of the labours of Bergman, Scheele, and Black, by making good translations of their works.

Claude Louis Berthollet, born at Talloire in Savoy in 1748, had his home in Paris from the year 1772 on, and showed a wonderful activity in the most various branches of chemistry, especially after the year 1780, when he was elected to the French Academy. He found vent for his great organising talents as a teacher in the Normal and Polytechnic Schools (after 1794), in Napoleon's historical expeditions to Italy and Egypt (in which he took part), and in undertakings for the public benefit. He attained to the highest honours both under the Empire and after the Restoration, and died at Arceuil, near Paris, in 1822. During the last years of his life, regular meetings, attended by eminent *savants*, were held at his house, their proceedings being published in the *Mémoires de la Société d'Arceuil* (1807-1817). At first a phlogistonist, Berthollet frankly declared for Lavoisier's doctrine in 1785.

His experimental researches were especially valuable and

¹ Thus sulphate of potash had five different names, most of which were unintelligible, viz. *sal polychrestum Glaseri*, *tartarus vitriolatus*, *vitriolum potassæ*, *sal de duobus*, and *arcanum duplicatum*.

fruitful during this period. Mention may be made here of those upon ammonia, prussic acid, sulphuretted hydrogen, and chlorate of potash, and upon the practical application of chlorine; he worked out the composition of the three hydrogen compounds just named with substantial correctness. But his researches and speculations upon chemical affinity were of more general and far-reaching significance; his *Essai de Statistique Chimique* exercised at that time and still exercises a most powerful influence upon this question. The cardinal points of his doctrine of affinity will be given in detail in conjunction with the results obtained by Proust (whose work arose from Berthollet's), which led to the knowledge of definite chemical proportions, and which, therefore, belong to the history of the development of the Atomic Theory.

Anton François Fourcroy (born 1755, died 1809) understood as a teacher how to inspire his pupils with enthusiasm, and worked in this way with quite remarkable vigour for the propagation of the antiphlogistic system, aiding the latter also by his writings. The chemical articles which he wrote (after 1797) for the *Encyclopédie Méthodique* contain panegyrics upon the antiphlogistic chemistry which, in his excess of patriotic zeal, and possibly not without an egotistical *arrière pensée*, he termed *chimie française*.¹ Fourcroy set forth the antiphlogistic doctrine in larger works also, among others in his *Système des Connaissances Chimiques*, and his *Philosophie Chimique*, etc.

Born one of an impoverished family, he had to earn the means for his studies under the most pressing circumstances. His work in medicine and natural history led to the honour of his inclusion in the French Academy in 1785, a year after he had succeeded Macquer as professor at the *Jardin des Plantes*. Later (especially after the Reign of Terror), when Fourcroy was on the Public Education Committee, he found an opportunity of utilising the rich experiences which he had collected as a teacher. Under Buonaparte he became himself minister of Public Education, which was reorganised,

¹ The well-known sentence with which Wurtz began his history—*La chimie est une science française*—is an echo of this.

for the most part, according to his views, special regard being had to scientific studies. It is certainly owing to him that chemistry bore such wonderful fruit in France during the succeeding decades. Lastly, he took the leading part in founding the Polytechnic and Medical Schools, the *École Centrale*, and the Natural History Museum.

Fourcroy's great merit lay in his activity as an organiser and teacher. And although his experimental investigations yielded no results of great general significance, they served as preparatory work in many branches, *e.g.* in those of physiological and pathological chemistry. His conjoint researches with Vauquelin, in which the latter undoubtedly had the principal share, were of special importance with regard to organic compounds, which had been but little worked with up to that time.

The results of most of these researches were published in the *Annales de Chimie*, which was founded, at Lavoisier's instigation, by Fourcroy, Berthollet, and Guyton de Morveau. This journal, which started into life during the first year of the Revolution (1789), lived through the storms of the latter, and formed the point of union for French chemists; it was at the same time the organ of the new doctrine, as opposed to the older *Journal de Physique*, in which the last adherents of the phlogiston theory endeavoured to uphold the latter. The *Mémoires* of the French Academy appeared in 1789 for the last time; the Academy itself ceased to exist four years after that date, to be replaced in 1795 by the *Institut National*, out of which the present *Académie Française* originated in 1815, shortly after the Restoration.

After Lavoisier's death the chief representatives of chemistry in France were the three men just named, together with Vauquelin the younger. The latter had won by his eminent researches the right of being numbered among those who gave effective aid in firmly establishing the antiphlogistic system. Vauquelin, born at Hebertot in 1763, was first brought into contact with chemistry as an apothecary's apprentice; a fortunate destiny led him to Fourcroy's laboratory, in which he found employment as

assistant. He soon became Fourcroy's collaborateur, and attracted the attention of chemists in general by his brilliant work. From 1793 on he filled various honourable posts, and laboured with success in many different directions, succeeding Fourcroy as Professor of Chemistry to the Medical Faculty after the death of the latter; he died in 1829. Vauquelin did not content himself with simply teaching chemistry by lectures, but gave systematic practical instruction in his laboratory to young men desirous of it, and thus trained many chemists who afterwards rose to distinction.

Vauquelin's work, which is characterised by great carefulness and exactitude, extended itself over the most various domains of chemistry. His investigations of minerals promoted the development of mineralogical chemistry, and led him to the discovery of new bodies, *e.g.* chromium and beryllia. His splendid gifts of observation likewise showed themselves in organic chemistry, in the discovery of quinic acid, asparagine, camphoric acid, and other substances. His papers are to be found, for the most part, in the *Annales de Chimie*, of which he was one of the editors after 1791, but some of them are contained in the *Annales des Mines* and other journals. An "Introduction to Chemical Analysis," which appeared in the *Annales de Chimie* in 1799, may be mentioned here; a German translation of this led to its becoming better known and appreciated than it would otherwise have been.

Fourcroy's contemporary and Berthollet's celebrated opponent, Joseph Louis Proust, belongs—in virtue of his chief work, which helped to found the doctrine of chemical proportions—to the succeeding period, under which he will therefore be spoken of. Other French chemists, *e.g.* Pelletier, Gengembre, Bayen, Parmentier, etc., who gave in their adhesion to the doctrine of Lavoisier during the lifetime of the latter, were also active in chemical research, but they produced no work of general significance; some of the observations made by them will be treated of in the special history of the chemistry of the time.

*The State of Chemistry in Germany at the End of the
Eighteenth Century.*

German chemists proved themselves much less accessible to the antiphlogistic doctrines than Lavoisier's countrymen. The more eminent among them only began to slacken in their warfare against the new views, and to accommodate themselves to these, during the last decade of the eighteenth century. Of those who lived during that period, and who were active both as investigators and teachers, Klaproth deserves the first mention. Richter likewise participated in the working out of a most important question for general chemistry, in that he was the originator of "stöchiometry"; his investigations are to be looked upon as valuable preparatory work for the chemical atomic theory, and they will be referred to under this. None of the other German chemists of that time produced work of general significance, although they laboured with success in particular departments of the science. Some of the most noteworthy of these efforts will find their place in the special history of certain branches of chemistry; Buchholz, Trommsdorff, Wiegleb, and Westrumb may be named here as having enriched pharmaceutical and technical chemistry by valuable observations. Hermbstädt and Girtanner were among the German chemists who first frankly recognised the antiphlogistic system, and they effectively aided in propagating it in their own country by means of their writings.

Martin Heinrich Klaproth, born at Wernigerode in 1743 (*i.e.* in the same year as Lavoisier), only began to teach chemistry when somewhat advanced in life, as he continued true to his apothecary's calling till 1787; but this did not prevent him from carrying out in his earlier years investigations of the utmost value, at first, under the guidance of Valentin Rose, and later on, independently. It was to these latter researches that he owed his reception into the Berlin Academy. When the University was founded in the Prussian capital, he was elected its first Professor of

Chemistry, and in this post he continued until the beginning of 1817—the year of his death.

Klaproth was distinguished by the care and thoroughness with which he carried out all his work; the quantitative method of research was materially developed and improved by him, and he thereby helped towards the recognition of the cardinal principles advocated by Lavoisier. After Klaproth had convinced himself of the correctness of the antiphlogistic doctrine, by thoroughly testing the reactions which took place in combustion and calcination, he became one of its truest adherents; and his example led many German chemists in the same direction. Other scientists, too, who were not precisely chemists, took a part in the contest regarding these theories; thus we find Alexander von Humboldt publicly declaring for Lavoisier's doctrine in 1793.

Klaproth's researches in analytical chemistry were rightly looked upon at that time as patterns for the younger generation of chemists. Like Vauquelin's efforts, they aimed at establishing the composition of minerals by means of improved analytical methods, and thereby laying the foundation for a chemical classification of these. His observations were so exact as to result in the discovery of various elements and earths—*e.g.* uranium, titanium, and zirconia—while, at the same time, he corrected and amplified results which had been arrived at by others upon many new substances. We shall frequently have occasion to refer to Klaproth's meritorious work, especially in the history of analytical and mineralogical chemistry. His conscientiousness further showed itself in the way in which he, contrary to the custom prevalent among chemists at that day, published the results of his analysis; instead of merely stating the conclusions presumably arrived at from his experiments, he gave the actual figures of these, and so made it possible to subject them to a minute criticism or correction.

Klaproth's experimental researches were published in various journals, *e.g.* in the "Memoirs" of the Berlin Academy and in Crell's *Chemische Annalen*; he himself collected these

scattered papers together into a five-volume work, entitled *Beiträge zur chemischen Kenntniss der Mineralkörper* (1795-1810), to which a sixth volume, *Chemische Abhandlungen gemischten Inhalts*, was added in 1815. His literary activity was further shown in the publication of the *Chemisches Wörterbuch* (1807-1810), and in the revision of the works of others, e.g. B. Gren's *Handbuch der Chemie* (1806).

That chemistry in general was carefully fostered in Germany during the last two decades of the eighteenth century is also proved by the fact that various journals were started during that period, whose main object was the publication of papers on Chemistry. Among these were L. von Crell's *Chemische Annalen* (already mentioned)—the editor of which merits our praise,—which were a continuation of the *Chemisches Journal*, begun in 1778; Scherer's *Allgemeines Journal der Chemie*, which was incorporated with Crell's *Annalen* after 1803; and the *Annalen der Physik*, founded by Gren and Gilbert in 1798, and which since 1825 have appeared as *Poggendorff's Annalen der Physik und Chemie*.

*The State of Chemistry in England and Sweden towards the
End of the Eighteenth Century.*

The most distinguished chemists in England and Sweden at the time of Lavoisier's attack upon the phlogiston theory, viz. Black, Cavendish, Priestley, Scheele, and Bergman, were avowed opponents of the new doctrine. Black alone, among them, frankly recognised its truth. Cavendish, who himself contributed to the downfall of the phlogistic view by his discoveries, could not bring himself fairly to renounce it. The others, who, by their brilliant observations, had likewise forged the best weapons for its overthrow, died without being convinced of its untenability. Other English chemists, for instance Henry, Kirwan, and Hatchett, also tried to retain the phlogistic hypothesis as long as it appeared possible to say anything in its favour. Kirwan especially, who was one of those who believed phlogiston to

be identical with hydrogen, continued the fight against the new doctrine till 1792, in which year he subscribed to it himself. Its first adherent in England was Lubbock, who concurred in Lavoisier's views so early as 1784. The chemists just named, being representatives of their science at that day, merit this brief mention; they advanced particular branches of chemistry by their work, but did not influence its general tendency. Thus their countryman, John Dalton, who soon after this made such a wonderful step in advance, showed only the greater individuality in pointing out the new path, by following which chemical research has discovered and conquered new domains.

After the deaths of Bergman and Scheele, Sweden had no scientist towards the close of the eighteenth century who enriched chemistry with facts of a general significance as these two had done. Ekeberg and Gahn worked energetically at analytical and mineralogical chemistry. But it was only at the dawn of this century that Berzelius' star arose, the light from which was to illumine nearly every branch of chemistry during its first four decades. A period rich in scientific facts for chemistry thus began with him, while in his contemporaries, Davy and Gay-Lussac, the science possessed two other powerful workers. Dalton's Atomic Theory, founded as it was upon the doctrine of chemical proportions, formed the basis of their efforts.

*Development of the Doctrine of Chemical Proportions ;
Dalton's Atomic Theory.*

The idea of atoms as forming the ultimate constituents of matter often arose of old in speculative minds, without, however, an exact chemical atomic theory following from it. Boyle's corpuscular theory was and remained only a product of ingenious speculation, which ended in the assumption of a primary material, and therefore remained unfruitful. Only after a series of acquired facts had led to the presupposing of atoms, and after this assumption had enabled those facts to be satisfactorily explained, could

there be any talk of founding a chemical atomic theory. The merit of establishing this is due without a shadow of doubt to John Dalton. But before it could be brought to completion, the meaning of the term "chemical proportions," according to which simple substances united to form compound ones, had to be firmly fixed; and an important share of this task was worked out by two chemists before Dalton, viz. Richter and Proust.

Richter, whose work was to all intents and purposes unknown to Dalton at the time when he conceived his atomic theory,¹ founded the doctrine of chemical proportions without, perhaps, seeing its great importance himself, while Proust proved that the proportion in which two elements combine chemically with one another is constant, or, if there is more than one compound of these elements, the proportion alters by definite increments. If we only consider that the atomistic hypothesis, from which the chemical atomic theory sprang, originated with an observation by Dalton which followed from Proust's demonstrations, and which was comprised within the law of multiple proportions, we see how intimate was the connection between the latter and these preparatory labours.

Jeremias Benjamin Richter, born at Hirschberg in Schlesien in 1762, became a mining official (*Bergsekretär*) in Breslau, and then chemist (*Bergassessor* and *Arkanist*²) in the porcelain manufactory at Berlin, in which city he died in 1807. His researches—from which the doctrine of proportions by weight was mainly established, and which showed that acids combined with bases to form salts, together with the conclusions which he drew from them—were published by him in his *Anfangsgründen der Stöchiometrie oder Messkunst Chemischer Elemente* ("Rudiments of Stöchiometry, or the Art of Measuring Chemical Elements"), (1792-1794), and in his work entitled, *Ueber die neueren Gegenstände in*

¹ Smith, *Memoir of John Dalton and History of the Atomic Theory*, p. 214.

² *Arkanist*, meaning literally "secret chemist," was the German title in use at that time.

der Chemie ("Upon recent Discoveries in Chemistry"), which was published in eleven parts at irregular intervals between 1792 and 1802; this latter was in great part a continuation of the first-mentioned book.

Many chemists before him had busied themselves with the same task—the determination of the amounts of acid and base in salts; in addition to Kunkel, Lemery, Stahl, and Homberg, special mention must be made here of Wenzel (who was born at Dresden in 1740, and died while director of the Freiberg foundries in 1793), who placed beyond a doubt the fact that acids and bases combine in constant proportions, grounding this conclusion upon the results of numerous and, for the most part, thoroughly serviceable analyses. Richter was in a position to deduce the important "law of neutralisation" (*Neutralitätsgesetz*) from his own researches upon the quantities of bases and acids which combine to form neutral salts—researches carried out with great circumspection. Translated from his writings, obscured as these were by much phlogistic verbiage,¹ into the chemical language of to-day, this runs somewhat as follows: "When equal amounts of one and the same acid are rendered neutral by different amounts of two or more bases, the latter are equivalent to one another, and *vice versa*." It follows quite clearly from his statements that he regarded those quantities of oxides which contain equal amounts of oxygen as equivalent to one another, *i.e.* as requiring like quantities of a given acid to neutralise them. Richter had come to the right conclusion as to the capacity of iron and quicksilver to unite with oxygen in two proportions, from the composition of the corresponding salts. With these weighty observations he thus anticipated the precisely similar ones of Proust.

Notwithstanding that Richter's work contained these far-reaching discoveries, they remained almost unnoticed, their value being manifestly not recognised. This was partly due to the peculiar phlogistic language—obscure and clumsy

¹ Although he had ceased to be a phlogistonist, Richter still made frequent use of phlogistic expressions, which often obscured his writings.

—in which he clothed the results of his researches. A curious speculation in which he indulged may also have caused his whole work to be unfavourably criticised,—his assumption, namely, that a definite arithmetical relation existed between the combining weights of the bases and acids.¹ Judicious as he was in other points, he believed that he had found a proof that the combining weights of the bases and acids form approximately regular series,—the former arithmetical, and the latter geometrical. The importance which he assigned to his “law of progression,” and his continuous efforts to furnish proofs in support of it, manifestly prevented him from perceiving the significance and range of his “law of neutralisation;” indeed, he held this speculation as being the more important of the two.

The chemical world was to a certain extent made acquainted with the truths lying dormant in Richter's papers by G. E. Fischer, who put his countryman's observations into intelligible language; he collected together in a clear manner the scattered numerical values which Richter had arrived at as representing the amounts of bases and acids which combined with one another, and thus prepared the first table of equivalent weights.² Notwithstanding that the attention of chemists was in this way drawn to Richter's researches, it was a long time before they became thoroughly known and estimated at their true value. It was thus that facts proved by him were rediscovered by others much later, e.g. the combination of bases which contain equal amounts of oxygen with equal quantities of acids, by Gay-Lussac, who was without doubt unacquainted with this portion of Richter's work. As Kopp pertinently remarks in his *Entwicklung der Chemie in der neueren Zeit*, s. 152 (“Development of Chemistry in Recent Times,” p. 152): “The history of our science affords few examples of important facts, whose

¹ Even before his scientific career had begun, Richter was animated with the conviction that “chemistry was a branch of applied mathematics.”

² This table was published by Fischer in his translation of Berthollet's *Recherches sur les Lois de l'Affinité*. The fact that the latter adopted Fischer's grouping in his work, *Essai de Statique Chimique*, vol. i. p. 134, made Richter's labours known in France also.

truth had been well proved, being overlooked for so long to such an extent; and, further, when these did come to be finally appreciated, of the merit of their discovery being minimised so far as regarded the discoverer himself, and wrongly ascribed in great part to another."

It was only long after his death that Richter's services were recognised to their full extent.¹ Starting from the observation that the neutrality is not disturbed by the mutual decomposition of two neutral salts, he created the doctrine of equivalents; he was the originator of "Stoichiometry,"²—"the art of chemical measurement, which has to deal with the laws according to which substances unite to form chemical compounds."

Josèphe Louis Proust.—The work of this investigator, who, independently of Richter, also partially proved the validity of the law of chemical proportions, fell later in point of time than the most important of Richter's researches. Born at Angers in 1755, Proust went through Rouelle's course of study, and then applied his knowledge of pharmacy and chemistry at first as manager of the apothecary's shop attached to the Salpêtrière Hospital in Paris, and later as a teacher in different Spanish universities. It was in Madrid, where he settled after 1791, that he carried out his most celebrated investigations. The war deprived him both of his post and of his splendidly equipped laboratory in 1808, and it was only towards the end of his life that his necessities were relieved by a pension, while he was at the

¹ Cf. especially C. Löwig's memoir, *Jeremias Benjamin Richter, der Entdecker der chemischen Proportionen* (Breslau, 1874) ["Jeremias Benjamin Richter, the Discoverer of Chemical Proportions" (Breslau, 1874)]. According to Fischer, Richter's work was particularly emphasised by Gehlen, Schweigger, and Berzelius. The discovery of the law of neutralisation was ascribed by Berzelius to Wenzel, in consequence of a misunderstanding on the part of the former; and it was left to H. Hess of St. Petersburg to point out this error, thirty-three years after Richter's death.

² Richter himself says that he was unable to devise a better name for this than the word "*Stoichiometrie*, from *στοιχεῖον*, signifying something which cannot be further divided, and *μετρεῖν*, which denotes the finding out of relative proportions."

same time received into the Paris Academy; he died at his native town of Angers in 1826.

His most significant work was induced by a series of questions which Berthollet had propounded. At the end of last century (*i.e.* from 1798 on), the latter's *Recherches sur les Lois de l'Affinité*, which he collected together in 1803 in his *Essai d'une Statique Chimique*, created an extraordinary sensation. Grounding his objections upon speculations apparently well founded, this gifted writer disputed the fact that constant proportion was the rule with regard to the constituents of chemical compounds. His ideas upon chemical affinity, by which the combination of substances with one another is regulated, will be treated of in detail in the special history of this part of our science. Suffice it to say here that, starting from the axiom that chemical processes are dependent upon the relative masses of the reacting bodies, he arrived at the conclusion that, in a chemical compound which results from the union of two substances, so much the more of the one substance must enter into it, the more there is of it available, always supposing that no exceptional circumstances stand in the way of this mass-action. Berthollet's great reputation may have been the reason why none of the other leading chemists of the day raised any objections, although they certainly did not concur in this doctrine. For, with respect to many compounds, salts especially, the constancy of the combining proportions of their constituents was a fact beyond all doubt to men like Richter, Klaproth, Vauquelin, and others.

Proust took up the cudgels against Berthollet, and, by means of exact experiment, overthrew one by one the theoretical conclusions of his opponent. This memorable controversy, which, beginning in 1799, was continued for eight years, and which was conducted on both sides with consummate ingenuity, ended in the conclusive proof of constant combining proportions.

To what extent Dalton was influenced by Proust's labours in his researches in a similar direction, it is hard to say; but

they were certainly not without some effect upon him, the dispute between Berthollet and Proust being followed with the keenest interest in scientific circles.

So early as the year 1799 Proust had proved the constant composition both of natural and of artificial carbonate of copper,¹ and had called special attention to the unvarying proportions by weight in true chemical compounds, as opposed to the varying ones in mixtures. Still more important than these were observations—to be supplemented later on by himself and others—upon the two stages of oxidation which tin shows,² and upon the two compounds which iron forms with sulphur;³ for he particularly emphasised the point that not only were the proportions between the metals and oxygen or sulphur constant in the individual compounds, but also that the combining proportions increase by leaps, and not gradually, when two elements unite to form more than one compound. Berthollet thought that he had proved exactly the opposite in his researches on the formation of oxides and salts⁴ (e.g. the nitrates of mercury), viz. that metals can form oxides with gradually increasing amounts of oxygen. But Proust⁵ showed that his experiments were wrong, and that he had deduced his conclusions from the analysis of mixtures and not of definite compounds. The superiority of Proust in experimental points was clearly manifested, since he proved to Berthollet that many of the substances which the latter regarded as oxides contained chemically combined water; it was Proust who classed the hydrates among chemical compounds. In fact, he succeeded by generalisation and by firmly establishing his conception—that combination between the other elements and oxygen or sulphur only takes place in one or, at most, in a few proportions—in completely routing the weak arguments of his opponent, many of which were advanced without any experimental proof to back them up.⁶

¹ *Ann. de Chimie*, vol. xxxii. p. 30.

² *Journ. de Phys.*, vol. li. p. 174. ³ *Ibid.*, vol. liv. p. 89.

⁴ Cf. *Essai de Statique Chimique*, vol. ii. p. 399 et seq.

⁵ *Journ. de Phys.*, vol. lix. pp. 260, 321.

⁶ *Ibid.*, vol. lxiii. pp. 364, 438.

Proust had repeatedly laid stress upon the validity of combining proportions without endeavouring to get clearly at the reasons for this. How near he was to recognising the law of multiple proportions which Dalton deduced from his researches—researches similar to Proust's, and scarcely excelling these in exactitude! One is led to the surmise that if Proust had calculated the results of his experiments on the composition of binary compounds otherwise than he did, he would have discovered that law. The happy idea occurred to Dalton of reckoning the amounts of one element, which combined in different proportions with another, in terms of a given chosen quantity of the latter; the result of this was that the multiple proportions became manifest, and these he explained by the aid of the atomic hypothesis.

DALTON'S ATOMIC THEORY.

John Dalton, the eldest son of a poor weaver, was born at Eaglesfield in Cumberland in 1766, and had to make his own living at an early age as an elementary teacher. Endowed with a strong bent towards mathematics and physics, he acquired a sound knowledge of these subjects, and was thus enabled to carry out independent investigations in them, and to take the post of mathematical and physical master in a college at Manchester in 1793. He soon included chemistry also in his studies, the most important problem of which he was destined to solve. In his modesty Dalton had no thought of acquiring for himself a brilliant position in life, the highest reward for his truly philosophic mind consisting in the elucidation of the truth. He died at Manchester in 1844.

Dalton's earlier researches on the physical behaviour of gases (their expansion by heat and absorption by liquids) were of great influence upon his later chemical labours. For it was through them that he acquired the experimental dexterity which stood him in such good stead when analysing those gases, whose composition led him to the law of multiple proportions.

The discovery of the latter, and the conception of the atomic theory which arose from it, date from about 1802-1803. After that time Dalton applied himself to the task of building up a firm foundation for these by amplifying his observations; he only published his discovery in 1808, when the first volume of his *New System of Chemical Philosophy* appeared. The outlines of the atomic theory had, with Dalton's concurrence, been made public by Thomson in his *System of Chemistry* a year before this, so that the first influence of this great scientific event upon the chemical world is to be dated from then. The second volume of Dalton's above-mentioned work, with material improvements, appeared in 1810, and the third volume so late as 1827, by which time its contents were mostly out of date.¹

The first of Dalton's observations which formed the starting-point for the setting up of the atomic theory consisted in the determination of the composition of oil-forming gas (ethylene), and light carburetted hydrogen (methane). From his analyses of these two gases he concluded that, for the same quantity of carbon, twice as much hydrogen was contained in the latter as in the former, *i.e.* that the proportions of hydrogen were as 2 : 1. This regularity induced him to investigate other compounds in the same direction; thus, in the case of carbonic oxide and carbonic acid, he found that, for the same amount of carbon, the proportions of oxygen present in these were again respectively as 1 : 2. His conviction that there must be a law underlying these so simple relations hardly required any further strengthening after he had met with like simple numerical proportions in the results of his analysis of nitrous oxide, nitric oxide, nitrous acid, and nitric acid (*i.e.* the anhydrides of the two last).² He had, therefore, proved that when different quantities of one element combined chemically with one and the same quantity of another, these amounts stood in a

¹ A German translation of the first two volumes by Fr. Wolf appeared in 1810.

² Dalton was, however, wrong in his analysis of nitric acid, which he made out to consist of nitrogen and oxygen in the proportions of 1 atom to 2.

simple relation to one another—a relation which could be expressed by whole numbers. The law of multiple proportions was thus discovered; it had, indeed, been deduced from experiments which were of necessity not very exact, in accordance with the state of chemical analysis at that time.

Dalton, however, did not remain content with this important result, but sought an explanation of the numerical relations which he had discovered. This was afforded him by the atomistic hypothesis, in the assumption, not new in itself, that substances consist of ultimate particles not further divisible—of atoms. This hypothesis gave a satisfactory explanation of the facts comprised within the law of multiple proportions, for one now only required to substitute absolute numbers for the relative ones, *i.e.* to assume that in carbonic oxide (for instance) one atom of carbon was combined with one of oxygen, and in carbonic acid one atom of carbon with two of oxygen, and so on. Upon the firm basis of this assumption Dalton erected his Atomic Theory, the essence of which is given in the two succeeding paragraphs:—

(1) *Every element is made up of homogeneous atoms whose weight is constant.*

(2) *Chemical compounds are formed by the union of the atoms of different elements in the simplest numerical proportions.*

His speculations upon the atoms themselves, which Dalton assumed for the sake of simplicity to be spherical in shape, and also the hypothesis that they do not come into direct contact with one another, but are separated by a heat zone, have but a merely subordinate significance as compared with the above two sentences; they exercised no influence on the development of the chemical atomic theory.

Dalton now sought to deduce the relative atomic weights from the proportions by weight in which the elements unite to form compounds, advancing to this task, which constituted the main feature of his *New System*, with a wonderful confidence. Since he had no certain means of arriving at these

numeric proportions of the combining atoms, assumptions had to be made, and these were of the simplest kind. The following statements by Dalton refer solely to compounds of two elements.

When only one compound of two elements A and B is known, we must assume that it is made up of one atom of the one and one atom of the other: $A + B$ (binary compound, or atom of the second order. Dalton spoke of an elementary atom as an atom of the first order).

If two compounds of two elements A and C are known, their composition is expressed by the symbols $A + C$ and $A + 2C$ (ternary compound, or atom of the third order).

When the composition of three compounds of two elements A and D had to be decided, then, according to Dalton, the following combinations were the probable ones: $A + D$, $A + 2D$, and $2A + D$. Atoms of the fourth order (*e.g.* $A + 3E$), etc., were also allowed by Dalton, although he favoured the more simple proportions. Compounds whose atomic numbers were as 2 : 3 or 2 : 5, he explained as resulting from two atoms of a higher order than the elementary atom (*e.g.* nitrous acid from one atom of nitric oxide and one of nitric acid).¹

Dalton's statement that the atomic weight of a compound is equal to the sum of the atomic weights of its constituent elements appears to us nowadays self-evident; but we

¹ Dalton's precise words, as given in his *New System*, vol. i., second edition, p. 213, are as follows:—

"If there are two bodies, A and B, which are disposed to combine, the following is the order in which the combinations may take place, beginning with the most simple, namely:

"1 atom of A + 1 atom of B = 1 atom of C, binary,

"1 atom of A + 2 atoms of B = 1 atom of D, ternary," etc.

Again, at p. 214:—

"1st, When only one combination of two bodies can be obtained, it must be presumed to be a *binary* one, unless some cause appear to the contrary.

"2d, When two combinations are observed, they must be presumed to be a *binary* and a *ternary*.

"3d, When three combinations are obtained, we may expect one to be a *binary* and the other two *ternary*.

"4th, When four combinations are observed, we should expect one *binary*, two *ternary*, and one *quaternary*, etc."

must not forget that at that period, as well as during Lavoisier's lifetime, the false idea of heat being material had by no means been discarded by all chemists, many still believing that a loss of matter occurred when heat was evolved from the combination of two elements.

Setting out with the above premises, Dalton endeavoured to determine the relative atomic weights of the elements as follows:—Starting with water, as the only compound of hydrogen and oxygen (peroxide of hydrogen being at that time unknown), he estimated the proportions in which both of these were present, and then took hydrogen as the unit to which oxygen and other elements were to be referred. The relative values of the latter, as deduced from the composition of their oxygen and hydrogen compounds, were, according to his conception, their atomic weights. In this way he determined the relative atomic weight of nitrogen from the composition of ammonia, which, as the only compound of hydrogen and nitrogen, consisted of one atom of each of those elements: and that of carbon from the analyses of carbonic oxide and carbonic acid, using in this case the value he had obtained for oxygen from the analysis of water.

As the analytical methods which he employed were liable to many sources of error, it was impossible that his results could be accurate; but the great merit belongs to Dalton of having propounded the principle of the determination of the relative atomic weights, or, to speak more correctly, of the combining weights of the elements. How far his first "atomic-weight numbers," as published by Thomson in 1805, differ from the values current to-day, is seen from the following table:—

"Relative Atomic Weights."				According to Dalton.	Their current Values.
Hydrogen.	.	.	.	1	1
Oxygen	.	.	.	6.5	7.98
Nitrogen	.	.	.	5	4.66
Carbon	.	.	.	5.4	6

Dalton published a greatly extended and, in part, improved table of "relative atomic weights" in the first volume of his work (1808), in which 7 is the value given for oxygen: the numbers which he obtained are too low throughout, and deviate from the true values by several units in the case of the elements of higher atomic weight.¹ His attempt to apply the atomic hypothesis to organic compounds must also be mentioned here, although it turned out unsuccessful, the results of his analyses of the latter being far from exact.

Nor must we forget Dalton's efforts to build up a system of notation which should illustrate atomic composition. The atoms of the elements were represented by various circular symbols, *e.g.* oxygen by an empty circle \bigcirc , hydrogen by \odot , nitrogen by \oplus , and sulphur by \oplus . These signs, placed conveniently near to each other, indicated the supposed constitution of chemical compounds: for water the symbol $\odot\bigcirc$ was used, for ammonia $\odot\oplus$.

for sulphuric acid ² $\begin{array}{c} \bigcirc \\ \bigcirc\oplus\bigcirc \end{array}$, and so on.

But the simpler and easily decipherable system of notation which Berzelius introduced, some time after this, prevented Dalton's from ever coming into general use.

Further Development of the Atomic Theory.

The reception which Dalton's atomic doctrine found among chemists was almost wholly favourable, although there were not wanting a few to depreciate the new theory, and even to ascribe the merit of originating it to others. In Great Britain it found from the beginning an enthusiastic

¹ This table of atomic weights shows his endeavours at rounding the numerical values, because of his perception of the insufficiency of the methods employed, as is seen in the following instances; the figures appended below in brackets, after those of Dalton, give the correct combining weights: sulphur 13 (16), iron 38 (56), zinc 56 (64.9), copper 56 (63.3), silver 100 (108), mercury 167 (200).

² Dalton did not know the compound SO_3 , but supposed that this formula gave the composition of sulphuric acid.

adherent in Thomas Thomson,¹ who, however, rather did it harm than good by his excess of zeal, a fatal tendency to speculation sometimes causing him to quit the sure ground of exact experiment. It was of particular importance, at the time a theory so far-reaching was set up, that the facts on which it rested (still few in number) should be amplified and deepened by reliable observations.

The estimations made by Thomson of the relative atomic weights of elements and compounds were still more defective than Dalton's, and became influenced later in an inexcusable manner by Prout's erroneous hypothesis,—and that, too, after Berzelius had begun his long series of classical labours with the accurate determination of atomic weights. On the other hand, Thomson's investigations of the potash salts of oxalic acid helped to confirm the atomic doctrine, since they showed that the quantities of potash which reacted with a given amount of oxalic acid were to each other as 1 : 2 : 4 by weight. An analogous observation was made by Wollaston,² who found that in the neutral and acid carbonates of potash the proportions of carbonic acid relatively to the same weight of potash were as 1 : 2. The applicability of the law of multiple proportions was thus also proved for some salts.

The position which from that time (about 1808) the most distinguished investigators of the day—Davy, Berzelius, and Gay-Lussac—took up with regard to Dalton's atomic

¹ Thomas Thomson (born 1773, died 1852) exercised no slight influence on the growth of theoretical chemical views, especially in England, both by his experimental researches in chemistry, and by his text-books. That it was he who first gave to the public the principles of Dalton's atomic theory has been mentioned already. As a historian of chemistry he was also active, his *History of Chemistry* appearing in 1830-31. Most of his papers were published in the *Annals of Philosophy*, which he himself edited. As a teacher in the University of Glasgow he was eminently successful, founding there the first chemical laboratory for general instruction in Great Britain.

² W. H. Wollaston was born in 1766 (the same year as Dalton), and died in 1828. Originally a physician, he soon gave himself up to the study of physics and chemistry, enriching the former especially by important observations. At the same time he became favourably known by his chemical researches, in particular by his work on the platinum metals. Most of his papers are to be found in the *Philosophical Transactions*, but a few of them in the *Annals of Philosophy*.

theory, renders an account of their most important work and general services appropriate at this point. The researches of Gay-Lussac upon gases, and even more the unrelenting endeavours of Berzelius to work out sure foundations for the determination of the true atomic weights, had the deepest influence on the development of the atomic doctrine, which is now the basis of chemistry.

Davy and Gay-Lussac: their life and work.—Davy was at first sceptical with regard to Dalton's rights as the originator of the atomic theory, and indeed, in 1809, he claimed for Higgins the priority for this doctrine, the latter having made use of the atomic hypothesis to explain chemical facts so early as 1789 (in his work, *A Comparative View of the Phlogistic and Antiphlogistic Theories*). Higgins certainly expressed opinions which, on a superficial glance, appeared similar to those of Dalton, stating as he did that the smallest particles combine in simple numerical proportions to form chemical compounds. But these views were brought forward without any internal organic connection, and, moreover, they were not founded upon experiment. It became clear to Davy later on that Higgins had no claim to be regarded as the originator of the atomic theory, and he then frankly recognised Dalton's service.

Humphry Davy, born at Penzance in Cornwall in 1778, was destined for a distinguished career, to be cut short by an early death, his creative genius being impaired during the last years of his life by a prolonged illness. So early as 1813, when only thirty-five years of age, he had to leave off work and seek renewed health on the continent, in Italy for the most part. After 1820 he lived and worked again in England, but left it in 1827, never to return, for he died in 1829 at Geneva on his homeward journey.

While only a surgeon's assistant, Davy acquired, by his own energy, such a wide knowledge of chemistry and the natural sciences, that at twenty years of age he was able to take the post of chemist in the newly-founded Pneumatic Institution at Bristol. The aim which this institution had

set before itself was to test the various artificially prepared gases for their physiological and medical action. It was here that Dalton carried out his researches on nitrous oxide, whose intoxicating and stupefying action he discovered, and on the effect of other gases (admixed with nitrogen) on the organism, *e.g.* hydrogen and carbonic acid; in this way he laid the foundation of his fame as a splendid experimenter. Already in 1801 we find him professor at the Royal Institution of London, and shortly afterwards a member of the Royal Society, whose president he became in 1820.

His most memorable work, which effected a complete transformation in many branches of chemistry, was accomplished during the first thirteen years of this century. We need only mention here the isolation of the metals of the alkalis and alkaline earths by the galvanic current, through which a whole series of hitherto undecomposed substances were recognised as compound. An almost still more important result of these observations was the discovery of the elementary nature of chlorine, which up till then was held to be a compound; this opened out entirely new stand-points, which led to a transformation of the views upon the constitution of acids. When it became proved that there were acids which did not contain oxygen, a material alteration in Lavoisier's theory became for the first time necessary. Discoveries of such range as this characterise the period in which Davy developed his wonderful activity. His most important experimental researches will be described partly in the further course of the general history of this period, and partly in the synopsis of the progress of particular branches of chemistry.

Davy contributed greatly by his popular lectures, especially by those given for the Board of Agriculture, to heighten the public interest in chemistry during the first decade of this century. He it was, too, who showed in what high degree chemistry could and should meet the requirements of technical industries and those of daily life; we have only to think in this connection of the miner's safety-lamp which he constructed.

Davy's genius in grasping chemical relations was especially shown in his efforts to discover the connection between electricity and chemical affinity, both of which he regarded as resulting from a common cause. He was the first to set up an electro-chemical theory grounded upon experiments, which were devised and carried out in a masterly manner, and in this way he opened out a province in which Berzelius was to work with such effect in the decade following.¹

Wherever Davy, with his aptitude for experiment and acuteness of mind, treated chemical problems, he achieved great results. Within the narrower limits of special research also, *e.g.* in his investigations on ammonium amalgam, phosgene, euchlorine, iodine, solid phosphuretted hydrogen, and the phenomena of combustion, the fruits of his labours were at once perceptible; his work always left a deep mark. After the year 1801 Davy published his most important papers in the *Philosophical Transactions*, but some are to be found in the *Annales de Chimie* and in the *Journal de Physique*. Of his few larger works,² the *Elements of Chemical Philosophy* (1810-12) became best known, especially as it was soon translated into German and French. After his death all his works were collected together and published by his brother John Davy.

In addition to the scientific interest which Davy's wonderful services call forth, there is to be added the purely human interest in his personality. The nobility and poetry of his nature are shown both in the journals which he kept during his extensive travels in France, Germany, and Italy.

¹ Davy's electro-chemical theory of affinity will be gone into along with that of Berzelius in one of the succeeding paragraphs.

² The judgment which Berzelius passed upon Davy's literary activity, in a letter written to Wöhler in 1831, is of much interest (cf. *Ber.*, vol. xv. p. 3166). The latter had been deploring that he was overwhelmed with literary work, whereupon Berzelius replied as follows: "Had Davy been forced to occupy himself as much with writing as you have to do at present, I am convinced that he would have advanced chemistry by a hundred years; but he remained only a 'brilliant fragment' (*glänzendes Bruchstück*), just because he was not compelled from the beginning to work himself thoroughly into all parts of the science as into one organic whole."

and in his beautiful relations to Faraday. The inventions made by him for the public good raise still higher our regard for this remarkable investigator.

Davy's historical-critical attitude towards Dalton's atomic doctrine has been already spoken of. But although he gave the latter credit for originating this theory, he continued sceptical with regard to Dalton's conclusions.¹ He would not admit that Dalton's *atomic weights* were really such : in his view these were merely the *proportion numbers* of the elements, for the determination of whose atomic weights there was no sure basis to go upon. Wollaston had before this given vent to a similar circumspect criticism of Dalton's bold speculations, having published in 1808 his opinion that the numbers arrived at by Dalton gave, not the atomic weights, but the *chemical equivalents* of the elements. Gay-Lussac, too, whose labours began at that time to exercise such a powerful influence on the development of chemistry, rejected the assumption of atomic weights, and merely allowed that the ratio (*rapport*) of one element (*e.g.* hydrogen, nitrogen, or iodine) to another (*e.g.* oxygen) was established by analytical and synthetical determinations.

Gay-Lussac, whose critical attitude to Dalton's atomic theory has just been touched upon, promoted the latter in a quite exceptional degree by his wide-reaching discovery of the so-called "Law of volumes"—more, indeed, than he was willing to confess.

Joseph Louis Gay-Lussac, born in 1778 at St. Léonard (in the old province of Limousin), after acting as Fourcroy's demonstrator became in 1808 professor of physics at the *Sorbonne*, and also in 1809 professor of chemistry at the *École Polytechnique*, at which he had been a pupil up to the year 1800. In 1832 he resigned his chair at the *Sorbonne* to fill that of general chemistry at the *Jardin des Plantes*; he died in 1850. While still very young, after his initiation into science by Berthollet, Gay-Lussac aroused the marked attention of his contemporaries by his physical investigations on the behaviour of gases—investigations which touched more or

¹ Cf. particularly his *Elements of Chemical Philosophy*.

less on the province of chemistry. Brief mention may also be made here of his bold balloon ascents in 1804-1805, undertaken at first along with Biot and afterwards alone. His researches after 1805, upon the laws deducible from the combining volumes of gases which unite chemically with one another, had most incisive results. How this yielded the richest fruits for chemistry as a whole, and not merely for the chemistry of gases, will be shown later on. Gay-Lussac's name is further associated with the discovery of the definite relation which exists between the volume of a gas and its temperature; it was only after this law, which supplemented that of Boyle and Mariotte, had been worked out, that reliable measurements of gases could be made.

In his work which bore upon special branches of chemistry Gay-Lussac likewise proved himself a masterly investigator; to exactitude in observing, and acuteness in explaining his observations, he added a wonderful lucidity in expounding his researches and the conclusions at which he arrived. His work on iodine and cyanogen and on their compounds would alone suffice to ensure him a place among the most distinguished chemists. How stimulating and full of matter were his papers! The one upon cyanogen, especially, was the basis on which the radical theory was afterwards developed, cyanogen having been characterised by Gay-Lussac as the first compound radical. Even his minor work bears the classical stamp: of it we may mention here his researches on the compounds of sulphur, and on the various stages of oxidation of nitrogen, and his conjoint work with Thénard¹ upon the alkali metals. Together with Liebig he

¹ L. J. Thénard, born in 1777, a pupil of Vauquelin and Berthollet, became professor at the *École Polytechnique* and in the *Collège de France*, and worked energetically for the promotion of the study of natural sciences in his country. His name is indissolubly united with that of Gay-Lussac, their conjoint work leading to a knowledge of many chemical processes, and contributing to the improvement of important methods. Thénard's *Traité de Chimie Élémentaire*, a text-book which was most widely used, thanks to the happy synoptical arrangement of its contents, was of great merit; the first French edition of it was published in 1813-16, and the first German edition (translated from the fifth French by Fechner) in 1825-33. Thénard died in 1857.

investigated the fulminates. Hidden in many of these pieces of work there lay germs which were to expand into important discoveries; for example, his observation on the action of chlorine upon wax laid the foundation for subsequent researches upon substitution reactions.

By his work on technical subjects, Gay-Lussac proved that he understood how to bring his results in analytical chemistry to bear upon these. He is to be regarded as the originator of volumetric analysis, and the improved analytical methods which he thus introduced, and which have since come into general use, have helped materially to advance chemical industries. We shall meet with his work in almost every important branch of chemical investigation,—in analytical, technical, physical, and pure chemistry.

Gay-Lussac published most of his experimental results in the *Annales de Chimie*,¹ but a few of them are to be found in the *Mémoires de la Société d'Arçueil* and in the *Comptes Rendus*. Of his papers which appeared separately, mention may be made here of a number upon methods of investigating and testing commercial products, silver ores, etc., which, as a member of various commissions, he worked out; also of the *Recherches Physiques et Chimiques* (1811), which he published conjointly with Thénard.

Prout's Hypothesis and its Effects.

During the period in which Davy and Gay-Lussac were carrying on their brilliant work, and before the star of Berzelius had attained to its full lustre, a literary-chemical event occurred which made a profound impression upon the chemists of that day, viz. the setting up of Prout's hypothesis. This was one of those factors which materially depreciated the atomic doctrine in the eyes of many eminent investigators. On account of its influence upon the further development of the atomic theory, this hypothesis must be discussed here, although it has happened but seldom

¹ After the year 1816 this journal was given out by Arago and himself under the title *Annales de Chimie et de Physique*.

that an idea from which weighty theoretical conceptions sprang originated in such a faulty manner as it did.

In the year 1815 a paper¹ appeared in which the relation between the atomic weights of elements and the specific gravity of their vapours was treated of; in this paper, and still more positively in a second,² published in the following year, the tenet was set up by their anonymous author that the atomic weights of the elements—taking that of hydrogen as unity—were expressible by whole numbers, *i.e.* that they were multiples of the atomic weight of the lightest element.³ From this there followed the hypothesis proper of Prout (who had, in the meantime, become known as the author of the two above papers),—that hydrogen may be regarded as the primary matter from which all other elements are formed by various condensations.

This idea, so lightly thrown out, and which adapted itself so usefully to the incomplete investigations of others,⁴ possessed both then and at various later periods a great charm for many distinguished chemists. Already, before these papers had been published, Dalton's friend Thomson had alluded to the fact that, according to his own experiments and those of others, the atomic weights of several of the elements were multiples of those of oxygen. He endeavoured, indeed, to establish the same point several years after this, without considering that the numbers which Berzelius had found, in the meantime, differed widely from his own, which had therefore become of very doubtful value. Thomson was the victim of this preconceived opinion; he went so far as to see in Prout's assumption a fundamental law of chemistry.

¹ *Annals of Philosophy*, vol. vi. p. 321.

² *Ibid.*, vol. vii. p. 111.

³ The author altered the numerical values of the atomic weights in a highly arbitrary manner, so that they should not merely be whole numbers, but should also show regular differences among each other, as is seen from the following examples:—

Calcium 20	Iron 28	Chlorine 36.
Sodium 24	Zinc 32	Potassium 40.

⁴ Prout himself was a physician, and his own investigations were few in number and anything but conclusive.

Although Berzelius and, later, Turner and others proved the untenability of Prout's hypothesis, many chemists nevertheless inclined towards the latter. In his text-book of 1827 L. Gmelin gave the "mixture weights" (*Mischungsgewichte*) as far as possible in whole numbers, which he was not justified in doing after Berzelius' classical researches. Later still, about the year 1840, Dumas and Stas, who had determined the atomic weights of carbon, oxygen, chlorine, and calcium with great exactness, and also Erdmann and Marchand in their numerous investigations in a like direction, betrayed a strong inclination to this hypothesis, the weakness of which was afterwards proved by Stas himself and by Marignac. The predilection shown by many chemists for such a conception, which led to such far-reaching deductions, helped to discredit the whole atomic doctrine in the minds of thoughtful investigators.

Like Davy and Gay-Lussac, who, it is true, did not specially occupy themselves with the problem of determining the atomic weights of the elements, Berzelius kept himself quite free from those prepossessions; and, since he devoted all his energies to the solution of questions allied to this, his opinions had therefore the very greatest value. Firm, and not led away by the alluring simplicity of Prout's hypothesis, he held fast to his aim,—the accurate, purely experimental determination of the atomic weights, and he firmly established by his masterly work the then unsteady edifice of the atomic doctrine.

BERZELIUS—A SURVEY OF HIS WORK.

The life of this investigator, who developed and enriched chemistry in its most important branches as no other man has done, was the quiet and uneventful one of a student. He was guided in his work by the great and comprehensive aims, — to investigate carefully the composition of chemical compounds, and to get at the laws according to which they are formed.

Jöns Jakob Berzelius was born at Westerlöv in Östergötland, Sweden, where his father was a school-

master, on the 29th of August 1779. A love for chemistry appears to have developed itself very early in him, but his desire to devote himself to its study at Upsala was only attained under many difficulties and disappointments (in 1798). The lectures and instruction given by his teachers Afzelius and Ekeberg were uninspired by the spirit after which Berzelius strove. We therefore find him turning to the study of medicine, without, however, losing sight of chemistry as a material aid to the latter. His early work, especially that which he carried out along with Hisinger upon the action of the galvanic current upon salts, made him known in his own country, so that in 1802 he was appointed assistant professor in medicine, botany, and pharmacy at Stockholm, and, five years later, professor of medicine and pharmacy. In 1815 he was called to the chair of chemistry in the newly-founded Chirurgico-Medical Institute there. His lectures, which were at first purely theoretical, according to the established custom, he began to enliven by judiciously chosen experiments. A very imperfectly equipped laboratory enabled him to carry through the most exact experiments for the more firm establishing of the doctrine of chemical proportions. In those modest rooms were accomplished numerous researches, most of them by himself alone, but some in conjunction with specially gifted pupils. The names of those latter are sufficient in themselves to show the wonderful results which he achieved by his teaching; among them we may mention here Heinrich and Gustav Rose, Mitscherlich, Wöhler, Chr. Gmelin, Magnus, and Mosander.

From the year 1818, when he was nominated permanent secretary of the Stockholm Academy of Sciences, of which he had been a member since 1808, and still more after 1832, when Mosander succeeded him in his chair, Berzelius devoted himself to literary work with an effectiveness which has hardly been equalled by any chemist before or after him. His energetic life came to a close on the 7th of August 1848. In 1818 he was ennobled by King Charles XIV., and in 1835 made a baron by the same monarch.

To give a short and at the same time succinct account of the scientific achievements of Berzelius is no easy task, for these not only touched upon the main points of chemistry, but penetrated deeply into them, and gave rise to weighty reforms. After occupying himself for the first seven years of his independent scientific work with researches in various branches of chemistry, especially physiological chemistry, and proving himself thereby to be an exceptional observer, his efforts rose—from 1807—to a higher level. For, from that date, his entire energy was devoted to one great aim; the minute investigation of chemical proportions and, with that, the development of the atomic doctrine he looked upon as his life-task. At the time when he began his work upon the combining proportions of the elements, the atomic doctrine was unknown to him. His first researches were inspired by J. B. Richter's papers and then by Davy's discoveries, before he was aware of the results of Dalton's labours, which had led to the setting up of the atomic theory. How Berzelius built up the doctrine of proportions by his own energy, by improving analytical methods, and by the clearsighted interpretation of his own researches and those of others, and how he created solid foundations for the determination of atomic weights, will be described in the following section.

But we must just mention here that he greatly enriched analytical chemistry by the discovery of new methods. These were, indeed, indispensable to him for the attainment of his great aim, for it was only by means of the most accurate possible analyses that the constancy of combining proportions could be definitely proved. This was, however, by no means the only branch of chemistry which was indebted to him, for analysis in his hands was made to open out other and larger domains. His first attempt to work out the composition of minerals on the basis of the atomic theory, *i.e.* with the aid of the law of multiple proportions, was made so early as the year 1812, and his setting up of a chemical mineral system created a wonderful excitement.

Of still wider value were his successful endeavours to

show that organic substances were likewise subject to the law of multiple proportions. After materially improving the methods of analysis of organic bodies, he was able to prove in 1814 that simple atomic relations prevail among the constituents of organic acids. The atomic theory thus became the guiding star both for Berzelius and for the whole science.

Berzelius assumed that atoms were electrically polarised, and looked upon this as the cause of the combination of elements in definite proportions. His electro-chemical theory, developed from this assumption, and his dualistic system, which was the immediate result of this theory, will be treated of in detail along with other similar attempts at explaining the phenomena of affinity.

Experiment formed the basis of his speculations. By connected observations on the chemical behaviour of simple and compound bodies, he enriched the most important branches of his science in a marvellous degree.

Of his numerous researches on inorganic substances, that upon selenium is a classical model, worthy to rank alongside of Gay-Lussac's upon iodine. We may also call to mind here his remarkable investigations upon ferro-cyanogen compounds, sulpho-salts, and fluorine compounds, among many others. All his experimental work shows the originality of a master mind; and although his inventive genius was not so great as that of Davy, his strict methods of procedure and conscientious observations led him to discoveries of the first importance.

The work of Berzelius in organic chemistry is less imposing than that which has just been sketched, but we have only to recall his discovery of racemic acid, and his important investigations on its isomerism with tartaric acid, to see that here also he made a deep mark. As he was the first to apply the principles of the atomic theory to organic substances, so he sought to introduce his electro-chemical and dualistic views here also. These efforts of his to simplify complicated relations were not, in this instance, permanently successful, for, although his radical theory had a fruitful

influence for a time, it could not hold its ground against the unitary conception. Much of his work in mineralogical and physiological chemistry was fundamental in its nature, and was even that of a pioneer, since it had as its immediate result (especially in mineralogical chemistry) the setting up of entirely new points of view and new aims.

The grand creative genius of Berzelius and the joy he had in his work are not only apparent in his experimental researches, but show themselves also in his activity as a teacher, whether as manifested in personal intercourse with his pupils or as finding expression in writing. In his little laboratory there assembled young men from far and near, most of them already well versed in chemical knowledge, to learn from his experiences and then to further propagate his doctrines. From Germany, in especial, where at that time there was no provision for practical chemical work, came aspiring students, who subsequently advocated the principles of his school and extended its influence.

Berzelius' literary activity is most strikingly shown in his *Lehrbuch der Chemie*,¹ of which five editions, each of them completely revised, appeared. Along with the absolute thoroughness which we also admire in his experimental work, clearness of description is united in this book with precision of expression. He did not merely confine himself to the simple exposition of known facts, but criticised the experiments from which these were deduced with perfect impartiality. His text-book remained a pattern for others during the succeeding decades. The many-sidedness of Berzelius and his power of work are further shown in the *Jahresbericht über die Fortschritte in der Physik und Chemie* ("Annual Report on the Progress of Physics and Chemistry"), twenty-seven volumes in all, which were published by him

¹ This book came out for the first time in 1808-1818 in three volumes (Swedish); the second Swedish edition (four vols., 1825-31) was translated into German by Wohler, while the subsequent editions were printed only in German. The third (four vols., 1833-35) and the fourth (four vols., 1835-41) were done into German by Wohler "from the Swedish MSS. of the author," while the fifth "original edition" (five vols., 1843-48) was written by Berzelius himself with Wöhler's co-operation.

in Swedish from the year 1810 until his death; these were also brought out in German by Gmelin and Wöhler (in Tübingen). He had undertaken to report to the Stockholm Academy upon the work published on those subjects, a task which he performed with diligence and perspicacity. With regard to work which came at all within his own province, he knew to perfection how to fill the rôle of critic, although on some occasions he was led by the characteristics of particular experimental researches to express a judgment which betrays a certain prepossession. Notwithstanding this, however, his *Jahresberichte* are and will remain indispensable sources of information for those who wish to understand the currents and changes of opinion in the chemistry of his day.

The experimental researches of Berzelius were as a rule first published in Swedish in the *Transactions of the Stockholm Academy*, but most of them were afterwards given out in German, and a few in English and French (in *Gilbert's*, *Poggendorff's*, and *Liebig's Annalen*, the *Annales de Chimie*, *Annals of Philosophy*, etc.). They are characterised by the same excellences as his text-book.

The above sketch of his main achievements is sufficient to indicate the eminent qualities which distinguished Berzelius as a classical investigator. Thoroughness and perseverance in everything which he undertook; exactness in all his observations, and the capacity for arranging these distinctly and explaining them clearly; inviolable adherence to the results of experience (which was his guide before everything else), and an equally firm adherence to results which, in his opinion, had been correctly arrived at from a number of data; these were the characteristics which distinguished this great man.

The desire to retain whatever of good the science possessed was developed in him in an exceptional degree: indeed, in sustaining this conservative attitude he went so far as to see a danger to the steady development of chemistry in every innovation which called in question views already proved and found useful. Hence his fervent opposition to many new hypotheses which he had in the end to recognise

as correct. His great services in furthering chemistry were, however, not lessened by this peculiarity, which had its real cause in a profound sense of justice; on the contrary, by a prudent adherence to approved opinions, Berzelius often prevented the confusion to which the views he combated might have given rise, had they been accepted without reservation. Not that he was averse to healthy reform. But against anything violent—to his mind revolutionary—he fought with all his energy; he did not shun even hot polemics¹ when anything that he regarded as sound was at stake.

His pupil Heinrich Rose gave a comprehensive review of his general character in the "Memorial Speech of Berzelius,"²—a speech of great beauty and with a pleasant warmth of tone running through it. At the close of it (p. 59) Rose says: "The irresistible captivation which Berzelius exercised over those who enjoyed the privilege of a lengthened intercourse with him was only partly due to the lofty genius, whose sparks flashed from all his work, and only partly to the clearness, the marvellous wealth of ideas, and the untiring care and great industry that gave everything with which he had to do the stamp of the highest perfection. It was also—and every one who knew him intimately will agree with me in this,—it was also those qualities which placed him so high as a man; it was his devotion to others, the noble friendship which he showed to all whom he deemed worthy of it, the great unselfishness and conscientiousness, the perfect and just recognition of the services of others,—in short, it was all those qualities which spring from an upright and honourable character."

We may close this section with the following words, in which the same chemist portrays in a few lines the wonderful

¹ His controversies with Dumas, Laurent, Liebig, and others have often been harshly and unfairly criticised, in that a false light has been thereby thrown upon his whole work. The younger generation of chemists, in especial, quickly forgot after his death the debt which was due to him for the imperishable services which he had rendered in the building up of the science. Indeed, derision and cheap ridicule of the mistakes he made are still to be found in recent works which treat of the development of chemical theories.

² Delivered at a public meeting of the Berlin Academy, 3d July 1851.

work of his master: "When a man who is endowed with exceptional talents as an investigator enriches every branch of his science with the most pregnant facts, distinguishes himself equally in empirical and speculative research, and grasps the whole subject in a philosophic spirit; when he arranges each detail systematically and clearly, and gives out the whole to the world in a doctrinal system, critically sifted and put in as perfect a form as possible; lastly, when he proves himself a noble example of a practical and theoretical teacher to a circle of pupils eager for knowledge,—that man so fulfils the highest demands of his science, that he will continue to shine forth as a brilliant model for ages to come."

The Firmer Establishment of the Doctrine of Chemical Proportions and the Development of the Atomic Theory by Berzelius; together with the share taken in these by Gay-Lussac, Dulong and Petit, and Mitscherlich.

It has been already stated in the preceding section that Berzelius regarded the investigation of chemical proportions, and of the laws which regulate these, as his life-task. Compounds of oxygen formed the starting-point for his researches and for the deductions which he made from them, this element being indeed, after the time of Lavoisier, the centre round which the whole of chemistry arranged itself. Even in the first investigations, which he began to publish in 1810 in Swedish, and in 1811 in German (in *Gilbert's Annalen*, vols. xxxvii., xxxviii., and xl.), Berzelius furnished powerful proofs of the existence of chemical and, more particularly, of multiple proportions in the oxygen compounds of the elements. If we consider that he carried out this great work and the subsequent investigations connected with it (for which entirely new methods had to be devised), without aid and relying on his own resources alone, we shall gain some idea of the wonder which such achievements created among his contemporaries.¹

¹ Many passages in the works of Berzelius prove that he looked upon the firm establishing of the doctrine of chemical proportions, and, in connection

A true scientist, Berzelius knew how to advance from the particular to the general; he first collected a number of significant facts which, taken together, rendered possible the gradual building up of the atomic theory. Among these were the proofs that the proportion of sulphur to metal in the metallic sulphides was the same as that in the corresponding sulphates; that the amounts of oxygen in the equivalents of bases were likewise the same; and that in salts of every kind the proportions between the quantities of base, acid, and water were simple ones,—and so on.

In the years 1812 to 1816 Berzelius investigated the stages of oxidation of most of the metals and metalloids then known, and, by determining the composition of these oxides, confirmed the law of multiple proportions. And, notwithstanding that he sometimes proceeded from erroneous premises, *e.g.* from the assumption that chlorine and ammonia contained oxygen, his grasp of the subject was so complete that he was able to keep the main conclusions from his experiments free from error.

Of special significance for the sound development of the

with this, the determination of the atomic weights of the elements and the constitution of chemical compounds, as his chief task. His own words may be quoted here to show how he, impressed as he was with the incompleteness of previous work on the subject, strove to improve upon it: "I soon convinced myself by new experiments that Dalton's numbers were wanting in that accuracy which was requisite for the practical application of his theory. I now perceived that if the light which had arisen upon the whole science was to be propagated, the atomic weights of as large a number of elements as possible, and, above all, of the most commonly occurring ones, must be determined with the greatest accuracy attainable; and, with this, the proportions according to which compound atoms (*zusammengesetzte Atome*) combine among each other, as, for instance, in salts, with the analysis of which I had been occupied for some time. Without work of this kind no day could follow the morning dawn. This was, therefore, the most important point for chemical research at the time, and I devoted myself to it with restless activity. Several of the more important atomic weights I subjected, after lengthened intervals, to a closer scrutiny, making use of improved experimental methods. After work extending over ten years, the results of which have been published in the scientific journals, I was able in 1818 to publish a table which contained the atomic weights, as calculated from my experiments, of about 2000 simple and compound substances."—*Lehrbuch der Chemie*, vol. iii. p. 1161, fifth edition.

atomic doctrine were his efforts (intimately connected with the work just mentioned) to deduce the relative atomic weights of the elements, and also of compound bodies, from the composition of chemical compounds as determined by analysis. He went about this with great circumspection, showing wonderful tact in the selection of proper footholds from which to approach the difficult task. Already in one of his earlier papers¹ we meet with the first statement of the "oxygen law," according to which the amount of oxygen in the acid of a salt stands in a simple numerical proportion to that in the base,—a statement which was the result of experience, and which Berzelius followed in many atomic weight determinations.

The propositions which Dalton had brought forward with a view of arriving at the atomic numbers of the constituents of chemical compounds were rightly designated by Berzelius as arbitrary. Among them, for example, was the assumption that the atomic proportion of two elements to one another, when only one compound of these was known, must be 1 : 1. He, Berzelius, set out indeed from simple premises, but had to exercise all his ingenuity in order to find further support for such assumptions. One of these latter (advanced at the beginning of his work on the subject) was—that 1 atom of one element *A* combines with 1, 2, 3, or 4 atoms of another element *B*. The less simple combining proportions $2A : 3B$ or $2A : 5B$ were first allowed by Berzelius about the year 1819, and without any reservation only in 1827.

With such propositions as a basis, even when including the definitely expressed "oxygen law" (which had been worked out in the meanwhile), Berzelius would have been hardly more successful in solving the question of the number of elementary atoms in a compound than Dalton and his immediate successors, had he not known how to appreciate the value of Gay-Lussac's important discovery of the "law of volumes" for the clearing up of the points in question. By making use of this, the simple combining proportions, in which many elements unite, became all at once apparent ;

¹ *Gilbert's Annalen*, vol. xxxviii. p. 161.

and, by applying it further, Berzelius was able to bring his experimental work to its first conclusion. His *Versuch über die Theorie der chemischen Proportionen und über die chemischen Wirkungen der Elektrizität* ("Essay upon the Theory of Chemical Proportions and upon the Chemical Action of Electricity") appeared first in 1814 in Swedish, in 1819 in French, and in 1820 in German. In this memorable work for the history of chemistry he developed his conception of the atomic doctrine, and his ideas upon the relations between chemical affinity and electric polarity. His dualistic views stood out clearly here, and at the same time he devised a new language and nomenclature for his system. Of special importance was the collection of the results of his arduous investigations in tables of the atomic weights of elements and compounds; he was able to give original figures for about 2000 bodies. In order to become thoroughly acquainted with the grounds which influenced Berzelius in his choice of these values, we must take into account the law of volumes above all other things, because, as has already been mentioned, he not only drew important inferences from it, but used it almost from the beginning of his researches as the basis of his atomic weight system.

Influence of the Law of Volumes upon the Atomic Theory.

Among the most striking of the services rendered by Gay-Lussac was the research which he published towards the end of 1808 in the *Mémoires de la Société d'Arçueil*, vol. ii. p. 207. Having three years previously, in conjunction with Alexander von Humboldt, observed that exactly two volumes of hydrogen unite with one volume of oxygen to form water, he showed by comprehensive investigations that similar simple volumetric relations exist between all gases which combine chemically with one another, and further, that the gaseous products formed also stand in a simple volumetric relation to their components. He proved this, for example, in the formation of two volumes of carbonic acid from two

of carbonic oxide and one of oxygen, and in the combination of hydrogen and chlorine and of ammonia and hydrochloric acid in equal volumes; he likewise showed that two volumes of ammonia were composed of three volumes of hydrogen and one of nitrogen, and two volumes of (anhydrous) sulphuric acid of two volumes of sulphurous acid and one of oxygen. Several of these proportions he was able to deduce from the results of other workers, *e.g.* Dalton, Davy, and Vauquelin, who had determined the volumes with fair accuracy in their experiments on gaseous compounds, without, however, recognising the underlying law.

Having concluded from their similar behaviour with regard to changes of pressure and temperature that all gases possess a like molecular constitution, Gay-Lussac deduced from his researches, just quoted, the following important law:—The weights of equal volumes of both simple and compound gases, and therefore their densities, are proportional to their empirically found combining weights, or to rational multiples of the latter. In this sentence the old idea—that certain definite relations exist in nature between the weight and mass (*pondus et mensura*) of compounds—first found distinct expression.

Gay-Lussac was himself inclined to connect his law of volumes with the atomic theory,—indeed, he recognised in it a support for the latter. But he was unable to set aside certain difficulties which, in spite of the simplicity of the known volume-relations, came in the way, and therefore adhered to his empirical standpoint.

The assumption obviously so closely related to the above, *viz.* that equal volumes of different gases contain equal numbers of smallest particles, and that, in the case of the simple gases, these are not undecomposable, but consist of several atoms, was made so early as 1811 by Avogadro.¹ From such an assumption it followed that the masses of these smallest particles, *i.e.* the molecular weights of the

¹ *Journ. de Phys.*, vol. lxxiii. p. 58. Amadeo Avogadro, born 1776, died while still professor of physics at Turin in 1856. It is through the treatise just mentioned that his name will always remain famous.

gases, were proportional to the vapour densities. The particles were termed by him *molécules intégrantes*, and their constituents (*i.e.* our atoms), *molécules élémentaires*. Notwithstanding the fruitfulness of those conceptions, and the ease with which by their aid the mutual relations between the volumes of gases and the atoms could be explained, they remained almost unnoticed. The reason for this may to some extent have been that Avogadro generalised too boldly, extended his hypothesis to non-volatile substances, and brought forward no new facts in support of it.

But although the conclusions drawn from the law of volumes by the scientist just named remained unheeded at the time, the law itself bore rich fruit for the atomic doctrine. Dalton himself showed a disinclination to agree with the results of Gay-Lussac's researches, indeed, he doubted their correctness. Thomson and Davy too did not perceive that the law of volumes had any special significance from the atomic point of view, as, although they frequently made use of the volume-relations of gases to arrive at their composition, they left these at other times quite out of account; thus they assumed that a volume of hydrogen contained only half as many atoms as an equal one of oxygen.

Berzelius, however, recognised in the law of volumes a welcome corroboration of the atomic theory, and allowed himself to be guided by it in his views upon the number of atoms in chemical compounds, and, consequently, upon the numerical values of the atomic weights. His "volume theory" (*Volumtheorie*) contained the attempt to conjoin Gay-Lussac's law with the atomic theory. He set forth the atomistic view, which he had himself shaped under the influence of the law of volumes, definitely and conclusively in two papers. He started with the assumption that in the case of every simple substance, when it was in the gaseous form, one volume corresponded with one atom, and therefore made use of the designation "volume-atoms" (*Volumatome*) for those smallest particles. He endeavoured, wherever it was practicable, to measure the

¹ *Ann. of Philos.*, vol. ii. pp. 359, 443 (1813).

volumes of the combining substances, and from these deduced the atomic numbers. The analysis of the compound, in which the volumes of the elementary constituents were known, led him to the true determination of the atomic weights of the latter. Thus, from the fact that water consists of two volumes of hydrogen and one of oxygen, he deduced the atomic composition of water which holds at the present day, together with the relative atomic weights of oxygen and hydrogen; and from the mode of formation of carbonic oxide and carbonic acid he arrived at the true composition of these compounds, and at the atomic weight of carbon, etc.

But, however much Berzelius was convinced at that date (1813) of the superiority of this conception over the "corpuscular theory," which took no account of volume-relations, he did not fail to recognise the limited application of his volume theory. To extend to non-volatile bodies the conceptions which he had gained from gases seemed to him hazardous; in fact, his doubts as to the possibility of regarding all elements and chemical compounds from the standpoint of the volume theory grew rapidly, as is easily seen in his *Essay upon the Theory of Chemical Proportions, etc.* (p. 217), which was published a few years after this. But he had, at any rate, found in the law of volumes a valuable aid towards the determination of the atomic composition of numerous substances, and the deduction from this of the atomic weights of many of the elements.

A glance at the table of atomic weights which he published in 1818 shows how reliable the values found by him are, comparing favourably as they do with those of other observers. A later table given out by him in 1827 contained marked improvements on the former one, and brought his atomic weights still nearer to those current at the present day. But great uncertainty still prevailed with regard to the proportional numbers of many of the atomic weights, as compared with that of hydrogen or oxygen. Berzelius took oxygen (as the most important element, the "pole of chemistry") as the basis of his other atomic weights, making that of

oxygen = 100. His ground for this preference¹ was that oxygen was capable of combining chemically with every other element; in fact, oxygen compounds were almost the only ones made use of for the derivation of the atomic weights.

If we calculate his values upon that of hydrogen, which is now the customary unit, we obtain numbers that can be compared with those in use to-day. The following selection of such atomic weights from the year 1818 will serve to corroborate what has just been said (the current values are those in brackets):—

Carbon 12·12 (12)	Lead 416 (207)	Sodium 93·5 (23)
Oxygen 16 (15·96)	Mercury 406 (200)	Potassium 157·6 (39)
Sulphur 32·3 (32)	Copper 129 (63·3)	Silver 433·7 (108)
	Iron 109·1 (56)	

The question now forces itself upon us—What were the grounds which led Berzelius to assume twice as high a value for many metals (*e.g.* iron, lead, mercury, copper, chromium, tin, etc.), and four times as high a one for potassium, sodium, and silver, as are now assigned to them? The reason lay in his presupposition of the simplest possible combining proportions, for at that time such proportions as 2 : 3, 2 : 5, 3 : 4, etc., appeared to him too complicated; only one atom of an element was, in his then view, present in (a molecule of) a compound. The compounds formed by the oxidation of iron, for example, in which the proportions of oxygen were as 2 : 3, and which we now express by the formulæ FeO and Fe_2O_3 , had for him the composition expressed by the formulæ FeO_2 and FeO_3 , whence the atomic weight of iron came out double what we now have it. An analogous composition was attributed to other metallic oxides corresponding to the protoxide and sesquioxide of iron, so that the atomic weights of their metals were doubled. In like manner Berzelius was led, by the assumption that the ratio

¹ In his text-book (first edition, vol. iii. p. 99) he expresses himself as follows: "To refer the other atomic weights to that of hydrogen offers not only no advantages, but has, in fact, many inconveniences, seeing that hydrogen is very light and is seldom a constituent of inorganic compounds. Oxygen, on the other hand, unites all the advantages in itself. It is, so to speak, the centre-point round which the whole of chemistry revolves."

of oxygen in potassic peroxide and oxide was as 3 : 2, to the erroneous conclusion that the latter contained one atom of potassium combined with two of oxygen, and the peroxide one of potassium combined with three of oxygen; hence for potassium and the analogous elements sodium, lithium, and silver, whose oxides have in reality the general formula Me_2O , atomic weights four times higher than the true ones were deduced.

Thus, in spite of Berzelius' gigantic labours, many points attaching to his system of atomic weights still remained uncertain; there were as yet too few reliable and comprehensive data to allow of the true relations of the values found to that of hydrogen or oxygen being firmly established. Berzelius himself was convinced of the insufficiency of the methods by which he had determined the atomic composition of compounds, and, from this, the atomic weights of the elements. Apart from his somewhat arbitrary suppositions, he had merely found in the physical behaviour of gases—in the relation of their specific gravities to the combining weights—a good basis upon which to work out the question of the magnitudes of the relative atomic weights.

The year 1819 brought with it two important discoveries in physical chemistry which helped to clear up the above uncertainties; attention was called almost simultaneously by Dulong and Petit¹ to the relations between the atomic weights of the elements and their specific heats, and by Mitscherlich to the connection between similarity of crystalline form and analogous constitution. The latter discovery and the doctrine of isomorphism which grew out

¹ P. L. Dulong, who was born in 1785 at Rouen, and died in 1838, while Director of the Polytechnic School at Paris, rendered imperishable service, in especial by his physico-chemical investigations. But, apart from these, his purely chemical labours, *e.g.* that upon chloride of nitrogen, in discovering which compound he lost an eye and several fingers (in 1811), that upon the oxygen compounds of phosphorus and nitrogen, and his fruitful speculations upon the constitution of acids, ensure him an honourable place in the history of the natural sciences.

T. A. Petit was born in 1791, and died while Professor of Physics at the Polytechnic School so early as 1820. To chemists he is known by his conjoint work with Dulong on the atomic heats of the elements (see above), his other researches being purely physical.

of it were largely made use of by Berzelius for determining relative atomic weights; but to Dulong and Petit's statement he paid much less heed, as it still required extension and corroboration. Both of those aids have exercised such a profound influence on the development of the atomic weight system that they must be shortly treated of here, in so far as they refer to the latter (cf. section devoted to the history of Physical Chemistry).

Dulong and Petit's Law.

From researches¹ carried out in part with substances not quite pure, and in part by methods upon which not much reliance could be placed, those two investigators drew the important conclusion that the specific heats of a number of the solid elements, the metals in particular, were nearly inversely proportional to their atomic weights. But, however bold those deductions were, deductions which they expressed in the sentence: "The atoms of simple substances have equal capacities for heat," their confidence in them was on the whole justified by later and more accurate experiments; at any rate most of the metallic elements fulfilled the Dulong-Petit law approximately. The exceptions to it, shown by many of the metalloids in a greater or lesser diminution of the atomic heats, have only in some measure been explained in recent years by the proof that the specific heats of such elements vary greatly with the temperature. In the case of simple chemical compounds, too, a relation was soon found between their specific heats and atomic weights (by Neumann, in 1831).

When once its validity had been proved, the significance of the Dulong-Petit law for the determination of the relative atomic weights of the elements became immediately apparent. One had merely to determine the specific heat of an element in order to arrive at its atomic weight from this, taken in conjunction with the atomic heat (which was assumed to be a constant), i.e. the product of the specific heat into the atomic

¹ *Ann. Chim. Phys.*, vol. x. p. 395 (1819).

weight. Dulong and Petit at once went on to apply their law to this problem, and came to the conclusion—a conclusion which was later recognised as correct—that the atomic weights ascribed by Berzelius to several of the metals must be halved.

There was, however, as yet no pressing reason why the latter, on a dispassionate review of Dulong and Petit's results, should at once agree to this demand. That those results were of great significance for theoretical chemistry he willingly admitted, but he maintained that they had not yet been proved to be of such general application that a law could be formulated from them. He especially opposed any fundamental alterations of his own atomic weights, as he held that, if this were done, improbable atomic proportions would have to be assumed for the compounds of some of the elements. This attitude towards the Dulong-Petit law was only gradually abandoned by Berzelius, after further proofs bearing on the point had been adduced.

Influence of the Doctrine of Isomorphism upon Atomic Weight Determinations.

After the development of crystallography by Romé de l'Isle and Haüy, various experimenters had observed that substances of different chemical composition crystallise together in one and the same crystalline form. As instances of this may be mentioned Gay-Lussac's observation that crystals of potash alum grow in a solution of ammonia alum, while still retaining their original crystalline form, and Beudant's, that copper vitriol is obtained in the same form as iron vitriol when a small quantity of the latter is added to a solution of the former, and so on. But neither this observation nor the definite statement by Fuchs upon the replacement of certain substances in minerals by others [his doctrine of "vicariating constituents" (*Vikariirenden Bestandtheilen*¹)] led to the recognition of the relation between crystalline form and chemical constitution.

¹ This means substitution without any accompanying change of crystalline form; thus, to give one or two examples, Fe'' can replace Ca'', and Al''' can replace Fe''' or Cr''' in this way.

This important discovery¹ was reserved for E. Mitscherlich,² who explained the occurrence of isomorphous crystals in substances of different nature by proving that they possessed a similar chemical composition. Thus he found, on examining the salts of phosphoric and arsenic acids, that only those of analogous composition and containing equal amounts of water of crystallisation were isomorphous. His subsequent investigations of selenates and sulphates, of the isomorphism of magnesium and zinc oxides, and of iron, chromium, and aluminium salts, confirmed the intimate connection existing between crystalline form and chemical composition. At first, after making those observations, Mitscherlich was of opinion that isomorphism depended chiefly on the number of the elementary particles (in the molecule), but he soon convinced himself that the chemical nature of these had also to do with it.

Berzelius, who regarded the discovery of isomorphism as "the most important since the establishment of the doctrine of chemical proportions," endeavoured to arrive at the atomic weights of the elements by the aid of isomorphous compounds. For, according to him, isomorphism meant similarity in atomic constitution; chemists only required to know the composition of one compound in order to deduce that of the remaining isomorphous ones from it. The quantities of the isomorphous elements which replaced one

¹ *Berl. Akad. Abhandlungen der phys. Klasse*, 1818-19, p. 426; also *Ann. Chim. Phys.*, vol. xiv. p. 172; vol. xix. p. 350.

² Eilhard Mitscherlich was born in 1794 in Oldenburg, and died in 1863 at Berlin, where he worked as Klaproth's successor in the University from the year 1821; he enriched chemistry by beautiful discoveries, and especially advanced it very greatly in the physical direction. At the beginning of his career he devoted himself to oriental and linguistic studies, only taking up the natural sciences incidentally; but his intercourse with Berzelius, to whom he went in Stockholm in 1819, was decisive as to his future course. His work will frequently be referred to in the special section of this book; but mention may be made here of his important investigation of manganic and permanganic acids, his work upon selenic acid, and that upon benzene and its derivatives. His successful attempts to prepare minerals artificially give further proof of the many-sidedness of the man, his greatest achievement of all being the discovery of isomorphism, mentioned above. His *Lehrbuch der Chemie* is marked by originality both of form and contents.

another, referred to a definite unit,—say oxygen or hydrogen—were regarded by Berzelius as the relative atomic weights. He made extensive use of this new aid to confirm the correctness of his atomic weight determinations.

The Atomic Weight System of Berzelius after 1821.

At first, in 1821, Berzelius did not consider that any change in the atomic weights was called for, as the new facts could be made to accord with his determinations and deductions. But five years later he resolved, after minute consideration, upon certain modifications, which consisted in halving the atomic weights of many of the elements. The grounds which weighed with him in this he set forth in a conclusive manner.¹ What mainly necessitated the giving up of his former assumptions was the composition of chromic oxide and chromic acid. The amount of oxygen in the latter (so he writes) was to that of the base as 3:1 in neutral salts, whence the composition CrO_3 followed for chromic acid; while in chromic oxide the proportion was as $\text{Cr}_2:\text{O}_3$. But, in order to concede this last, he had to give to ferric and aluminic oxides (oxygen compounds isomorphous with and capable of replacing chromic oxide) the analogous compositions Fe_2O_3 and Al_2O_3 , and to their metals, as a consequence, only half as large atomic weights as he had previously done. Iron protoxide received the simplified formula FeO , and the oxides of magnesium, zinc, nickel, cobalt, etc., which were isomorphous with it, were regarded as similarly constituted. The result of all this was, as already stated, the halving of the atomic weights hitherto in use, so that these now conformed to Dulong and Petit's law. With the atomic weights of sodium, potassium, and silver, which Berzelius likewise halved, the circumstances were peculiar. He had arrived at the conclusion, with respect to basic oxides, that the strong bases (such as oxide of potassium) contained metal and oxygen in the

¹ *Pogg. Ann.*, vol. vii. p. 397; vol. viii. pp. 1, 177.

proportion 1 : 1, and therefore gave potassium, sodium, and silver double their proper atomic weights; for, according to our present ideas, two atoms of the metal are combined in these with one of oxygen. The following list by him of atomic weights of some of the more important elements, with hydrogen as the unit, shows the approximation of the numbers to those in use to-day, and also the amendment¹ which some of them had undergone during the years 1818-26 (cf. table, p. 205)—

Carbon	.	12.25	(12)	Lead	.	207.4	(207)
Oxygen	.	16	(15.96)	Mercury	.	202.8	(200)
Sulphur	.	32.24	(32)	Copper	.	63.4	(63.3)
Nitrogen	.	14.18	(14)	Iron	.	54.4	(56)
Chlorine	.	35.47	(35.4)	Sodium	.	46.6	(23)
Phosphorus	.	31.4	(31)	Potassium	.	78.5	(39)
Arsenic	.	75.3	(75)	Silver	.	216.6	(108)

The figures in brackets indicate the current values.

In this table of 1826 we find for the first time the atomic weights of nitrogen and chlorine as simple substances. Berzelius held longer than any other chemist to his assumption that they contained oxygen; the grounds which necessitated his giving up this hypothesis are gone into below.

If we review those efforts of Berzelius at determining the atomic weights of the elements, we see that he was mainly guided, in the case of non-volatile bodies, by the composition of the oxygen compounds, *i.e.* by the determination of the proportion of element to oxygen, and then, secondly, by the doctrine of isomorphism, while to the

¹ Berzelius, who had devoted his whole energies to perfecting analytical methods and amending the atomic weight numbers, had to suffer harsh language from others who, by reason of improvements in such methods, attained to still more exact results; this applied in an especial degree to Dumas (cf. *Ann. Chem.*, vol. xxxviii. p. 141 *et seq.*), who determined the equivalent of carbon "with every imaginable precaution," and found its value to be 6. The difference between this number and that which Berzelius had found, viz. 6.12, caused Dumas to utter the most severe reproaches against the great master of analysis (cf. Berzelius' mild reply, *Lehrb. d. Chem.*, vol. iii. p. 1165, and Liebig's admirable protest against Dumas' procedure, *Ann. Chem.*, vol. xxxviii. p. 214 *et seq.*)

Dulong-Petit law he allowed only a slight influence. In those cases where the elements or simple compounds of them were known in the gaseous state, his volume theory came in as an aid in deducing the desired values. Berzelius still held fast to the idea that the amounts of the elements contained in equal gaseous volumes were proportional to their atomic weights. But this assumption was soon overthrown by the remarkable results of an investigation which exercised such a profound influence on the views of many chemists that it must be described at this point.

Dumas' Attempt to alter the Atomic Weights.

In the year 1827 a young chemist, J. B. A. Dumas, who had already made himself favourably known by other work, published a research,¹ the great merit of which lay in the working-out of an admirable method for the determination of vapour densities. By this method he succeeded in estimating the specific gravity of the vapours of several elements; and the relation existing between these comparable values was, according to Dumas (who took up here the same standpoint as Berzelius), that of the relative atomic weights. The elements which he adduced were iodine and mercury, and to these he added phosphorus and sulphur a little later.² The result of this was the obtaining of different numerical values from those assumed by Berzelius for the atomic weights of the above elements, which had been held for a year past. Taking the atomic weight of hydrogen as 1, and that of oxygen as 16 (Berzelius' numbers), the above vapour densities gave the values 123 for iodine, 101 for mercury, 62·8 for phosphorus, and 96 for sulphur. Further, Mitscherlich determined the vapour density of arsenic in 1833, and calculated from this the atomic weight 150. True, these numbers bore a simple relation to the atomic weights of Berzelius, that of the latter for mercury (200) being double, those for phosphorus and arsenic (31 and 75) half, and that for sulphur one-third as

¹ *Ann. Chim. Phys.*, vol. xxxiii. p. 337.

² *Ibid.*, vol. xlix. p. 210; vol. l. p. 170.

great as the values deduced by Dumas, and held by him to be the correct ones. The result of this alteration of the atomic weights by the latter was great confusion. While Berzelius remained true to his own numbers, holding mercuric oxide, for example, to be composed of mercury and oxygen in atomic proportions, Dumas assumed in it two atoms of mercury to one of oxygen, and gave it the composition and formula which Berzelius ascribed to mercurous oxide, viz. Hg_2O . Again, to phosphuretted hydrogen, in which Berzelius quite rightly assumed the proportions of three atoms of hydrogen to one of phosphorus, on account of its analogy to ammonia, Dumas gave twice as many atoms of hydrogen, and therefore the formula PH_6 .

In making the above alterations Dumas' procedure was quite without method, and only helped to complicate matters further. He drew a theoretical distinction between smallest physical and chemical particles, bearing Avogadro's speculations in mind; but this attempt at separating molecule from atom remained not only unfruitful, but resulted in confusion. The manner in which Dumas spoke of half an atom of oxygen, and of hydrochloric acid as composed of half atoms of hydrogen and chlorine, must have been unintelligible at that time,¹ and was sharply criticised by Berzelius.

A comparison of the atomic weights of Berzelius and Dumas with those of to-day shows us how fully justified the former was in adhering to his own, which he had arrived at after the most mature consideration; Berzelius' values have proved to be the right ones. In view of recent experiences, however, he became more cautious in the use of his volume theory, and from henceforth only applied the law—that the atomic weights of the elements are proportional to the densities of their vapours—to the permanent gases.

The mighty reform which Dumas aimed at in this section of theoretical chemistry remained without result; and there is justification for the reproach brought against

¹ If Dumas had been fully acquainted with Avogadro's ideas, he would have cleared up the opposing points which remained unsolved.

him by many, and more especially by Berzelius, of having introduced obscurity and disorder into the atomic weight system of the latter. For the sake of an unproven hypothesis Dumas neglected the most striking chemical analogies (*e.g.* that between ammonia and phosphuretted hydrogen), and frequently confused things which were perfectly clear. In consequence of the objections which he raised to Berzelius' atomic weights of the elements, the distrust of these latter by contemporary chemists grew in extent, so that we find even the most distinguished investigators like Gay-Lussac and Liebig doubting whether it was possible to determine the relative weights of the atoms with certainty. They would have satisfied themselves with establishing the equivalents, and leaving the atomic weights quite out of account. The opposition to the atomic weight system of Berzelius was at its height towards the end of the third and beginning of the fourth decade of the century. In Germany, in especial, L. Gmelin advocated the establishing of the simplest "combining weights"; but the certainty of being able to determine the true equivalents of the elements was not in itself sufficient, although Faraday's discovery of the electrolytic law in 1834 appeared to guarantee a solid basis for this (see second paragraph below).

Michael Faraday, who was born in London in 1794, was endowed with such exceptional inclination and aptitude for the study of the natural sciences that he worked his way up from humble circumstances, although he had received no systematic training previous to his connection with Davy. Davy immediately recognised the extraordinary talents of the youth, and got him to assist him in his work, this connection soon developing into a close friendship. Faraday's most important discoveries belong to the domain of physics (his investigations on induction currents, electromagnetism, and diamagnetism). His electrolytic law, which was of such significance for the electro-chemical theory, is touched upon below. He made himself known to the chemical world particularly by his beautiful investigations on the liquefaction of gases, by his work on the hydro-

carbons from oil-gas (when he proved the isomerism of butylene with ethylene), and by that on the chlorides of carbon. He was one of the earliest to promote the study of physical chemistry, which received from him its first great advance since the investigations of Dulong and Petit on specific heat, and those of Mitscherlich on isomorphism. The results of most of his experimental work were published in the *Philosophical Transactions*, but some in *Poggendorff's Annalen* and other journals. During the greater part of his life (he died in 1867) he worked at the Royal Institution, in which he became professor in 1828. In addition to his wonderful gifts as an investigator, Faraday possessed in an exceptional degree the power of clear and pleasant exposition; the memory of his "Lectures to Children" at the Royal Institution still survives (see his delightful little book, *The Chemical History of a Candle*). In private life the simplicity and amiability of his character made him greatly beloved.

Faraday made the memorable observation (see above) that the same galvanic current decomposed electrolytes, *e.g.* water, hydrochloric acid, and metallic chlorides, in such a manner that equivalent amounts of hydrogen or metal were separated at the negative pole, and the corresponding quantities of oxygen or chlorine at the positive.¹ He grouped those facts together under the title of "The Law of definite Electrolytic Action." In the determination of electro-chemical equivalents he saw a sure auxiliary means for fixing chemical atomic weights in doubtful cases. Berzelius, however, did not recognise any necessity in this case either for departing from his own atomic weights, but disputed the correctness of the numbers obtained by the electrolytic method.

The time for a clear grasp of the terms *equivalent*, *atom*, and *molecule*, and for drawing a sharp distinction between these, was not yet come. Berzelius was, therefore, perfectly justified in adhering to his relative atomic weights, the best proof for which was to be furnished later on. But, as already remarked, he now only made use of his volume-theory

¹ *Phil. Trans.* for 1834, or *Pogg. Ann.*, vol. xxxiii. p. 301.

in a greatly modified degree, in consequence of the results obtained by Dumas and Mitscherlich. With regard to vapours, he foresaw (in 1835) the possibility of the relation between volume and atomic weight being a variable one (he drew a distinction between gases and vapours, and only strictly applied the law of volumes to the latter).

How, in the course of the 'succeeding decades, Gmelin's combining weights became gradually replaced by the atomic weights now in use (most of which had been brought forward by Berzelius), will be detailed later on. The reader's attention will be chiefly directed in the following sections to Berzelius' energy in a speculative direction, as shown in the setting up of his dualistic system; this last was the fruit of an electro-chemical theory which, along with Davy's, now falls to be briefly considered.

*The Electro-Chemical Theories of Davy and Berzelius.—
The Dualistic System of Berzelius.*

The perception that a close relation existed between electrical force and chemical reactions spread rapidly at the beginning of the century, after the decomposition of water into its constituents by the galvanic current had been proved by Nicholson and Carlisle, and that of salts into their bases and acids by Berzelius and Hisinger (in 1803). The first fruit of the many and varied observations on the action of the current on chemical compounds, and on the accompanying electro-motive force in chemical reactions, was Davy's Electro-Chemical Theory,¹ which he thought that he had founded firmly by his ingeniously devised researches, begun in the year 1800. He took as his starting-point the proved experimental fact that different substances, capable of combining chemically with one another, *e.g.* copper and sulphur, became oppositely electrified upon contact when insulated. Heating intensified the resulting difference of potential, until it vanished in consequence of the chemical combination of the substances. This latter,

¹ *Phil. Trans.*, 1807, p. 1; cf. also his *Elements of Chemical Philosophy*.

Davy then reasoned, is simultaneous with the equalisation of the potentials. The greater the difference between these before combination, the greater must also be the chemical affinity of the different substances for one another. By the addition of electricity to the compounds, their constituents receive the same electric polarities which they possessed before combination; the positive constituents go to the negative pole, and the negative ones to the positive.

Davy inclined to the assumption that electrical processes and the phenomena of chemical affinity arose from a common cause. His electro-chemical theory was characterised by the axiom that the small particles of substances which have an affinity for one another only become oppositely electrified upon contact. But later researches, especially those of Berzelius, led to the abandonment of this principle, while, otherwise, many of Davy's original ideas were retained.

Berzelius brought forward the main outlines of his electro-chemical theory in 1812,¹ after having already at various times expressed his views upon the indissolubility of chemical and electrical processes, upon combustion as an electro-chemical phenomenon, and on the probability of the small particles being polarised. But the theory as a whole, with its far-reaching conclusions, was first published in his *Versuch über die Theorie der Chemischen Proportionen*, etc., already mentioned at p. 201. In this we see clearly how he deduced his theory from facts, and then how, from the standpoint so obtained, he succeeded in penetrating and dominating with it the whole domain of chemistry. His doctrine, developed in this way from the electro-chemical point of view, continued the prevailing one for the next twenty years, until it had to yield to the pressure of facts with which it could not be reconciled.

Berzelius started with the primary assumption that the atoms of elements were in themselves electric; electric polarity, therefore, was an essential property of these smallest particles, which further possessed at least two poles, whose quantities of electricity were in most cases

¹ *Schweigger's Journ.*, vol. vi. p. 119.

different, so that either positive or negative electricity predominated in the particle as a whole. Thus elements were divided into positive and negative, according to whichever of those electricities prevailed; and this last point was easily solved by noting whether the element in question was separated at the negative or positive pole of the galvanic battery upon electrolysis.¹ In like manner Berzelius assumed a polarity for compounds as well as for elements, although, in consequence of the neutralisation of the opposite electricities by one another in the formation of compounds, this polarity was thereby weakened. The *intensity of the polarity* was, according to him, a measure of the excess of one or the other kind of electricity. The dissimilar polar intensity of the small particles was regarded as the cause of their various affinities (*der verschiedenen affinitätswirkungen*). And, as the forces of affinity were found to be dependent on the temperature, so was also polarity to be regarded as a function of heat.

Chemical combination of the elements or compounds consisted, according to Berzelius, in the attraction of the dissimilar poles of the small particles, and in the consequent neutralisation of the different electricities. If positive electricity predominated in the original substance, then an electro-positive compound resulted, and *vice versa*. If the electricities neutralised one another, then an electrically indifferent product was the result. Oxygen, as the most electro-negative element, served Berzelius here (as it had done in his atomic weight estimations) as the standard by which to determine the kind of polarity of the various elements. Those elements which yielded basic compounds with oxygen, even although only their lowest oxides were basic, were classed as electro-positive, and those whose oxides were acids as electro-negative. Following this principle he arranged the simple substances in a series, in which oxygen as the first member was followed by the other metalloids, while hydrogen formed the bridge between the

¹ At first Berzelius designated the elements after the poles at which they were separated, *i.e.* he called the metals negative, and the metalloids positive.

latter and the metals, the whole ending with sodium and potassium. In referring to this, Berzelius frequently stated that many elements which were positively polar with regard to some were negatively polar with regard to others, *e.g.* sulphur was positive to oxygen, but negative to the metals and hydrogen,—and so on. Oxygen alone he held to be an absolutely negative element, because in no case did it behave as a positive one with respect to any other.

By the aid of such conceptions, which formed the substance of his electro-chemical theory, Berzelius was enabled to interpret satisfactorily the facts which were at that time considered of greatest moment. The electrolytic processes, *i.e.* the separation of the positive and negative constituents of compounds at the negative and positive poles respectively, were explained in a simple manner by the assumption that the galvanic current reinvested the small particles of compound bodies with their original polarity. The many and various manifestations of affinity could in this way be referred back to a common cause.

Proceeding from this one hypothesis,—that electric polarity was a property of the atoms of substances, Berzelius was able to bring light and order into the province of inorganic chemistry, which was at that time (1819) almost the only branch of the science to be considered. His electro-chemical theory led him, in the first instance, to a perfectly definite conception of the “constitution or rational composition of chemical compounds,” and then to a nomenclature and corresponding system of formulæ developed from this. His efforts in this direction were crowned with the greatest success. Even at the present day we cannot do without the chemical language which he introduced, although, on the other hand, his dualistic views on the composition of chemical compounds have not lived so long. He was, however, the first to draw a precise distinction between the *empirical* and *rational* composition of chemical compounds. The *constitution* of the latter was, according to him, arrived at by investigating their proximate constituents (such being, for instance, Cu_2O , CuO , and $(\text{C}_2\text{H}_5)_2\text{O}$ in copper

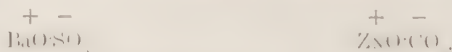
salts, ethers, etc.), and this task he regarded as one of the most important which falls to the lot of the chemist. He himself devoted his whole energies to its solution, the electro-chemical theory serving as a means whereby he might attain to this great end.

The Dualistic System of Berzelius.

The necessary consequence of the electro-chemical view was the assumption that every compound body consisted of two parts, which were electrically different: without such difference a chemical compound could not be formed. Further, the constitution of the latter was known when its positive and negative constituents were demonstrated. It was again compounds of oxygen,—acids, bases, and salts, by means of which Berzelius developed this, his dualistic doctrine. The elements which were combined with oxygen were the positive constituents, *e.g.* the metals in oxides, and the metalloids in acids. The electro-chemical antithesis was illustrated by the following formulæ:—

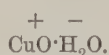


The anhydrous bases are the positive constituents of salts, and the acids—in which negative polarity predominates—the negative ones, as is shown by the formulæ—



Berzelius considered that the strongest proof of the correctness of this theory lay in the electrolytic decomposition of compounds, especially of salts, into the above-mentioned two portions, which were separated at the poles of opposite electricity to their own. He further sought to explain the composition of double salts according to the dualistic hypothesis, giving, for example, sulphate of potash as the positive, and sulphate of alumina as the negative constituent of alum.

In the year 1818, when Berzelius published a detailed exposition of his electro-chemical theory, he was convinced that all acids contained oxygen. In his view water played in hydrated acids the part of a weak electro-positive constituent, and in metallic hydroxides that of a weak electro-negative one; the hydrates of sulphuric acid and of cupric oxide therefore received the formulæ—



The binary conception, which had already been applied by Lavoisier to acids and bases, and even by Rouelle to salts, thus received the strongest support from the electro-chemical theory. It will be shown in the next section how Berzelius was obliged to give up Lavoisier's one-sided theory of the oxygen acids.

The efforts of Berzelius at introducing a rational and generally applicable nomenclature go back to the year 1811.¹ His nomenclature is a continuation of that of Lavoisier, de Morveau, and Berthollet, which, however, he greatly extended and amplified, his first efforts in this direction having been published in the *Versuch über die Theorie der Chemischen Proportionen*, etc., already frequently mentioned. The division of the elements into metalloids and metals, according to their electro-chemical character; that of the positive oxygen compounds into suboxides, oxides, and peroxides; and the corresponding division of the acids (which were designated according to their degree of oxidation), have been found to be so convenient that only very trifling alterations have had to be made in them. In like manner he designated the chlorine compounds corresponding to the oxides by adding different final syllables or prefixes, *e.g.* sub-chloride (*Chlorur*), chloride, perchloride, etc. In the nomenclature of the oxygen salts the name of the acid constituent preceded that of the basic, *e.g.* sulphate of oxide of copper.

He also endeavoured to apply similar principles in

¹ *Journ. de Phys.*, vol. lxxiii. p. 257.

naming organic compounds, whose constitution had been determined on his own lines. But the time had not yet come when it was possible to devise a rational nomenclature for these.

Berzelius next established a system of chemical notation, connected in the most intimate manner with his chemical nomenclature, which had given expression in clear language to the electro-chemical views on the composition of bodies: this notation was to attain to the same end in a more concise manner. In doing this he rendered an immense service, for it thus became possible, by the aid of simple symbols, not merely to express the composition of chemical compounds, but to show even complicated reactions of these in an easily intelligible manner. He gave to each element a symbol, which was usually the first or the two first letters of its Latin name, less often of the Greek one; thus the symbol *H* stands for hydrogen (*hydrogenium*), *S* for sulphur (*sulphur*), *O* for oxygen (*oxygenium*), *C* for carbon (*carbon*), and so on. These symbols denote at the same time the atomic weights of the elements in question.

By placing these symbols alongside of one another, and adding a figure to indicate the number¹ of atoms when the latter amounted to more than one, the formulæ of chemical compounds was obtained: *e.g.* H_2O for water, SO_2 for sulphurous acid, CO_2 for carbonic acid, $\text{Na}_2\text{O} \cdot \text{CO}_2$ for carbonate of soda, etc.

What an advance upon Dalton's attempts towards the same end, his figures being only of use for illustrating the simplest of compound substances! Dalton's notation was soon forgotten, never having, indeed, met with general approval, while that of Berzelius became an indispensable aid to chemists of all nations.

Berzelius attached a special meaning to the symbols with a bar drawn across them, these being employed by him to indicate that the elements in question were in the state

¹ Berzelius also denoted the number of oxygen atoms by dots, and that of sulphur atoms by commas, *e.g.* calcium oxide, Ca ; iron bisulphide, Fe ; this system remained longest in use among mineralogists.

of double atoms, or, as he put it,¹ that "they remained connected together;"² this applied, for example, to the hydrogen in water, HO , to the chlorine in anhydrous perchloric acid, ClO_7 , and to the iron in the sesquioxide, FeO_3 . This manner of notation, which had exceedingly bad results, arose from Berzelius taking oxygen as his unit, and using it to some extent as the standard for the saturation-capacities of other elements.³ He was thus led to the assumption of the double atom constituting a chemical unit, and the above symbols with bars served him to give expression to this; at a later period, however, he gave up using them, and reverted to the true atomic weights. This idea, which Berzelius cherished for a time, of the atoms of certain elements being only present as pairs in compounds, was not acquiesced in by many chemists; these latter assumed simple instead of double atoms, and therewith equivalents instead of atoms. Blomstrand, who has shown in his admirable work, *Die Chemie der Jetztzeit* ("The Chemistry of the Present Time"), the close connection which exists between the views of Berzelius and those held to-day, describes the results of the system of notation and of the views just mentioned in the following eloquent words: "This erroneous conception was without doubt the almost sole reason why Berzelius' atomic theory found so little acceptance; it prevented, like a restraining curb, the free development of the latter, and led little by little to a peculiar confusion with regard to the fundamental principles of chemistry, in that the distinction between atomic weight and equivalent was by degrees nearly effaced, until at last the volume-atomic weights and the whole atomic theory of Berzelius became almost entirely forgotten by the great majority of the chemists of his school."

Like every innovation, the admirable system of notation which Berzelius recommended met with most violent opposi-

¹ *Lehrb. d. Chemie*, fifth edition, vol. i. p. 121.

² "... dass sie zusammenhängend bleiben."

³ Berzelius designated oxygen as "the measure of the relative weight according to which an element entered into combination" (*dass Mass der relativen Gewichtsmenge, nach welcher ein Grundstoff corzugsweise Verbindungen eingeht*).

tion from many chemists, especially in England. People spoke of "abominable symbols" which were more calculated to introduce confusion than clearness.

Thus in 1820 the dualistic system, with the electrochemical theory for its basis, stood fully equipped, and was soon made use of by the vast majority of chemists as a guide in the confusion which resulted from the daily accumulation of new facts. Berzelius further attempted to apply the dualistic hypothesis in organic chemistry, which, from the third decade of the century, was more and more attracting the attention of chemists. How it came into collision here with the unitary theory, and had finally to succumb to the latter, will be treated of further on.

*Manifestations against Dualism—Theory of the Hydrogen-
and of the Polybasic Acids.*

The tenet which was set up by Lavoisier, and which Berzelius defended with all his might,—that the character of acids depends upon their containing oxygen, and that consequently this element is an unfailing constituent of their salts,—this *theory of the oxygen acids* was already greatly shaken towards the end of the first decade of the present century, and was abandoned by most chemists during the second, as a knowledge of facts opposed to it increased. Finally Berzelius, who remained longest true to the older idea, convinced himself of the existence of acids free from oxygen. The gradual transformation of chemistry, which resulted from the setting aside of this dogma (that all acids contained oxygen), was a thorough one, for the unadaptable dualistic system was thereby battled with, and its fall prepared for.

In order to comprehend this change of views thoroughly, a clear light must be thrown upon the facts which brought it about. The discovery of the alkali metals by Davy and the allied researches on the nature of chlorine must be regarded as the starting-points from which the light of the new knowledge radiated. Before Davy, who had recognised

in the galvanic current a powerful means for decomposing chemical compounds, isolated potassium and sodium from the alkalies by its aid,¹ the latter were regarded as undecomposable, even although, from the time of Lavoisier, it was considered probable that they were constituted analogously to the metallic oxides, and were therefore oxygen compounds. The many fruitless experiments which Davy had made with the alkalies in solution were finally crowned with success when he exposed those substances, only slightly moistened, to the action of a strong current. His correct assumption, that the metals separated at the negative pole were true elements, did not indeed find immediate acceptance; in fact he himself was temporarily in doubt as to whether they did not contain hydrogen, especially after the presence of the latter element in the alkalies had been proved by Gay-Lussac and Thénard, both of whom from this point took an active part, by their researches,² in the solution of the problems in question. The idea that the alkali metals might be hydrogen compounds had crept in from an analogy drawn between them and ammonia; at that time the latter was supposed to contain oxygen, which was withdrawn from it in the formation of ammonium amalgam. The erroneous conclusion that the above metals contained hydrogen, which resulted from this false interpretation, was, however, put right by Gay-Lussac and Thénard, who explained the point correctly. (Cf. *below*. It was mainly upon the three reactions specified towards the end of the next paragraph that Gay-Lussac and Thénard relied here; from these the elementary nature of the alkalies, as well as of chlorine, followed.) Consequently, from the year 1811, potassium and sodium were regarded as metals and therefore as elements.

With the clearing up of the above points, the question as to whether chlorine was really a compound substance, and not rather a simple one, rapidly approached its solution. According to the assumption of Berthollet and Lavoisier,

¹ *Phil. Trans.* for 1808, p. 1.

² *Ann. de Chimie*, vol. lvi. p. 205; vol. lxx. p. 325.

hydrochloric acid contained oxygen combined with a *radical muriatique*, and the chlorine which was liberated by its oxidation was looked upon as oxidised hydrochloric acid. At the time when Davy¹ and Gay-Lussac and Thénard² began their memorable investigations, hydrochloric acid gas was generally held to contain chemically combined water. But even with the most powerful reducing agents these chemists were unable to prove the presence of oxygen either in hydrochloric acid or in chlorine, and this of itself made them incline to the belief that chlorine was an element and hydrochloric acid its hydrogen compound. The idea, however, of oxygen being a necessary constituent of all acids had taken such firm root that numerous fresh investigations were required before it could be driven away. The most important of the observations which led to this were the following:—Hydrogen and chlorine unite to form anhydrous hydrochloric acid, which is decomposed by sodium with the liberation of half its volume of hydrogen and the formation of sodium chloride, while the latter also results directly from the combination of sodium and chlorine.

Upon the ground of those facts Davy was the first to express the distinct opinion that chlorine was an element, suggesting for it the name³ by which it has since been known. At first Gay-Lussac and Thénard had misgivings about agreeing to this, fearing to disturb the uniformity of the chemical system. But, after the former had completed his famous investigation upon iodine, both he and Thénard, as well as other French chemists, were obliged to concur in Davy's view. Iodine and fluorine now received a place among the elements, next to their analogue chlorine.

Berzelius did not allow himself to be convinced at once of the necessity for this thorough innovation, which entailed the abandonment of the oxygen theory. The unity of chemical theory went with him before everything else; he saw in the reform aimed at an overthrow of the principles

¹ *Phil. Trans.* for 1810, p. 231.

² *Mémoires de la Société d'Arçueil*, vol. ii. p. 339.

³ *Phil. Trans.* for 1811, p. 1.

which had governed the older chemical system. After having given eloquent expression to his ideas on the subject in letters to Marcet, Gilbert, Thomson, and others, he collected together the arguments in favour of the older conception in a treatise¹ entitled: *Versuch einer Vergleichung der älteren und der neueren Meinungen über die Natur der oxydierten Salzsäure, zur Beurtheilung des Vorzuges der einen vor der anderen* ("An Attempt to compare the Old and New Opinions with regard to the Nature of the Oxidised Muriatic Acid, and to estimate the Advantages of the One over the Other"). His standpoint is clearly set forth in the following words: "I decline to give in my adhesion to the new doctrine until it has been made perfectly consistent and uniform with the new theoretical science which its authors claim to have built upon the ruins of the chemical theory which they have demolished. For I demand uncompromisingly from any chemical theorem that it shall agree with the rest of chemical theory and be capable of incorporation in it; if this be not the case, then I must reject it, unless, indeed, the evidence in its favour is of such an incontrovertible nature as to necessitate a revolution in the chemical theory with which it does not coincide."

In one point, however, Berzelius soon gave up the opinion that every acid must contain oxygen, by recognising sulphuretted and telluretted hydrogens as hydrogen acids; this latter nomenclature (*hydracides*) was first made use of by Gay-Lussac. At that time Berzelius still held that oxygen was present in chlorine, iodine, and fluorine, even after Gay-Lussac's admirable research upon the salts of hydrocyanic acid had proved that these were free from it. It was only after he had been able to make the results of his own investigations on ferro-cyanogen and sulpho-cyanogen compounds agree with the theory of non-oxygenated acids that he resolved to include chlorine and iodine among the elements. About the same time (1820) he gave up the idea that nitrogen and ammonia contained oxygen; but it was not until 1825 that he abandoned what remained of his old view, by

¹ *Gilbert's Annalen*, vol. I. p. 356.

including fluorine with chlorine and iodine among the salt-forming elements or halogens;¹ he drew a sharp distinction between the *haloid salts*, i.e. the salts produced by the combination of the above elements with the metals, and the *amphid salts*, or those containing oxygen.

Theory of the Hydrogen Acids.

Even before Berzelius had given up the oxygen-acid theory, Davy,² and almost at the same moment Dulong, made the attempt to bridge over the gap between the oxygen and hydrogen acids by a uniform interpretation of their constitution. In these efforts we see the beginnings of the hydrogen-acid theory, which was to become of such great importance a few decades later on. From his observation that iodic anhydride was devoid of acid properties, but acquired them after combination with water, Davy drew the conclusion that hydrogen and not oxygen was the acidifying principle in the latter compound; hydrogen, in his opinion, was an essential constituent of all acids. The assumption, that hydrated acids and salts contained water or metallic oxides together with acid anhydrides, he held to be unproven. Dulong expressed himself in a similar sense after an investigation of oxalic acid and its salts; the former he regarded as a compound of hydrogen with carbonic acid, while in the latter he assumed an analogous combination of the metals with the elements of carbonic acid. In these discussions a dualistic conception of acids and salts was still apparent, hydrogen and the metals being placed opposite salt-forming radicals; but the way was now opened for a unitary theory of acids and salts.

Berzelius' criticism of those attempts to explain the constitution of important classes of compounds was unusually mild; but at the same time he adhered to his dualistic view, since he laid especial weight upon the possibility of preparing the immediate constituents (of the acids),

¹ *Jahresber.*, vol. vi. p. 185; also in his *Lehrb. d. Chemie.*

² *Phil. Trans.* for 1815, p. 203.

the radicals of the hydrogen-acid theory being but seldom capable of isolation.

As his electro-chemical theory became better known, and was received with approbation, the opposing views of Davy and Dulong lost ground; it was only in the thirties that they reappeared, with fresh arguments to back them up, after which they were gradually accepted. The following observation by Daniell upon the electrolysis of salts was brought forward as an argument in their favour: "When galvanic currents are passed through different electrolytes, *e.g.* acidified water, fused chloride of lead, or a solution of sulphate of potash, amounts of hydrogen, lead, and potassium are set free at the negative pole, which stand to one another in the proportions of their chemical equivalent-numbers." This is in accordance with Faraday's "Electrolytic Law," excepting that in the case of the sulphate of potash an equivalent of hydrogen is liberated in addition to an equivalent of the base. The current therefore appears to do double work here, in spite of the law just mentioned; for, if it be assumed that the immediate constituents of one equivalent of the salt are potash and sulphuric acid, then only one equivalent of potash—as the electro-positive portion—should result, and not one of potash + one of hydrogen. But this apparent contradiction is done away with by adopting the view of Davy and Dulong, *i.e.* by assuming potassium as the positive, and the radical SO_4 (*oxy-sulphion*) as the negative constituent. The two equivalents of potash and hydrogen are then seen to be secondary products of the decomposition of one equivalent of water by the potassium originally separated at the negative pole. The conclusion drawn from this observation on the constitution of salts was then of course extended to that of acids, in which hydrogen was assumed as the one constituent, and a radical—either containing oxygen, or free from it—as the other.

The theory of the hydrogen-acids became still more clearly defined after Liebig had brought forward his

*Doctrine of the Polybasic Acids.*¹

This we shall consider here, because of its close connection with the above views of Davy and Dulong. Many chemists at that time, in especial Gay-Lussac and Gmelin, inclined to the assumption that the atoms of the various metallic oxides contained one atom of oxygen to one atom of metal, and combined with one atom of acid to form neutral salts; Berzelius too, after 1826, was of opinion that this combining proportion was the rule. But a view of such simplicity as this, according to which almost every acid was regarded as monobasic, could no longer hold its ground after Graham's² famous investigation of the phosphoric acids.³ For this chemist showed that ordinary, pyro-, and meta-phosphoric acids contained different amounts of "basic water" to 1 atom of P_2O_5 , viz. 3, 2 and 1 atoms of water, these latter being replaceable by metallic oxides. The different saturation-capacities of those acids were in this way demonstrated, being held to depend upon the amounts of basic water which entered into their constitution.

Liebig built upon the ground which Graham had prepared, and with such success that, by the aid of his own admirable and comprehensive researches upon a large number of acids, he was able firmly to establish his *theory of polybasic acids*. By his investigations on citric, tartaric, cyanuric, comenic, and meconic acids, he convinced most chemists that these resembled phosphoric acid in basicity (*i.e.*

¹ *Ann. Chem.*, vol. xxvi. p. 113.

² Thomas Graham, born in Glasgow in 1805, became in 1830 Professor of Chemistry at Anderson's College of that city, and then in 1837 at University College, London. In 1855 he resigned this post on being appointed Master of the Mint; he died in 1869. His admirable text-book, *Elements of Chemistry*, was used not only in England, but was recast and translated into German by J. Otto and H. Kolbe. Graham's originality was shown by his valuable physico-chemical investigations on the diffusion of gases, osmose, etc., which opened out new paths in the science, while at the same time he enriched general chemistry, especially inorganic, by his purely chemical work. His collected researches have been published in one large volume, entitled *Chemical and Physical Researches* (Edinburgh, 1876).

³ *Phil. Trans.* for 1833, p. 253; or *Ann. Chem.*, vol. xii. p. 1.

were polybasic). He distinctly and definitely resisted the application to them of the arbitrary tenet that the atoms of all acids are equivalent to one another, and he gave as the criterion of a polybasic acid its capability of forming compound salts with different metallic oxides (*e.g.* such a salt as $\text{PO}_4 \begin{pmatrix} \text{Na}_2 \\ \text{K} \end{pmatrix}$). Liebig was the first to distinguish between mono-, di-, and tri-basic acids.

In order to express the facts, he still made use of the definition of acids in the dualistic sense, according to which these were regarded as compounds of one atom of acid anhydride with one, two, or three atoms of water. But this remained unsatisfactory to him, since it did not allow of regarding acids and salts from a uniform standpoint. He showed with great acuteness the contradictions which were involved in the retention of this view, summing up his criticism as follows: "In order to explain one and the same phenomenon we make use of two different modes. We are obliged to ascribe to water the most various properties, calling it basic water, water of hydration, and water of crystallisation, while at the same time we see it enter into compounds in which it assumes no one of these forms. And all because we have chosen to draw a sharp line of demarcation between haloid and oxygen salts—a line not observable in the compounds themselves, seeing that in all their relations they show similar properties."

Liebig was led to the theory of hydrogen acids from grounds of probability, and still more from grounds of convenience. The sentences in which he enunciates this doctrine explain his standpoint so clearly and tersely that they must be quoted here.

"Acids are particular compounds of hydrogen, in which the latter can be replaced by metals."

"Neutral salts are those compounds of the same class in which the hydrogen is replaced by its equivalent in metal. Those substances which we at present term anhydrous acids only become, for the most part, capable of forming salts with metallic oxides after the addition of water, or they are

compounds which decompose those oxides at somewhat high temperatures.”¹

These sentences distinctly show us the influence which the accumulating observations on the substitution of hydrogen by other elements had exercised upon Liebig. This inclination of the latter to a unitary hypothesis was keenly felt by Berzelius, who to the end of his life described Liebig's theory of the polybasic acids as one which “has led to the confusion of ideas, and has stood in the way of a more perfect knowledge.” But in thus criticising views of such great importance, and which served in quite an especial degree to clear up the uncertain notions with respect to the term “equivalent,” Berzelius stood almost alone.

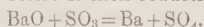
Development of the Dualistic Doctrine in the domain of Organic Chemistry—The Older Radical Theory.

During the second, and still more during the third decade of our century, organic chemistry emerged from its modest beginnings, to play an important part even so early as during the forties. It was destined to be the medium for the development of important views, and of doctrines evolved from these, thereby reacting beneficially upon its elder sister inorganic chemistry. At first it remained on pretty much the same lines as the latter, the dualistic hypothesis, which had kept its place so well with inorganic, being applied to organic compounds also. Here again Berzelius struck in as a reformer with his accustomed energy, and guided for a time the fortunes of organic chemistry. A glance at the earlier history of the latter will show us how imperfect was the knowledge of this branch of our science before the second decade of this century.

The Growth of Organic Chemistry previous to 1811.

So early as at the close of the seventeenth century mineral substances were classed apart from vegetable and

¹ Liebig here formulates sulphates as $\text{SO}_4 + \text{M}^{\text{II}}$. The decomposition of the metallic oxides to which he refers is their reduction, thus—



animal, the three being treated separately in text-books of chemistry, for instance, in that of Lemery; this division was in agreement with the classification of natural substances according to the three "kingdoms of nature," which was even then in vogue. It was from this empirical standpoint that the chemistry of organic compounds developed itself, after Lavoisier had proved qualitatively that the main constituents of these were carbon, hydrogen, oxygen, and sometimes nitrogen, occasionally together with sulphur and phosphorus. How he sought to utilise this quantitatively also, by working out a method of organic analysis, will be described under the history of analytical chemistry. He it was at all events who laid the foundation for a thorough knowledge of the subject; for, before scientific investigation in this branch could become possible, the composition of organic compounds had to be established. Notwithstanding that but very little was known at that time about the chemical constitution of these, Lavoisier tried to form an opinion on the point in particular cases. Worthy of especial mention was his view—a view which exercised great influence for a long time—that the organic acids were oxides of compound radicals, while he supposed that most of the mineral acids contained oxygen united with an element; this had indeed a distinct resemblance to the conceptions of the radical theory adopted at a later period.

While Lavoisier and other chemists after him remained true to the old classification of substances, Bergman began about the year 1780 to distinguish organic from inorganic bodies. The line, however, which remained drawn between vegetable and animal substances, in spite of the simplicity which this proposal had to recommend it, only gradually came to be removed as the knowledge increased that the same chemical compounds occurred both in vegetables and animals, as proved, *e.g.*, in the case of several fats, formic acid, benzoic acid, etc. But it was generally felt to be necessary to strictly separate organic from inorganic bodies, it being represented as an infallible distinction that the former could not be prepared directly from their elements. But even

this barrier was destined to fall before very long, and both classes of compounds to be regarded from the same stand-points.

The Position of Berzelius with regard to Organic Chemistry.

At the beginning of this century chemists of such eminence as Dalton, de Saussure, Proust, and especially Gay-Lussac and Thénard, exercised all their ingenuity in trying to work out a reliable method for determining the quantitative composition of organic compounds, but the results of their experiments only partly approximated to the truth. Before Berzelius (1811), no one had attempted to give a definite answer to the question whether the composition of organic substances was, like that of inorganic, subject to the law of multiple proportions; whether, therefore, the former were to be looked upon as chemical compounds in the sense of the atomic theory. He himself had so far elaborated a method of analysing the salts of organic acids that he was able to deduce with tolerable certainty from his results the existence of simple chemical proportions between the elementary constituents of an acid and the oxygen of the base.¹ This first successful attempt to bring organic compounds under the atomic theory, like inorganic, was followed in 1813 and 1814 by investigations² carried on with improved processes, which strengthened his conviction that the law of multiple proportions applied in the fullest degree to organic compounds also. In determining these atomic weights, he recommended, as a principle to be followed wherever possible, that the substances in question should be analysed in the form of their compounds with inorganic bodies.

But even although these researches—the first made in this direction—led to the recognition of an analogy between the two classes of substances, still Berzelius did not immediately make up his mind to look upon organic compounds as constituted exactly like inorganic (*i.e.* with respect to the arrangement of their constituent elements).

¹ *Gilbert's Annalen*, vol. xl. p. 247.

² See especially *Annals of Philosophy*, vols. iv. and v.

On the contrary, he considered it necessary to draw a sharp distinction between the latter as binary, and organic compounds as ternary and quaternary; for these, as he stated in 1813, contain more than two elements. As a consequence of this, compounds like marsh gas, cyanogen, and the hypothetical oxalic anhydride were classified as inorganic, an arrangement which was long retained (and still is, to some extent) on grounds of convenience, (Gmelin being especially strong in his recommendation of it. But this empirical separation of the two series of substances soon proved to be quite insufficient, particularly after various oils had been recognised as binary compounds of carbon and hydrogen of complicated composition.

Berzelius himself made the attempt, in his treatise¹ referred to above, to bridge over the gap between inorganic and organic bodies by assuming that the latter, like the former, are constituted binarily, but contain compound radicals in place of elements.

Gay-Lussac's beautiful researches on cyanogen had without doubt a powerful effect in reviving this idea, which had already been advanced by Lavoisier, for they proved the important fact that cyanogen, as a compound radical, can play the part of an element perfectly. This in its turn gave rise to further efforts to search for like atomic complexes (*Atomkomplexe*) in organic compounds. Gay-Lussac himself expressed the opinion that alcohol consisted of ethylene and water, and, as its vapour density proved, of equal volumes of these; while he assumed carbon and water as the immediate constituents of sugar. Hydrochloric ether was regarded by Robiquet as a compound of ethylene with hydrochloric acid, and anhydrous oxalic acid by Döbereiner as one of carbonic acid with carbonic oxide.

These efforts to look upon compound radicals as the immediate constituents of organic substances may be regarded as the beginnings of the radical theory. The above attempts at a solution were, however, not approved by Berzelius, who raised a warning voice and declared them

¹ *Versuch über die Theorie der chemischen Proportionen*, etc. (Dresden, 1820).

incompatible with the electro-chemical conceptions. In accordance with the latter, the electro-negative oxygen was placed opposite to a compound radical as the positive constituent of a compound, this showing that at that date Berzelius did not believe in radicals containing oxygen. At that time also he conceded the variability (*Veränderlichkeit*) of radicals, but went back from this later on, thereby putting an obstacle in the way of the healthy development of the radical theory.

The time for the completion of this doctrine was not yet come; but the theorising upon the proximate constituents of organic compounds was of much benefit, in that it gave a stimulus to the study of the latter. To the first task of determining their empirical composition was added the far higher one of investigating their chemical constitution by getting at the proximate constituents, as these were understood by Berzelius. The discovery of the first case of isomerism in the third decade of the century gave a powerful impetus to this, and caused the great importance of the task to be better appreciated, and a more correct idea of it to be formed. If we try to picture to ourselves the standpoint of the chemists of that day, we see how such observations of compounds having the same chemical composition, but differing totally in their properties, forced them of necessity to the conclusion that the cause of this phenomenon (termed isomerism) was to be sought for in a dissimilarity of the proximate constituents of the compounds in question. What a powerful and continually renewed charm was thereby given to the search for those different radicals of organic compounds!

*Isomerism and its Influence on the Development of
Organic Chemistry.*

Up to about the year 1820 it was considered an axiom in chemistry that substances of the same qualitative and quantitative composition must possess the same properties. Even then, it is true, cases were known which appeared to

contradict this so natural assumption, viz. the different modifications of chromic oxide and of silicic acid, and, in especial, the proof given by Berzelius of the two varieties of tin dioxide. But little weight, however, was placed upon those observations; they were simply looked upon as exceptions to the general rule, and considered merely as indicating physical differences, as in cases of dimorphism, of which a number were known.

So little were chemists prepared for the existence of substances of the same composition, but of different chemical and physical properties, that most of them considered the first observed case of isomerism in organic chemistry as an error. In 1823 Liebig had found, on comparing his analysis of silver fulminate with that of silver cyanate, which Wöhler had investigated a year before, that the results of the analyses of both salts were alike.¹ Satisfied of the correctness of his own work, he thought that Wöhler had probably made some mistake, but became convinced that this was not the case, upon repeating the investigation himself. From that date, accordingly, two compounds, which differed as widely as possible from one another chemically, were recognised as having the same composition.

While Berzelius attached full significance to the above observation, he did not immediately give in his adhesion to it,² but rather waited for further confirmation of the point; Gay-Lussac, on the other hand, felt no doubt whatever as to the correctness of the discovery, and explained the differences in the above salts by assuming a difference in the manner in which their constituent elements were combined. After Faraday's discovery,³ in 1825, of a hydrocarbon in oil gas which had the same composition as ethylene, but which showed a totally different behaviour, and after Wöhler in 1828 had obtained urea from the transformation of the similarly composed cyanate of ammonium, chemists became more conversant with

¹ *Ann. Chim. Phys.*, vol. xxiv. p. 264.

² At first Berzelius was of opinion that an error had probably been made on one side or the other (cf. *Jahresbericht*, vol. iv. p. 110; vol. v. p. 85).

³ *Annals of Philosophy*, vol. xi. pp. 44 and 95.

the existence of *isomeric compounds*. Berzelius only accepted those facts after hesitation, but ultimately convinced himself of their absolute correctness by experiments of his own. He proved that racemic acid had the same composition as tartaric,¹ and therewith proposed the term *isomeric* for those substances which, with the same chemical composition, possess different properties. The general designation *isomerism* has since then been retained. Berzelius soon saw himself necessitated to define more strictly the meaning to be attached to this word:² he distinguished between *polymerism* and *metamerism*, as special cases of isomerism, in essentially the same manner as we still do to-day.³ His power of generalising, even with but a scanty number of facts to go upon, was shown here in a very high degree.

The ideas of Berzelius as to the probable cause of isomerism in organic compounds are clearly shown in many of his utterances; in his view isomeric compounds are those in which the atoms of the elementary constituents have grouped themselves differently into compound radicals. "The isomerism of compounds in itself presupposes that the positions of the atoms in them must be different." To conclude from this sentence that Berzelius looked upon the problem of elucidating the relative positions of the atoms in space as one which was soluble, is certainly not justifiable; what he no doubt had in his mind was the determining of the mutual relations of atoms in their compounds, and, in especial, the establishing of the mode in which atoms are combined to form the proximate constituents or compound radicals of compounds. The accumulating observations of cases of isomerism quickly brought the question of chemical constitution in this sense to the stage at which an experimental solution of it was deemed possible, and this was attempted by grouping together a number of organic compounds on the basis of the hypothesis of definite common radicals. The outcome of

¹ Berzelius' *Jahresber.*, vol. xi. p. 44.

² *Ibid.*, vol. xii. p. 63.

³ Berzelius regarded the different modifications of elements as a particular case of isomerism; the designation *allotropy*, now employed for this, only dates from 1841.

this attempt was the Radical Theory, in the shaping of which Berzelius and Liebig had the greatest share. To distinguish it from the more recently revived form of views of a similar character, it is known as the *older* Radical Theory.

The older Radical Theory.

Prior to 1830, as has already been stated, efforts were not wanting to explain the constitution of particular compounds by the assumption of compound radicals. The chief incitement to those efforts lay in the proof that cyanogen acted like an element in its numerous compounds, besides being known in the free state itself. The observation that alcohol is easily transformed into ether and ethylene may have given rise to the supposition that ethylene was a constituent of both of these.

This idea, which was held by Gay-Lussac, had new life imparted to it for the time being by Dumas and Boullay's attempt¹ to generalise it by extending it to derivatives of alcohol and ether. The radical "etherin" (*Ætherin*),² C_2H_4 , was assumed by them to be present in what became afterwards known as ethyl compounds, and was compared with an inorganic compound, ammonia. Like the latter, etherin was regarded as a base, capable of forming a hydrate with water, and ethers (analogous to salts) with acids. The following table will help to explain the endeavours to establish an analogy between organic and inorganic compounds (some of the latter having, as a matter of fact, no existence):—

Etherin, C_2H_4	Ammonia, H_3N
Alcohol, $C_2H_4 + H_2O$	
Ether, $2C_2H_4 + H_2O$	
Hydrochloric ether, $C_2H_4 + HCl$	{ Chloride of ammonia, $H_3N + HCl$.
Acetic ether, $2C_2H_4 + C_8H_6O_3 + H_2O$	{ Acetate of ammonia, $2H_3N + C_8H_6O_3 + H_2O$. ³

¹ *Ann. Chim. Phys.*, vol. xxxvii. p. 15 (1838).

² The radical C_2H_4 had, at Berzelius' suggestion, received the name *Ætherin*.

³ Dumas' atomic weights, taking $H=1$, were $C=6$, and $O=16$.

This attempt, which is known under the name of the *etherin theory*, was so far the precursor of the true radical theory in that it had the comparison of organic with inorganic substances in common with the latter. In criticising it Berzelius was thoroughly justified in laying emphasis on the point that it was quite admissible to group the above compounds in tabular form alongside of one another, while at the same time he expressed the opinion that their presumed constitution was highly doubtful.

But the real development of the existing idea that organic compounds owe their characteristics to the radicals which they contain, was mainly brought about by Liebig and Wöhler's memorable research, entitled *Ueber das Radikal der Benzoesäure* ("Upon the Radical of Benzoic Acid").¹ In this they proved incontestably that in numerous transformations of oil of bitter almonds, and of chlorine and bromine compounds prepared from it, a radical of the composition $C_{14}H_{10}O_2$,² which they termed *Benzoyl*, remained unaltered. They showed by convincing experiments that this radical may be assumed as present in benzoic acid, benzoyl chloride, and bromide, benzamide, benzoic ether, and benzoyl sulphide, and that it comports itself in these compounds like an element. This piece of work was not only of profound significance for the radical theory, but it has also exercised a most powerful influence on the development of organic chemistry generally, the new methods given in it for the preparation of particular compounds having proved applicable to whole classes. The authors laid greatest stress upon the proof of a "compound element, benzoyl, in a series of organic compounds."

Berzelius was so convinced by these astonishingly clear results of the correctness of their interpretation, that he concurred enthusiastically in the assumption of the radical benzoyl;³ the facts were so strongly in its favour that he

¹ *Ann. Chem.*, vol. iii. p. 249 (1832).

² Berzelius' atomic weights were: H=1, C=12, O=16.

³ In his letter to Liebig and Wöhler (*Ann. Chem.*, vol. iii. p. 282), Berzelius proposed the name *Proïn* or *Orthrin* (from $\pi\rho\omega\iota$ and $\epsilon\rho\theta\rho\omega\varsigma$ respectively, meaning "morning blush"), because with this research a new day had dawned for organic chemistry.

felt himself compelled to give up his axiom,—that oxygen cannot be a constituent of a radical. But unfortunately this was only for a short time, as he soon reverted to the opinion that the existence of oxygenated radicals was absolutely incompatible with his electro-chemical theory.

Most chemists of that day held that the radicals which were proved to be present in several compounds were to be regarded as atomic groups capable of existing separately, and that their isolation should therefore be striven after. Although benzoyl itself had not been isolated, as little doubt was felt with respect to its separate existence as with respect to that of calcium or of nitric anhydride, neither of which had yet been obtained. The natural result of Liebig and Wöhler's investigation was a strong incitement to chemists to search for the atomic groups peculiar to different series of compounds, whose modes of formation and behaviour pointed to a probable connection between them.

The radical theory proper, in the establishing of which Berzelius and Liebig took part during the ensuing years, arose out of such endeavours. A series of organic compounds, closely related to alcohol, furnished the most suitable object for such a view, these compounds being among the most carefully investigated of organic substances. In 1833 Berzelius¹ emphasised the necessity of assuming a binary structure for all organic as for all inorganic compounds, renouncing at the same time the idea of oxygenated radicals. Benzoyl he explained as being the oxide of the complex, $C_{14}H_{16}$, the peroxide of this being anhydrous benzoic acid. Ether he regarded as the sub-oxide of ethyl, and he gave it the formula $(C_2H_5)_2O$; this last corresponded to the inorganic bases, and was combined with acids in ethers exactly as the metallic oxides were in salts. Alcohol, on the other hand, which is so nearly related to ether, was looked upon by him as the oxide of a radical C_2H_5 , a view which entirely effaced the connection between the two compounds.²

¹ *Jahresber.*, vol. xiii. p. 190 *et seq.*

² Berzelius conceived himself obliged to take this view of the atomic composition of alcohol and ether on account of their vapour densities; from these

Liebig,¹ noting this error, published in the following year his opinion that alcohol, as well as ether and its derivatives, were compounds of one and the same radical *ethyl*, to which, however, he gave the formula C_4H_{10} (in place of C_2H_5 by Berzelius). His view is apparent from the following table:—

Ether, $C_4H_{10}O$	Ethyl iodide, $C_4H_{10}I_2$
Alcohol, $C_4H_{10}O.H_2O$.	.	.	{	Nitrous ether (<i>Saltpeteräther</i>),
					$C_4H_{10}O.N_2O_3$
Ethyl chloride, $C_4H_{10}Cl_2$	Benzoic ether, $C_4H_{10}O.C_{14}H_{10}O_3$.

He accordingly designated ether as ethyl oxide, and alcohol as hydrate of ethyl oxide, comparing the former with potassic oxide, and the latter with potassic hydrate. Notwithstanding, however, his recognition of the fact that the same radical is common to both, he fell into an error which Berzelius had avoided, viz. he attributed to alcohol and the corresponding compounds twice the atomic weight that they really possess. But apart altogether from these mistakes of Liebig and Berzelius, the advantages of their *ethyl theory* were at once apparent. A broad pathway was opened out for the conception that organic compounds were constituted analogously to inorganic. Ethyl played in a large number of compounds the same part as potassium or ammonium² did in others. Liebig finally extended this comparison to mercaptan and ethyl sulphide, then just discovered. It was due in a high degree to his eloquent advocacy of the assumption of "compound elements" that the radical theory found such wide recognition.³

he deduced the correct molecular formulæ, without however being able to arrive at the true constitution of alcohol, as he did at that of ether.

¹ *Ann. Chem.*, vol. ix. p. 1, *Ueber die Konstitution des Aethers und seiner Verbindungen* ("On the Constitution of Ether and its Compounds").

² In the place of the assumption that ammonia itself is combined with acids in its salts, the view—originally held by Ampère (in 1816) and which had now the authority of Berzelius to back it—gradually spread, that in those salts ammonium, NH_4 , acts analogously to the metals.

³ We must not omit to state here that Kane, independently of Berzelius and Liebig, pointed out the analogy between a radical *Äthercum*, i.e. ethyl, which was to be assumed in ether, alcohol, etc., and the hypothetical ammonium; the paper, however, in which he expressed this view (which was published in 1833 in *The Dublin Journal of Medical and Chemical Science*, vol. ii. p. 348) remained quite unnoticed.

The influential chemists of that day held firm to their expressed opinions regarding radicals:—Dumas to the assumption that etherin was the radical of alcohol, etc.; Berzelius to the view that alcohol and ether had different constitutions, although he did not absolutely deny the admissibility of the extended ethyl theory; while Liebig remained true to the latter. He differed most from Berzelius upon the question of oxygenated radicals, which were in his opinion indispensable; thus he had no doubt that carbonic oxide was a constituent of carbonic and also of oxalic acid. But in one point those chemists were all agreed, viz. that compound radicals existed as distinct constituents in their compounds.

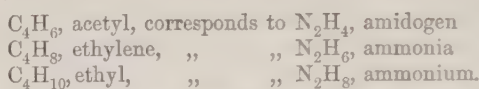
Liebig by degrees took up another and wider view of the nature of radicals than Berzelius, who inclined more and more to the opinion that they were unalterable. In Liebig, on the other hand, we get frequent glimpses of the idea that the grouping of the elements to radicals must prove of essential service to the better understanding of the modes of decomposition and formation of compounds. This conception appears to have forced itself upon him from the result of an investigation¹ which Regnault² had undertaken at his suggestion. The latter had obtained a substance of the composition $C_4H_6Cl_2$, which he termed chloro-aldehyde, by decomposing ethylene chloride with alcoholic potash. Liebig³ thereupon expressed his opinion that the radical C_4H_6 was a constituent of this chloride and of numerous other compounds; this radical he named *acetyl*, and he placed it parallel to the hypothetical *amidogen* (*Amid*), and its

¹ *Ann. Chem.*, vol. xv. p. 60.

² H. V. Regnault, who was born at Aix la Chapelle in 1810 and died at Auteuil near Paris in 1878, was a pupil of Liebig. Up to 1840 he gave his attention to organic chemistry, which he enriched by valuable work, but after that devoted himself to physico-chemical researches which will ensure him a distinguished place in the history of the science. His many-sidedness is shown in his admirable investigations on the respiration of animals, undertaken conjointly with Reiset. His *Cours Élémentaire de Chimie* 1847-49 became well known and appreciated in other countries besides France, by means of translations.

³ *Ann. Chem.*, vol. xxx. p. 229.

hydrogen compounds, ethylene and ethyl, to ammonia and ammonium, thus :—



Liebig laid especial weight upon the finding of an expression for the constitution of aldehyde and acetic acid: these he looked upon as the protoxide and hydrated oxide of the acetyl radical, and he gave them the formula $C_4H_6O \cdot H_2O$ and $C_4H_6O_3 \cdot H_2O$. This conception paved the way for the explanation of the conversion of alcohol into aldehyde and acetic acid, while at the same time it raised up doubt as to the rigid unchangeability of a radical.

The year 1837 may be looked upon as that in which the older radical theory attained to its zenith and stood out most securely, in spite of the many attacks which it had to undergo. Liebig and Dumas, who were convinced of the untenability of the etherin theory, joined together to make a thorough investigation of organic compounds with respect to the radical theory. In a paper¹ given out jointly in his own name and Liebig's, Dumas set forth his altered opinions and described the problems to be solved. Organic chemistry was regarded by both as the *Chemistry of Compound Radicals*, and was defined accordingly.² These radicals were compared with the elements, *e.g.* *ethyl*, *methyl* (whose existence in wood spirit was deduced from Dumas and Peligot's memorable research), and *amyl*³ with the metals, *acetyl* with sulphur, and so on; and their compounds with the corresponding compounds of the elements, *e.g.* ethyl sulphide ($C_2H_5)_2S$, with sulphide of potassium, K_2S , etc.⁴

¹ *Comptes Rendus*, vol. v. p. 567.

² Cf. Liebig's *Handb. d. organ. Chemie*, p. 1.

³ Cf. Cahours' investigation of fusel oil, *Ann. Chem.*, vol. xxx. p. 228.

⁴ The following quotation from the paper cited above (note 1) shows the then standpoint of Dumas and Liebig: "Organic chemistry possesses its own elements, which sometimes play the part of chlorine or oxygen, sometimes that of a metal. Cyanogen, amidogen, benzoyl, and the radicals of ammonia, of the fats, and of alcohol and its derivatives, constitute the true elements of organic nature, while the simplest constituents, such as carbon, hydrogen, oxygen, and nitrogen, only appear when the organic substance is destroyed."

The chemists of that day did not, however, remain content with simply contrasting organic with inorganic compounds as an aid to getting at their formulae; on the contrary, they applied in the happiest manner to the investigation of organic compounds the principles which they knew to hold good in inorganic chemistry, faithful to the axiom enunciated by Berzelius in 1817: "The application of what is known regarding the combination of the elements in inorganic nature, to the critical examination of their compounds in organic, is the key by which we may hope to arrive at true ideas with respect to the composition of organic substances."

As the presence of such atomic complexes in organic compounds came to be assumed with more confidence, the term *radical* became more sharply defined. Liebig himself enunciated in 1838 three characteristics by which a compound radical was distinguished. In bringing forward his view he made use of cyanogen as an instance, and spoke as follows:¹ "We term cyanogen a radical because (1) it is the unchanging constituent of a series of compounds; (2) because it is capable of replacement in these by simple substances; and (3) because, in those cases where it is combined with one element, this latter can be exchanged for its equivalent of another element." At least two of the conditions here adduced had to be fulfilled in order that an atomic complex might be stamped as a radical. The existence of these conditions, moreover, could only be established by the most minute investigation of the chemical behaviour of organic bodies. That is to say, the nature of the radicals assumed in the latter could only be arrived at from the study of their reaction- and decomposition-products.

The radical theory gave such a powerful impulse to the science that its influence, even when it fell into error, cannot be too highly prized. Chemists of the highest eminence were attracted to the task of investigating the constituents of compounds which were related to one

¹ *Ann. Chem.*, vol. xxv. p. 3.

another. Among the most fruitful of those efforts were a series of admirable researches upon the cacodyl compounds¹ by Robert Bunsen, begun in the year 1839 (*see below*).

Robert Wilhelm Bunsen, born at Göttingen on 31st March 1811, became assistant professor at the university there, then succeeded Wöhler at Cassel, and was appointed professor in the University of Marburg in 1838. His next post (only occupied for a short time) was at Breslau, after which he was called (in 1851) to Heidelberg, of whose university he remained a bright ornament till his resignation in 1889. Chemistry is indebted to him for a vast number of the most important researches in every branch of the science: his name will therefore be very often referred to in the special history of its various sections. Beginning with work in inorganic chemistry, he soon turned his attention to the organic compounds of arsenic, by investigating which he raised up a powerful support for the radical theory. His work upon gases led him to devise new methods, by sifting and combining which he created the gas analysis of to-day. The discovery of spectrum analysis by him and Kirchhoff—one of the grandest and most fruitful of the last fifty years—is fresh in every one's recollection. His labours in other branches of physical, analytical, inorganic, and mineralogical chemistry will be referred to in the detailed description of these. Throughout he has shown himself an investigator of the most marked originality and a pioneer in the science; while his career as a teacher, extending over more than half a century, has been singularly successful in its results.

Bunsen's researches on the cacodyl compounds resulted in the proof that the so-called *alkarsin*, the product of the distillation of acetate of potash with arsenious acid, contained the oxide of an arseniuretted radical $\text{As}_2\text{C}_4\text{H}_{12}$ ($\text{H}=1$, $\text{C}=12$, $\text{As}=75$), this radical remaining unchanged in a long series of reactions of that oxide, and being even itself capable of isolation. This "compound element" con-

¹ *Ann. Chem.*, vol. xxxi. p. 175; vol. xxxvii. p. 1; vol. xlii. p. 14; vol. xlv. p. 1.

taining arsenic (an unusual constituent of organic bodies) was thus shown to be a true radical.

The investigations of Gay-Lussac upon cyanogen, of Liebig and Wöhler upon benzoyl compounds, and of Bunsen upon the compounds of cacodyl, have been justly termed the three pillars of the radical theory. The assumption of radicals gained so immensely in probability from the results of those researches, that the hypothesis which lay at the root of the theory might now be regarded as well established. In any case the older radical theory formed an indispensable link in the chain of theoretical views, and marked an extraordinary advance upon the previous unconnected opinions. And even although this theory (as it then stood) exercised no very permanent effect, being soon overthrown by opposing currents, it showed itself capable of further development in a high degree. For, shortly after the catastrophe which came upon it, it was able to throw off a few restraining fetters, and start again into fresh life.

Before proceeding to describe the development of the hypotheses directed against the older radical theory, it will be convenient to give a short account here of the lives and chief labours of the three chemists who were mainly instrumental in changing the direction of organic chemistry during the third and fourth decades of this century, and who furthermore exercised a powerful influence upon our science up to a very much more recent date.

Liebig, Wöhler, and Dumas—A Survey of their more important Work.

Liebig and Wöhler, who were guided by similar scientific aims, and were at the same time close personal friends, must be spoken of together in the history of the science; the portrait of the one is incomplete unless supplemented by the characteristic features of the other. The fruit of their common labour is among the richest in the whole of chemistry.

Justus Liebig,¹ whose influence in shaping the radical theory and upon organic chemistry in general has just been touched upon, earned by his scientific work the right to be regarded as one of the most distinguished investigators of the century. Born at Darmstadt on 12th May 1803, his early years did not seem to give any special promise of the fiery spirit which he later developed, although it was not long before he felt himself drawn towards chemistry with irresistible power. He soon forsook the calling of apothecary, through which alone it was possible at that time to gain a practical knowledge of chemistry, in order to devote himself to academic studies. Relying on himself alone, he continued his early-begun investigations upon fulminate of silver, which he hoped would give him a certain definite position in science. But however independent the youth thus showed himself in this direction, he was unable to resist the influence of the natural philosophy current at that day. At a later period we find him speaking with bitterness of the two years that he had lost by it, during which time he studied under Schelling at Erlangen.² But he rescued himself from this by going in search of his science to where, at that time, it flourished most brilliantly,—to Paris, where Gay-Lussac, Thénard, Dulong, Chevreul, Vauquelin, and others were hard at work. With recommendations from Alexander von Humboldt to Gay-Lussac and other influential chemists, he recovered himself in those surroundings, and soon became closely associated with Gay-Lussac, the result of which was their important investigation of the fulminates. This piece of work paved a way for him; in 1824 he was called as professor to Giessen, where he remained for twenty-eight

¹ Cf. the Memoirs by H. Kolbe, *Journ. pr. Chem.* (2), vol. viii. p. 428; and by A. W. Hofmann, *Ber.*, vol. vi. p. 465.

² In an essay entitled *Ueber das Studium der Naturwissenschaften* ("On the Study of the Natural Sciences"), published in 1840, Liebig expressed himself as follows: "I myself spent a portion of my student days at a university where the greatest philosopher and metaphysician of the century charmed the thoughtful youth around him into admiration and imitation; who could at that time resist the contagion? I too have lived through this period—a period so rich in words and ideas and so poor in true knowledge and genuine studies; it cost me two precious years of my life."

years, but where at first he had to fight hard and continuously in order to maintain his position, his youth being a source of offence to the older professors. In 1852 he accepted a call to the University of Munich, being led thereto by the desire to throw off the fatigues of laboratory teaching and to live all the more ardently for research. His magnificent labours were brought to a close there by death, on 18th April 1873, but the genius which inspired them, and which had acted with such powerful effect upon his contemporaries, continued to influence mankind. How powerful an influence he exercised—as shown in his greatness as a teacher, in the transformation of whole branches of the science, and in the setting aside of firmly rooted views which in his opinion were erroneous—we shall now attempt shortly to describe.

As a teacher Liebig stands almost alone. Berzelius, the great master, only drew around himself pupils who had already a considerable knowledge of the subject, and worked (directly) in a comparatively narrow circle. Liebig, on the other hand, founded a real school of chemistry, by sparing no pains in instructing his pupils individually from the commencement of their course of study. He was the first to give systematic teaching in chemistry, for up to that time there was no laboratory in existence which was solely devoted to that purpose. And he was also the first to recognise the necessity for having chemical institutes which should further not merely the science itself, but also the many other branches dependent upon it. His laboratory in Giessen served as a pattern upon which numerous others were in the course of years modelled, at first slowly but afterwards in more rapid succession. By the charm of his own personality Liebig stimulated his pupils and animated them with enthusiasm, especially when the solution of a scientific question came up. Kolbe has described for us his unique character as a teacher in the following striking sentences:—"Liebig was not a teacher in the ordinary sense of the word. Scientifically productive himself in an unusual degree, and rich in chemical ideas, he imparted the latter to his more advanced pupils, to be put

by them to experimental proof; he thus brought his pupils gradually to think for themselves, besides showing and explaining to them the methods by which chemical problems might be solved experimentally."

In addition to this Liebig gave a new form and meaning to his experimental lectures, so that here also he set up a standard. His pupils were legion; many of them afterwards spread abroad the doctrines of their master in universities, polytechnic institutes, technical schools, etc. Out of a long list of them which might be given here, the following may be mentioned:—A. W. Hofmann, H. Kopp, Regnault, Strecker, Fresenius, Will, Fehling, Henneberg, Schlossberger, Rochleder, Schlieper, Scherer, Redtenbacher, v. Bibra, Varrentrapp, Playfair, Muspratt, Stenhouse, Brodie, Gerhardt, Williamson, Wurtz, Frankland, and Volhard.

The mental vigour which was shown in the results of Liebig's teaching is also seen in his literary activity, which awakens a feeling of astonishment by its many-sidedness, embracing as it does the most various branches of the science. Throughout it all we see the capacity of the true investigator to state points correctly and clearly, to grasp the connection between different processes distinctly, and to draw able and ingenious conclusions. These merits impart to Liebig's writings a great and ever-renewed charm. His numerous experimental researches, together with the conjoint ones with Wohler, were mostly published in the *Annalen*,¹ which he began to give out in 1832. His extended investigations in physiological chemistry, which were begun in 1837, led him on to the grand achievement of setting forth the applications of chemistry to agriculture, physiology, and pathology in three separate works.² In these he combated

¹ Till 1840 this journal was termed *Annalen der Pharmacie*, and after that date (with Wöhler as joint editor) *Annalen der Chemie und Pharmacie*.

² *Die Chemie in ihrer Anwendung auf Agrikultur und Physiologie*, 1840 ("Chemistry in its Application to Agriculture and Physiology," 1840); *Die Thierchemie oder organische Chemie in ihrer Anwendung auf Physiologie und Pathologie*, 1842 ("Animal or Organic Chemistry in its Application to Physiology and Pathology," 1842); *Der chemische Prozess der Ernährung der Vegetabilien und die Naturgesetze des Feldbaues*, 1862 ("The Chemical Processes in the Nutrition of Vegetables, and the Natural Laws of Tillage," 1862).

the erroneous doctrines which were held with regard to the nutrition of plants and animals, basing his arguments upon exact experiments. Notwithstanding the great excitement which those publications produced, Liebig found leisure to write his *Chemische Briefe* ("Chemical Letters," 1844), by which he proved that chemistry might be treated popularly, and yet at the same time scientifically. It is almost inconceivable how he still found time remaining to devote to the *Handwörterbuch der reinen und angewandten Chemie* ("Dictionary of Pure and Applied Chemistry"), founded by Wöhler, Poggendorff, and himself, and, after the death of Berzelius in 1848, to the *Jahresbericht der Chemie*. In addition to all these there are still to be mentioned his occasional papers,¹ some of which exercised a powerful effect; this applied in an especial degree to the two essays upon the state of chemistry in Austria and Prussia. In these, as in other papers devoted to questions of theoretical chemistry (*e.g.* in his writings directed against the views of Dumas, and of Laurent and Gerhardt), is shown the sparkling critical vein of this gifted man, who, from his rectitude and love of truth, never palliated what he felt to be erroneous or insincere. Occasionally Liebig may have gone too far in his critical utterances upon particular men; but the mainspring of his decided attitude with respect to them was always the love of the science and of truth.

As an investigator Liebig shows all his individuality. To organic chemistry he had devoted the fullest attention from the very beginning, without, however, neglecting any important part of inorganic. His very first work—that upon the fulminates—led to valuable results; for, through it, the isomerism of cyanic and fulminic acids became recognised, a new field for investigation being thereby opened up. Another result of this laborious research upon these easily decomposable substances was the perfecting of organic analysis, to which Liebig gave its present form. By the aid

¹ These were published by M. Carriere under the title *Reden und Abhandlungen* ("Speeches and Essays"), by Justus von Liebig. (In 1845 he was made a baron by the Grand Duke of Hesse.)

of methods improved by himself, he established the composition of numerous organic compounds, in especial, of various acids. His work upon these last led him to a clear conception of the term *basicity*; from this he developed his doctrine of polybasic acids (already touched upon), doing more to clear up the points involved here than any other chemist before him.

His previous admirable researches upon compounds closely related to alcohol and acetic acid, *e.g.* ethyl-sulphuric acid, aldehyde, acetal, chloral, etc., rendered him specially capable of developing the radical theory and infusing fresh life into it. The work which he did upon sulphocyanogen compounds and upon the decomposition products of ammonium sulphocyanide showed him as a brilliant experimenter in all his many-sidedness.

But his most remarkable achievements were the researches carried out conjointly with Wöhler, which show them both in their full freshness and power, and which will long continue to call forth the admiration of chemists. Wöhler's work upon cyanic acid and Liebig's upon the fulminates drew them together; their friendship is beautifully shown by the investigations which they undertook in common, during which each animated the other, while striving at the same time to do his best himself.¹ And how strikingly was the one man the complement of the other! Liebig—fiery, restless, and always advancing, able to utilise his rich experiences gained in the preparation and analysis of organic compounds for overcoming difficulties. Wöhler, on the other hand, quiet, almost prosaic, but not less conscious of his aim than Liebig himself, exercising patience in clearing up obscure points to which too little attention had been paid. The memorable research upon the radical of benzoic acid has been already detailed. The investigations upon amygdalin cleared up the difficult point as to how bitter almond oil was formed, and those upon uric acid, published in the same year (1837), enriched organic chemistry to an undreamt-of extent with

¹ Cf. the letters of both quoted in A. W. Hofmann's Memoir of Wöhler, *Ber.*, vol. xv. p. 3127 *et seq.*

a wealth of the most remarkable compounds,—compounds which have quite recently proved objects of the greatest interest to chemists. We are indeed not wrong in asserting that the organic chemistry of to-day is grounded mainly upon the pioneering labours of Liebig, and of Liebig and Wöhler together.

In addition to all this, inorganic chemistry was anything but neglected by Liebig, who enriched it by valuable observations on the most various subjects; we have only to recall his work upon the compounds of alumina, antimony, and silicic acid, and many analytical methods which he worked out, *e.g.* the separation of nickel from cobalt. The results obtained by him in the laboratory were frequently of great service for technical chemistry; for instance, the improved preparation of cyanide of potash for the galvano-plastic process, and the reduction of a solution of silver by aldehyde for the production of mirrors.

Liebig's share in the development of organic chemistry, especially with regard to the views which had come to be accepted in it, became less towards the end of the thirties, as from that time he gave all his energies to the solution of a great question which had only an indirect bearing upon chemistry. The nutrition of plants and animals, the transformations of matter in animated nature—these were the grand problems which he strove to solve by experimental researches in an entirely new direction. The influences which emanated from him, the setting right of erroneous views, the ingenious interpretation of natural processes investigated by himself and his pupils, and the stimulus which invariably accompanied his labours and the deductions drawn from them,—all these can but be referred to here. The most important results of those researches will be spoken of under the history of physiological chemistry. Liebig's experiments on the nutrition of animals led him to distinguish clearly between nutrient substances among themselves, and between these and other substances which, though not directly nutrient, bring about metabolic changes in the organism.¹ By getting at the relative nutritive values of these

¹ “. . . Unterscheidung der Nährstoffe unter sich und von den Genussmitteln.”

materials he was enabled to introduce improved systems of feeding, and so to further the laws of health; we have only to recall here his extract of meat and his "children's food."² He was thus in this respect a general benefactor of mankind.

We may close this attempt at depicting within narrow limits the scientific achievements of Liebig with the following eloquent words of A. W. Hofmann:—"If we sum up in our minds all that Liebig did—in industries, in agriculture, and in the laws of health—for the good of mankind, we may confidently assert that no other man of learning, in his course through the world, has ever left a more valuable legacy behind him."

Friedrich Wöhler,¹ whose work blended so happily with that of Liebig, also proved himself by his own individual researches a master in his science. By far the greater portion of his work lay in the domain of inorganic chemistry, which he furthered in a remarkable degree.

Wöhler's life may be sketched in a few sentences. Born in the village of Eschersheim, near Frankfurt on the Main, in the year 1800, he received in the latter city a splendid education at the hands of such eminent teachers as Karl Ritter and F. C. Schlosser. There, too, he first made acquaintance with chemistry, to which he remained faithful, thanks to the influence of L. Gmelin, notwithstanding that he went through the medical curriculum at Marburg and Heidelberg. It was Gmelin, too, who recommended the young doctor of medicine to Berzelius, the latter receiving him with open arms. After barely a year's stay in Stockholm,—a year, however, rich in experiences and un-effaceable impressions, and of which he himself has given us such a clear picture²—Wöhler returned to Germany in the autumn of 1824, to become shortly afterwards a teacher in the Technical School (*Gewerbeschule*) at Berlin. In 1831 he had to leave the pleasant and stimulating society of his friends there (among whom we may mention Mitscherlich, the brothers Rose, and Poggendorff), to fill the post of pro-

¹ Cf. A. W. Hofmann's Memoir of Wöhler, *Ber.*, vol. xv. p. 3127 *et seq.*

² *Ber.*, vol. viii. p. 838 *et seq.*

fessor in the newly-founded Higher Technical School at Cassel; while in 1836 he accepted a call to Göttingen, where, till his death on 23d September 1882, he remained a bright ornament of the *Georgia-Augusta* (the university of that town).

Wöhler's influence as a teacher, especially after his removal to Göttingen, may be described as enormous. Like his friend Liebig, he laid the greatest weight upon a thorough grounding in the rudiments of chemistry. The advantages which he had gained from his analytical work under Berzelius he now imparted to his pupils. Out of a long list of these, a few may be named who themselves subsequently continued to teach in the spirit of their master:—Th. Scherer, H. Kolbe, Henneberg, Knop, Städeler, Geuther, Limpricht, Fittig, Beilstein, Hübner, and Zöller.

Wöhler was especially active in a literary sense during the earlier portion of his life, as is shown by his co-operation in the *Dictionary of Chemistry*, already mentioned, and his translations of the *Text-Book* and *Annual Reports* (*Jahresberichte*) of Berzelius. The first edition of his *Grundriss der anorganischen Chemie* ("Outlines of Inorganic Chemistry") occupying about 150 pages, appeared in 1831, the *Organic* following in 1840; both of these went through numerous editions.¹ His results in the investigation of minerals he collected together in 1853 in the valuable work, *Praktische Übungen in der chemischen Analyse* ("Practical Exercises in Chemical Analysis)."² His experimental researches—most of which he published in the *Annalen der Chemie*, but some of the earlier ones in *Poggendorff's* and in *Gilbert's Annalen*—embrace almost every branch of inorganic chemistry. Some of them also led to the opening up of important branches of organic, e.g. his splendid work upon cyanic acid and its salts, the discovery of urea, and also the investigations carried on along with

¹ From its sixth edition the *Organic Chemistry* has been edited by Rudolf Fittig; the fourteenth and fifteenth (the last editions of the *Inorganic*) were given out by H. Kopp.

² The second edition appeared in 1861 under the title *Die Mineralanalyse in Beispielen* ("The Analysis of Minerals, illustrated by Examples").

Liebig. In all of them, as also in his later labours, his remarkable gifts as an observer are apparent.

We cannot enter into details at this point either with regard to his work in analytical chemistry, which he enriched by admirable methods, or to that in inorganic. But a few investigations in the latter branch must just be mentioned, viz. those upon aluminium, boron, silicon, and titanium, and their remarkable compounds, by which the resemblance between the two last-named elements and carbon was clearly brought to light.

The papers in which Wöhler describes the results of his experiments are written in a clear, forcible, and simple manner, and attract our attention not merely by those characteristics—nowadays somewhat rare,—but above all by the depth of their contents. That he had plenty of humour at command is proved by his letters to Liebig, and by the delicious satires¹ which he wrote when Dumas allowed himself to be carried too far by the deductions which he drew from the doctrine of substitution. Wöhler never rushed of his own accord into discussions upon important questions of theoretical chemistry,—a trait characteristic of his quiet disposition, and one which distinguishes him from Liebig, the born reformer.

Jean Baptiste André Dumas,² who was born at Alais in 1800, and died at Cannes in 1884, rendered to his science extraordinary services, to which we shall frequently have occasion to refer. Beginning life as apprentice to an apothecary in his native town, he found this calling un congenial, and set out on foot for Geneva in the autumn of 1816. Coming in contact there with such distinguished men as Pictet, Décaudolle, de la Rive, and others, he was stimulated to scientific researches which quickly attracted the attention of the *savants* just named. He made himself known particularly by the active part which he took in the physiologico-chemical investigations of Prévost. With the versatility which distinguished him, he soon began to

¹ *Ann. Chem.*, vol. xxxiii. p. 309 (see also *below*, p. 265).

² Cf. A. W. Hofmann's *Memoir*, *Ber.*, vol. xvii. *Ref.* p. 629 *et seq.*

take up problems in organic as well as in physical chemistry. In 1823, acting on A. v. Humboldt's advice, Dumas betook himself to Paris, finding there the most friendly reception at the hands of the eminent chemists of that city. At Paris he spent the rest of his life, filling various posts as a teacher and also other offices; he lectured with striking effect at the *Athenaeum*, the *École Centrale des Arts et Manufactures*, the *Sorbonne*, and the polytechnic and medical schools.

No laboratory having been placed at his disposal, he established one at his own expense in 1832. After the year 1848 Dumas was frequently called into the public service, being for a long time minister, besides having to fill other offices, so that his work as a teacher was often interrupted. The keen interest which he felt in public affairs was shown in many cases by his active co-operation, *e.g.* in furnishing Paris with a water supply, and in devising means to remedy the diseases of the silkworm and vine, etc. In 1868 he was further nominated permanent secretary of the Academy, of which he had long been a member.

We have still to make mention of the more important of Dumas' literary labours. The first of the larger works by which he became known was his *Traité de Chimie appliquée aux Arts* (1828); in its treatment of the matter, and especially its arrangement, this remained a model for many subsequent text-books on technology. The whole individuality of the man comes out in his *Leçons sur la Philosophie Chimique* (published in 1837 by Bineau from Dumas' lectures), in which he treats the development of chemical theories with great clearness and with a rare charm of style; this work, however, cannot be regarded as a strictly historical one. The numerous panegyrics which Dumas delivered are in their form, down to the minutest detail, carefully elaborated works of art; among them may be mentioned those upon Pélouze, Balard, Regnault, and Faraday.

The *Essai de Statique Chimique des Êtres Organisés*, par

MM. Dumas et Boussingault (1841), became especially well known; in this the life of plants and animals, and, in particular, the processes of metabolism, were treated from the chemical point of view. The opinions expressed here were in part instigated by the pioneering work of Liebig, whose influence, however, was not sufficiently recognised by the authors, so that he felt himself called upon to draw attention to his perfectly justifiable claims in very distinct language.¹ A debt of gratitude is due to Dumas for the pious service which he rendered in editing the reissue of Lavoisier's works.²

Most of the numerous experimental researches which we owe to Dumas were published by him in the *Annales de Chimie et de Physique*, of which he was one of the editors after 1840. In recalling his most important and productive labours, emphasis must be laid upon the great service which he rendered in working out various methods of general application. His mode of determining vapour densities and that of estimating nitrogen have found universal appreciation. His admirable investigations in organic chemistry shed a brilliant light over wide branches of it, and guided many chemists for a time as to the direction in which they should work. Mention must be made too of his conjoint researches with Peligot³ upon wood spirit and upon aethyl (from spermaceti),—compounds whose analogy to alcohol he proved; and of his discovery and investigation

¹ *Ann. Chem.*, vol. xli. p. 351. In this as well as in other instances Dumas unfortunately did not show in a very favourable light. The historian is bound to notice such facts, since they cannot be erased from the scientific character of so eminent an investigator. Liebig criticised those peculiarities of Dumas with great severity (cf. *Ann. Chem.*, vol. ix. pp. 47, 129; also Kolbe's claim of priority, *Journ. pr. Chem.* (2), vol. xvi. p. 30). Such occurrences are, to quote Liebig, "black leaves in the book of chemical history, black, because they absorb the rays of light without thereby becoming luminous themselves." Dumas was unable to disprove or even to minimise the heavy charges which Liebig brought against him.

² Cf. p. 152, note.

³ E. M. Péligot, born in 1811, was for a long time Professor of Chemistry at the *Conservatoire des Arts et Métiers*, and distinguished himself by much admirable work in inorganic, organic, and technical chemistry; he died in April 1890.

of trichloroacetic acid, which crowned the edifice of the substitution theory. The general character of his work naturally led Dumas to take an active share in the discussion of problems in theoretical chemistry. His participation in the question of the values of the atomic weights has been already noticed. The determinations which (partly in conjunction with Stas) he made of the atomic weights of carbon, oxygen, and other elements deserve to be recorded as experimental work carried out with the utmost care and circumspection.

Apart from the shadow thrown upon Dumas' achievements by some of the incidents in his scientific life, his services will long continue to excite the highest admiration as evidences of a powerful and comprehensive mind. The immense influence which he exercised upon the form assumed by organic chemistry, and, in particular, upon the development of general views opposed to dualism, will be detailed in the following section.

*The Development of Unitary Views in Organic
Chemistry.—Substitution Theories.*

At the time when Dumas brought forward his own as well as previous observations upon the substitution of hydrogen by chlorine and other elements as a basis for theoretical statements, the electro-chemical doctrine of Berzelius, and the radical theory which fitted in with it, were in high repute. The idea (deduced as it was from numerous facts) that electro-positive elements like hydrogen could be replaced by electro-negative ones like chlorine, oxygen, and others, was bound to become a stumbling-block for the dualistic hypothesis, which could not after this be maintained in its integrity. The various attempts to explain the phenomena of substitution from general standpoints, which now fall to be detailed, were at the same time the significant utterances of aspiring unitarism against the binary view.

One has to recall to mind that, according to the position

of Berzelius' dualistic doctrine at that time, the radicals were looked upon as unalterable atomic-complexes. The consequence of the electro-chemical view was the assumption that negative elements like chlorine, bromine, and oxygen could not enter into the composition of a radical. That the observations on the substitution of hydrogen atoms in organic compounds by those other elements was in direct contradiction to this assumption, appears to us now self-evident.

Dumas' Laws of Substitution.

Some isolated facts, which proved that a substitution of this kind could go on among the elements, were already known when Dumas turned his whole attention to the subject. Thus Gay-Lussac had established the formation of cyanogen chloride from hydrocyanic acid, Faraday that of sesquichloride of carbon (C_2Cl_6) from ethylene chloride, and Liebig and Wöhler the conversion of bitter almond oil into benzoyl chloride. It had not escaped those chemists that, when the above compounds were subjected to the action of chlorine, an amount of hydrogen, equivalent to the chlorine which entered into them, was separated: indeed, the opinion was expressed (by some, if not all, of them) that the one element had replaced the other.

In the year 1834 Dumas,¹ *à propos* of an investigation on the mutual action between chlorine and oil of turpentine, but more especially of his work upon the production of chloral from alcohol, condensed into two empirical rules the facts with regard to substitution, for which he proposed the designation *metalepsy* (i.e. exchange, *μετάληψις*). These were not intended to comprise a theory of substitution, as his first utterances on the subject show, but only to give expression to the facts. They were as follows:—

“When a compound containing hydrogen is exposed to the dehydrogenising action of chlorine, bromine, or iodine, it takes up an equal volume of chlorine, bromine, etc., for each atom of hydrogen that it loses.

¹ Cf. *Ann. Chim. Phys.* (2), vol. lvi. pp. 113, 140.

"If the compound contains water, it loses the hydrogen of this without replacement."

The second of these rules was deduced from the transformation of alcohol into chloral, and was thus intended to explain the mode of formation of the latter, and at the same time to support Dumas' view of the constitution of alcohol, the latter being regarded by him as a compound of ethylene and water.

Dumas soon extended his statement to one of great significance, viz. that in chemical reactions generally an exchange of equivalents of one element for equivalents of others takes place. It was from this standpoint that he regarded the oxidation of alcohol to acetic acid, and that of bitter almond oil to benzoic acid, etc. etc., and he emphasised the point that each atom of hydrogen was here replaced by half an atom of oxygen. Those views, which gave evidence of great clearness of vision, were, however, obscured by certain additions which could not fail to create confusion with regard to the constitution of the compounds in question; thus, to give one instance only, formic acid was looked upon as a "metaleptic product" of alcohol.

Laurent's Substitution or Nucleus Theory.

Dumas limited himself at that time (1835) to condensing the known facts into the two above-mentioned laws. But his countryman Laurent went further, in that he took into consideration the nature of the compounds produced by substitution, and compared them with the original ones. He was thus led to the proposition¹ that the structure and chemical character of organic compounds are not materially altered by the entrance of chlorine and the separation of hydrogen. This law, when conjoined with the view that chlorine takes over the rôle of the substituted hydrogen, is the kernel of the Substitution Theory proper, of which Laurent must be regarded as the author; for Dumas denied at that time the

¹ Laurent frequently enunciated this (cf. *Ann. Chim. Phys.* (2), vol. lx. p. 223; vol. lxi. p. 125; vol. lxvi. p. 326).

analogy between substitution derivatives and the original compounds, and in reply to Berzelius, who attacked him for this assumption, threw the responsibility for it upon Laurent.¹

The latter then strove to erect a system by developing the above doctrine, the result of his efforts being the so-called Nucleus Theory,² which was published in the year 1836; a short account of this must be given here, even although it never met with very hearty approval.³ According to Laurent, organic compounds contained nuclei (*radicaux*), and he distinguished between original nuclei (*radicaux fondamentaux*), composed of carbon and hydrogen in simple atomic proportions, and derived nuclei (*radicaux dérivés*), which were produced from the first-named either through the substitution of hydrogen by other elements or by the taking up of additional atoms. He further stated that compound radicals like amidogen or nitroxyl might substitute in place of elements. This attempted classification of organic compounds, under the name of the nucleus theory, shows a distinct connection with the radical theory; but the one-sided view of the latter—that the radicals were unalterable—has here disappeared. While this change in principle marks an advance, the giving up of the relation between organic and inorganic compounds was undoubtedly a great defect, since it involved the loss of a support indispensable for a natural classification of organic substances.

It was not difficult for the chief exponents of the radical doctrine to prove the insufficient basis of the nucleus theory, the more so that Laurent laid himself open to criticism not merely as a theoriser but also as an experimenter. His work was severely criticised by Liebig, who came to the conclusion that Laurent's theory was unscientific and pernicious. Berzelius likewise raised his voice energetically against it, and indeed went so far as to say that he con-

¹ *Comptes Rendus*, vol. vi. pp. 647, 695. Laurent stood up for his own view (*Ann. Chim. Phys.* (2), vol. lxvii. p. 303).

² Cf. *Ann. Chim. Phys.* (2), vol. lxi. p. 125.

³ L. Gmelin did, it is true, make use of the subdivision of organic compounds, according to different nuclei, as a basis in his well-known text-book.

sidered a detailed criticism of it superfluous. But, as a matter of fact, Laurent was too much depreciated from this side; for, however much we may dissent from many of his untenable speculations, his effort to classify organic compounds on uniform principles, and to show their connection with one another, was not without merit. In addition to this he had effectively aided in overthrowing the dogma of the unchangeability of radicals. And, finally, we are indebted to him for the proof that Dumas' empirical rules of substitution are by no means always applicable.

Before Laurent, in conjunction with Gerhardt, had again brought forward his ideas in a more perfect form, Dumas¹ entered the lists to do battle against the radical theory, and, with this, against the dualistic idea in general. His beautiful discovery of "chloracetic acid" afforded him the immediate occasion for this, and he now gave in his adhesion to Laurent's opinions, which formerly he would have nothing to do with. The substituting atoms, *e.g.* the halogens, take up the rôle of the expelled hydrogen atoms, and the resulting halogen compounds must therefore show an analogy to the original ones,—this was for Dumas the clear result of his work upon trichloroacetic acid; and he drew the same conclusion from the similar relations existing between aldehyde and chloral. To put his ideas into a more permanent form, he referred such related compounds to definite types, from which they were derivable.

Dumas' Type Theory.

This effort, which reminds us strongly of Laurent's nucleus theory (since in this case also whole series of compounds were referred to fixed atomic complexes), bears in the history of chemistry the name of the Older Type Theory, to distinguish it from the newer one of Laurent and Gerhardt. Dumas was led to establish his theory of types² from the behaviour of trichloroacetic acid, as observed by him—

¹ *Ann. Chim. Phys.* (2), vol. lxxiii. p. 73 *et seq.*

² *Ann. Chim.*, vol. xxxiii. p. 259; cf. also M. Berthelot's recent work, *Introduction à l'Étude de la Chimie des Anciens et du Moyen Âge* (1889).

self; he laid stress upon the fact that, in spite of the entrance of six atoms of chlorine in place of six atoms of hydrogen,¹ the character of this derivative remained essentially the same as that of acetic acid itself. Both compounds are monobasic acids, and both yield products of analogous composition with alkalis. From all this he concluded that "there are in organic chemistry certain types which remain unchanged, even when their hydrogen is replaced by an equal volume of chlorine, bromine, or iodine." Acetic and trichloroacetic acids, aldehyde and chloral, marsh gas and chloroform, belong severally to the same chemical types. One such type embraced, according to Dumas, compounds which contained the same number of equivalents combined in a like manner, and whose properties were in the main similar. We see here that the mutual relations of compounds belonging to one chemical type are the same as those which Laurent assumed between his original and derived nuclei.

But the term "chemical type" did not satisfy Dumas; he allowed it to merge into that of "mechanical type,"² this latter comprising all compounds which might be supposed to be formed from one another by substitution, even if they differed totally in properties. Acting on this idea, Dumas quite rightly classified alcohol and acetic acid under the same mechanical type; but, on the other hand, he brought together compounds which had no sort of connection with one another, *e.g.* formic acid and methyl ether. The ultimate result was that an empty scheme of formulation carried the day over what was really good in this doctrine—a doctrine developed from Laurent's nucleus theory. The endeavour to arrange organic compounds upon certain types outweighed and pushed aside the higher problems which Berzelius had sketched out for chemical science. The idea of definite atomic complexes or radicals, which was meant to pave the way for a knowledge of the chemical

¹ Dumas assigned to acetic acid the formula $C_4H_5O_4$, and to (tri)chloroacetic acid that of $C_4H_2Cl_6O_4$.

² Regnault had already (in 1838) spoken in a similar sense of *molecular types*, which remain unchanged in chemical reactions.

constitution of compounds, was superseded by the setting up of mechanical types, and thereby the link intended to connect organic with inorganic compounds was completely snapped.

This complete relinquishing of the principles set up by Berzelius, and found by him to be so serviceable, could not fail to arouse his liveliest opposition. Dumas had characterised Berzelius' electro-chemical doctrine as erroneous; the unitary conception was to step into the place of the dualistic which the latter theory involved. *Every chemical compound forms a complete whole, and cannot therefore consist of two parts. Its chemical character is dependent primarily upon the arrangement and number of the atoms, and in a lesser degree upon their chemical nature.* These propositions of Dumas stood in the sharpest opposition to the doctrine of Berzelius; they proclaimed a one-sided unitarism, which was therefore combated by Berzelius with every force at his command.

The Overthrow of Berzelius' Dualistic Doctrine.

Dumas did not scruple to say plainly that the dualistic doctrine was harmful and retarded the development of organic chemistry, and he made every effort to set it aside and to replace it by the unitary theory. His attack upon Berzelius' doctrine (at that time held in high repute by most chemists) was vigorously answered both by the latter and by Liebig. Liebig¹ indeed admitted many points which were disputed by Berzelius, *e.g.* the fact of substitution, but he protested against Dumas' wide extension of this principle (of substitution). The assertion of the latter that every element of a compound might be replaced by another, and yet the type be retained, was characterised by Liebig as entirely unproven, and met with an ironical rejoinder.² Berzelius, who saw his whole system based upon the electro-chemical theory threatened, directed his criticism in the

¹ *Ann. Chem.*, vol. xxxiii. p. 301.

² Cf. *Ann. Chem.*, vol. xxxiii. p. 308. The satirical letter given here was composed by Wöhler and published by Liebig.

Jahresberichten for 1838 and the next five years or so against the theory of types. In opposition to Dumas' unitary view he set up, as sharply as it was possible to do, the electro-chemical and therefore dualistic as the fundamental principle; he adhered indeed essentially to his former standpoint, according to which electro-negative elements could in no case enter into the composition of radicals.

Berzelius sought to get over the difficulties which the substitution of hydrogen by chlorine and other elements involved, by arguing that compounds formed in this manner must have a constitution different from that of the original ones. But here he entered upon dangerous ground, and was thereby led, prudent investigator as he was, into the most utter contradictions of the principles which he had formerly held to be inviolable.

Berzelius first expressed himself upon the composition of acetic and trichloroacetic acids. While the former (*i.e.* the anhydrous acid)¹ was regarded by him as the oxide of the radical acetyl, and given the formula $C_4H_3 + O$, he looked upon trichloroacetic acid as a so-called "copulated compound" or "conjugate compound" (*gepuartte Verbindungen*²) of quite different constitution, *viz.* as a chloride of carbon copulated with oxalic acid, of the formula $C_2Cl_3 + C_2O_3$. But he could not at that time make up his mind to follow this to its logical conclusion, and to ascribe to acetic acid an analogous composition (*i.e.* to write it down as methyl copulated with oxalic acid), manifestly from the apprehension that he would in so doing surrender a principle of his electro-chemical doctrine. He attempted similarly to explain the constitution of other chlorine organic derivatives, by assuming copula (*Paarlinge*) containing chlorine, with the result that

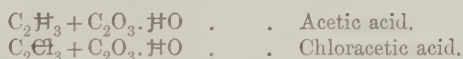
¹ Berzelius formulated acetic acid as hydrate, $C_4H_3.O_3 + H_2O$, *i.e.* as a compound of the anhydride (at that time unknown) with water.

² The idea that certain organic compounds are copulated or conjugated (*gepuart*) was definitely expressed for the first time in one of the earliest of Gerhardt's papers (*Ann. Chim. Phys.* (2), vol. lxxii. p. 184). In this paper he used the word copulation (*accouplement*) to signify the combination of organic substances with inorganic. The one portion of such compounds he termed the copula (*copule*), *e.g.* the organic substance which is copulated with an inorganic acid.

a different rational formula was assigned to the mother substance from that given to its derivatives.

These unfortunate attempts to explain by the speculative method the constitution of chemical compounds, that problem which, in his own opinion, was the most important in the science, led Berzelius completely astray. In order to carry through his doctrine of copulae, he had to assume arbitrary radicals in organic compounds, without being able to adduce the least evidence in favour of such assumptions. Above all, he did not see what these really led to, for he overlooked the fact that his chlorinated copulae could only be formed by the substitution of the hydrogen atoms of the radical by chlorine.

Melsens'¹ important observation, made in the year 1842, that chloracetic acid is reconverted into acetic by the action of potassium amalgam, convinced Berzelius² that his view of the two acids having different constitutions was no longer tenable. He therefore decided to regard acetic acid in the same way as its chlorine derivative, *i.e.* as a copulated oxalic acid with the copula C_2H_3 , formulating the two compounds thus—



But in doing this he made the important admission of the substitution of hydrogen by chlorine in the copula. And even although he did emphasise the point that the latter exercised no particular effect upon the compound to which it belonged, he none the less recognised hereby a fundamental principle of the doctrine of substitution.

But, notwithstanding this admission, Berzelius remained to the end of his life an opponent of the theory of types, and tried to uphold the dualistic view by every means in his power. He had to undergo the pain, however, of finding his hitherto faithful adherents no longer able to follow him in this, and indeed of hearing them dissent publicly from his treatment of the question as to how the constitution

¹ *Ann. Chim. Phys.* (3), vol. x. p. 223.

² *Lehrb. d. Chemie* (fifth edition), vol. i. p. 709.

of organic compounds was to be explained. Liebig, who had already before this taken the facts of substitution into account,¹ declared openly against Berzelius' artificial attempts at explanation,² the more so since the chlorine and bromine derivatives of aniline had been investigated in the Giessen laboratory by A. W. Hofmann, and had been accepted as evidence that the chemical character of a compound depends to a not inconsiderable extent upon the arrangement of its atoms. Liebig therefore turned himself to the unitary theory. The following words³ show us the attitude taken up by Liebig, and we may be sure by others also, towards Berzelius at that time: "During the last years (of his life) Berzelius ceased to take an experimental share in the solution of the problems of the time, and turned the whole force of his mind to theoretical speculations: but these, not being the result of his own observations or supported by them, found no echo or approval in the science."

This much is certain, that, by carrying his speculations too far, Berzelius had not only shaken the edifice of his own doctrine, but had also greatly injured the radical theory, more particularly by the heaping up of unproven hypotheses. His opponents went so far as to assert that he had by his arbitrary assumptions "made a theory regarding substances which had no existence" in organic chemistry. It almost seemed as if his whole system was doomed to fall. One result of all this was that many chemists became visibly discouraged, and, holding all speculation as dangerous, either applied themselves to the empirical side of the science, or turned to other subjects. And yet, in spite of the slight regard in which the radical theory was held in many quarters, it soon became evident that, for the investigation of chemical constitution, the assumption of radicals, which had displaced the theory of types, could not be dispensed with. In the course of the forties a fusion of the radical theory with the older doctrine of types took place on the

¹ *Ann. Chem.*, vol. xxxi. p. 119; vol. xxxii. p. 72.

² *Ibid.*, vol. i. p. 295 ("Berzelius und die Probabilitätstheorien").

³ *Ibid.*, vol. i. p. 297.

unitary side; from the joint work of Laurent and Gerhardt there resulted the new theory of types. Upon the other side, at the same time, the much-derided copulæ were brought back to fresh life by H. Kolbe; with Frankland's aid a clearer notion of the meaning of copulated compounds was arrived at, and thus the way was smoothed for the setting up of the new radical theory and the doctrine of valency.

*Fusion of the older Theory of Types with the Radical Theory
by Laurent and Gerhardt.*

Of the two investigators who by their common work effected a transformation of the old into the new theory of types, Laurent had been already active as the originator of the substitution theory proper. Although both of them were resolute opponents of the dualistic view, they had, nevertheless, no objection to make use of the conception of radicals, though to these latter they attached a meaning of their own. In addition to Laurent and Gerhardt other chemists contributed materially to the establishment of the new theory of types, both by the ideas to which they gave expression and by the observations which they made. The stimulus thus given by Wurtz, Hofmann, and Williamson falls to be recorded here.

Laurent and Gerhardt exercised a strong mutual influence upon, and undoubtedly supplemented one another. Gerhardt was endowed with the special gift of bringing together isolated facts under one common point of view, and of drawing general conclusions therefrom. Laurent too was happy in perceiving the great importance involved in particular ideas, and he kept himself freer from prepossessions upon many points than his colleague.

A few sentences may be added here with regard to the lives of these two men. Auguste Laurent, born at La Folie near Langres in 1807, was initiated into chemistry by Dumas, thus acquiring a special knowledge of the organic part of it, to which with a certain one-sidedness he sub-

sequently remained faithful. His work upon naphthalene and carbolic acid, together with their derivatives, is evidence of this. After filling various posts, the last of which was a chemical professorship at Bordeaux, Laurent became Warden of the Mint at Paris, where he remained in intimate connection with Gerhardt until his early death in 1853.

Charles Gerhardt was born at Strasburg in 1816, and began his scientific career well equipped with a wide general education; he studied chemistry at various places in Germany, finally under the stimulating guidance of Liebig, to whom he, like so many others, was so greatly indebted. After working for several years in Paris, he became Professor of Chemistry at Montpellier from 1844 to 1848, and after another prolonged residence in the first-named city (where he opened a school for chemistry, which, however, was not commercially a success), was called in 1855 to fill the chemical chair in the Faculty of Sciences at Strasburg, where he died in the following year. His important services for the development of organic chemistry, together with the joint theoretical views of Laurent and himself, are detailed below.

Gerhardt's Theory of Residues.

At the time that Gerhardt brought out his first scientific work, the fight between the radical and substitution theories was at its height. The latter found pronounced expression in Dumas' theory of types, and was opposed not merely to the dualistic views upon which the older radical theory was based, but to radicals in general. Gerhardt doubtless felt the disadvantages which the giving up of the proximate constituents of organic compounds involved. Without forsaking the strict unitary standpoint of Dumas, he attempted to reintroduce the disdained radicals into chemistry under another name and with an altered meaning.—he set up the theory of residues (*théorie des résidus*).¹

Residues were, according to him, atomic complexes which

¹ *Ann. Chim. Phys.* (2), vol. lxxii. p. 184.

remain over from the interaction of two compounds, as the result of the stronger affinity of particular elements for one another, and which combine together because they are incapable of existing separately. Thus Gerhardt explained the formation of nitro-benzene from benzene and nitric acid, and, generally, the production of those bodies which he termed "copulated compounds" (*gcpaarte Verbindungen*) in the following simple manner:—"When two substances react with one another, an element (hydrogen) present in the one combines with another element (oxygen) present in the other to produce a stable compound (water), while the residues unite together." The latter he did not look upon as being actual atomic groups present in the compound in question, but as imaginary quantities; they were in his view absolutely distinct from the compounds of the same composition which were known in the free state, *e.g.* sulphurous acid (SO_2) or nitrogen peroxide (NO_2). Gerhardt gave expression to this difference by assuming the residues as being present in the "substitution-form." Further, the supposition of different residues in one and the same compound, according either to its mode of formation or decomposition, was likewise brought forward by him at that time.¹

If we examine this conception of Gerhardt's more closely, we see that his views upon substitution are expressed in the same breath with those upon radicals as variable atomic complexes. He endeavoured, in fact, to explain the processes of substitution by the aid of this notion, in teaching that an eliminated element is replaced by an equivalent of another element or residue of the reacting substance.

Dumas and Laurent too had already said the same thing in a different way. But Gerhardt knew how to draw important conclusions from his theory with regard to the chemical nature of "copulated compounds"; it did not escape him that the saturation-capacities of the latter with

¹ It must be mentioned here that the founders of the radical theory, Berzelius and Liebig, had expressed at one time the former in 1834, and the latter in 1838, perfectly similar views as to the possibility of assuming different radicals in the same compound (cf. *Berzelius' Jahresbericht*, vol. xiv. p. 348; *Ann. Chem.*, vol. xxvi. p. 176).

respect to bases were quite different from those of the original acids before these had been "copulated" with an alcohol or a hydrocarbon. Thus nitro-benzene, an indifferent substance, was produced from nitric acid and benzene, and the monobasic ether-sulphuric acids from sulphuric acid and the alcohols. Gerhardt concluded from these and similar observations that "the basicity of a copulated compound is equal to the sum of the basicities of the copulating substances minus 1." By means of this, his "Law of Basicity" (*Basizitätsgesetz*),¹ he was able to determine the chemical nature of acids about whose saturation-capacities doubt still prevailed at that time. With absolute definiteness he stated acetic acid to be monobasic, although it forms an acid sodium salt, and the same with regard to hydrochloric and nitric acids, because all these yield only neutral ethers: while sulphuric and oxalic acids were dibasic because, on copulation with an alcohol, they yield in the first instance monobasic ether-acids.

Gerhardt's first Classification of Organic Compounds.

Even before Gerhardt had attained to such clearness in this so important question, he had directed his endeavours to the classification of organic compounds. His first attempt at this is contained in the *Précis de Chimie Organique* (1842). Here we find him strongly influenced by Dumas and his type theory; like the latter, he avoided the use of any formulæ which might appear to indicate the proximate or rational composition of chemical compounds. These he arranged in an ascending series according to their empirical formulæ, in such a manner that substances containing equal amounts of carbon constituted one group. Inclined to express himself in figurative language, he compared this classification of organic compounds to a ladder, whose lowest steps were formed of the substances of simplest, and whose highest of those of most complicated composition. And since, from the oxidation of compounds rich in carbon,

¹ Cf. *Comptes Rendus*, vol. xvii. p. 312; *Comptes rendus des Travaux Chimiques par Laurent et Gerhardt* (1845), p. 161.

others which contain fewer atoms of that element are produced, he gave his arrangement the name of "combustion ladder" (*échelle de combustion*).

There was nothing of an unconstrained and natural classification here; on the contrary, the most different substances were collected into one class, provided only they fulfilled the necessary condition of containing the same number of carbon atoms. Not the slightest heed was paid to their chemical nature; acetic ether was placed alongside of butyric acid, and ethyl-oxalic acid alongside of succinic, solely for the reason given above. We note distinctly here the influence of Laurent, who not long before had made a mechanical classification of organic substances in a precisely similar manner (this, however, had made no impression).

Indeed, it is hardly conceivable to imagine how the older radical theory could have received a more severe blow than it did by the undue exaggeration of Dumas' theory of types. Gerhardt himself quickly felt this; his attempt at classification, which found its final and most definite expression in the new theory of types, showed distinctly that he had found a point of connection with the views of the radical theory, and that he strove to amalgamate the latter with the doctrine of substitution.

Before setting forth these works of Gerhardt, the efforts which he made—partly in conjunction with Laurent—to bring about uniformity of view with regard to the atomic weights of elements and compounds must be touched upon. In particular, the great and lasting service which those two men rendered in clearly defining what is meant by the term "molecule," and therewith reviving Avogadro's hypothesis, deserves our fullest recognition.

Gerhardt's "Equivalents."

At the beginning of the forties the uncertainty as to what atomic weights should be ascribed to the elements, and what atomic (*i.e.* molecular) weights to chemical compounds, had become one of great moment. The doubt which Gay-Lussac, Davy, and others had previously urged against the

assumption of definite atomic weights was brought forward again by Gmelin and his school. The atomic weight system of Berzelius, that work which he had accomplished after such great labour, came very near to being given up. In place of his atomic weights, based as they were upon solid foundations, "combining weights" were to be introduced, *i.e.* those values which were expressed by the simplest proportions of the substances entering into combination. All speculations upon relative atomic values were to be banished, and only the most sober possible formulation of chemical compounds attempted. The immediate result of this reaction was the halving of a large number of the atomic weights which Berzelius had introduced into the science. In place of the values assumed by him for carbon, oxygen, sulphur, and most of the metals, other values only half as great were taken; these *equivalents* were: C = 6, O = 8, S = 16, Ca = 20, Mg = 12, and so on.

Gerhardt began to oppose those equivalents in the year 1842, and was able to prove by cogent arguments that their assumption was inadmissible.¹ He showed, namely, that the amounts of water, carbonic acid, carbonic oxide, and sulphurous acid, which were separated during the reactions of organic compounds, were never expressible by what was known as one equivalent, but by two or some multiple of two. The smallest equivalent formulæ for those compounds, according to Gmelin's view, were H_2O_2 , C_2O_4 , C_2O_2 , and S_2O_4 . But, argued Gerhardt, there must be an error underlying this: "the symbols H_2O_2 and C_2O_4 either express one equivalent, or they express two." If we assume the former of these, then the formulæ of the inorganic compounds must be doubled; if the latter, then the "organic formulæ" must be halved. Gerhardt did away with the contradiction which existed in the formulation of organic and inorganic compounds by reinstating Berzelius' atomic weights for the elements carbon, oxygen, and sulphur, which were the ones of greatest moment here; *i.e.* taking

¹ Cf. *Journ. pr. Chem.*, vol. xxvii. p. 439; also his *Précis de Chimie Organique*, vol. i. p. 49.

H=1, then C=12, O=16, and S=32.¹ But he carried this reform only half-way; for, while he gave the proper values for the elements just named, he was led by special considerations to assume values for most of the metals which were only half as great as those proposed by Berzelius. Unlike the latter, who began by assuming that most of the metallic oxides had the composition indicated by the general formula MeO, Gerhardt compared those oxides with water, giving them therefore the formula Me₂O. He thus arrived at the correct atomic weights of the monovalent metals, but at incorrect ones for the divalent: *e.g.* for calcium the value 20 instead of 40, for lead that of 103.5 instead of 207, and so on.

Apart from this incompleteness there was an obscurity in Gerhardt's views with respect to atomic weights which could not fail to produce confusion; he both called the atomic weights just mentioned *equivalents*, and at the same time made use of this term for those amounts of chemical compounds which corresponded to their molecular weights, *i.e.* speaking generally, for quantities which are by no means necessarily equivalent chemically. Thus the quantities of hydrochloric, sulphuric, and acetic acids represented by the formulae HCl, H₂SO₄, and C₂H₄O₂, were in his mind equivalent to one another. We must here emphasise the point that Gerhardt attached another meaning to this word to what we now do; *equivalents* of chemical compounds were for him merely the comparable quantities of these.

Absolute clearness in the above points was only arrived at through Laurent's assistance. The latter definitely grasped the distinctions between molecular, atomic, and equivalent weights, the correct determination of whose values constitutes the basis of our present views upon molecule and atom; it was he who brought Avogadro's hypothesis back to life again, and prepared the way for its development, so significant for chemical science.

¹ Cf. *Journ. pr. Chem.*, vol. xxx. p. 1. It is very extraordinary that Gerhardt should have made no reference here to the identity of the atomic weights which he proposed with those of Berzelius.

The distinguishing between the terms Molecule, Atom, and Equivalent by Laurent and Gerhardt.

Gerhardt's most memorable efforts had for their aim the expression of the composition of all chemical compounds by means of formulae based upon one common standard, *i.e.* formulae comparable with one another. The formulae of volatile compounds ought, according to him, invariably to indicate those quantities which occupy two volumes when in the gaseous state, taking the volume of one atom of hydrogen as equal to 1. This sound principle has ever since been fully recognised.

Acting upon this, he altered the *four-volume* formulae of many organic compounds into *two-volume* ones by halving them. The false conception, much current at that time, according to which acetic acid (for example) received the formula $C_4H_8O_4$, alcohol that of $C_4H_{12}O_2$, and ethylene that of C_4H_8 , had grown up as the result of the dualistic views upon the composition of organic compounds, and also of the use of several incorrect atomic weights.¹ It was precisely to organic compounds—most of them volatile without decomposition—that Gerhardt's law could be most extensively applied, the law, namely, that their formulae depend upon the amounts by weight which are contained in equal volumes.

Much of Gerhardt's indistinctness, *e.g.* that produced by his using the word "equivalent" in a totally mistaken sense, was put right by Laurent. The latter pointed out with emphasis and clearness² that Gerhardt's equivalents were not even comparable with those of compounds, let alone of equal value with them; he showed that Gerhardt's equivalents of the elements must be regarded as their atomic

¹ For deducing the atomic composition of organic acids, the silver salts of the latter were chiefly made use of; for acetate of silver Berzelius had arrived at the formula $C_4H_6O_3 \cdot AgO$ ($Ag=216$), from which the composition of the acid, as given above, followed. Alcohol was regarded by Liebig as the hydrate of ethyl ether, and consequently formulated as $C_4H_{10}O \cdot H_2O$, whence the composition C_4H_8 was ascribed to ethylene, and so on.

² *Ann. Chim. Phys.* (3), vol. xviii. p. 266.

weights, and those of compounds as their molecular weights. Laurent's merit consisted in clearly grasping the meanings to be attached to those terms.

Laurent understood the molecular weight of an element or chemical compound as meaning that quantity which, under like conditions, occupies the same volume as two atoms of hydrogen; the quantity represented by the latter he looked upon as a molecule of hydrogen. For him, therefore, the molecular weights of chlorine, oxygen, nitrogen, and cyanogen were expressed by the formulæ Cl_2 , O_2 , N_2 , and $(\text{CN})_2$, and the molecular weights of hydrochloric and acetic acids by the formulæ HCl and $\text{C}_2\text{H}_4\text{O}_2$, because the quantities indicated by those symbols filled the same space as two parts by weight of hydrogen. The agreement between his ideas and those of Avogadro is plainly evidenced here; but to Laurent belongs the further merit of developing these in a high degree. He defined the molecule as "the smallest quantity which can be employed in order to produce a compound." And he saw a proof of the correctness of this view in the fact that the atoms of chlorine, bromine, hydrogen, etc., always come forward in pairs to act chemically.

The atom, according to Laurent, is the smallest quantity of an element which can be present in a compound; for atomic weights he adopted the values proposed by Gerhardt, which agreed in great part with those of Berzelius. Equivalents, lastly, signified for him the "equivalent amounts of analogous substances" (*die gleichwerthigen Mengen analoger Körper*). This last definition led logically to the assumption that one and the same element may have various equivalents, if it reacts with others in varying combining proportions.¹

¹ "The idea of an equivalent includes in itself the view of a similar function; we know that one and the same element can fill the rôle of two or of several others, whence it must follow that different weights also correspond to those different functions. On the other hand, we find different weights of the same metal, e.g. iron, copper, mercury, etc., replacing the hydrogen of acids, and thus forming salts which contain the same metal but possess different properties. These metals have therefore various equivalents" (cf.

The joint work of Laurent and Gerhardt upon this question—so excessively important for theoretical chemistry—found very little acceptance amongst chemists; indeed, many actively opposed such a conception as that of variable equivalent values. The sound but not yet sufficiently grounded views of Laurent upon the magnitude of the molecules (*i.e.* molecular weights) of elements and compounds did not, however, succeed in making their way at that time (towards the end of the forties). Gmelin's combining weights were still for the most part adhered to, and at the date of the appearance of Gerhardt's *Lehrbuch der Chemie* (1853) were in such general use that the author, against his better judgment, used Gmelin's numbers for the chemical symbols in his first three volumes, *i.e.* he employed equivalent formulæ.¹ Stronger proof than that given by Laurent and Gerhardt had to be adduced in order to convince people that the atomic and molecular weights which they employed were the correct ones. It was the researches of Williamson, published at the beginning of the fifties, which were especially instrumental in leading to this. The true perception was again arrived at here from experience collected in the field of organic chemistry.

Influence of the Researches of Wurtz, Hofmann, and Williamson upon the Development of the Theory of Types (1848-51).

The discovery by Wurtz² and Hofmann of organic derivatives of ammonia was of great significance for the firm

Comptes rendus des Travaux Chimique par Laurent et Gerhardt (1849), p. 1 *et seq.*).

¹ Gerhardt gave his reasons for using this notation in the preface to his book (vol. i. pp. 1, 2) as follows: "*J'y ai même fait le sacrifice de ma notation, pour m'en tenir aux formules anciennes, afin de mieux démontrer par l'exemple, combien l'usage de ces dernières est irrational, et de laisser au temps le soin, de consacrer une réforme, que les chimistes n'ont pas encore généralement adoptée.*"

² C. A. Wurtz, who was born at Strassburg in 1817 and died at Paris in 1884, was a pupil of Liebig, Balard, and Dumas; his life and works have been described very fully by A. W. Hofmann (*Ber.*, vol. xx. p. 815 *et seq.*), and by Friedel (*Notice sur la Vie et les Travaux de Wurtz*). From the year 1845 onwards, Wurtz filled the post of professor at various teaching institu-

establishment of the views finally comprised in Gerhardt's theory of types. In 1849 Wurtz observed the remarkable decomposition of cyanic ether by means of caustic potash, whereby he discovered methylamine and ethylamine, compounds so closely resembling ammonia.¹ Before this Berzelius had already expressed the view that the organic nitrogenous bases in general might be looked upon as substances which were copulated with ammonia. Liebig, on the other hand, regarded these as amido-compounds analogous to the ethers. Wurtz fluctuated between those two opinions, besides also suggesting the possibility of the organic bases being substitution products of ammonia, *e.g.* of "methyliak" (our methylamine) being ammonia in which one hydrogen atom was replaced by methyl. At first, however, he appears to have given the preference to the older view of Berzelius, according to which ethylamine, for example, was "ammonia copulated with ætherin."

The "typical" view of these bases was first arrived at through A. W. Hofmann's² brilliant researches upon amine bases;

tions in Paris (including the *École de Médecine* and the *Sorbonne*), his influence becoming very great as time went on. From 1866 to 1875 he was Dean of the Medical Faculty, and in this position materially aided in raising the standard of instruction in practical chemistry and physiology for medical students. Among his writings were the *Léçons de Philosophie Chimique* (1864) and *La Théorie Atomique* (1879), works which treated of questions in theoretical chemistry and which found much acceptance because of their clearness and the charming style in which they were written; also his *Traité Élémentaire de Chimie Médicale* (1864), and the *Dictionnaire de Chimie Pure et Appliquée* (edited by him). His admirable experimental researches, by which he acted as a pioneer in opening up particular branches of organic chemistry, will be spoken of under the special history of the subject. Most of these works were published in the *Annales de Chimie et de Physique*, of which he became one of the editors in 1852, and in the *Comptes Rendus*.

¹ *Comptes Rendus*, vol. xxviii. p. 223 et seq.

² August Wilhelm von Hofmann, born at Giessen on 8th April 1818, after several years of philosophical and juridical studies devoted himself to chemistry under the guidance of Liebig, whose assistant he soon became. After filling for a short time the post of assistant-professor (*Privatdozent*) at Bonn, he accepted in 1845 a call (made at Prince Albert's instigation) to the newly founded College of Chemistry in London, which became a government institution in 1853; in 1855 he was also made a non-resident Assayer of the Mint (these appointments, which were held by eminent chemists, otherwise unconnected with the Mint, were abolished in 1870). In 1864 he removed back again to

the production of these from ammonia and haloid compounds of the alkyls furnished a splendid proof of the correctness of the view that those compounds were formed from ammonia by the exchange of one or more hydrogen atoms for alcohol radicals.¹ The constitution of the *imide* and *nitrile* bases, like that of di- and tri-ethylamine, could scarcely be explained in any other way than by their derivation from ammonia, through the substitution of hydrogen atoms by alkyl radicals.² It was only after those important investigations by Hofmann that Wurtz clearly perceived that this relation of all these bodies to ammonia was the only conclusive explanation of them. He condensed the result of the above researches into the words: "It was thus that the ammonia type was created."

To this Williamson,³ by his distinguished researches,⁴ Bonn, and in 1865 was called to Berlin, as successor to Mitscherlich,—and there he still labours, an ornament to the university of that city.

His work as a teacher has been in every respect extraordinarily fruitful, while his organising talent has found scope in the building and fitting up of two admirable laboratories for general instruction at Bonn and Berlin. To success as a teacher there has also been added, in a marked degree in his case, that of an author; here he has shown his power of representing facts, and chemical doctrines founded upon them, in a clear and perspicuous manner. As an instance of this we may mention his *Einführung in die moderne Chemie* ("Introduction to Modern Chemistry"). The Obituary Memoirs (upon Liebig, Wöhler, Dumas, Sella, and Wurtz) by him are characterised by the loving care with which he enters into the life and works of the men whom he extols, besides being written in a most fascinating style.

As an investigator in experimental chemistry Hofmann meets us at every step; organic chemistry, more especially the field of nitrogen and phosphorus compounds, has been thoroughly studied by him, and in part exhausted. Finally, reference must be made here to the wonderful influence which he has exerted upon the coal tar colour industry, an industry which partly arose out of his scientific studies. Most of Hofmann's papers have been published in the *Annalen der Chemie* and in the *Berichte* of the German Chemical Society (at Berlin), which was founded by him in 1868.

¹ *Ann. Chem.*, vol. lxxiv. p. 174.

² *Ann. Chim. Phys.* (3), vol. xxx. p. 498.

³ A. W. Williamson, born in 1824, was a pupil of Liebig, and filled for a long time the chair of chemistry in University College, London, retiring from this post only a year or two ago. Especially in the years between 1850-60 did he enrich organic chemistry with valuable observations, which led to deductions of general application. His work upon the formation and constitution of ethers, in particular, was of the first importance.

⁴ Cf. especially *Ann. Chem.*, vol. lxxvii. p. 37; vol. lxxx. p. 73. Or *Journ. Chem. Soc.*, vol. iv. pp. 106 and 229.

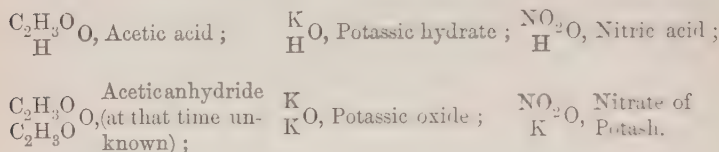
added the water type, thereby with Wurtz and Hofmann providing the foundation for Gerhardt's theory of types. In his experiments Williamson started with the idea of replacing hydrogen in known alcohols by hydrocarbons, so as to obtain homologues of the former. The action of ethyl iodide upon potassium ethylate yielded him, however, ethyl ether, and not the expected ethylated alcohol. This result induced him to try whether, by the action of potassium ethylate upon methyl iodide, a mixture of ethyl and methyl ethers or only one homogeneous compound would be produced. The latter was found to be the case; methyl-ethyl oxide, a "mixed ether," was obtained, and with this the much-discussed and at that time burning question of the molecular weights of ether and ethyl alcohol, and also that of the atomic weight of oxygen, were solved.¹ Liebig's idea that alcohol was the hydrate of ether had to be given up; on the other hand, Williamson's researches proved that the molecular formulæ of both compounds which had been assumed by Berzelius were the correct ones. The formation of ether by the interaction of alcohol and sulphuric acid, a process which had so greatly exercised the minds of the most eminent chemists, was thus now made perfectly clear by Williamson. Alcohol and ether he regarded as analogous to and built up on the type of water, as his definitions and formulæ show:—



This view (a view of which Laurent and other chemists had previously spoken favourably as being an admissible one) Williamson then proceeded to extend to many other substances, organic and inorganic, endeavouring at the same time to make its advantages evident. Thus he compared the acids, ketones (of whose true composition he had furnished beautiful experimental proof by a process similar to that mentioned above), and other compounds with water, *i.e.* he derived from water the compounds just named, by the

¹ Chancel arrived in a similar manner at the same result, independently of Williamson (cf. *Comptes Rendus*, vol. xxxi. p. 521).

substitution of one or both atoms of hydrogen by compound radicals or elements. The following examples will serve to illustrate his "typical" theory:—



Williamson expressed himself as follows with regard to the applicability of the typical view:¹ "The method here employed of stating the rational constitution of bodies by comparison with water, seems to me to be susceptible of great extension; and I have no hesitation in saying that its introduction will be of service in simplifying our ideas, by establishing a uniform standard of comparison by which bodies may be judged of."

His confidence in the possibility of extending the "typical" idea came out still more strongly upon another occasion,² when he expressed the opinion that reference to the one type of water sufficed for all inorganic and for the best-known organic compounds; only that in the case of many substances, *e.g.* dibasic acids, the formula of water must be taken doubled. The views expressed here are also to be found for the most part in Gerhardt's theory of types. The most important result of Williamson's researches consisted, however, not in the one-sided typical mode of explaining the constitution of chemical compounds, but rather in the determination of the true molecular values of organic substances. The methods which he made use of in order to attain to this very soon proved themselves exceptionally productive; they led Gerhardt to the discovery of the acid anhydrides, and Wurtz to that of mixed hydrocarbon radicals, the investigation of both of which has finally settled the controversy as to the molecular formulae of whole series of organic compounds.

¹ *Journ. Chem. Soc.*, vol. iv. p. 239.

² *Ibid.*, p. 350 (1851).

*Gerhardt's new Theory of Types.*¹

What has just been said is sufficient to show how effectively the "typical" view of organic compounds was furthered by the experimental researches of Wurtz, Hofmann, and Williamson. Numerous nitrogenous compounds were referred to the ammonia type, and a still larger number of oxygenated ones to the water type. Gerhardt consummated his work by adding to those two the hydrogen and hydrochloric acid types, and then he made the attempt to include all organic compounds under those few forms.

The endeavour to compare organic with inorganic bodies, which was already so strongly marked in the radical theory, was again distinctly apparent here; and here again it was ethyl compounds which mainly gave rise to the setting up of inorganic types as models for organic compounds. So early as 1846 Laurent² had thrown out the suggestion which was established in full detail by Williamson later on,—that alcohol and ether might be looked upon as derivatives of water, thus—



The inorganic acids and oxides too might be viewed (according to Laurent) as substitution-products of water. These compounds, so various in their natures, were regarded as built up after the same pattern.

In and after 1848 the American chemist Sterry Hunt published several papers,⁴ in which he gave a wide extension to the typical view by showing how numerous oxygenated compounds, inorganic as well as organic, might be pictured as derived from water, and how hydrocarbons belong to the hydrogen type. But his work, being unknown in Europe, did not in any way quicken the growth of the similar ideas

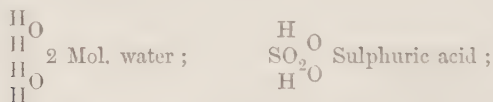
¹ Cf. *Ann. Chim. Phys.* (3), vol. xxxvii. p. 331; also *Traité de Chimie*, vol. iv. (1856).

² *Ann. Chim. Phys.* (3), vol. xviii. p. 266 *et seq.*

³ Cf. also Berzelius' view with regard to Ether, p. 241.

⁴ *Amer. Journ. of Science* (2), vols. v., vi., vii., and viii.

then running through many other minds. On the other hand, the above definite utterances of Williamson upon the reference of many organic compounds to water (as the form of compound of most general application) undoubtedly brought about a more rapid development of the doctrine of types. Not merely oxygenated bodies, but also non-oxygenated ones like amines, were without any hesitation taken as derived from water. Although Williamson thus lost his firm standing ground in consequence of the all too great elasticity of his formulæ, he gained, on the other hand, marked advantages from the extension of the type idea. He referred many compounds to the double or triple water type, and thereby introduced the notion of polyatomic radicals into chemistry. Sulphuric acid, for example, he referred to two molecules of water in which two atoms of hydrogen are replaced by sulphuryl (SO_2)—



while phosphoric acid was derived in a like manner from three molecules of water, and so on.

Stimulated especially thereto by his own important discovery of the anhydrides of monobasic organic acids,¹ Gerhardt collected the accumulated mass of "typical" ideas and condensed them into uniformity. Before everything else he desired to classify the large number of organic compounds in a synoptical manner, and for this the water, ammonia, hydrogen, and hydrochloric acid types were to serve as models. In addition to this he made use of a principle for the arrangement of organic substances, which had indeed been already applied by other chemists, but never in such a general manner, viz. he arranged them in different series, the members of each series belonging to the same type.

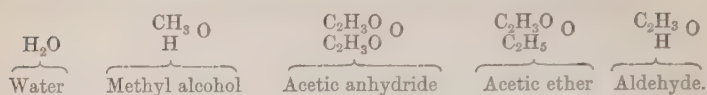
¹ *Ann. Chem.*, vol. lxxxii. p. 128. Those bodies, whose existence had been predicted by Williamson, were formerly supposed by Gerhardt to be incapable of preparation.

His first classification of organic compounds (cf. p. 272) did not possess the advantages which such a grouping in series offered. Since then Schiel¹ had established the conception of homology by directing attention to the equal differences in the composition of analogous bodies, in particular, of the alcohols, while Dumas had proved the same thing for the acids. And the researches of Kopp had further shown, with the utmost clearness, not only the chemical but also the physical resemblance of homologous compounds.

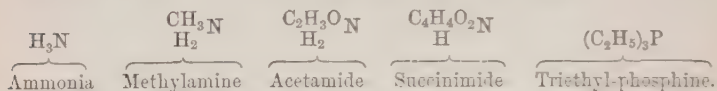
Gerhardt now collated the results of those preparatory labours with great ingenuity, and associated with the series of homologous bodies, which differed in composition by the increment $(\text{CH}_2)_n$, other series of *isologous* and *heterologous* compounds. The former of those were, according to him, chemically analogous substances which show another composition-difference from homologous ones, *e.g.* ethyl alcohol, $\text{C}_2\text{H}_6\text{O}$, and phenol, $\text{C}_6\text{H}_6\text{O}$; propionic acid, $\text{C}_3\text{H}_6\text{O}_2$, and benzoic acid, $\text{C}_7\text{H}_6\text{O}_2$ —compounds which differ from one another by the increment C_4 . Heterologous series contain such substances as are chemically dissimilar, but show a close connection with one another in their modes of formation. To such a series belong, for instance, ethyl alcohol, $\text{C}_2\text{H}_6\text{O}$, and acetic acid, $\text{C}_2\text{H}_4\text{O}_2$; amyl alcohol, $\text{C}_5\text{H}_{12}\text{O}$, and valeric acid, $\text{C}_5\text{H}_{10}\text{O}_2$.

As already mentioned, Gerhardt looked upon the members of such series as derivatives of one of his four types, resulting from these by the partial or complete substitution of their hydrogen atoms by *residues*. From the water type were derived (as Williamson had already taught) most of the organic compounds, including the alcohols, acids, simple and compound ethers, acid anhydrides, ketones, aldehydes, and salts. Alongside of water the analogously constituted sulphuretted hydrogen was placed as an auxiliary type, and from it the sulphur compounds corresponding to the oxygen compounds just mentioned were derived, *e.g.* the sulphides, mercaptans, thio-acids, etc. The following examples will serve to illustrate what has just been said:—

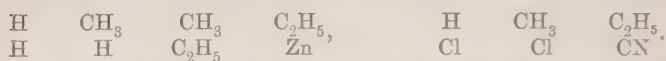
¹ *Ann. Chem.*, vol. xliii. p. 107 (1842).



Under the ammonia type were classified the amines, acid amides and imides, phosphines, arsines, etc., thus—



The hydrogen type included the hydrocarbons, together with the organo-metals; and the analogous hydrochloric acid type the chlorides, iodides, cyanides, etc., thus—



Gerhardt was quite justified in terming this classification of organic compounds according to types a *système unitaire*, for all assumption of an opposite within the chemical compounds themselves, or of a binary structure, was here entirely eliminated. Each compound was looked upon as a complete whole; even in those cases where the dualistic conception appeared to be indicated (especially in that of salts), derivatives of water alone were seen.

The question now arises,—did Gerhardt himself believe that he would get nearer to the solution of that problem, which Berzelius had designated as being of supreme importance to chemistry, by setting up those types and referring organic compounds to them? Did he consider that he had thereby materially advanced the solution of the chemical constitution of organic bodies? The answer to this must be in the negative, if we mean “constitution” in Berzelius’ sense. Gerhardt repeatedly expressed the opinion that it was impossible to arrive at the true constitution of these compounds, meaning by this the arrangement of their atoms (*l’arrangement des atomes*). In his view no strictly rational formulæ for organic compounds could be brought forward which would satisfy this demand, since several formulæ showing different proximate constituents or residues might

be looked upon as equally correct, according to the modes of formation or decomposition of the compounds. Grounds of expediency alone must decide whether one formula was to be preferred to another; that formula which explained the larger number of methods of formation and decomposition of the particular compound in question was to be chosen. This elastic view was brought prominently forward by Gerhardt at every opportunity, especially in the fourth volume of his text-book, and he emphasised the point that the constitution of compounds, according to the type theory, was not the same thing as their rational composition in Berzelius' sense.

Formulae were for Gerhardt merely pictures of the changes which chemical compounds underwent; they simply illustrated the modes of formation and decomposition of the latter. Types, on their part, even when their composition is exceedingly simple, "do not in any respect show how the atoms are grouped, but only the analogies of their metamorphoses. The type is the unit with which are compared all those compounds which show analogous decompositions, or which are the products of analogous decompositions."

After this exposition of Gerhardt's system in its main points, it will be intelligible why it has been spoken of as resulting from the fusion of the type theory of Dumas with the older radical theory. Gerhardt had made use of particular parts in both of these, and had recast them for incorporation into his *système unitaire*. The idea that organic compounds are constructed on certain models, to which they can be referred, originated essentially in the older type doctrine, but, although hidden, it was also contained in the radical theory; in the latter, groups of organic substances had been directly compared with analogously constituted inorganic ones. Now it was of fundamental importance for the success of the new type theory that it borrowed from the radical theory the conception of atomic groups which behaved like simple substances; these groups could not, however, exist in the free state, as had formerly been supposed, but could only

act in place of elements in compounds. This conception, coupled with that of the alterability (by substitution) of those atomic complexes, has proved to be absolutely correct, and at the same time of the greatest value. The question of the proximate composition of the above groups was left unanswered by Gerhardt, the key to its solution being supplied from quite another quarter, *i.e.* by Kolbe and Frankland.

While the older type theory of Dumas ascribed no appreciable influence to the chemical nature of the constituents of a compound upon the character of the latter, Gerhardt showed his greater insight in this point also by recognising certain principles of Berzelius' school, even when he appeared mainly intent on opposing their spirit. He pointed out that the elements or atomic groups, which take the place of hydrogen in his types, determine according to their electro-chemical nature the nature of the result-

ing compounds. Thus he represented potash, $\frac{\text{K}}{\text{H}}\text{O}$, as a

basic, and nitric acid, $\frac{\text{NO}_2}{\text{H}}\text{O}$, as an acid body, because the

hydrogen of the neutral water was replaced respectively in these by an electro-positive and an electro-negative radical;

but alcohol, $\frac{\text{C}_2\text{H}_5}{\text{H}}\text{O}$, as an almost neutral compound, ethyl

being of pretty much the same nature as hydrogen itself. This return to views, which had formerly been combated so vigorously by that side, deserves to be especially noted.

The criticisms passed upon Gerhardt's type theory at that time varied very much. Many chemists, especially the younger ones, greeted it as an important conquest on the part of research. But, as a matter of fact, the favourable reception given to the typical view was due to grounds of a practical nature; men gave it as their opinion quite frankly, that the chief advantage which the reference of organic compounds to a few inorganic types brought with it, consisted in its thereby simplifying the study of organic

chemistry. Liebig, who had criticised Gerhardt's earlier efforts at classifying organic compounds most severely,¹ acknowledged later on² the "utility of the so-called type theory"; but at the same time he laid stress upon the point that it left the weighty question of the formation of organic compounds untouched. Kolbe took up a more drastic attitude than this; he designated the grouping of organic compounds into the above four types a mere playing with formulæ. His own efforts he directed to replacing these purely formal types by real ones, which should stand in a natural connection to the compounds derived from them. Indeed, there was a serious danger that a door would be opened for empty formulation. We have only to recall that Odling and also Wurtz³ endeavoured to simplify Gerhardt's types by referring those of water and ammonia to the double and triple hydrogen ones. With this the momentous question of the chemical constitution of organic compounds was diverted appreciably from the direction which had been given to it by the school of Berzelius and Liebig. The term "constitution," already very elastic in Gerhardt's theory, threatened to lose all meaning by formulation so exaggerated.

Extension of the Type Theory by Kekulé.

Gerhardt did not live to enjoy the cordial reception which was given by many chemists to the opinions laid down by him in the fourth volume of his text-book. His type theory underwent a not inconsiderable extension the year after his death (*i.e.* in 1857), by the assumption of the so-called *mixed types*, which aimed at making clear the relations of many organic compounds to two or more types. The more general application of this by Kekulé⁴ was

¹ *Ann. Chem.*, vol. lvii. p. 93, *Herr Gerhardt und die organische Chemie*.

² *Ann. Chem.*, vol. cxxi. p. 163.

³ Cf. *Ann. Chim. Phys.* (3), vol. xlv. p. 305.

⁴ *Ann. Chem.*, vol. civ. p. 129.

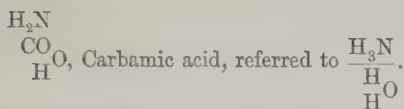
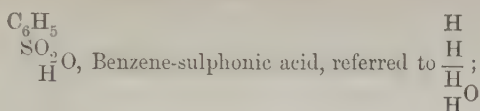
preceded by Williamson's idea that certain organic compounds might be derived from *multiplied* or *condensed types*. Just as chemical compounds proceeded from these through the substitution of several hydrogen atoms by polybasic radicals, so were different types like water and ammonia, or water and hydrogen, etc., conjoined in order to derive from them those substances which had previously been known as copulated compounds (*gepaarte Verbindungen*), to distinguish them from others which were readily classified under one type.¹ Kekulé² recognised in the removal of this barrier the main advantage which was to be derived from the assumption of mixed types, as is apparent from the following extract: "The so-called copulated compounds are not constituted differently from other chemical compounds; they can in like manner with these be referred to types in which hydrogen is replaced by radicals; and, in respect to formation and saturation-capacity, they follow the same laws which hold good for all chemical compounds."³

A few examples of formulæ will serve to make the use of the mixed types intelligible:—

¹ The same idea which Kekulé generalised later on had indeed occurred to Gerhardt, in so far that he had referred the aminic acids (for example) to the mixed ammonia-water type.

² August Kekulé, born at Darmstadt on 7th September 1829, became assistant professor of chemistry at Heidelberg in 1856, and then professor at Ghent from 1858 to 1865; in the latter year he was called to the University of Bonn, where he is still the *doyen* of his science. By his *Lehrbuch der organischen Chemie* ("Text-Book of Organic Chemistry," Erlangen, begun to be published in 1859), in which he endeavoured to work out the typical view and subsequently the structural doctrine to their logical conclusions, he has exercised an immense influence upon the chemists of his time. More especially, by his happy conception of benzene (the basis of the "aromatic" hydrocarbons) as a hexa-methine, has he furnished the direction for an exceptionally large and widespreading branch of chemical research, and this still holds at the present moment. His researches on fulminate of mercury, unsaturated dibasic acids, and the condensation of aldehyde (to name only a few) have proved him an admirable investigator. Mention may also be made here of his share in the editing of a former journal, the *Kritische Zeitschrift für Chemie*, etc., and of the present *Annalen der Chemie*, in the latter of which most of his experimental work has been published.

³ *Ann. Chem.*, vol. civ. p. 139.



Almost simultaneously with the above extension of the type theory, a suggestion was made by Kekulé which, thanks to special circumstances, was destined to guide this doctrine into other and higher paths. *À propos* of his researches upon fulminate of mercury,¹ he had expressed the opinion that the methyl compounds and the numerous bodies derived from them might be referred to the type of marsh gas, to which he gave the equivalent formula C_2H_4 . He illustrated the connection of several compounds to the new type by the following examples:—



Kekulé's formulation here is noteworthy, in that he makes use of atomic weights which he had formerly regarded as incorrect, *i.e.* $\text{H} = 1$, $\text{C} = 6$, and $\text{O} = 8$. And a remark that he made strikes one as strange, *viz.* that the new type was not to be taken in the sense of Gerhardt's unitary theory but in that of Dumas'. From this one might infer that marsh gas was not intended to be placed alongside of Gerhardt's four types; but, notwithstanding this, to give it a place by itself does not seem to have been meant by Kekulé, since he adds, quite in the spirit of the newer type theory, that what he mainly wishes to indicate by his formulæ are the relations in which the compounds enumerated stand to one another.

Already in the following year (1858), the meaning which he attached to methane as the mother substance of a large number of compounds became more clear. But a detailed account of his views upon this must be reserved for a later

¹ *Ann. Chem.*, vol. ci. p. 200.

section of the book, when the transition of the type theory into the structure theory will come to be treated of.

Before, however, this development of chemical hypotheses could be consummated, much work had to be done in order to get nearer to a knowledge of the chemical constitution of organic compounds. The types themselves could not aid in the solution of this problem without their own nature being first cleared up. The key to the explanation of these relations was forged by the labours and speculations of Frankland and Kolbe. To these two investigators is primarily due the more profound conception of the constitution of organic substances as opposed to the typical (*typisch schematischen*). Their researches contributed more than any others to bring about the change in direction taken by the type theory; they were, in fact, the indispensable preliminary to that transformation of theoretical views which completed itself towards the end of the fifties. The correctness of this statement will be seen from what follows in the succeeding sections.

It is true that the typists estimate the services of Frankland and Kolbe quite otherwise. The influence exercised by those two men on the remodelling of the type theory has not only been greatly minimised, but even the exact contrary has been asserted, viz. that "typical" hypotheses influenced them.¹

*Development of the Newer Radical Theory by Kolbe—
A Survey of his Principal Work.*

Before speaking of Kolbe's scientific labours, which produced a deep and lasting effect on the development of theoretical chemistry, a short sketch of his life may be fitly appended here.²

¹ Such erroneous conceptions are always long of being dispelled. Thus, in the description of "the theories of to-day" in Wurtz's *Histoire des Doctrines Chimiques*, the influence of the above two scientists is very much neglected. It seems hardly credible that Frankland, the real originator of the doctrine of valency, should scarcely be mentioned in this publication.

² Cf. the memoir by E. v. Meyer, which appeared shortly after Kolbe's death (*Journ. pr. Chem.* (2), vol. xxx. p. 417); Voit, *Bayer. Acad.*, 1885; and A. W. Hofmann, *Ber.*, vol. xvii. p. 2809.

Hermann Kolbe, son of the Pastor of Elliehausen, near Göttingen, was born in 1818, and applied himself to the study of chemistry under Wöhler's stimulating guidance in 1838. The results of his first research were published in 1842, and for the next forty-two years he continued to enrich his science with a long succession of the most valuable experimental and theoretical work. His outward life, if we except perhaps the first few years immediately following his university curriculum, was that of a German scientist. From 1842-47 he was assistant to Bunsen at Marburg, and then to Playfair in London, during which time he occupied himself mainly with practical chemical work; after this came the years of his literary apprenticeship (1847-51) in Brunswick, where he had gone at the request of the well-known publishers, Fr. Vieweg and Son, to take up the editorship of the *Dictionary of Chemistry* started by Liebig. This work not being of such a nature as to satisfy him permanently, he willingly accepted in 1851 a call to Marburg, where, as Bunsen's successor, he developed exceptional powers as a teacher, especially in the years following 1858. In 1865 he was called to the University of Leipzig, and worked there with marked success until his death on 25th November 1884.

The great influence which Kolbe exercised upon chemical science depended to an unusual degree upon his experimental work, which will be treated of later on, but at the same time also upon his eminence as a teacher, in which respect he may be spoken of along with Liebig. His method of teaching was very like that of the latter and had the best results; the student of practical chemistry was taught to observe and think for himself. Kolbe's gifts as a teacher were greatly aided by his sound common-sense and organising talent, which showed themselves in a marked degree in the building and fitting up of the new Leipzig laboratory.

In addition to his work as a teacher, based as this was upon oral instruction, Kolbe was also extremely active in a literary sense. Apart from his numerous scientific papers, valuable articles for the *Handwörterbuch der Chemie* ("Die-

tionary of Chemistry"), and occasional pamphlets, he published a large *Lehrbuch der organischen Chemie* ("Text-Book of Organic Chemistry," Braunschweig, 1854-65), and smaller text-books both of inorganic and organic chemistry (1877-83). These books are distinguished by clearness in arrangement, precision of expression, a delightful style, and perspicacity and acuteness in discussion.

In his writings upon questions of theoretical chemistry, published for the last fourteen years of his life in the *Journal für praktische Chemie* (of which he succeeded Erdmann as editor in 1870), Kolbe gave play to a keen criticism, which became intensified as time went on, upon the defects and extravagances which he considered were due to the direction taken by modern chemistry. If those critiques were often strongly polemical and did not altogether avoid the danger of introducing the personality of many a brother chemist, still his only aim in them was the welfare of his beloved science, which he believed to be in serious danger. His efforts at exposing error were often wrongly interpreted by many of his contemporaries.

*The Re-animation of the Radical Theory by Kolbe—
Frankland's Co-operation.*

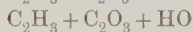
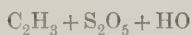
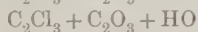
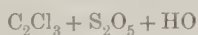
At the time when Kolbe published the first of his more important researches,¹ the doctrine advocated by Berzelius, that organic compounds contain definite radicals which act similarly to elements in inorganic compounds, had been strongly driven back by the attack of unitarism. Many chemists were of opinion that the partly arbitrary supposition of hypothetical radicals could not advance the science any further. The assumption of copulae (*Paarlinge*) in the so-called copulated compounds satisfied very few. In short, the old radical theory in its original form was held to be no longer capable of existence. The preference given by the school of Gmelin to the simplest views which were possible is sufficient evidence of this sense of discouragement. Facts

¹ *Ann. Chem.*, vol. xlv. p. 41; vol. liv. p. 145.

alone were to decide; any intelligent grouping of those facts together was deemed useless.

Kolbe now united the conclusions deduced from his first researches with the declining theory of Berzelius; he endued the latter with new life by throwing aside whatever of it was dead and replacing this by vigorous principles. From his own and other investigations he came to the conclusion that the unalterability of radicals, as taught by Berzelius, could no longer be maintained, since the facts of substitution had to be taken into account. He did, indeed, take up again Berzelius' hypothesis of copulæ, but attached another meaning to these, since he allowed that they exercised a not inconsiderable influence upon the compounds with which they were copulated.¹

If we desire to sum up the main results of his labours just cited, and of his synthesis of trichloroacetic acid, so immediately connected with them, we may do so as follows: Trichloro-methyl-hyposulphuric acid (our present trichloro-methyl-sulphonic acid), discovered by him, and trichloroacetic acid, together with the compounds free from chlorine obtained from these by reduction, were analogously constituted acids, copulated respectively with trichloromethyl and methyl, thus—



True, the mode in which these two radicals were combined with the acids was not yet known, but the germ of the correct explanation with regard to the constitution of carboxylic and sulphonic acids, which was given by Kolbe at a later date, was already present in these beginnings.

This germ was soon to undergo further development by investigations carried out at first by Kolbe alone, and afterwards together with Frankland in London. From their beautiful researches on the transformation of the alkyl cyanides into fatty acids,² they concluded with perfect precision that methyl, ethyl, and similar radicals were

¹ Cf. *Ann. Chem.*, vol. liv. p. 156.

² *Ibid.*, vol. lxx. p. 288.

immediate constituents of acetic acid and its homologues. Kolbe himself was led to the same conclusion by his important work upon the electrolysis of salts of the fatty acids;¹ he saw in the methyl and butyl, separated at the positive pole from acetic and valerianic acids respectively, the proof of the correctness of this assumption. He believed, indeed, that he had isolated the radicals themselves: and even although he was wrong in so thinking (the hydrocarbons obtained by him having double the molecular weight that the radicals would possess), this affected but little the question of the constitution of the carboxylic acids. The chief goal of his endeavours, *i.e.* the finding out of the true composition of the above and similar acids, was still kept in view by him, notwithstanding this mistake.

The outcome of the work of his which has just been mentioned was that the view previously held with regard to these organic acids no longer satisfied him. He did not, however, abandon this all at once, but rather developed from it a theory which approximated to the truth, and which soon showed itself capable of further improvement. Even so early as when writing the articles upon *Formulæ* and *Copulated Compounds* for his Dictionary (in 1848), he expressed and gave reasons for the view that the fatty acids were oxygen compounds of the radicals hydrogen, methyl, ethyl, etc., combined with the double carbon equivalent C_2 .²

Acetic acid contained as its immediate constituent an atomic complex constituted similarly to that of the cacodyl compounds. *Cacodyl* itself, which was here, for the first time, interpreted as being arsenic copulated with two methyl radicals, corresponded to the so-called *acetyl* of acetic acid, *i.e.* $C_2H_3C_2$ (not to be confounded with the radical acetyl of to-day, which at that time was known as *acetoxyl*).

Even at this early date Kolbe expressed the important

¹ *Ann. Chem.*, vol. lxxix. p. 258.

² Kolbe, like many others, made use at this time of Gmelin's equivalent weights, in which $H=1$, $C=6$, $O=8$, $S=32$, etc. His formulæ were, notwithstanding this, molecular formulæ; thus he gave carbonic acid, acetic acid, alcohol, aldehyde, and acetone the same atomic (*i.e.* molecular) weights as we employ for these substances to-day.

opinion that in the acetyl ($C_2H_3C_2$) of acetic acid, "the last C_2 alone forms the connecting link for the oxygen, the methyl being in some sort only an appendage." This idea, which recalls Berzelius' doctrine of copulæ, was based upon the point that it was unessential for the nature of the acids whether hydrogen or methyl, ethyl, etc., was copulated with the C_2 .

He went into those important ideas in detail in a treatise entitled, *Ueber die chemische Konstitution und Natur der organischen Radikale* ("Upon the Chemical Constitution and Nature of the Organic Radicals").¹ Taking his stand upon the basis of the older radical doctrine, he built this up into a living theory by eliminating from it those principles which stood in contradiction to the facts. But at the same time he did not remain stationary upon the point of vantage which he had thus gained.

Under the influence of the admirable researches of Frankland² upon the alcohol radicals and the organo-metallic compounds, which were begun at that time, Kolbe advanced step by step. With regard to this period, he stated definitely himself,³ that "the want of clearness in (my) conception of the mode in which the so-called copulæ were combined, was a great weakness in the hypothesis of copulated radicals. . . . It is Frankland's merit to have

¹ *Ann. Chem.*, vol. lxxv. p. 211; vol. lxxvi. p. 1.

² Edward Frankland, born in 1825, studied chemistry with Liebig and Bunsen, and also under Kolbe's stimulus in Germany, and afterwards filled successively the chairs of chemistry in the Owens College, Manchester, and in the Royal School of Mines, London, retiring from the latter of these posts only a few years ago. He attracted the attention of chemists even by his earliest work, which led him to the discovery of the organo-metals, and also by his joint researches with Kolbe. The chief share which he took in the development of our present views upon the valency of the elements will be treated of in detail later on, while his other memorable investigations in organic chemistry will often have to be referred to under the special history of this branch. Frankland's papers have mostly been published in the English journals and the *Annalen der Chemie*; in 1877 they were collected into one volume, entitled *Researches in Pure, Applied, and Physical Chemistry*. He is also the author of the text-book, *Lecture Notes for Chemical Students*.

³ Cf. *Das chem. Laboratorium der Universität Marburg, etc.* (Braunschweig, 1865), p. 32.

been the first to throw light upon this, and therewith to have completely done away with the idea of copulation, by recognising the fact that the various elements possess definite saturation-capacities."

Kolbe readily embraced his friend's views, and copulae thus received a totally different meaning from what they had formerly done; henceforth they were to be regarded as integral parts of organic compounds and not as mere appendages.

This change in his opinions was not long of bearing fruit. And here it was again the fatty acids whose constitution he undertook to work out. In 1855¹ he first gave definite expression to the view that the acids, considered as anhydrous, were derivatives of carbonic acid: for instance, acetic was methyl-carbonic acid, *i.e.* C_2O_4 , in which one oxygen-equivalent was replaced by methyl, C_2H_3 . The hydrated acids he still regarded dualistically as compounds of the anhydrides with water.

The assumption that those acids were substitution-products of carbonic acid had developed itself from the views held regarding the organo-metallic compounds. Just as Frankland explained cacodylic acid as arsenic acid with two methyls in the place of two equivalents of oxygen, and stannio-ethyl oxide as the corresponding tin derivative, so did Kolbe happily interpret the constitution of other organic compounds. He soon advanced beyond the domain of the organic acids, and developed the idea, similar to that mentioned above, that many organic substances are to be regarded as derivatives of carbonic acid, and many others as derivatives of sulphuric. How this idea expanded into a perfect whole is seen from his writings in the years 1857-58,² and also from those portions of his text-book written both at that time and shortly before it. These

¹ *Handwörterbuch der Chemie*, vol. vi. p. 802.

² *Ann. Chem.*, vol. ci. p. 257; this paper is a joint one with Frankland, *i.e.* Kolbe lays emphasis on the point that he is here giving utterance both to his own and Frankland's views. Cf. also Kolbe's pamphlet (1858), *Ueber die chemische Konstitution organischer Verbindungen* ("On the Chemical Constitution of Organic Compounds").

theoretical considerations and, with them, the revived radical theory attained to their completed form in a treatise published in 1859, entitled, *Ueber den natürlichen Zusammenhang der organischen mit den anorganischen Verbindungen, die wissenschaftliche Grundlage zu einer naturgemässen Klassifikation der organischen chemischen Körper* ("Upon the Natural Connection existing between Organic and Inorganic Compounds, being the Scientific Basis of a Rational Classification of Organic Chemical Substances").¹

The main outcome of Kolbe's speculations is given in the following sentence: "Organic compounds are all derivatives of inorganic, and result from the latter—in some cases directly—by wonderfully simple substitution-processes." This idea runs through the whole treatise, and is illustrated with the most convincing clearness by numerous examples taken from the wide field of organic chemistry.

The alcohols, carboxylic acids, ketones, and aldehydes were derived, according to Kolbe, from carbonic acid, $(C_2O_2)_O_2$, and its hydrate, $C_2O_2 \begin{smallmatrix} OHO \\ OHO \end{smallmatrix}$, respectively. The polybasic carboxylic acids proceeded in the same way from two or three molecules of the hydrated carbonic acid, through the entrance of polyatomic radicals, just as the monobasic did from one molecule. Similar definite views were expressed by Kolbe with regard to other classes of organic compounds, *e.g.* the phosphinic and arsenic acids, amines and amides, and the organo-metals, which he derived in the simplest manner from inorganic compounds. He laid the utmost emphasis upon his formulæ being the unambiguous expression of precise opinions; with Gerhardt's assumption, that various constitutional formulæ might, with equal justice, be set up for one and the same compound, he had absolutely nothing in common.

Kolbe himself gave a striking proof in the treatise above mentioned of the capacity for development of his views respecting the constitution of organic compounds. He comprised in his survey not merely those classes of

¹ *Ann. Chem.*, vol. cxiii. p. 293.

compounds which were known, but advanced beyond them to others at that time unknown. From the relations so clearly recognised by him as existing between the alcohols and the carboxylic acids, he deduced the possibility of preparing new varieties of alcohols; he predicted the existence both of secondary and of tertiary alcohols,¹ and even went so far as to indicate a probable method for preparing and decomposing the first of these. No such brilliant deductive treatment of chemical questions had as yet been seen in organic chemistry. And the discovery of those classes of compounds which he had prognosticated had not to be waited for long; Friedel isolated secondary propyl alcohol in 1862, and Butlerow tertiary butyl alcohol in 1864.

Kolbe's most important Experimental Researches (1857-63).

The comprehensive speculations of Kolbe upon the constitution of organic compounds could not have attained to the firm hold and the wide significance which they did, had they not been conjoined throughout with admirable experimental work. We shall frequently have occasion, in the special history of organic chemistry, to refer to those labours, through which the rational composition of important classes of compounds was first arrived at with certainty. Thus it was his researches upon lactic acid which showed it to be oxy-propionic, and the corresponding alanin to be amido-propionic acid. Glycollic acid and glycollin were likewise shown by Kolbe to belong to the same class, the one being proved to be oxy-, and the other amido-acetic acid; he also recognised salicylic acid as oxybenzoic, and the so-called benzoic acid (*Benzaminsäure*) as amido-benzoic. He was thus in a position to clear up the constitution of compounds upon whose investigation chemists of such eminence as Kekulé and Wurtz had laboured in vain. Numerous substances, the names (*Trivialbezeichnungen*) given to which showed how little was known with respect to their constitution, received from Kolbe their proper place among other

¹ Cf. *Ann. Chem.*, vol. cxiii. p. 307.

compounds. The conversion of malic and tartaric acids into succinic, which was carried out by Schmitt at his suggestion, revealed at one stroke the hitherto unknown relations existing between the two first of these acids and the last. By his researches upon taurine, which he taught how to prepare artificially, he proved how both it, and the isethionic acid produced from it, were constituted analogously to alanin and lactic acid. And the same clearness shed itself over the rational composition of asparagine and aspartic acid, which he was the first to interpret correctly.

The above are merely the results of labours performed within a short period of time, but they are amply sufficient to prove what undying services he rendered in investigating the chemical constitution of organic compounds. And no mention has been made here of a large number of other researches carried out at his suggestion and with his co-operation; among these were the work of Griess upon the class of diazo-compounds, Oefele's discovery of the sulphines, and Volhard's synthesis of sarcosine.

In order to round off in some degree this short record of Kolbe's achievements, we ought further to recall several investigations made in the years following, in which he was guided throughout by the aspiration to gain as close an insight into the constitution of organic compounds as possible. Among these we may refer to his proof of malonic acid resulting from cyan-acetic, the discovery of nitro-methane, the series of memorable researches upon salicylic and para-oxybenzoic acids, and lastly, that upon isatoic acid, which was cut short by his death.

Kolbe's Attitude towards the older and the newer Chemistry.

In all Kolbe's investigations, whether speculative or experimental, we feel the salutary historic method by which they are characterised. He built upon the edifice already existing, and remained in his scientific efforts in continuity with the chiefs of the classical school. He was always glad to acknowledge that his success as a chemist was due

primarily to Berzelius, and, after him, "to the great exemplars Liebig, Wöhler, and Bunsen, who, to use a phrase of Berzelius, were true workers in chemistry" (*wahre Bearbeiter der Chemie gewesen sind*).

The criticisms passed upon Kolbe by his contemporaries, in so far as regarded his attitude to organic chemistry, differed very greatly. The exponents of the earlier period appreciated his services better than the disciples of the type theory,—a theory which he himself did not value at its true worth. A few remarks upon the relation between Kolbe's views and those of the typists will be in place here. As already stated, he spoke of the type theory as being unscientific; he saw in it not a real theory but merely a play upon formulæ. In spite of his definite utterances upon this point, however, it has frequently been asserted that he took Gerhardt's doctrine of types as his basis, and that therefore his derivation of organic compounds from carbonic acid, carbonic oxide, sulphuric acid, sulphurous acid, etc., coincided with that from the three types of hydrogen, water, and ammonia. Kolbe did indeed connect organic with inorganic compounds, but he repeatedly emphasised the point¹ that these latter were real types, as opposed to the formal ones of the type theorists. His most ardent wish was to fathom the chemical constitution of organic compounds; but to merely classify the latter upon certain models, or to go so far as to force them into arbitrary types, was in the highest degree distasteful to him. Kolbe attached especial weight to the relations actually existing between organic and inorganic bodies, whence the emphasis laid in the title of his treatise, above spoken of, upon the "natural connection between these as forming a scientific basis of a rational classification of organic substances." Hence, also, his attempts, begun at an early date, to prepare the latter artificially from simple inorganic compounds, with the object of thus gaining an insight into their chemical constitution.

We thus see Kolbe pursuing his own way, and, not led aside by the criticisms of his contemporaries, working with

¹ Cf. (*e.g.*) *Journ. pr. Chem.* (2), vol. xxviii. p. 440.

wonderful effect, more particularly in advancing a knowledge of the rational composition of organic compounds. The older radical theory acquired through him new life, and the radicals themselves received a more profound meaning. While in the type theory the latter were looked upon as residues whose nature could be no further investigated, Kolbe devoted his whole energies to breaking up the radicals into their immediate constituents. To give but a few examples,—he showed cacodyl to be arsene-dimethyl, acetyl to be a compound of methyl and carbonyl, and the alkyls to be derivatives of methyl. These and other results of his investigations, together with the rich fruits of Frankland's labours, were undoubtedly of the first importance, indeed indispensable, for the development of the new type doctrine into the structure theory.

These two men, the workers of greatest originality in the field of organic chemistry during the storm-and-stress period of the fifties, contributed most essentially by their labours to the recognition of the fact that the peculiarity of Gerhardt's types rested upon the different saturation-capacities of the elements which they contained. The chief merit of having worked in this direction belongs to Frankland.

THE FOUNDING OF THE DOCTRINE OF THE SATURATION-
CAPACITY OF THE ELEMENTS BY FRANKLAND.

In the foregoing section the influence exercised by Frankland on the views developed by Kolbe with regard to the constitution of organic compounds has been already distinctly emphasised. It was Frankland who, in his memorable paper,—*On a New Series of Organic Compounds containing Metals*¹—furnished the proof that the copulation of radicals with elements (*e.g.* carbon, arsenic, and sulphur), as taught by Kolbe, depended upon a property inherent in the elementary atoms of the compounds just named. The notion of copulation was recognised by Frankland as being one-sided, and the misconception which had crept in from its use was done away with by him,—the idea, namely, that the radicals present as so-called copula in organic substances exercised no appreciable influence upon those compounds with which they were supposed to be copulated.

From his experiences gained from the organo-metallic compounds, Frankland developed the doctrine of the valency of the elements. If, freeing our minds from all prepossession, we turn our glance backward, we recognise the germ of this doctrine as being already contained in the law of multiple proportions, which stated that the elements show different, but at the same time perfectly definite stages in their combinations. Among the facts known at a very early period was, for instance, that of one atom of phosphorus combining with three and five atoms of chlorine to definite compounds; but the expression for this and other similar observations, *viz.* that phosphorus and many other elements were possessed of more than one valency, *i.e.* could manifest varying saturation-capacities, had yet to be found. Further, no one had any clear conception of a limit to the saturation-

¹ *Phil. Trans.*, vol. cxlii. p. 417; *Ann. Chem.*, vol. lxxxv. p. 329. This paper was read before the London Chemical Society in 1852.

capacities of elements, and, what was of the first importance, a sharp distinction between the terms "atom" and "equivalent" was still wanting. With regard to this latter point, the experiences gained respecting the substitution of the hydrogen of organic compounds by chlorine, oxygen, etc., and the deductions drawn from these had tended to elucidate matters. So early as 1834 Dumas had pointed out that 1 atom of hydrogen was replaced by 1 atom of chlorine, but only by $\frac{1}{2}$ atom of oxygen; those quantities were therefore equivalent to 1 atom of hydrogen. The idea of the "replaceable value" of certain metals also came more distinctly into prominence through the doctrine of polybasic acids, already spoken of; this was exemplified, for instance, in Liebig's statement that 1 atom of antimony was equivalent to 3 atoms of hydrogen, but 1 of potassium only to 1 atom of hydrogen. Notwithstanding this, however, a precise expression for such facts as these had not yet been found. In the course of the forties the conception of a chemical equivalent as distinguished from an atom, a conception which had been arrived at after so much labour, completely died out; the growing influence at that time of the Gmelin school affords us eloquent testimony of this backward step.

It is a remarkable fact that, for the founding of the doctrine of valency, not simple compounds of inorganic chemistry, but the more complicated ones of organic were called into service. The relations which in the former found clear expression, and were easily read in the law of multiple proportions, had to be first laboriously deciphered here from organic compounds.

As stated already, it was the organo-metals from which Frankland deduced the results which constituted the kernel of the theory of valency. He acted as pioneer in this branch more than any other man, and distinguished himself by his admirable investigations. Before him (more particularly) Bunsen had accomplished his memorable work on the cacodyl compounds, and cacodyl itself had been designated by Kolbe as arsene-dimethyl. Relying upon his own observa-

tions on the stanno-ethyl compounds, and on the behaviour of the cacodyl derivatives and other bodies, Frankland proved with convincing clearness the untenability of the theory of copula. Frankland's train of reasoning was somewhat as follows:—If we start with the latter theory, we must assume that the power of the metals to combine with oxygen is not altered by their being copulated with radicals. But facts tell against such an assumption, as is seen at a glance from the following examples:—Tin-ethyl (SnC_4H_9 ; $\text{C} = 6$) ought, according to that theory, to unite with oxygen in two proportions, but in reality it is only capable of taking up one equivalent of this element, and not two, like tin itself. Cacodyl, which is arsenic copulated with two methyls, does indeed form two oxides, from which it might be argued that the one with one equivalent of oxygen corresponded to arsenic sub-oxide, and the other with three equivalents to arsenious acid; but this hypothesis affords no explanation whatever of the fact that the latter compound is very readily oxidisable, whereas its supposed analogue cacodylic acid cannot be oxidised by any means.

These and similar contradictions were done away with by Frankland in the simplest manner, by the assumption that the so-called copulated compounds were derivatives of inorganic bodies in which oxygen had been replaced by its equivalent of hydrocarbon radicals. Stanno-ethyl oxide was explained as tin dioxide, SnO_2 , in which one equivalent of oxygen was replaced by ethyl, and cacodyl oxide as arsenious acid, in which two equivalents of oxygen had been substituted by two methyls. Frankland then proceeded to extend this conception to other compounds in the most felicitous manner, and thus brought the laws which are shown in the composition of organic and inorganic substances into relation with the fundamental properties of the elements which these contain.

He expressed his views upon this point in the following sentences,¹ which have a claim to a special place in a history of chemistry: "When the formulæ of inorganic chemical

¹ *Phil. Trans.*, vol. cxlii. p. 417; *Ann. Chem.*, vol. lxxxv. p. 368.

compounds are considered, even a superficial observer is impressed with the general symmetry of their construction. The compounds of nitrogen, phosphorus, antimony, and arsenic, especially, exhibit the tendency of these elements to form compounds containing 3 or 5 atoms of other elements; and it is in these proportions that their affinities are best satisfied: thus in the ternary group we have NO_3 , NH_3 , NI_3 , NS_3 , PO_3 , PH_3 , PCl_3 , SbO_3 , SbH_3 , SbCl_3 , AsO_3 , AsH_3 , AsCl_3 , etc.; and in the five-atom group, NO_5 , NH_4O , NH_4I , PO_5 , PH_4I , etc. Without offering any hypothesis regarding the cause of this symmetrical grouping of atoms, *it is sufficiently evident, from the examples just given, that such a tendency or law prevails, and that, no matter what the character of the uniting atoms may be, the combining power of the attracting element, if I may be allowed the term, is always satisfied by the same number of these atoms.*"

In this way was set up the doctrine that a varying, but at the same time, within certain limits, definite saturation-capacity appertains to the atoms of the elements. For the ones which have just been named this was expressed by the numbers 3 and 5; Frankland did not assume any higher stage of saturation for them. By this treatise of his, so rich in ideas and facts, he opened up a new field in theoretical chemistry, which, assiduously cultivated ever since, has served both as the centre- and the starting-point for all chemical investigations. Under the influence of the theory of valency all theoretical chemical views thenceforth developed themselves, as will be clearly seen from the following sections. The happy interpretation of the constitution of the so-called copulated compounds was the immediate cause of this great advance, in so far that Frankland proved copulation to be a consequence of saturation-capacity.

After the definite valency of particular elements had been established by Frankland, it might have been imagined that every chemist could have deduced for himself the saturation-capacities of other elements from their behaviour. Frankland's pioneering work did not, however, produce fruit with such rapidity. How slowly his views found acceptance

among chemists is proved by a paper of Odling's, published in 1854, and entitled *On the Constitution of Acids and Salts*.¹ The latter chemist still adhered firmly to the type theory. He argued that salts and acids, especially those containing oxygen, can be referred to the simple or multiple water type in such a way that the hydrogen of the latter is partially or completely substituted by elementary or compound radicals of definite *replaceable value*. This latter term was used by Odling to express what Frankland had done by the word *atomic*. Iron and tin had, according to Odling, two replaceable values, whose magnitudes he indicated by the dashes which have since then been so largely employed, thus: Fe'' and Fe''', Sn' and Sn''. Thus far he followed Frankland's conception of the saturation-capacity of the elements. For the polybasic acids he accepted Williamson's views, in that he assumed in them oxygenated radicals of definite replaceable value, which were introduced into the type (H₂O)_n. Just as sulphuric acid was built up on the double water type by the entrance of the diatomic radical SO₂, so he derived phosphoric and arsenic acids from the triple water type (3H₂O) by introducing the atomic groups (PO)''' and (AsO)'''; while in the carbonates the radical CO, with a replaceable value of 2, was assumed, and so on. But mischievous obscurations now began to be mixed up with this. As a result of his one-sided typical conception, Odling did not hesitate to assume that the diatomic radical SO₂ acted as monatomic in dithionic acid,² and the diatomic radical CO as monatomic in oxalic acid; and this last (for example) he referred to the double water type, thus: (CO) $\frac{(CO)'}{H^2}$ } 2O''. But, with all this, Odling deserves credit for being instrumental in causing a constant replaceable value to be ascribed to particular elements, to hydrogen and oxygen in especial, whereby the atomic weights of those two latter served as standards for fixing the

¹ *Journ. Chem. Soc.*, vol. vii. p. 1.

² This he formulated:—(SO₂)'(SO₂)' } 2O.

replaceable values of other elements and compound radicals. Williamson thereafter aided most materially to clear up the meaning of Odling's formulæ, and to bring about a more intelligent conception of the constitution of chemical compounds.¹

The utterances of Wurtz² and of Gerhardt³ upon the saturation-capacity of the nitrogen atom also showed that Frankland's ideas acted but slowly; for the last named had expressed himself on this point in almost exactly the same sense three years previously. In many cases chemists were content merely with the notion of compound radicals, without investigating the influence of the contained elements upon the saturation-capacities of these complexes; this applied in an especial degree to the radicals composed of carbon and hydrogen, with whose replaceable value (that of the radicals) various eminent investigators occupied themselves.

The Recognition of the Valency of Carbon.

A considerable time elapsed before any definite utterance was made with regard to the valency of the carbon of alcohol radicals—the *organic element* in the true sense of the term. Instead of deducing this fundamental property from its oxygen compounds, carbon monoxide and dioxide, a more tedious method was adopted; it was the investigation of carbon-containing radicals which led to the final solution of this question. Among the researches which were of effective service here, we must first mention that by Kay,⁴ made at Williamson's suggestion, upon "tribasic formic ether"; this compound, which resulted from chloroform and sodium ethylate, was regarded as a derivative of three atoms of ethyl alcohol, in which the three atoms of basic hydrogen had been replaced by the "tribasic radical of chloroform, CH." Ranking alongside of this important

¹ Cf. *Journ. Chem. Soc.*, vol. vii. p. 137; or *Ann. Chem.*, vol. xci. p. 226.

² *Ann. Chim. Phys.* (3), vol. xliii. p. 492 (1855).

³ *Traité de Chimie*, vol. iv. pp. 595 and 602 (1856).

⁴ *Journ. Chem. Soc.*, vol. vii. p. 224.

piece of work came that of Berthelot upon glycerine.¹ Aided materially by Wurtz's expositions, Berthelot characterised this compound as a triatomic alcohol, since he assumed in it a *tribasic radical*, C_6H_5 ($C \equiv 6$), replacing three atoms of hydrogen in the triple water type. To the alkyls which took the place of three atoms of hydrogen, diatomic ones were soon added, ethylene being designated as such an one by H. L. Buff.² The brilliant discovery by Wurtz of the first known diatomic alcohol, glycol,³ served as a corroboration of this view.

Chemists were, it is true, upon the track of the cause of the different replacing values of those radicals $(CH)''$, $(C_6H_5)'''$, and $(C_2H_4)''$, for we find utterances by Gerhardt and Wurtz to the effect that ethylene was dibasic, because one atom of hydrogen had been withdrawn from the monobasic ethyl, and glyceryl tribasic, because it contained two atoms of hydrogen less than the corresponding propyl. But no one had attained to a complete explanation of these radicals; their saturation-capacities had never been distinctly referred back to that of carbon.

In a paper entitled, *Ueber die Konstitution und die Metamorphosen der chemischen Verbindungen und uiber die chemische Natur des Kohlenstoffs* ("On the Constitution and Metamorphoses of Chemical Compounds, and on the Chemical Nature of Carbon"),⁴ which was published in 1858, Kekulé drew the following nearly allied conclusion. He applied to carbon what had already for a long time been recognised with regard to other elements, — to nitrogen and its chemical analogues in the first instance. The reasons given by him for carbon being tetravalent are contained in the following sentences: — "If we look at the simplest compounds of this element, CH_4 , CH_3Cl , CCl_4 , $CHCl_3$, $COCl_2$, CO_2 , CS_2 , and CHN , we are struck by the fact that the quantity of carbon

¹ *Ann. Chim. Phys.* (3), vol. xli. p. 319.

² *Ann. Chem.*, vol. xvi. p. 302.

³ *Comptes Rendus*, vol. xliii. p. 199.

⁴ Couper, too, independently of Kekulé, and shortly after the appearance of the paper just cited, expressed the view that the atom of carbon was tetravalent (cf. *Comptes Rendus*, vol. xlvi. p. 1157).

which is considered by chemists as the smallest amount capable of existence—the atom—always binds four atoms of a monatomic or two of a diatomic element, so that the sum of the chemical units of the elements combined with one atom of carbon is always equal to four. We are thus led to the opinion that carbon is tetratomic.” This train of thought is almost the same as that which led Frankland to deduce the tri- and penta-valence of nitrogen, phosphorus, arsenic, and antimony,¹ the latter having also arrived at the saturation-capacities of these elements from a study of their simplest compounds. It follows from this that the above utterance of Kekulé cannot be regarded as a thoroughly original achievement or scientific feat, all the more since the tetravalence of carbon had already been recognised both by Kolbe and Frankland, and especially since it formed the basis of the latter’s statements upon the constitution of organic compounds.² In curious contrast with the high value which most chemists have placed upon

¹ Cf. p. 307.

² Cf. Kolbe’s publication entitled *Zur Entwicklungsgeschichte der theoretischen Chemie* (“Contribution to the History of the Development of Theoretical Chemistry”), Leipzig, 1881, p. 26 *et seq.*, especially p. 33. Others, too, have claimed for Kolbe the merit of being the first to perceive the tetravalence of carbon, e.g. Blomstrand, who thus expresses himself in his *Chemie der Jetztzeit* (“Chemistry of the Present Time”), p. 110: “No other chemist can lay the same claim as Kolbe to being regarded as the originator of the doctrine of the saturation-capacity of carbon. Alongside of him must be placed Frankland, whose uninterrupted researches, conceived and carried out with equal felicity, continually furnished new supports in aid of the doctrine mentioned above,—a doctrine which comprises in itself everything that relates to saturation, and which has found in Kolbe’s carbonic acid theory by far its most important application.” A. Claus (*Journ. pr. Chem.* (2), vol. iii. p. 267) has written in a similar sense. Kekulé is therefore not justified in claiming for himself the merit “of having introduced the idea of the atomicity of the elements into chemistry” (“den Begriff der Atomigkeit der Elemente in die Chemie eingeführt zu haben”), (cf. Kekulé, *Ztschr. Chem.* for 1864, p. 689). This was without doubt primarily due to Frankland, who expresses himself clearly and unequivocally on the point in his *Experimental Researches* (1877), p. 145, as follows: “This hypothesis which was communicated to the Royal Society in the second of the following papers” (cf. p. 307 of this book), “on 10th May 1852, constitutes the basis of what has since been called the doctrine of atomicity or equivalence of elements; and it was, so far as I am aware, the first announcement of that doctrine.”

this service of Kekulé's is the depreciatory way in which he talks of it himself.¹

Kekulé's service in this point must be sought for in the fact that he endeavoured to get at the root of the problem as to how two or more carbon atoms combine with one another, and how their mutual affinities are satisfied. The immediate result of these speculations was the doctrine of the "linking of atoms" (*Verkettung der Atome*) in chemical compounds. Indirectly, Kolbe's and Frankland's views had a most material share in developing this crowning edifice of the structure theory.

¹ Thus Kekulé says, at the close of his above-mentioned treatise, p. 109: "Lastly, I feel bound to emphasise the point that I myself attach but a subordinate value to considerations of this kind. But since in chemistry, when there is a total lack of exact scientific principles to go upon, we have to content ourselves for the time being with conceptions of probability and expediency, it appeared appropriate that those views should be published, because they seem to me to furnish a simple and tolerably general expression precisely for the latest discoveries, and because therefore their application may perhaps conduce to the finding out of new facts."

DEVELOPMENT OF CHEMISTRY UNDER THE INFLUENCE OF
THE DOCTRINE OF VALENCY DURING THE LAST THIRTY
YEARS.

The chemical atomic theory had been in existence for nearly fifty years before the natural inference was drawn from it with sufficient exactitude that each elementary atom possesses a definite saturation-capacity, and that this is expressible in some cases by a constant factor, but in most cases by a varying one. In recognising this a great advance was made,—an advance which showed itself in the fact that, after the founding of the valency theory by Frankland, a more definite conception came to be held with regard to the chemical constitution of inorganic, and more especially of organic compounds. From thenceforth continuous efforts were made to solve this problem, first recognised in its fullest signification by Berzelius, by the aid of the ideas which Frankland had either himself expressed or had induced in others. Chemists endeavoured, by breaking up compound bodies (in part actually, and in part on paper only) and distributing the elementary atoms according to their supposed saturation-capacities, to work out the mutual relations of these ultimate constituents. In this way there shone forth from valency a light which now illumines the whole field of chemistry.

The theory of the linking of atoms was considered by most chemists as the necessary result of the idea that a saturation-capacity (with respect to other elements), expressible by figures, belonged to the atoms of each individual element. With the development of this view, in organic as well as in inorganic chemistry, many brains have been busily engaged. The idea of a definite saturation-capacity for each element has formed a necessary aid in the solution of numerous important points which have come up during this period, *e.g.* the question of the nature of valency, the reasons for many cases of isomerism hitherto un-

explained, etc., and it still remains an indispensable guide in all scientific chemical investigations.

Beginnings of the Structure Theory—Kekulé and Couper.

The theory of types, according to which all organic compounds were referred to a few simply constituted bodies, had been rendered objectless by Frankland's conception of that property of elements which we now term valency. The types now presented themselves as hydrogen compounds of mono-, di-, tri-, and tetra-valent elements. Had Frankland's ideas at once received the attention which they merited, the detailed development of the theory of types, as given by Gerhardt in the fourth volume of his text-book, could have been entirely dispensed with.

Out of Frankland's idea of saturation-capacity there grew the further notion that the elementary atoms were combined among themselves by one or more affinities, according to their nature, and that a disappearance of individual affinities took place as the result of this. This idea was first given out by Kekulé, and shortly after by Couper, in the treatises already referred to (in 1858). These therefore contain the beginnings of the structure theory.¹

After having deduced the "tetratomicity" of carbon from the composition of a number of simple compounds of that element, Kekulé expressed himself upon the constitution of compounds which contain more than one atom of carbon as follows: "In the case of substances containing several carbon atoms we must assume that some of the (other) atoms at least are held bound by the affinities of the carbon atoms, and that the latter are themselves united together, whereby a part of the affinity of the one is

¹ The term "structure" (*Struktur*) was first introduced by Butlerow (*Ztschr. Chem.* for 1861, p. 553); through it he quite unintentionally awakened the erroneous idea that the actual spacial arrangement of the atoms could be arrived at by the aid of the above hypothesis.

² *Ann. Chem.*, vol. cvi. p. 154.

necessarily bound by an equally large part of the affinity of the other."

"The simplest and therefore the most probable case of such a combination (*Aneinanderlagerung*) of two carbon atoms is that in which one affinity of the one atom is tied by one affinity of the other. Of the four affinity-units of each of the two carbon atoms, two are thus taken up in keeping both atoms together; six consequently remain over, to be available for atoms of other elements."

Here, therefore, there was set up the hypothesis that the carbon atoms join together,¹ and lose in consequence a portion of their affinities. Starting with the assumption that more than two atoms of carbon can coalesce in the same manner, Kekulé generalised this particular case by establishing the value $2n + 2$ for the saturation-capacity of the complex C_n . He did not, however, remain stationary at this point, but represented further that "a more compact combination of the carbon atoms" might be assumed in other organic compounds, *e.g.* benzene and naphthalene. As the "next most simple coalition of carbon atoms" he conceived the case of the mutual interchange of two affinity-units. The relations, too, of other polyvalent elements to the carbon atoms were taken into account by him, and he gave illustrations to show that these were bound either by all their affinities or by a portion of them to the affinities of the carbon.² The main features of the doctrine of the "Linking of Atoms" (*Bindung der Atome*) were contained in those sentences of Kekulé's.

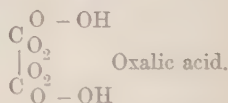
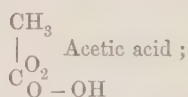
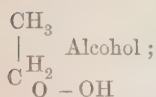
Couper³ arrived, independently of Kekulé, at similar views with respect to the mutual linking of several carbon atoms. Being definitely of opinion that Gerhardt's doctrine of types did not satisfy the claims required by a theory, he made the attempt to get at the constitution of chemical compounds by falling back upon the elementary atoms. He

¹ *Sich aneinander lagern.*

² Cf., for instance, *Ann. Chem.*, vol. cvi. p. 155.

³ *Comptes Rendus*, vol. xlv. p. 1157; *Ann. Chim. Phys.* (3), vol. liii. p. 469.

laid stress upon the point that, in addition to the affinity proper (*Wahlverwandtschaft*), the degree of that affinity (*Gradverwandtschaft*) of the small particles came into play in the formation of chemical compounds. For the atom of carbon the highest power of combination was expressible by the number 4. In general he adopted Frankland's doctrine of the varying saturation-capacities of the elements. Couper further laid great emphasis upon the capacity of the carbon atoms to unite with one another, and this in such a manner that a part of their own individual powers of combination were thereby neutralised. This linking of the atoms he illustrated by bars drawn between the chemical symbols of the combining particles; he thus laid the foundation of the so-called "structural formulae."¹ The following examples will serve to illustrate this:—



Both Kekulé and Couper expressed with absolute distinctness the axiom that the "atomicity of the elements" was to be made use of for arriving at the constitution of chemical compounds. The idea of the term "atomicity" had without any doubt been introduced by Frankland six years previous to this. The further development of the above axiom and its utilisation in the theory of the linking of atoms was worked at mainly by Kekulé, and in the succeeding years also by Butlerow and Erlenmeyer.

Before an absolutely certain knowledge of the atomicity or, better, the valency of the elements could be attained, perfect clearness has to be arrived at with respect to the magnitudes of the atomic weights; in particular, the distinction between the atom and equivalent of polyvalent elements had to be clearly grasped. That was, however, by no means

¹ Wurtz manifestly forgot Couper's paper in the *Ann. de Chim. et de Phys.*, for he took credit to himself as being the first to make use of these linking-bars (see his *Atomic Theory*, fourth English edition, p. 214, note).

the case at this time. In writing the formulæ of chemical compounds, most chemists employed Gmelin's equivalents from force of habit; but, in making use of these, the true chemical values of the atoms remained indistinct, and only became apparent after the conversion of the equivalents into atomic weights. For instance, the functions of the simple atoms C and S were ascribed to the double equivalents C_2 and S_2 in the formulæ employed by Kolbe, while for hydrogen, chlorine, nitrogen, and other elements, the equivalents were identical with the atomic weights.¹ And the disorder was increased by many chemists, Couper among the number, giving to carbon its correct atomic weight (12), while retaining the equivalent (8) for oxygen. It is true that Gerhardt had already attempted to bring order into the prevailing confusion, but his mode of procedure had not been logical enough.²

Thanks to the efforts of the Italian chemist Cannizzaro, a way was prepared in 1858 for the clearing up of this unsatisfactory state of matters, although those efforts only came slowly into recognition. It was he who, by his criticism in a paper entitled *Sunto di un Corso de Filosofia Chimica* ("Outlines of a Course of Chemical Philosophy"),³ threw light upon the methods employed for arriving at the relative atomic weights of the elements. He recognised, as especially reliable, the deduction of these values from the vapour densities of chemical compounds,—a method now in universal use. And he further showed to what extent the specific heat of the metals might be regarded as a trustworthy aid in the determination of their atomic weights, wrong values for many of these having come to be accepted as the result of Gerhardt's statements.

After the correct atomic weights of the elements had been established in this way, it became possible to build up

¹ The meaning of this is at once apparent if we take Kolbe's old formula for acetic acid, $C_2H_3.C_2O_2.OHO$, and convert it into our present formula, $CH_3.CO.OH$, by changing the double atoms C_2 and O_2 into the single ones C and O.

² Cf. p. 274.

³ *Nuovo Cimento*, vol. vii. p. 321

the doctrine of the chemical values of the elements from a more general point of view than before. First it was applied to the compounds of carbon, whose constitution became the subject of the most ardent investigation. Kekulé in his text-book (begun to be published in 1859), and Butlerow and Erlenmeyer in various papers, and subsequently in text-books also, endeavoured to explain the connection existing between the elementary atoms within the molecules, by setting out with the conception that a definite atomicity appertained to each element; carbon, hydrogen, oxygen, and nitrogen came primarily into question here.

Butlerow was the first to express himself clearly upon the principle which underlay these efforts, and, with this, upon the nature of the Structure Theory (which received its name from him).¹ We must premise here that he took up his position on the valency theory founded by Frankland, according to which many of the elements possess a varying saturation-capacity. Butlerow defined the *structure* of a chemical compound as the "manner of the mutual linking of the atoms in a molecule;" he decisively rejected the idea that it afforded any information as to the position of the individual atoms in space. He advanced the opinion that the chemical character of a compound depended first upon the nature and quantity of its elementary constituents, and then upon its chemical structure. The latter had, to his mind, but one meaning; he could not agree with Gerhardt that several rational formulae might be proposed for one and the same chemical compound, one formula only appearing possible to him.

The more that the former adherents of the type theory came to feel the necessity of abandoning it, and, free from the yoke of this doctrine, of basing all considerations with respect to chemical constitution upon the "atomicity" of the elements, the more definitely ought the views upon the nature of this property of the elements to have shaped themselves. The conclusion, deduced from numerous experiments, that the atoms of certain elements show a con-

¹ *Ztschr. Chem.* for 1861, p. 549 *et seq.*

stant combining value and the atoms of others a varying one, came at that time into opposition with the opinion that this capacity of the elements was invariable.

Controversies respecting constant and varying Valency of the Elements.

Frankland, the originator of the doctrine of the saturation-capacity of elementary atoms, held aloof from the lively discussions to which it gave rise, more especially after the year 1870. This may perhaps account for his service in developing such an important doctrine having been forgotten by many chemists, and precisely by those who have taken the most active share in the above discussions.¹ About the year 1860 Frankland's views regarding a saturation-capacity peculiar to the elements, which, under certain circumstances, might be a varying one, were accepted by most chemists of standing. Even so early as 1856 Gerhardt had stated in his text-book that nitrogen was sometimes triatomic, sometimes pentatomic,—a view which coincided exactly with that of Frankland. Wurtz, Williamson, and Couper also held this opinion, and not for nitrogen and its analogues alone, but also as being characteristic of many other elements; that Kolbe likewise agreed with Frankland on this point has been stated already. In the assumption that a constant valency was characteristic of a few elements and a varying one characteristic of many more, Kolbe merely saw another expression for the law of multiple proportion; this conception, as corresponding with facts, he considered necessary, because nothing was known of the real cause of valency.

This view, then, which had so many observations to support it, led to the conclusion that each element possessed a

¹ Kekulé, for instance, in the section of his text-book entitled, *Historische Entwicklung der Ansichten über die Constitution der organischen Verbindungen* ("The Historical Development of the Views regarding the Constitution of Organic Compounds," vol. i. p. 59 *et seq.*), makes no mention of Frankland's opinions which laid the foundation of the doctrine of valency, while he goes into the type theory in detail.

maximum saturation-capacity, but that lower stages of saturation might coexist along with this ; Kolbe had expressed himself in this sense so far back as the year 1854.¹ Towards the beginning of the sixties, several chemists who took an active part in developing the structure theory gave utterance to the same opinion in a more definite manner. Erlenmeyer, in particular, maintained in various papers,² and afterwards in his *Lehrbuch der organischen Chemie*, that each element possesses a maximum valency, or that each is furnished with a definite number of *Affinivalenzen* or affinity-points (*Affinitätspunkten*), only part of which, however, are in many cases combined with the affinity-points of other elements. In ammonia, for instance, only three of the five equivalents of the nitrogen atom come into play, while in chloride of ammonium all five are satisfied. Following this out, Erlenmeyer distinguished between saturated and unsaturated compounds. Strictly speaking, this is nothing else than Frankland's view.

At about the same time a lively discussion with respect to the atomicity of the elements went on between Wurtz and Naquet³ on the one hand, and Kekulé⁴ on the other. The two former declared for the assumption of a varying valency in the case of many of the elements, while Kekulé, on the other hand, expressed his opinion more definitely than before that the "atomicity of the elements is a fundamental property of the atoms, quite as unalterable as their atomic weights."

In order to confirm this theorem of absolute or constant valency, and to reconcile it with conflicting facts, Kekulé was obliged to have recourse to hypotheses which laid themselves strongly open to criticism. A few examples may be given here to illustrate his view of the valency of each element being constant. According to him, nitrogen and its chemical analogues acted only as trivalent, sulphur, like

¹ Cf. *Lehrbuch der organischen Chemie*, vol. i. p. 22.

² *Ztschr. Chem.* for 1863, pp. 65, 97, and 609 ; for 1864, pp. 1, 72, and 628.

³ *Ibid.*, p. 679.

⁴ *Ibid.*, p. 689 ; *Comptes Rendus*, vol. lviii. p. 510.

oxygen, only as divalent, and chlorine, bromine, and iodine as monovalent. In order, therefore, to explain the constitution of compounds, in which, upon the assumption of a varying valency, the elements just named had a higher saturation-value than he assigned to them, Kekulé had to presuppose a fundamental difference as existing between compounds of one and the same element. To his first hypothesis of absolutely constant valency he added the further one, that those compounds, in which the elements are present in their supposed normal values, are distinguished from the others by a more compact structure; the former he termed atomic, and the latter molecular compounds. The components of the latter, *e.g.* ammonia and hydrochloric acid in salmiac, phosphorus trichloride and chlorine in phosphorus pentachloride, were, according to his view, held together by forces of another kind to those which acted in the atomic compounds. In order to give expression to the looser connection between the molecules of those substances, he placed their components dualistically alongside of one another in writing the formulae; thus he gave $\text{PCl}_3 \cdot \text{Cl}_2$ as the formula of phosphoric chloride, and $\text{H}_3\text{N} \cdot \text{H}_2\text{S}$ as that of ammonium hydrosulphide. He would not admit a variation in the saturation-values of nitrogen and phosphorus in compounds like those just named.

Other chemists were thus justified in asking what his grounds were for assuming such a distinction between the forces by which chemical constitution was conditioned; for, in both kinds of compounds the same atomic laws held good. Kekulé regarded the breaking up of compounds into their components at a somewhat high temperature as a criterion of their being *molecular compounds*, while *atomic compounds* were those which could be converted into the gaseous state without decomposition. But this distinction between the two categories could not be maintained in the face of known facts; it soon became evident that such an artificial partition only served to introduce confusion and bring about contradictions which were irreconcilable.

This theory of the constant valency of the elements

could not therefore long withstand the critical examination to which it was subjected by Kolbe,¹ and more especially by Blomstrand,² not to mention others. The known facts could not by any possibility be brought into accordance with the assumption of saturation-capacity being invariable, and this helped more than anything else to cause the theory to be abandoned by its most zealous adherents. How, for instance, could the existence and behaviour of the organic ammonium bases, the sulphones and sulphoxides, perchloric and periodic acids, and many other compounds be explained by the aid of the above hypothesis? Other weighty arguments have recently been brought forward which must be regarded as incompatible with those urged shortly after the setting up of Kekulé's theory; to take compounds of one element only, we may refer here to the discovery of the isomeric triphenyl-phosphine oxides, in one of which the phosphorus must be pentavalent, and also to the proof given of phosphorus pentafluoride existing in the gaseous state. Such facts are not to be reconciled with the assumption of phosphorus being only trivalent.

We may assert that in the course of the last twenty-five years most chemists have adopted the opinion that the atoms of the elements possess a varying saturation-capacity, varying according to the conditions. The idea prescribed as essential at the time the theory of an unchanging valency was set up, viz. that this was a fundamental property of atoms, may be fully recognised without our being thereby

¹ Cf. *Journ. pr. Chem.* (2), vol. iv. p. 241.

² In his work, *Die Chemie der Jetztzeit*, Blomstrand has gone carefully into the doctrine of the saturation-capacity of the elements, and by his comprehensive treatment of the question has materially lightened the labours of other critics as to the share taken by different workers in its development.—C. Wilhelm Blomstrand, born in 1826, has filled the chair of chemistry in the University of Lund in Sweden since 1854. His eminent researches in various branches of mineralogical and also of organic chemistry are distinguished by their thoroughness, and show the influence of Berzelius, whose doctrines Blomstrand has endeavoured, in his book mentioned above, to reconcile and bring into close connection with the more recent views. From the electro-chemical basis, in especial, he has been able to throw light upon the valency question, and to gain for it new points of view.

forced to the conclusion that the valency of the elementary atoms must therefore be constant.

In connection with these weighty discussions upon the nature of valency, reference may be made here to a problem nearly related to it, which has given rise of recent years to frequent debate, and also to important experimental work, viz.—the question has been raised as to whether the individual affinity-units or valencies of one element are alike or different. If we only took into consideration some particular facts, such as the dissimilar functions of the two atoms of oxygen or sulphur in carbonic acid and carbon disulphide respectively, we might be inclined to favour the assumption of a difference in two affinities of the carbon atom with respect to the other two. But the numerous investigations which have been made by Popoff, Schorlemmer, L. Henry, Röse, and others, with the object of deciding this point so far as regards carbon, have led to the conclusion that its four affinities are alike.

The equality or inequality of the affinities of the sulphur and nitrogen atoms is still undecided, in spite of many facts bearing on the point having been collected together; among other researches we may mention the work of Krüger, which appeared to prove a difference in the valencies of sulphur, while that of Klinger and Maassen led to the opposite conclusion. The remarkable isomerism in the derivatives of hydroxylamine, which has been worked out by Lossen, seems quite compatible with the assumption of the affinities of nitrogen being different; more recent researches by Lossen, Beckmann, and Behrend, however, point to another solution of the question.

The main directions which chemical investigation has taken, since these discussions with regard to valency came up, are characterised by the endeavour to gather from the chemical behaviour of compounds an insight into their constitution, by the aid of the assumption that the elements have a definite saturation-capacity; while at the same time efforts are being made to arrive at the mutual relations between the physical properties of compounds and their constitution as determined

by chemical means. To this problem, which has only recently been assiduously attacked, although it has been projected for a long time, an analogous one has been added, viz. the clearing up of the connection which obviously exists between the relative atomic weights of the elements and their chemical and physical properties.

The further Development of the Structure Theory—The chief Directions taken by Organic Chemistry during the last twenty-five Years.

At a first glance it strikes one as strange that organic chemistry in particular should have been made the field for speculations as to the composition of chemical compounds, speculations which had the valency theory as their basis. The reason for this preference is undoubtedly to be sought for in the peculiarity of that element which is never wanting in the so-called organic compounds, carbon, even if we allow for the fact that it was from compounds of carbon—the organo-metallic ones—that the idea of the saturation-capacity of elements developed itself.

From the tendency of the atoms of carbon to unite with one another according to different degrees of affinity (*Gradverwandtschaft*), i.e. by the interchange of one, two, or three affinities, the production of the variously composed carbon compounds could be explained without difficulty. The addition of elements like hydrogen, oxygen, sulphur, nitrogen, and chlorine to the complexes of carbon atoms was rendered intelligible in a similar manner by assuming that the individual affinities of the elements named were satisfied by a like number of affinities of carbon. The combination of the carbon atoms among themselves or with other elementary atoms, as illustrated in this way, was termed "linking" (*Verkettung*). From thenceforth the adherents of the structure theory came to grasp more clearly the problem of chemical investigation. They sought to combine the atoms of the various elements in question suitably with one

another, according to their saturation-capacities, directing their efforts mainly to investigating the structure of the compounds of carbon, since inorganic substances, as being of much simpler composition, seemed to offer few or even no difficulties to the application of the above principle. The conceptions thus gained of the structure of organic substances were then tested with more or less minuteness by actual experiment, with the object of seeing whether the modes of formation and decomposition of the compounds in question, and their chemical behaviour generally, agreed with the theoretical hypotheses.

The readiness with which many chemists took to the construction of formulæ which were meant to express the mutual relations existing between the atoms of a compound, *i.e.* the structure of the latter, may in some cases have given rise to the belief that by the aid of such symbols an insight into the actual arrangement of the atoms in space might be obtained. Some investigators of eminence may have incited to such daring hopes and expectations by indistinct modes of expression and by unhappily chosen comparisons and illustrations. In the minds of younger chemists, in especial, it was easy for erroneous ideas regarding the problems of chemistry to effect a lodgment. We may recall here that Kekulé spoke of the carbon atoms as sliding over and adhering to one another,¹ and of the *other* side of a molecule, etc.; that in his text-book he brought forward graphic formulæ, in which the elementary atoms have different forms according to their saturation-capacities; and, further, that the smallest particles of an element have been pictured by Naquet and Baeyer as being furnished with small hooks by which they catch hold of one another. Metaphors such as these tended, at any rate, to an over-estimation of the capabilities of the structure theory.

The more prudent advocates of the latter, with Butlerow at their head, dissented all along from the idea that such formulæ could furnish any picture of the arrangement of

¹ "Von einem Zusammenschieben oder Aneinanderleimen der Kohlenstoffatome."

the atoms in space. On the other side Kolbe, in particular, protested with all his critical acumen and polemical force against such exaggerations, which might easily lead to error. He remained true to the point of view which he had laid down in 1854,¹ believing that no clear conception could ever be arrived at as to how the atoms of a compound were thus arranged.

*Constitution of Organic Compounds according to the
Structure Theory.*

Although the structure theory was unable to realise the highly-pitched expectations which aimed at a knowledge of the spacial arrangement of the atoms, still it possessed a great practical value. The development of organic chemistry since the middle of the sixties shows, in fact, that, through the aid of the structural hypothesis, the discovery of new modes of formation and decomposition of compounds, the recognition of the relations existing between various classes of bodies, and, especially, the interpretation of the constitution of numerous organic substances, became possible. Kekulé's theory of the aromatic compounds (see below) forms the most striking proof of this.

The working out of the constitution of the so-called *saturated* compounds offered fewer difficulties than that of the compounds poorer in hydrogen,—the *unsaturated* ones. Kekulé was the first to express the definite opinion that in all fatty compounds the carbon atoms were united to one another by an affinity of each, a point which might have been deduced from Couper's and also from Kolbe's rational formulae, had the equivalents used by them been converted into the atomic symbols. The expositions given by Kekulé and also by Erlenmeyer, Butlerow, Claus, and others, in text-books of organic chemistry and occasional papers, with regard to the constitution of such compounds, soon became the common property of nearly all chemists.

¹ *Lehrb. d. organ. Chemie*, vol. i. p. 13.

More difficult was the question—What was the function of the carbon atoms in organic compounds poorer in hydrogen? With respect to the constitution of these, Kolbe, Couper, and Wurtz had already expressed the view that in them—*e.g.* ethylene, acrylic acid, acetylene, etc.—one or several atoms of carbon acted as divalent. Kekulé hesitated at first between two opinions. He was, on the one hand, inclined to assume a “more compact,” *i.e.* a double or treble, linking of particular pairs of carbon atoms in the substances in question; while, on the other, his experimental researches upon unsaturated organic acids led him to prefer the idea that the affinities of certain of their carbon atoms were not completely saturated, and that these therefore show gaps (*Lücken*), by means of which the capability of further combination which such compounds possess can be explained. The latter of the two views coincided in the main with the one mentioned above, in which divalent carbon atoms were presupposed. Kekulé, it is true, never definitely admitted that the saturation capacity of carbon might be a varying quantity. Of recent years preference has been given to the conception of a more compact linking of the carbon atoms, although the other view does not want for eminent adherents. Thus Fittig,¹ from his work upon unsaturated acids, has expressed himself in favour of the assumption of carbon being divalent in some of these compounds.² But the question of the constitution of such compounds has not yet been conclusively answered; for numerous observations have been made which appear to

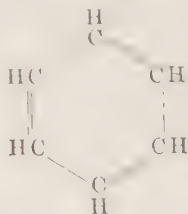
¹ Rudolf Fittig, born 6th December 1835, after working for several years on the teaching staff of the University of Göttingen, became Professor of Chemistry at Tübingen in 1869, and was called from thence to the University of Strassburg in 1876, where he still continues; the beautiful laboratory there was planned by him. His name will often be mentioned in the special history of organic chemistry, which he has greatly enriched by most admirable researches, more especially upon aromatic and unsaturated compounds. Wöhler's *Grundriss der organischen Chemie* (“Outlines of Organic Chemistry”), entirely recast by him and published under the same title, has run through numerous editions; he supplemented it in 1872 by the companion volume, *Grundriss der anorganischen Chemie*.

² Cf. *Ann. Chem.*, vol. clxxxviii. p. 95.

show that the complete solution of this problem by the aid of structural-chemical hypotheses alone is impossible.

Theory of the Aromatic Compounds.

In Kekulé's hands the structure theory has scored by far its greatest victory, in the deciphering of the constitution of the so-called aromatic compounds.¹ These were defined by him as derivatives of benzene; his first task therefore consisted in elucidating the structure of this long-known hydrocarbon, *i.e.* in explaining how the six carbon and the six hydrogen atoms were combined together. Here Kekulé took up again his previously expressed idea of a more compact linking of the carbon atoms, and discussed the possible cases of how the six in benzene could be connected together, setting out with the assumption that the carbon acted as tetravalent and the hydrogen as monovalent. While the compounds of the fatty series contained—in the language then and now current—an *open chain*, Kekulé assumed in benzene a *closed* one, and pictured each of the six carbon atoms present in the molecule as being united to two others. The structural formula which followed from this was the *hexagon*, since then so widely made use of, whose angles were formed of carbon atoms linked alternately to each other by one and two bonds, and also combined in every case with one atom of hydrogen, thus—



Kekulé and his pupils, together with many other chemists who had busied themselves with the derivatives of benzene after this view had been published, now directed their efforts

¹ *Bull. Soc. Chim.* for 1865, p. 104; *Ann. Chem.*, vol. cxxxvii. p. 129.

to comparing all the known and rapidly increasing observations bearing upon this class of bodies with the deductions drawn from the above formula, and therewith to proving by actual experiment the admissibility of the assumptions on which the formula was based. An almost boundless number of facts were thus collected together, which, taken as a whole, were found to agree readily with Kekulé's hypothesis. The first inference to be drawn from it, viz. that the six hydrogen atoms which were distributed similarly among the six carbon ones were in every respect equal to one another, was confirmed by the observation, made over and over again, that only one and the same product resulted from the replacement of any one of the hydrogen atoms of benzene by a monovalent radical or element, and never a second isomeric compound. When two or three atoms of hydrogen became substituted, the case was otherwise. From his formula Kekulé deduced the number of isomers which were then to be expected; he stated his opinion that three isomeric compounds, and not more, would result in both cases through the replacement of two or three of the hydrogen atoms of benzene by the same substituent. If two dissimilar radicals took the place of two atoms of hydrogen, the number of possible isomers was not increased; those did augment, however, to a definite number when three hydrogen atoms were replaced by two different substituents. The truth of these and of other prognostications by Kekulé has since been verified in the most brilliant manner by a vast number of observations.

This happy interpretation of the constitution of benzene shed a great light over a hitherto neglected branch of the science. Not merely the immediate derivatives of benzene, but also substances much more distantly related to it, like naphthalene and anthracene, and, more recently, phenanthrene, fluorene, and many other hydrocarbons, together with their numberless and often important derivatives, had their chemical constitution successfully investigated by the aid of Kekulé's hypothesis.

This hypothesis did not, however, completely satisfy

a number of chemists, who considered modifications in it necessary. We need not enter here into the reasons which led to such modifications, but may just mention Ladenburg's ¹ *prism formula* and Claus's ² *diagonal one*, which were brought forward by those investigators as explaining more completely than Kekulé's hexagon formula the chemical behaviour of benzene. The discussions upon this point still continue; thus, the results of a recent admirable investigation by A. Baeyer³ have, he considers, given him grounds for disputing all the above hypotheses on the constitution of benzene, while Claus⁴ maintains—and not without cause—that Baeyer's view is identical with his own.

But, notwithstanding all this, the fact must be fully recognised that Kekulé's conception, even although it by no means affords a complete picture of the constitution of benzene, has borne many and rich fruits. Through the stimulus which was given by his theory of the aromatic compounds, the work of numberless chemists with this class of substances, work extending over a long period of time, received a particular stamp of its own.

The meaning of the term *Aromatic Compounds* has of late years received a wide extension since the near relation of pyridine and quinoline and their derivatives to benzene has come to be recognised. The ardour shown in the investigation of those nitrogenous bodies, with their endless derivatives, has gone on increasing in proportion with the increasing surmise of a close connection existing between them and the vegetable alkaloids, and with the actual proof of this in some particular cases. Körner was the first to propound the important idea that pyridine may be regarded as benzene in which a methine (CH^m) is replaced by the

¹ *Ber.*, vol. ii. p. 140; also his pamphlet, *Theorie der aromatischen Verbindungen*.

² *Theoretische Betrachtungen und deren Anwendung zur Systematik der organischen Chemie* (1867), ("Theoretical Considerations and their Application to the Systematising of Organic Chemistry").

³ *Ann. Chem.*, vol. cexlv. p. 103.

⁴ *Journ. pr. Chem.* (2), vol. xxxvii. p. 455.

trivalent nitrogen atom.¹ The inferences drawn from this with respect to the derivatives of pyridine, like those deduced from the structure of benzene, have formed the subject of numberless experimental researches and theoretical discussions, which are still proceeding. Reference will be made to some of the more important results of these investigations in the special history of organic chemistry.

The efforts at gaining a clear conception—in the widest sense of the word—of the structure of benzene and its derivatives have also been of use in the case of other classes of compounds, especially for those analogous substances furfuran, thiophene, and pyrrol, which are now universally regarded as being characterised by a closed ring containing four carbon atoms. Victor Meyer's² splendid and thorough researches on thiophene and its derivatives³ have before all others led conclusively to the recognition of the analogous composition of the above substances, and also to a more precise conception of the term *aromatic compounds*. According to Meyer,⁴ it is the chemical behaviour of a substance with regard to nitric acid, sulphuric acid, bromine, and acid chlorides (in the presence of chloride of aluminium), which decides whether it has a claim to be ranked among those compounds. He lays here the greatest weight upon facts, whereas in previous determinations of the nature of this class of substances the existence of a closed ring of six carbon atoms was held to be a fundamental condition.

¹ Dewar was the first to publish this view (*Journ. Chem. Soc.*, vol. xxiv. p. 145; or *Ztschr. Chem.* for 1871, p. 117), Körner having however already given utterance to it in his lectures.

² Viktor Meyer, born 8th September 1848, after filling the post of professor of chemistry at Stuttgart and at Zurich, was called to the chief chemistry chair at Gottingen on Wohler's death in 1885, and removed from there in 1889, to succeed Bunsen at Heidelberg. His comprehensive researches upon nitro-compounds of the fatty series, upon iso-nitroso compounds, and upon thiophene are among the very first of our time, and have contributed largely to increase our knowledge of organic chemistry. The method devised by him for vapour-density determinations has become a standard one, and has also been successfully applied to the solution of important theoretical questions (*c.g.* to that of the valency of aluminium).

³ Cf. his work, *Die Thiophengruppe* ("The Thiophene Group"), Braunschweig, 1888.

⁴ *Ibid.*, p. 276.

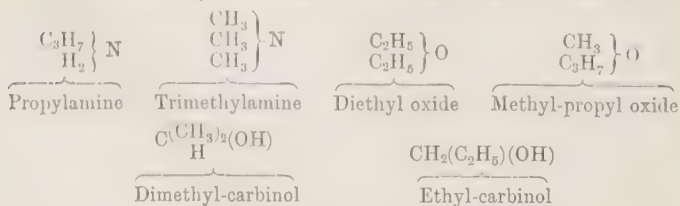
*Application of Structural-chemical Conceptions to the
Investigation of Isomerism.*

Detailed reference has already been made to the significance which the investigation of the isomeric relations of organic compounds has for the question of their chemical constitution.¹ Indeed, the efforts made during the last twenty-five years to prepare as large a number of isomers as possible, and to establish their *structure*, is a main feature of the mode in which organic chemistry has been and still is being studied. Before the derivatives of benzene had acquired that predominating interest for chemists which they afterwards came to do, the constitution of metameric substances was held to be sufficiently explained by a difference in the grouping of the atoms of the radicals. We have only to recall here the proof given of the rational composition of trimethylamine as opposed to that of propylamine: the reason assigned for the metamerism of diethyl oxide and methyl-propyl oxide; and, lastly, to think of the secondary and tertiary alcohols or acids, whose constitution was predicted with perfect definiteness before they had been discovered (*i.e.* of the metamerism of dimethyl-carbinol with ethyl-carbinol, and that of trimethyl-carbinol with propyl-, isopropyl-, or methyl-ethyl-carbinol),² etc.

To such quite satisfactorily explained cases of metamerism as these, the investigation of the aromatic compounds now added numerous others which, however, unlike the former, could not be referred back to a different grouping of the atoms in the radicals. Kekulé therefore sought to

¹ Cf. p. 237.

² A few formulæ may serve to illustrate the above cases of metamerism—



explain the like composition of various benzene substitution-products (*e.g.* of the three dibromo-benzenes, the three phenylene-dicarboxylic acids, etc.) from his conception of the structure of benzene, by assuming different *relative positions* of the substituents to one another. Such compounds were termed *position-isomers*. The question of the relative positions occupied by the entering substituents, or, as it was also called, the *determination of the chemical position* of the latter, was ardently studied from different sides, after the problem had been started by Kekulé. Among the investigations which helped in an especial degree towards the solution of this was that of Baeyer upon the constitution of mesitylene and its derivative isophthalic acid, that of Graebe upon naphthalene and phthalic acid, and that of Ladenburg on terephthalic acid. By the ingenious conclusions drawn from these and many other researches, the structure of the so-called Ortho-, Para-, and Meta-compounds was arrived at with considerable certainty. Some errors, however, did creep in here,—for instance, the wrong interpretation of the constitution of quinone from theoretical considerations, a point which gave rise to very great confusion before the mistake was finally put right.

The investigation of these metameric relations among the derivatives of benzene materially lightened that of the still more complicated phenomena among the pyridine and quinoline bases which were referable to similar causes. The metamerism of the pyridine-carboxylic acids and other derivatives, which had been predicted on theoretical grounds from conceptions as to the structure of pyridine, was beautifully confirmed later on by the comprehensive researches of Weidel, Skraup, and others; while considerations of the same kind have proved equally fruitful in the investigation of the derivatives of thiophene and pyrrol, and also of indole and other aromatic compounds.

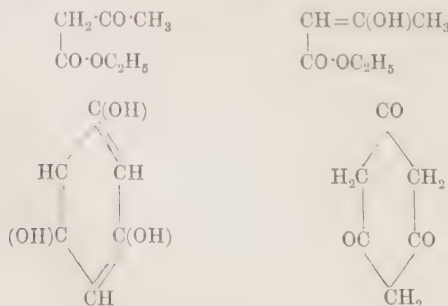
But the certainty with which the constitution of metameric substances was supposed to have been established left much to be desired in many cases. The symbols employed to express the structure of such compounds were

intended to have but one definite meaning; Gerhardt's view, that several formulæ might be indifferently used to picture the reactions of the bodies in question, was entirely abandoned. On the other hand, more organic compounds became known whose constitution could be illustrated equally well by two totally different formulæ, according to their chemical behaviour in different circumstances. Many of the reactions of aceto-acetic ether, for instance, cause us to give to it the constitution which is apparent in its name, but in others it behaves like the ether of an oxy-crotonic acid. Phloroglucin, which has been for long, and justly, looked upon as trioxy-benzene, may also be indicated, from some of its reactions, as a metameric tricarbonyl compound.¹

The constitution of these, as well as of certain other compounds, *e.g.* isatin, oxindole, carbostyryl, cyanamide, etc., is therefore capable of two explanations. Opinions are still divided among chemists who have busied themselves with this question as to which of the two possible structural formulæ is the correct one for such compounds. Baeyer distinguishes between a stable (*stabile*) modification and an unstable (*labile*) one, the latter being termed the *pseudo-form*; for isatin, *e.g.*, the formula containing hydroxyl is the stable modification, while pseudo-isatin is unknown in the free state, only derivatives of it being capable of existence.

C. Laar,² who has discussed this question minutely, applies

¹ The *tautomerism* of the above compounds is seen from the following formulæ:—



² *Ber.*, vol. xviii. p. 648; vol. xix. p. 730

the name *tautomerism* to these phenomena. A "change in combination or position of hydrogen atoms"¹ is always involved here, as is readily seen in what is doubtless the simplest case of such a tautomerism—in hydrocyanic acid. The chemical behaviour of this acid leads on the one hand to the structural formula $\text{H}-\text{C}\equiv\text{N}$, and on the other to that of $\text{C}=\text{N}-\text{H}$ (in which the carbon is divalent); in the former case the hydrogen is linked with carbon, and in the latter with nitrogen. Laar imagines oscillatory conditions within the hydrocyanic acid molecule, which cause the hydrogen atom to take up the one and the other position alternately; he therefore presupposes the simultaneous existence of both modifications. Since all cases of tautomerism depend upon a change in the linking of the atoms of carbon, nitrogen, and oxygen with respect to hydrogen, Victor Meyer has proposed to replace the above indefinite term by the better one of *desmotropism*.

Here then we have an instance of the constitution of one and the same compound being expressible by two structural formulæ, either one of them apparently as correct as the other. In another group of metamers we find just the opposite conditions, *i.e.* one and the same structural formula applying to two totally different chemical compounds of the same composition. J. Wislicenus² was the first to establish such an *identity in structure* (*Strukturidentität*) for two different substances—the fermentation- and para-lactic acids.³ The

¹ Ein "Bindungs- oder Platzwechsel von Wasserstoffatomen."

² Johannes Wislicenus, born at Klein-Eichstedt, near Querfurt in Thuringen, on 24th June 1835, became in 1885 professor of chemistry and head of the chief chemical laboratory in the University of Leipzig, after filling from 1872-85 the corresponding post at Würzburg, before which he taught at Zürich. After the death of Strecker, whom he succeeded at Würzburg, he re-edited the former's text-book of chemistry. His experimental researches, most of which have been published in the *Annalen der Chemie*, pertain almost exclusively to the domain of organic chemistry, in the special history of which we shall frequently have occasion to refer to them. The very important work which he did on the lactic acids impelled him even so early as 1873 to the conclusion that the cause of the difference between two of them must be sought for in the spacial relations of the atoms in the molecule. His quite recent speculations upon geometrical isomers are referred to above.

³ *Ann. Chem.*, vol. clxvii, p. 343.

structure theory is therefore insufficient to explain such cases of metamerism as this. Further instances of the same kind are found in crotonic and iso-crotonic, fumaric and maleïc, and mesaconic and citraconic acids. Wislicenus has designated this species of metamerism *geometrical isomerism*, and Michael, who has likewise occupied himself for a long time with the study of this branch, *allo-isomerism*.

Wislicenus¹ has attempted to explain phenomena of this kind by the aid of an hypothesis propounded by van t'Hoff and Lebel.² According to this hypothesis, which was designed with the object of explaining the optical activity of isomeric compounds, the centre of gravity of an atom of carbon is supposed as in the middle of a tetrahedron, and its four affinities as at the four corners. When two atoms of carbon become linked together, with the consequent neutralisation of one affinity of each, then van t'Hoff and, after him, Wislicenus assume that both are capable of rotating in opposite directions about a common axis; and the possibility of such rotation is supposed to cease with the double or triple linking of the carbon atoms. Wislicenus has made this hypothesis the basis of his discussions and experimental researches. An important aid to this conception is added in the supposition that, in the rotation of systems with carbon atoms linked together by one affinity of each, "specially directed forces, the affinity-energies," come into play, which regulate the spacial relations of the atoms to one another. Wislicenus believes that in those suppositions he possesses a means whereby "the establishing of the spacial arrangement of atoms in particular cases may be arrived at by experiment."

This idea of referring the causes of many cases of isomerism to the different geometrical arrangement of the atoms has at any rate had a most stimulating effect, and has led to the discovery of many hitherto overlooked

¹ Cf. *Die räumliche Anordnung der Atome in Organischen Molekülen* (Leipzig, 1887), ("The Spacial Arrangement of the Atoms in Organic Molecules"); also the *Tageblatt der Naturforscherversammlung zu Wiesbaden*, 1887 ("Journal of the Assembly of Scientists at Wiesbaden, 1887").

² Cf. van t'Hoff's pamphlet, *Dix Années dans l'histoire d'une Théorie* (1887).

relations existing between isomeric substances.¹ Since they were first brought up, the question has been much discussed whether the hypotheses indicated above are really sufficient for the explanation of such phenomena. Thus Victor Meyer² has called attention to the fact that the existence of several isomeric benzil-dioximes, proved by himself, renders necessary a modification of this view with respect to the free rotation of two carbon atoms linked together by one affinity of each. By bringing electro-chemical conceptions to their aid, he and Riecke have propounded some ingenious speculations as to "the constitution of the carbon atom."³ Closely allied to this problem we have the further ones of how the affinities of the carbon atom are directed, and of the nature of valency in general. The views expressed upon those points in the papers just mentioned, and also in one by Baeyer,⁴ can merely be recalled here.

It is impossible to give an answer to the question whether the spacial arrangement of the atoms within a molecule actually corresponds with the configurations assumed by the above-named scientists, for no proof can be furnished of the absolute correctness of these conceptions. The expectations raised by them—of obtaining a deeper insight into the mode in which the atoms are arranged in a compound—were doubtless pitched too high. Criticism has not as yet sufficiently tested the admissibility of the "spacial" (stereochemical) view.⁵ Attempts have been made quite recently to establish a "stereo-chemistry" of nitrogen and its compounds;⁶ these must however be regarded as premature.

¹ We need merely refer here to the researches of Wislicenus and his pupils on the tolane dichlorides, acetylene-dicarboxylic acid, the butylenes, and many of the derivatives of crotonic acid, and to those of C. A. Bischoff on dimethyl- and diethyl-succinic acids.

² *Ber.*, vol. xxi. pp. 536 and 788 (1888). With regard to these important and brilliant researches, as well as to the more recent investigations on stereochemical questions, cf. the work of K. Auwers, *Die Entwicklung der Stereochemie* (Heidelberg, 1890) ("The Development of Stereo-chemistry").

³ *Ber.*, vol. xxi. p. 951.

⁴ *Ibid.*, vol. xviii. p. 2277.

⁵ A. Michael, *Journ. pr. Chem.* (2), vol. xxxviii. p. 6 *et seq.* (1888); vol. xl. p. 29; Anschütz, *Ann. Chem.*, vol. ccliv. p. 173.

⁶ Cf. Willgerodt, *Journ. pr. Chem.* (2), vol. xli. p. 291; Werner u. Hantzsch, *Ber.*, vol. xxiii. pp. 11 and 1243; Behrend, *ibid.*, vol. xxiii. p. 454.

*The Development of Important Methods for investigating
the Constitution of Organic Compounds.*

The above-mentioned discussions upon isomers are sufficient to show us how materially these have aided the development of organic chemistry since the subject was zealously taken in hand. Hardly any other group of phenomena has furthered the solution of the question of chemical constitution in a more lasting manner, for the attempts to establish the constitution of isomeric bodies have coincided with those to fathom the cause of isomerism. The methods followed during the last decade for investigating the rational composition of organic compounds have in great part developed themselves from others previously in use. The paths which have led towards the wished-for goals were smoothed by the preparatory labours of Liebig, Wöhler, Bunsen, Kolbe, Frankland, Dumas, Williamson, Gerhardt, Hofmann, Wurtz, and others.

Synthetic Methods.

The mode of attaining to a knowledge of the constitution of organic compounds, which had been least worked out of any, was their artificial preparation from others of simpler composition. After Wöhler had published his memorable observation on the production of urea from its elements, and had therewith furnished a complete synthesis of it, years elapsed before any further successful labours in this direction fell to be recorded. Referring the reader to the special history of organic chemistry, we need merely recall here the important discoveries during the fifties by Kolbe and Frankland, — the synthesis of acetic acid by the former, and the building up of hydrocarbons from substances of simpler composition by the latter.

The importance of synthetic research was from thenceforth recognised in an increasing degree; indeed, it was from artificial modes of preparation that the constitution of

many organic substances could first be deduced with certainty. Thus (to give only one or two instances) the rational composition of acetic acid was arrived at from its production from the methyl compounds—methyl cyanide and sodium-methyl. The constitution of hydrocarbons was inferred from their synthesis from halogen-alkyls with zinc or sodium, and that of the ketones through their formation from acid chlorides and zinc-alkyls. Light was thrown upon the true composition of the oxy-acids by their synthesis from aldehydes or ketones and hydrocyanic acid, and also from phenates and carbonic acid. And to what a wealth of synthetic reactions and discoveries of new compounds have not the sodium derivatives of certain acid ethers, *e.g.* aceto-acetic and malonic ethers, led!¹

In every section of the wide field of organic chemistry, great success has followed the application of synthetic methods; and the worth of these latter is not to be measured merely by the vast number of new compounds to which they have given rise, but by their own intrinsic value, which has shown itself in the knowledge thereby gained of the chemical constitution of organic compounds. The so-called *condensation* syntheses have proved themselves of especial value in this direction. This term “condensation” has, since Baeyer’s explanations on the subject, been employed generally for those reactions in which several similar or dissimilar molecules coalesce together, with elimination of water, in such a manner that the carbon atoms become linked to one another. A classical instance of it (observed a long time ago by Kane, but first explained by Baeyer, as above) is given in the transformation into mesityl oxide and then into mesitylene which acetone experiences under the influence of sulphuric acid. Similar reactions go on in the case of other ketones and of aldehydes, *e.g.* the condensation of acetic to crotonic aldehyde (Kekulé). Through these and other processes a bridge was thrown over the gap between the saturated and unsaturated

¹ With reference to these and other syntheses, cf. the special history of organic chemistry.

compounds, while at the same time light was shed upon the constitution of the latter. The reaction, called after W. H. Perkin sen., which depends on the condensation of aldehydes with fatty acids, likewise aided in clearing up the rational composition of unsaturated acids.

A. Baeyer,¹ in conjunction with a large number of his pupils (E. and O. Fischer, V. Pechmann, Königs, Knorr, etc.), has minutely investigated this subject of condensation in the most admirable manner, as have also Kekulé, Fittig, Ladenburg, Victor Meyer, and, in fact, almost all chemists who have occupied themselves with organic chemistry of recent years; indeed, this study seemed for a time to give to organic chemistry its most characteristic stamp. The ardour for carrying out such syntheses increased more especially after it was seen that the processes going on in plant organisms, *viz.* the formation of compounds rich in carbon from carbonic acid, water, and ammonia, were for the most part based upon condensation. The history of organic chemistry could relate many results of efforts to imitate such natural processes, or at least to prepare products which occur in the vegetable kingdom (acids, colouring matters, alkaloids, carbohydrates, etc.) from substances of simpler composition. The most important of those vegetable acids which had long been known were prepared synthetically,—oxalic acid from carbonic, succinic acid from ethylene, malic and tartaric acids from succinic, and citric acid from acetone (which,

¹ Adolf von Baeyer, born at Berlin on 30th November 1835, became a pupil of Bunsen and of Kekulé, and applied himself under the stimulating influence of the latter to organic chemistry, which he has enriched by a wealth of admirable and important work. His untiring study of condensation reactions has led him to results of the highest value, which will frequently be referred to in the special history of organic chemistry. From his laboratory there has come forth much work of a fundamental nature; we need only recall here that of Graebe and Liebermann on alizarin, and that of E. and O. Fischer on rosaniline, etc. Since 1860, in which year Baeyer became assistant professor in Berlin, he has continued energetic as a teacher,—first at the Berlin Technical College, then from 1872-75 in Strassburg, and lastly, from 1875 in Munich, where, as head of the university laboratory, which was built after his own plan, he has found a brilliant sphere of activity.

like ethylene, could be built up from its elements); further, benzoic acid from benzene, cinnamic acid from benzaldehyde, and so on. By those observations, the list of which might be extended by numerous others on the artificial formation of acids, the chemical constitution of these substances was arrived at with absolute precision. Similarly from the synthesis of vegetable colouring matters—*e.g.* alizarin, purpurin, and indigo blue,—trustworthy conclusions have been drawn with respect to their rational composition. The important problem of preparing the sugars and vegetable alkaloids artificially has been taken in hand with success,—witness the beautiful researches of E. Fischer upon carbohydrates, which have quite lately led to the artificial formation of grape sugar, and the ingenious synthesis of conine by Ladenburg.¹

One may safely express the opinion that a clear idea of the chemical constitution of these and other difficultly accessible classes of compounds, whose proximate composition has as yet been but imperfectly worked out, will only be arrived at after they have been synthetised from simpler ones of known structure. The history of the synthesis of organic compounds has already proved the truth of this axiom in numerous instances.

The chemical behaviour of organic compounds is in every case regarded as an aid of the first importance to the working out of their constitution, and has been valued accordingly ever since organic chemistry began to flourish. A short sketch only can be given here of a few of the more important methods which have been applied during these last decades, with the object of getting at the chemical constitution of organic compounds from their reactions, transformations, and decompositions.

The general principle of such methods consists, in contradistinction to that of the synthetic, in investigating the products obtained by the chemical alteration of the compounds in question, and in deducing the constitution of the latter from these. In many cases of transformation the

¹ Cf. the special history of organic chemistry.

chemist keeps his attention fixed upon particular elements or atomic groups united to carbon, the carbon framework itself undergoing no change; in many others, on the contrary, carbon is separated as carbonic acid, carbonic oxide, or even in a more complex form. For those classes of substances which are among the best investigated, special reactions have been discovered which make it possible to decide whether a hitherto unknown compound belongs to this or that group. Of recent years great attention has been paid to the refinement of such reactions. To mention only one or two important steps in this direction:—Phosphorus pentachloride, acetic anhydride, and hydriodic acid have been found of inestimable value for determining whether an organic compound contains hydroxyl, and, if so, what function that hydroxyl performs. Further, the transformation of nitro- into amido-compounds by reduction, and that of the latter into oxy-derivatives by oxidation, the conversion of cyanides into carboxylic acids, and of hydrocarbons into acids, have all become typical reactions, which, when rightly interpreted, lead very quickly to the explanation of the constitution of such bodies. Lastly, we may recall here the beautiful method of V. Meyer and E. Fischer by which the presence of the carbonyl group in a compound can be proved by means of hydroxylamine or phenyl-hydrazine. All the above and other similar reactions have for their aim the definite recognition of the *rôle* of elementary atoms or compound radicals in organic molecules, and therewith the partial solution of the constitution of the latter.

The decompositions of organic substances into others poorer in carbon, which may be made use of for deciding the same point, are legion, and will just be touched upon here, in order to illustrate the principle of the method. This plan is the direct opposite of the synthetic; while by the latter the constitution of an organic compound is deduced from that of its components, the former leads to the same conclusion through a study of the resulting decomposition-products. To give only one or two examples:—Let us recall

the important inferences drawn by Baeyer from the decomposition of derivatives of uric acid into simpler bodies; the constitution of those compounds thus arrived at by him was confirmed later on by direct synthesis. The significant researches by Frankland, Geuther, Wislicenus, and others on the modes of decomposition of aceto-acetic ether must also be mentioned, researches which, hand in hand with other synthetic ones, cleared up the constitution of the latter. Further, carbonic acid, formic acid, etc., are very often eliminated from organic compounds, whose decomposition-products thus furnish a clue to their rational composition. The changes produced by oxidation in the case of numerous substances, such as the ketones, quinoline bases, and unsaturated compounds, furnish most excellent proof of the invaluable aid given by researches of this nature towards solving the question of chemical constitution. For further details on this point, the reader is referred to the special history of organic chemistry.

*The Main Currents in Inorganic and General Chemistry
during the last Thirty Years.*

The doctrine of the saturation-capacities of the elements which has proved of such extraordinary importance for the development of organic chemistry, has not by any means found the same rapid and general application in inorganic. After Odling, so early as 1854, had applied Frankland's idea of valency to the oxides of a large number of the elements, remaining, however, at the same time enchained by the type theory (cf. p. 308), various chemists attempted by degrees, either in text-books or in their experimental researches, to engraft upon inorganic compounds the conceptions which had so quickly found acceptance with respect to the linking of carbon atoms among themselves or with other elements. The gain which arose from this was first apparent in the systematising of those compounds, which became classified into natural families according to the valencies ascribed to the individual elements. Similarity in saturation-capacity formed

the common link which held the different members of such groups together. Thus Frankland had already recognised the analogy between nitrogen, phosphorus, arsenic, and antimony, from the fact that they were all capable of acting either as tri- or as pentavalent. Alongside of carbon were ranged silicon, titanium, and zirconium, as being in the main tetravalent elements, whereas boron, which had formerly been ranked along with carbon, was seen to be trivalent and was relegated to another group. These and similar efforts to introduce clearness into the systematising of the elements, by classifying them according to their chemical values, soon led to the setting up of the important *Natural System of the Elements* (cf. p. 346).

The problem of interpreting the constitution of inorganic compounds similarly to that of organic, by getting at the relations which exist between their component elements, has not been treated with the same care as in the case of the latter. For substances of simple composition the difficulty of the point was usually under estimated: this showed itself more particularly in the arbitrary attempts at explaining the constitution of inorganic compounds on the supposition that the valencies of the elements were invariable. Thus it was often overlooked that the chemical behaviour of a substance was not in accordance with the structural formula assigned to it. Sulphur chloride, for example, was given



the formula, $\begin{array}{c} | \\ \text{S} - \text{Cl} \end{array}$, without any heed being paid to the fact

that one of its atoms of sulphur behaved quite differently from the other. And the constitution of phosphorus oxychloride could only be illustrated by the adherents of constant



valency by the formula, $\text{P} \begin{array}{l} \swarrow \text{Cl} \\ \searrow \text{Cl} \end{array}$, a formula which in-

dicated an (unproven) difference between one chlorine atom and the other two.

And how the ordinary rules were strained in order to indicate the composition of more complicated compounds!

According to Wurtz,¹ the constitution of bodies rich in oxygen could usually be explained by assuming the oxygen atoms to be linked to one another; take, for example, periodic anhydride, in which seven atoms of oxygen were linked together in a chain, with the two supposed monovalent iodine atoms at either end. This very one-sided assumption of a constant valency of the elements became, however, gradually superseded, a sounder view taking the place of such artificial explanations. But trustworthy methods of arriving at the constitution of complicated compounds are scarcely yet developed in inorganic chemistry, although in organic much has already been done in this direction.

The researches of greatest value for inorganic chemistry which have been made during the last few decades are those upon particular elements, more especially upon such as had hitherto been imperfectly or even not at all investigated. Thus the work of Roscoe on vanadium, of Marignac on niobium and tantalum, and of Zimmermann, Krüss, von der Pfordten, and others on uranium, gold, titanium, etc., have enabled those elements to be put in their proper place among the others; this of course only became possible after their chemical character had been examined with great care. The same applies to the more recently discovered elements—thallium, indium, gallium, scandium, germanium, etc., which have likewise been investigated by their discoverers in a masterly manner.

All these researches, which will be referred to again in the special history of inorganic chemistry, have had the same ends in view, viz. the establishment of the chemical character, and, in particular, of the combining relations of the element in question, and the most careful determination of its relative atomic weight. In addition to all this, an increasing value has come to be laid upon the observation of its physical properties. Such investigations upon individual elements became more systematised after it was clearly seen that a close connection existed between their

¹ *Leçons de Philosophie Chimique*, p. 157.

chemical and physical properties on the one hand and the magnitudes of their atomic weights on the other. Of course, when it came to a question of proving this intimate relation, the first thing was to determine the relative atomic weights as accurately as it was possible to do.

The efforts of many chemists had already for a long time been directed to improving as far as practicable the methods of determining atomic weights, before the importance of this question for the systematising of the elements had come to be recognised. The memorable labours of Berzelius were followed during the forties by those of Turner, Dumas, Marignac, Erdmann, Marchand, and Pérouze, and were crowned by the classical researches of Stas upon the atomic weights of oxygen, chlorine, bromine, iodine, nitrogen, sulphur, silver, etc. In Stas's case the extreme limit of accuracy was reached which was possible with the means at command. But this certainty with respect to the magnitudes of the relative atomic weights only extended to some of the elements, the values hitherto assigned to many (e.g. molybdenum, antimony, platinum, osmium, iridium, etc.) being exceedingly inaccurate. Much has, however, been accomplished in this direction of late years.¹

The Periodic System of the Elements.

Prout's hypothesis, according to which the atomic weights of all the elements stand in a simple relation to that of hydrogen, acted for a long period like a ferment, in that it gave rise to continually renewed speculations upon the connection which existed between the elements and their atomic weights. The observed fact that chemically analogous elements possessed either nearly equal atomic weights, or atomic weights separated from one another by definite numerical increments, afforded food for such theorising. During the last sixty years attention has frequently been drawn, with more or less emphasis and ability, to regularities of this kind; the discussions of the point by Dobereiner, L. Gmelin,

¹ Cf. the special history of inorganic chemistry.

Pettenkofer, Dumas, Kremers, Odling, and others may be recalled here.¹ But it is only of recent years that a systematic classification has followed from those efforts to discover a connection between the atomic weights and the natures of the elements.

In the year 1864 Newlands² in England and Lothar Meyer³ in Germany—independently of one another—arranged a number of the elements according to the magnitudes of their atomic weights,⁴ and thereby observed that while, at a superficial glance, the elements following one another showed apparently no regularity in properties, after the lapse of a certain *period* the chemical and physical behaviour of the elements now succeeding each other strongly recalled that of the previous group, in fact repeated it. The elements which resembled one another were therefore united into groups or *natural families*, and these in their turn were distinguished from the *periods*, which comprised the elements whose atomic weights lay between those of two successive members of a natural family. This attempt to classify the elements according to the magnitude of their atomic weights, and to deduce therefrom an important connection between the latter and the properties of the former, called forth at first more astonishment than recognition. Indeed,

¹ Cf. L. Meyer's *Moderne Theorien* (fifth German edition), p. 133.

² *Chem. News*, vol. xxxii. pp. 21 and 192; also Newlands' pamphlet, *The Discovery of the Periodic Law* (London, 1884).

³ Lothar Meyer, born 19th August 1830, has filled since 1876 the first chair of chemistry in the University of Tübingen, after having previously worked as an academic teacher in Breslau, Neustadt-Eberswalde, and Karlsruhe. His first experimental researches dealt with questions of physiological chemistry, but he afterwards turned his attention more to theoretical and physico-chemical problems. The outcome of this was his valuable work, *Die Modernen Theorien der Chemie* (fifth edition, 1884), which has recently been translated into English by Professors Bedson and Carleton Williams under the title, *Modern Theories of Chemistry*. The efforts mentioned above, which he made with the object of firmly establishing the periodic system of the elements, led him to make a careful review of all that had been written on their atomic weights (cf. his and K. Seubert's meritorious work, *Die Atomgewichte der Elemente aus den Originalzahlen neu berechnet*, 1883), ('The Atomic Weights of the Elements newly Recalculated from the Original Numbers').

⁴ Cf. *Moderne Theorien* (first German edition, 1864).

Newlands did not escape banter on the subject, being asked whether he would not try, with a similar result, to classify the elements according to the initial letters of their names.

After the year 1869 these beginnings were soon greatly extended and perfected by Mendelejeff¹ and Lothar Meyer,² the atomic weights of various elements having in the meantime been determined with greater accuracy than before. Mendelejeff made what was for that time the bold attempt to classify all the elements according to the magnitudes of their atomic weights, the correctness of some of which was extremely doubtful. He was thus able to show that the elements which belonged to a natural family, *i.e.* those which were chemically similar, followed one another in regular periods. In this way the elements were brought together into a *natural system*, as it was termed, in which, however, there was much that was arbitrary because of the inaccuracy of many of the atomic weights. But the fundamental idea developed by the above-mentioned investigators, *viz.* that the elements arrange themselves on the one hand into periods, and on the other into natural families, and that all their properties are periodic functions of their atomic weights, has been strengthened and verified in every direction by many subsequent investigations.

These efforts, so significant for the systematising of the elements, have led to many important deductions. Thus, in virtue of the periodic system, definite values could be assigned to the hitherto uncertain atomic weights of various elements; for each element claims a place of its own in this system and an atomic weight corresponding with this place, the magnitude of the latter being calculable within certain limits. When, for example, only the equivalent of an element was known, the atomic weight could be deduced from its behaviour and from the position thus accruing to it in the natural system, as was actually done, *e.g.*, for

¹ *Ztschr. Chem.* for 1869, p. 405; and, more fully, *Ann. Chem.*, Supplement, vol. viii. p. 133.

² *Ann. Chem.*, Supplement, vol. vii. p. 354; and also in the recent editions of his *Moderne Theorien*.

beryllium and indium. Further, a choice could be made between different definite values for one and the same element, and the more suitable one taken, to be afterwards verified, however, with the utmost care. In this way the periodic system has been applied in the happiest manner to correcting the atomic weights of molybdenum, antimony, caesium, etc.

Other conclusions of a speculative nature have likewise been drawn with the best results from this classification of the elements into periods and natural families. The gaps shown by the system at the time it was brought forward, and in a lesser number to-day, were and are intended to be filled up by new and hitherto undiscovered elements. Mendelejeff sought to predict from the positions of such blanks, not merely the existence of elements and their approximate atomic weights, but also their properties and chemical behaviour, together with that of some of the compounds which they would form. His prognostications have been fulfilled in the most striking manner by the discovery of gallium, scandium, and germanium, and by the verification of their behaviour as foreseen by him.

The perception of the fact that the physical and chemical properties of all the elements show a periodic dependence upon their atomic weights is therefore a result of this natural classification. But the discovery of the common cause which underlies those peculiar relations, and its formulation into a law, still remain tasks of the future. Some chemists have thought to lift this veil already by assuming that all the various elements, or at least those belonging to a natural family, may be referred back to still simpler ones. We perceive clearly here a reapproach to Prout's hypothesis, which threatened to exercise such an unfavourable influence on the rational development of the atomic doctrine, had not the most eminent chemists of the time raised up a protest against its admissibility. During the last few years Crookes has again brought up this ticklish question, whether the so-called elements are to be

regarded as simple, or not rather as compound.¹ According to him all the elements have resulted, by gradual condensation, from a primary material which he terms *protyle*, this view having been arrived at by him from his observations on the phosphorescence spectra of the yttrium earths.

But until the transformation of one element into another has been incontestably proved by experiment, chemists cannot give up the idea of indivisible elementary particles, *i.e.* the present atomic theory.

The General Significance of Physico-chemical Investigations.

The relations thus discovered between the atomic weights of the elements and their physical properties have materially contributed to enlarge our knowledge of the boundary land between physics and chemistry. Many investigators had already before this followed the example of H. Kopp² (who began his stimulating labours in the forties), in assiduously tracing out the connection existing between the chemical constitution of compounds and their physical behaviour. The advances made in this direction fall to be treated of in the special history of physical chemistry. Here it need

¹ Cf. *Chem. News*, vol. liv.; also Crookes' Presidential Addresses to the Chemical Society in 1888 and 1889, published in the Society's *Journal* for those years.

² Hermann Kopp, born 30th October 1817 at Hanau, at which place his father was an esteemed physician, was drawn through Liebig's attraction to Giessen, where he became assistant professor of chemistry in 1841 and professor at a later date; at Giessen he remained until his removal to the University of Heidelberg in 1864, where he still continues to work in full vigour. His services as a historian of chemistry have already been frequently referred to. All his historical works [*Geschichte der Chemie* ("History of Chemistry"), four vols. 1843-47; *Die Entwicklung der Chemie in der neueren Zeit*, 1873 ("The Development of Chemistry in Recent Times"); *Beiträge zur Geschichte der Chemie* ("Contributions to the History of Chemistry"), etc.] are distinguished by their comprehensiveness and thoroughness. He possesses in a remarkable degree the gift of sympathetically tracing out the development of important ideas and hypotheses. The stimulus produced by his physico-chemical researches was a highly gratifying one (cf. the special history of physical chemistry). In addition to all this, he has taken a share in the editing of Liebig's *Jahresbericht* and of the *Annalen der Chemie und Pharmacie*, besides writing his *Lehrbuch der theoretischen Chemie* (1863) for the Graham-Otto series.

merely be said that it has come to be more and more recognised, especially within the last thirty years, that chemical investigation, without a free use of physical aids, runs the danger of becoming one-sided. Chemists have perceived the necessity for their science of physico-chemical methods.

Thus, what an extensive application have not the latter found in the estimation of the molecular weights of elements and compounds! The determination of vapour density has proved its value for the attainment of this end in an infinite number of cases, and has been applied to the solution of the most important theoretical questions; thus, of recent years the saturation-capacities of numerous elements, *e.g.* tungsten, vanadium, beryllium, thorium, germanium, aluminium, etc., have been established by the aid of this process. The constant relations between the molecular weight of a substance on the one hand and its point of solidification and the vapour pressure of its solutions on the other, first observed by Raoult and de Coppet (independently), have rapidly become the basis of easily-carried-out methods for the determination of molecular weights. We need only refer here to the deduction of the atomic weights of elements from their specific heats, and of equivalents from the electrolysis of salts, in order to emphasise the significance of physical methods for establishing the most important of chemical values. Of the wealth of work which has been accomplished in the branches of spectrum analysis, thermo-chemistry, upon the doctrine of affinity (*Verwandtschaftslehre*), and in the investigation of the connection between optical properties and chemical constitution, an account will be given in the special section. The position of chemistry to physics will there come out more clearly than is possible at this point.

Similarly the relation of chemistry to other branches of science can only be properly represented by going into details. This will show itself in the case of mineralogy, which is joined to inorganic chemistry by a firm band. The connection with physiology is proved by the fact that

organic chemistry is indispensable to the latter. To whatever quarter we may turn in the extensive range of the natural sciences, we find that chemistry is to most of the latter an indispensable aid, and to the remaining ones useful in a high degree. The history of the different branches of natural science shows in the most distinct manner this constantly recurring reciprocal action.

SPECIAL HISTORY OF THE VARIOUS BRANCHES
OF CHEMISTRY FROM LAVOISIER TO THE
PRESENT DAY

CHAPTER VI

SPECIAL HISTORY OF THE VARIOUS BRANCHES OF CHEMISTRY FROM LAVOISIER TO THE PRESENT DAY

Introduction.—In the general history of this period the attempt has been made to set forth the more important ideas and new points of view which have led to the development of particular influential doctrines, and at the same time to give a description of the latter. In conjunction with these objective discussions, short sketches have been appended of the lives of those investigators who have exercised a permanent effect upon the development of chemistry, and more especially upon the shaping of it into a system.

Up to the years of the fourth or fifth decade of our century, the leading chemists were able to cover in their work a very large part of the ground which was either occupied by chemistry, or in which it was an indispensable aid; we have but to think of Berzelius and Liebig, and of their labours, which were at the same time both pioneering and fundamental, in analytical and pure chemistry, physiology, and mineralogy. But during the later decades the tremendous growth of the science has rendered necessary a large subdivision of work, indeed an almost one-sided specialisation in research. This may even give rise to the apprehension that, with increasing specialisation, a danger is run of losing sight of general guiding principles. Organic chemistry may serve as an example of this subdivision of

labour, particular branches of it having been opened up which in themselves alone are sufficient to absorb the full devotion of a large number of talented investigators; take, for instance, the chemistry of the aromatic compounds, more especially that portion of it comprising the pyridine and quinoline bases.

In the following special section of this book, which deals with the different branches of chemistry in succession, there are recorded such facts and investigations as have contributed to the true advancement of the various parts of our science.

The history of analytical chemistry is placed first in order, since the latter is an indispensable aid to all chemical research, and therefore to all the other branches of chemistry, pure as well as applied. Following it comes the history of pure chemistry, which divides itself into inorganic and organic, although there is no natural partition between the two. Next to pure chemistry stands physical, with whose history that of the doctrine of affinity (*Verwandtschaftslehre*) is bound up in the most intimate manner. It was the endeavour to discover relations between chemical and physical properties which led to the establishment of this important middle kingdom between chemistry and physics.

That chemistry is necessary for the healthy growth of other sciences is particularly shown in the history of mineralogical, physiological, and pathological chemistry, which are likewise treated of according to their historical development. The opening up and cultivation of the fields of mineralogy, geology, and vegetable and animal physiology are indissolubly connected with the names of such distinguished chemists as Lavoisier, Vauquelin, Klaproth, Berzelius, and Liebig.

Last in order comes the history of technical chemistry, which illustrates in the most brilliant manner the influence of chemical research upon the development of chemical industry. To give a historical account of the penetration of the scientific spirit and of chemical methods into this

branch, a branch hitherto worked empirically, is a task which repays itself in a special degree.

As an appendix to the whole, an attempt has been made to picture within short space the growth which chemical instruction and the aids towards it have undergone during the course of the present century.

HISTORY OF ANALYTICAL CHEMISTRY IN RECENT TIMES.

The main problem of chemistry, the investigation of the true composition of compounds, necessarily carries along with itself the constant endeavour to elaborate and perfect the means employed for arriving at this end. Thus, since the time of Lavoisier, analytical methods, which constitute the tools for the solution of this problem, have been and are being improved in a continuously increasing degree.

Qualitative Analysis of Inorganic Substances.

Even so early as during the phlogistic period, men like Boyle, Hoffmann, Marggraf, Scheele, and especially Bergman, had collected together a large number of valuable observations, by the aid of which it was possible to test with certainty for many inorganic compounds. In a knowledge of the various reagents which served for this end Bergman was the furthest advanced; he it was who first attempted to publish a system for the qualitative analysis of substances in the wet way (cf. p. 136). From the analytical course of procedure which he proposed, and which had for its aim the separation of different substances into particular groups by converting them into insoluble compounds, the methods in use at the present day have developed themselves. To the perfecting of this (previous to the time of Berzelius, who worked with the greatest effect in this branch also), Lampadius and Gottling materially contributed; the former published in 1801 his *Handbuch zur chemischen Analyse der Mineralien* ("Text-Book on the Chemical Analysis of Minerals"), and the latter his *Practische Anleitung zur prüfenden und zerlegenden Chemie* ("Practical Introduction to the Chemistry of Testing and Decomposing"),—works in which the best analytical methods of the time are given.

The many and varied observations collected by Klaproth, Vauquelin, Berzelius, Stromeyer, and others in their analyses of minerals further helped to strengthen the qualitative method. The text-books of analytical chemistry by C. H. Pfaff and Heinrich Rose enable us to judge of the rate of its continuous development; alongside of the latter of those works, which became celebrated in an especial degree and ran through numerous editions, must be placed the well-known and highly prized *Anleitung zur qualitativen chemischen Analyse* ("Introduction to Qualitative Chemical Analysis") of R. Fresenius, which covers the whole ground on the subject, and is a marvel of thoroughness and accuracy. The procedure in qualitative analysis has undergone no material alterations since Fresenius first published his book, and is treated of in numerous works, most of which are intended to instruct the beginner in its principles.¹

Qualitative analysis in the dry way has been perfected by the more general and improved use of the blowpipe, which Berzelius² and Hausmann were in a high degree instrumental in introducing into chemistry and mineralogy; this valuable instrument has been employed with the greatest success, more especially for the detection of the constituents of minerals. Bunsen's important flame-reactions³ have, however, enabled it to be dispensed with in a number of cases. Among the most noteworthy of dry reactions are the spectroscopic, which, thanks to their extraordinary delicacy and certainty, serve for the detection of the most minute quantities of many metals, and have rendered possible the discovery of a number of new elements. Spectrum analysis, through which we are able to deduce the nature of a glowing substance by examining the light which it emits, was founded by the masterly researches of Bunsen and Kirchhoff;⁴ Talbot.

¹ Out of the large number of such text-books, those of Beilstein, Birnbaum, Classen, Drechsel, Genther, Medicus, Rammelsberg, Stadel, Kolbe, Will, Odling, Harcourt and Madan, Thorpe, Clowes, and Jones may be mentioned.

² His pamphlet, *Ueber die Anwendung des Lothrohrs* ("On the Application of the Blowpipe"), was first published in 1820.

³ *Ann. Chem.*, vol. cxxxviii. p. 257; also in a much extended form as a separate pamphlet.

⁴ *Pogg. Ann.*, vol. cx. p. 161.

Miller, Swan, and others had before this investigated the spectra of coloured flames, without however applying their results with a definite aim to the analysis of substances. The first proposal to utilise the different flame colorations for distinguishing potash from soda salts, was made so long ago as by Marggraf.¹

Quantitative Analysis of Inorganic Substances.

The exact investigation of the behaviour of bases, acids, and salts towards different reagents, especially towards such as yield with them either sparingly soluble or insoluble precipitates, constituted the basis of the gravimetric estimation of individual substances. Before the time of Lavoisier but few attempts had been made at quantitative analysis, but the path which it was bound to follow had been already clearly indicated by Bergman; for, from him originated the principle of converting the substance to be analysed into a convenient form of known composition, and deducing from the weight of the compound thus precipitated or otherwise obtained that of the substance in question. At that date chemists either already knew or became acquainted with the precipitation of silver solutions by hydrochloric acid, of solutions of lime salts by oxalic or sulphuric acid, of lead salts by liver of sulphur or sulphuric acid, and many similar reactions. It was Klaproth who taught the ignition of precipitates before weighing them, in those cases where they did not suffer decomposition through this, and he also co-operated largely with Vauquelin in developing the quantitative analysis of minerals. The observations of both of these chemists, especially of Klaproth (who directed his efforts to ascertaining correctly the composition of those compounds into which the constituents of the substances to be analysed were usually transformed), attained to a fairly high degree of accuracy, which is also to be ascribed to those analyses of salts carried out by Wenzel at an earlier date, although to these hardly any attention had been paid.

¹ Cf. p. 135.

Richter's endeavours to establish the quantitative composition of salts, and the success by which they were followed, have been sufficiently treated of in the general history of this period; in spite of the fact that his analyses were not particularly accurate, he understood how to draw significant and correct deductions from them.

Lavoisier, who had clearly grasped the importance of proportions by weight, and therewith, that of quantitative analysis, from the beginning of his scientific career, examined more particularly the composition of oxygen compounds. Thus he established with tolerable correctness, for example, the relation of carbon to oxygen in carbonic acid, but only approximated to that of hydrogen to oxygen in water, and was wide of the mark in the relation of phosphorus to oxygen in phosphoric acid. The values which he obtained for the composition of water and carbonic acid, he sought ingeniously to apply for establishing the composition of organic substances. Lavoisier, however, introduced no original methods for the quantitative analysis of inorganic bodies and their separation from one another.

Proust effected infinitely more in this branch, his analytical work leading, as has already been stated, to a clear grasp of the law of constant proportions, and of the alteration by definite increments in combining proportions. Quantitative analysis was also strengthened and extended by the establishment of stoichiometry (which found its perfect support in Dalton's atomic theory), since a check upon the results obtained was thereby rendered possible.

Endeavours were at that time mainly directed to the determination of the relative atomic or, more correctly, combining weights. The splendid results obtained by Berzelius from his pioneering labours in this direction have already been detailed. He devised a large number of new gravimetric methods of estimation, and tested those already in use for the separation of substances, working out better modes for attaining to this end. His researches on the composition of chemical compounds extended over all the elements which were at all well known. Berzelius, far more than any

other man, developed the principles by which their atomic weights could be established; and the degree of accuracy at which he arrived in his analyses is seen from the tables of atomic weights published by him after the year 1818 (cf. pp. 205 and 211).

The great task of determining the atomic weights, the constants of the atomic theory, with the utmost possible accuracy, has led ever since the time of Berzelius to the development and improvement of gravimetric methods; for, what was required here was to establish by various procedures an unalterable value for each element, a value which should form the basis for the composition of all the compounds of that element. The endeavours and the speculations at rounding off these numerical values in accordance with Prout's hypothesis were replaced by exact quantitative determinations. Among the latter, the researches of Dumas, Erdmann and Marchand, Marignac, and Stas deserve especial mention.¹

The systematic development of quantitative analysis was thus mainly promoted by the investigation of mineral substances, since the chief requirement here was to find out modes for separating their constituents from one another. After the valuable preparatory labours of Bergman (with whom, for instance, the fusion of silicates with alkaline carbonates originated), and the researches of Klaproth, Vauquelin, and Proust it was Berzelius who worked out entirely new methods; we need only recall here his plan of decomposing silicates by hydrofluoric acid, and that of separating metals from one another by means of chlorine. He it was, also, who employed far smaller quantities of substances rather than the large amounts recommended by Klaproth, who introduced the spirit-lamp which bears his name, thus facilitating the ignition of precipitates, and who taught how to incinerate the filter-paper and determine its ash; in fact, to speak generally, he was the first to try a large number of practical contrivances and apparatus for the carrying out of

¹ Cf. Lothar Meyer and K. Seubert, *Die Atomgewichte der Elemente* (1883).

analyses. His greater analytical researches, such as those upon platinum ores and on various mineral waters, show Berzelius as a master in devising good methods of separation.

His pupils, more especially H. Rose¹ and Fr. Wöhler, worked up the valuable experiences of their teacher, extended them largely by wide-reaching observations of their own, and made analytical methods public property by their admirable works² on the analysis of minerals and chemical bodies generally. R. Fresenius,³ our present chief exponent of analytical chemistry, has likewise perfected and strengthened this branch of the science in all its various parts by collating and sifting the methods formerly in use, and, more especially, by working out numerous new ones. It is impossible to enumerate here what other workers (among whom Liebig, Thomson, Stromeyer, Bunsen, Turner, Scheerer, Rammelsberg, Gibbs, Blomstrand, R. Schneider, and Pérouze

¹ The brothers Heinrich and Gustav Rose belonged to a Berlin family which has produced distinguished chemists for several generations. Their grandfather, Valentin Rose the elder, a pupil of Marggraf, and also their father, Valentin Rose the younger, were energetic pharmacists and chemists. Gustav Rose, who was born in 1798 and died in 1873 as Professor of Mineralogy at Berlin, was only connected with chemistry indirectly. But Heinrich Rose, born 1795, died 1864, was an ardent exponent of the science, and enriched it by most important work, especially in analytical and inorganic chemistry (see special history of these). He reciprocated fully and truly the affection of his master Berzelius, as is vividly shown in the beautiful memorial address which he gave of the latter (cf. p. 197). In his two-volume *Handbuch der Analytischen Chemie*, H. Rose collected together in a masterly manner the best of the then known methods in qualitative and quantitative analysis.

² H. Rose, *Ausführliches Handbuch der analytischen Chemie* ("Detailed Text-Book of Analytical Chemistry"); Fr. Wöhler, *Die Mineralanalyse in Beispielen* ("The Analysis of Minerals, illustrated by Examples").

³ C. Remigius Fresenius, born at Frankfurt on the Maine in 1818, became assistant to Liebig in Giessen in 1841, and assistant professor there in 1843; in 1848 he opened his now universally known laboratory at Wiesbaden, which has undergone a continuous extension, and been frequented by students from all parts. His text-books of chemical analysis, of which the qualitative appeared for the first time in 1841, and the quantitative in 1846, have had an extraordinarily wide distribution, as their numerous editions in different languages prove. By founding in 1862 the *Zeitschrift für analytische Chemie* ("Journal of Analytical Chemistry", Fresenius furnished a centre-point for the analytical branch of the science.

may be named) have done for the development of quantitative analysis.

We may, however, mention here that the galvanic current has recently been called in to the service of analysis, the quantitative determination of many metals being rendered possible by its aid. After Gibbs (in 1865) had worked out the electrolytic determination of copper, and other chemists later on had busied themselves with similar investigations. Alexander Classen¹ rendered especial service in the development of the method. This branch of chemical analysis is of the utmost use for metallurgy, in which even already it forms an important part of docimacy. The latter, originally confined to the determination of the noble metals in the dry way, has become expanded into an important branch of analytical chemistry, particularly since C. Fr. Plattner's comprehensive researches and the publication of his classical book, *Die Probierekunst mit dem Lotiröhr* (Leipzig, 1835). ("Docimacy by Means of the Blowpipe").²

Volumetric Analysis.

Besides the analytical methods which have been touched upon above, volumetric ones have become developed within the last sixty years or so; these are of great use, particularly in manufacturing chemistry and pharmacy, and have therefore the widest application. Since in volumetric methods no weighing is required to be done after the standard solutions have once been made up, and the wished-for results are arrived at simply by reading off the amounts of the solutions used, much time is saved and at the same time sufficient accuracy attained to, the requirements of technical analysis (more particularly) being thereby met.

Gay-Lussac must be regarded as the man who introduced volumetric methods into the science, and rendered them

¹ Cf. his work, *Handbuch der chemischen Analyse durch Electrolyse* ("Text-Book of Chemical Analysis by means of Electrolysis").

² Cf. Kerl's *Metallische Probierekunst* ("Metallic Docimacy"), (1886); Balling's *Probierekunde* ("Docimacy"), (1879), and his *Fortschritte im Probiervwesen* ("Advances in Docimacy"), (1877).

available for chemical industries; before him various investigators—of whom Deseroizille and Vauquelin must be especially mentioned—had attempted to apply such methods empirically to comparable determinations of chemical products.

Gay-Lussac worked out with the greatest care his methods of *chlorimetry* (1824), of *alkalimetry* (1828), and of the determination of chlorine and silver (1832).¹ Notwithstanding the good results which those volumetric processes yielded, they received but slowly the recognition which was their due. The application of permanganate of potash to the estimation of iron by Margueritte in 1846, and, more particularly, Bunsen's process with equivalent solutions of iodine and sulphurous acid (by the aid of which a large number of different substances can be accurately estimated by one and the same reaction) are landmarks in the history of "titrimetry," which soon after this began to rank alongside of gravimetric analysis. One of the chief promoters of volumetric methods was Friedrich Mohr, who both improved old processes and introduced many new ones; he rendered great service by the publication of his *Lehrbuch der chemischen Titrimethode* ("Text - Book of Volumetric Analysis").² Among the many investigators who have enriched this branch of the science, we may name J. Volhard,³ who devised an exact method (the determination of silver by means of ammonium sulphocyanide) capable of numerous applications.

In organic chemistry volumetric analysis has not been able to take up anything like the same position that it has in inorganic, the methods as yet introduced being wanting in precision. Among the most noteworthy processes here are Fehling's for the determination of grape sugar, Liebig's for that of urea, the volumetric estimation of phenol by means of bromine,⁴ etc.

¹ Cf. his *Instruction sur l'Essai des Matières par la Voie Humide* (1833).

² The latest edition of this is edited by A. Classen. Among other valuable books on volumetric analysis are those of Cl. Winkler, Medicus, and Fleischer in Germany, and of Sutton in England.

³ Cf. *Ann. Chem.* vol. xc. p. 1 *et seq.*

⁴ Cf. Degener, *Journ. pr. Chem.* (2), vol. xvii. p. 390; Koppeschaar, *Ztschr. Anal. Chem.* for 1876, p. 223.

Development of Methods of Gas Analysis.

The history of the volumetric analysis of liquids naturally leads us on to a description of the efforts at analysing gases qualitatively and quantitatively. It is worthy of note here that the systematic qualitative analysis of these was much later of being developed than their quantitative determination. The first attempts in this direction were made by Priestley, Cavendish, and Lavoisier, to be followed by those of Dalton, Gay-Lussac, Henry, Saussure, and others at the beginning of this century. But it has been through Bunsen's fundamental researches¹ that the quantitative analysis of gases has been brought to such perfection that those methods which depend upon the absorption or combustion of the gas under investigation are among the most exact of our science, having required but trifling alterations since he first published them.

In addition to Bunsen's methods, others have been worked out with a special view to technical gas analysis; although the same as the former in their main principle, these allow of determining the composition of the so-called industrial gases by the aid of simple apparatus within a short time, and with sufficient accuracy. Cl. Winkler and W. Hempel have rendered great service here by materially simplifying the apparatus required and by generalising methods.² Among others who have done good work in gas analysis of recent years may be mentioned Frankland, Pettersson, Orsat, Coquillion, and Bunte.

The qualitative analysis of gases has only quite recently become developed scientifically, and here, too, Winkler has

¹ These researches of Bunsen's began about the year 1838, and were collected together under the title of *Gasometrische Methoden* (Brunswick, 1857; second edition, 1877); this most valuable work was translated into English by Rosecoe.

² Cf. Clemens Winkler, *Anleitung zur chemischen Untersuchung der Industriegase*, Freiberg, 1876-77 ("Methods for the Chemical Examination of Industrial Gases"); the same author's *Lehrbuch der technischen Gasanalyse* (1885), ("Text-Book of Technical Gas Analysis"); and W. Hempel's *Gasanalytische Methoden* (1890).

laboured with success; by the systematised use of absorptives, he has divided gases into different groups, thus proceeding in the same manner as is done in the analysis of substances in the wet way. The improvements in methods of gas analysis have drawn the attention of chemists to gases in an increasing degree, and have proved of the greatest benefit to theoretical as well as to practical chemistry.

The Analysis of Organic Substances.

The fact that animal and vegetable products, which came to be comprised under the term "organic," always contain carbon, usually hydrogen and oxygen, and frequently also nitrogen, was, as already stated, a long time of becoming recognised. Here again we have a brilliant proof of Lavoisier's far-seeing glance, and of his power of drawing general conclusions from detached observations. It had indeed struck previous experimenters, *e.g.* van Helmont and Boyle, that spirit of wine, wax, etc., form water when burned, while Priestley perceived that carbonic acid was produced at the same time; in fact Scheele stated in 1777 that both of these compounds were products of the combustion of oils. After it had become clear to Lavoisier that carbonic acid consisted of carbon and oxygen, and water of hydrogen and oxygen, he went on to deduce the composition of organic substances. Thus, with the finding out of what were the most important elements of organic compounds, the first step in qualitative organic analysis was reached. The principle of arriving at the constituents of organic bodies by transforming them into compounds of known composition has ever since been retained. In the same way nitrogen, which Lavoisier himself recognised as being peculiar to many organic substances,¹ was detected by conversion either into ammonia or sodium cyanide (Berthollet, Lassaigne), and

¹ How uncertain the tests for the elements of organic substances were at the beginning of this century is shown by the fact that Proust believed he had proved nitrogen to be an integral constituent of acetic acid.

phosphorus and sulphur by conversion into phosphoric and sulphuric acids respectively.

While the elementary constituents of organic compounds are thus easily arrived at, the detection of the compounds alongside of one another can only be carried out in a few cases; only small beginnings have as yet been made at a systematic course of qualitative organic analysis, in the sense in which we apply the term to inorganic.¹ In many instances one has to depend upon isolated characteristic reactions of organic substances, *e.g.* in the investigation of colouring matters, alkaloids, etc.

The quantitative analysis of organic compounds has developed itself from the observation that carbonic acid and water are products of their combustion: the method, therefore, which served for the detection of the constituents carbon and hydrogen was applied in a perfected form to their exact determination. Lavoisier was again the first to point out the right path here; he attempted to burn the organic compound in question completely, and to estimate the resulting carbonic acid and water—the latter indirectly. In order to be able to deduce the amounts of carbon and hydrogen themselves, it was necessary to know the quantitative composition both of carbonic acid and of water; but, since the values obtained by him for these were not very accurate,² it was impossible that the results of his analysis of an organic substance could turn out correct, and this all the more from the method of the combustion being such as to involve errors in itself.

Lavoisier's process for easily combustible substances was to burn a weighed quantity in a known volume of oxygen, contained in a receiver closed by mercury, and then to esti-

¹ Cf. Barfoed's *Qualitative Analyse organischer Körper*; also Allen's *Commercial Organic Analysis*.

² The following are Lavoisier's figures for the composition of carbonic acid and water (the correct values being given in brackets):—

Carbonic Acid	{	Carbon	28	per cent (27·2)
					{	Oxygen	72	„ (72·8)
Water	{	Hydrogen	13·1	„ (11·1)
					{	Oxygen	86·9	„ (88·9)

mate the resulting carbonic acid together with the residual oxygen; from these data the amounts of carbon, hydrogen, and oxygen were calculated. For difficultly combustible bodies, such as sugars and resins, Lavoisier (as we now learn from his recently published journals)¹ used, instead of the free gas, substances which yield up their oxygen upon being heated, *e.g.* red oxide of mercury and red lead; he thus adopted the plan which later on became the standard one, while at the same time he estimated the weight of the carbonic acid produced by this oxidation by means of a solution of caustic potash.

Had those researches become known at that time, organic analysis would, without doubt, have undergone a more rapid development than it actually did. The efforts of Dalton (1803), Saussure (about 1800-1803), and Thénard (1807) to get at the composition of organic compounds by exploding their vapours with oxygen and analysing the products would never have been made. Gay-Lussac and Thénard² endeavoured to solve this problem in a more felicitous manner by the combustion of the organic substance with chlorate of potash; from the amounts of resulting carbonic acid and residual oxygen they calculated the percentage of carbon, hydrogen, and oxygen in the substance under analysis, and attained in some instances at any rate to serviceable results. Compared with this method, uncertain as it was on account of the violence of the combustion, the one followed by Berzelius showed a marked improvement;³ for here the organic substance, mixed with chlorate of potash and sodium chloride, was gradually decomposed, and then not merely the resulting carbonic acid but also the water determined directly—the latter by means of chloride of calcium. A further advance was made by Gay-Lussac⁴ in 1815, in the use of black oxide of copper as the oxidising agent. The rounding off of the whole procedure by the introduction of a convenient bulb-shaped

¹ *Œuvres de Lavoisier*, vol. iii. p. 773.

² *Recherches Physico-chimiques*, vol. ii. p. 265.

³ *Annals of Philosophy*, vol. iv. pp. 330, 401.

⁴ *Schweigger's Journ.*, vol. xvi. p. 16; vol. xviii. p. 369.

apparatus, and the consequent simplification of the manipulation required, is due to Liebig.¹ Since this last, elementary organic analysis has not altered essentially, the modifications introduced having had reference to the combustion furnaces and to the mode of carrying out the combustion; with respect to the latter, Koppfer's method² must be mentioned, a method by which the substance is burnt in a current of oxygen, with the aid of platinum black. Plans for the combustion of organic compounds in a stream of oxygen had before this been proposed by Hess, Erdmann and Marchand, Wöhler, and others.

The exact determination of nitrogen in organic compounds first became possible after Dumas³ (in 1830) had devised his admirable method. For many nitrogenous organic substances the process worked out by Will and Varrentrapp⁴ at a later date, in which the nitrogen is estimated as ammonia, has proved itself thoroughly applicable. In addition to these, the quite recent method of Kjeldahl⁵ must be mentioned, a method which is found to be of great use, especially for agricultural-chemical analyses.

Analytical methods have found the most extended application in judicial cases, in questions of hygiene, and in all the branches of technical chemistry; a short historical account of them must therefore be given here. Forensic chemistry, whose task consists in the absolutely certain detection of poisons, could only reach its present stage of development after analytical methods in general had been put upon a firm basis. Fresenius admirably depicted in 1844 the position and duties of a forensic chemist at that date.⁶ The great progress which has since been made in the certainty with which poisons can be detected is distinctly seen by an examination of the various works on legal-chemical analysis which

¹ *Pogg. Ann.*, vol. xxi. p. 1; also his pamphlet, *Anleitung zur Analyse organischer Körper* ("The Analysis of Organic Compounds").

² *Ber.*, vol. ix. p. 1377.

³ *Ann. Chim. Phys.*, vol. xlv. pp. 133, 172; vol. xlvii. p. 196.

⁴ *Ann. Chem.*, vol. xxxix. p. 257.

⁵ *Ztschr. analyt. Chem.*, vol. xxii. p. 366; vol. xxiv. p. 199.

⁶ *Ann. Chem.*, vol. xlix. p. 275.

have been published from time to time.¹ In addition to Fresenius—J. and R. Otto, Marsh, Graham, Stas, Mohr, Husemann, Dragendorff, and others have rendered especial service in working out good methods. The Stas-Otto process for the detection of individual alkaloids has proved of great importance for the development of this branch; since the discovery of the ptomaines,² it has had to undergo some modifications, as the resemblance between many of the reactions of these products and those of the vegetable alkaloids may give rise to most serious mistakes, and in fact has already done so.

A special branch of chemical analysis is represented by the methods of testing and investigating used in industrial chemistry. Since these have for their aim the attainment of a fair degree of accuracy within the shortest possible time, volumetric processes are the ones most frequently employed. A glance into the most recent text-books of technico-chemical methods³ is sufficient to allow us to recognise the high degree of development to which these have been brought. A large number of processes have in the course of time been devised, more especially for the commercial analysis of organic products; we may recall here the estimation of sugar by polarisation, the determination of the heating power of combustibles, and the valuation of coal-tar dyes by test-colorations and by specific reactions, not to speak of numerous other methods which have become standard ones in chemical technology.

For technical chemists, and in an equal degree for medical officers of health, the development of the analysis of articles of food and drink has been of great importance; the pharmacist, too, frequently finds it needful to apply the methods which have approved themselves in such cases. By their aid the analyst is able to decide whether the products are what they pretend to be, or, if they are adulterated, the

¹ Reference may be made here to Otto's *Anleitung zur Ausmittlung der Gifte* ("Methods for the Detection of Poisons"), sixth edition, 1881; also to Stevenson's new edition of *Taylor on Poisons*.

² Cf. the special history of physiological chemistry.

³ The works of Post (Brunswick, 1882), of Bockmann (Berlin, 1888), and of Sutton may be mentioned here.

nature of such adulteration. The reader has but to recall to mind the quickly executed methods for analysing milk, butter, meal, wine, beer, coffee, etc., in order to appreciate the true blessing of these applied analyses. The gradual work of numerous investigators has rendered possible the development, within a comparatively short period of time, of the processes which have now become standard ones here. We cannot now refer in detail to the services rendered by single individuals in this branch. Full particulars are to be found in König's admirable work, *Die menschlichen Nahrungs- und Genussmittel* (Berlin, 1883), a book which furnishes a complete review of the subject, and at the same time indicates clearly the share which different chemists have taken in it.¹

¹ In C. Flügge's *Lehrbuch der hygienischen Untersuchungsmethoden* ("Text-Book of Methods of Hygienic Research"), hygiene possesses a splendid guide for investigations of this kind.

THE PROGRESS IN PURE CHEMISTRY FROM LAVOISIER TO THE PRESENT TIME.

While only the main currents of chemistry have been depicted in the general history of this period, we have now in the following section to pick out, from the endless number of experimental researches made, those which have materially contributed to the extension of our chemical knowledge. This rich material is divided between the two great branches of inorganic and organic chemistry. If we glance back over the labours of the last fifty years, we recognise that organic chemistry has gone on preponderating more and more over inorganic; the former has outgrown the latter,—its elder sister. But inorganic remains nevertheless the basis upon which organic chemistry rests, although on the other hand we must not forget that important fundamental principles and doctrines (*e.g.* the doctrine of valency and the true conception of chemical constitution) were first fruitfully developed in the domain of organic chemistry.

SPECIAL HISTORY OF INORGANIC CHEMISTRY.

The great revolution in ideas with regard to the constitution of many substances, which was brought about by Lavoisier's system, has been described in detail in the special part of this book. A large number of bodies, which had formerly been looked upon as compound, belonged from thenceforth to the elements; while many, which had been considered simple substances, were either proved to be compounds, or regarded as such from their analogy to others. The clarifying process which Lavoisier had commenced went vigorously forward, thanks to the efforts of Klaproth, Vauquelin, Davy, Berzelius, Gay-Lussac, and others. But we are still far from having attained to a clear and definite knowledge of the nature of all the elements and their compounds, new elements being from time to time added to the

long series already known: and the relations of those to the others have to be established by an accurate study of their chemical behaviour. Emphasis has already been laid upon the great effect which the so-called periodic system has had on the classification of the elements.

Historical Notes on the Discovery of Elements — The Determination of their Atomic Weights.

The knowledge of the elements was increased to a very large extent soon after the death of Lavoisier (who had not himself discovered any), and this exactly in proportion as methods of chemical analysis became more perfect. While Lavoisier was able to bring forward twenty-six elements in his *Traité de Chimie*, the number of those whose existence has been definitely established has now extended to at least sixty-six.

To the aid which was rendered by improved methods of analysis, other means especially effective for the discovery of new elements soon came to be added. Among these was the application of the galvanic current to the decomposition of chemical substances, and the breaking up of haloid compounds by means of the alkali metals: in spectrum analysis, lastly, chemistry now possesses a most admirable implement, which has already led to the isolation of a number of most important elements.

After the establishment of the atomic theory, the original task of acquiring a qualitative knowledge of a new element and its compounds was supplemented by the further and higher one of determining its relative atomic weight and explaining, on the basis of the atomic hypothesis, the constitution of the compounds which it forms with other elements.

For oxygen, which Lavoisier was the first to claim as a simple substance, the elementary nature was always afterwards maintained. Nitrogen, on the other hand, was temporarily regarded by Davy (1808) and by Berzelius¹

¹ Cf. Kopp, *Gesch. der Chemie*, vol. iii. p. 218.

(1810) as a compound of an unknown element *nitricum* with oxygen, because only in this way could they find an explanation of the basic properties of ammonia, in which they likewise assumed the presence of oxygen. Davy was the first of the two to give up this hypothesis in favour of the simpler one of nitrogen being an element, Berzelius only doing this in 1820.

Hydrogen, too, was for a short time looked upon by Berzelius as being compound, *i.e.* as containing oxygen, and the same applied to sulphur and phosphorus, in which the presence of hydrogen and oxygen, besides that of other unknown elements, was conjectured. That many distinguished chemists were inclined to regard chlorine as the oxide of a hypothetical element has been already detailed, as has also the profound influence which this view exercised upon important sections of chemistry.¹ Before this idea had been abandoned by Berzelius, iodine—discovered by Courtois in 1811 in the ashes of marine plants—was shown to be an element analogous to chlorine, through the splendid researches of Davy, and still more those of Gay-Lussac.² Bromine, isolated by Balard³ in 1826 from the mother liquor of sea-salt, and whose investigation was materially promoted by Lowig's labours in 1829, completed for a long time to come the group of Berzelius' "halogen" elements. Fluorine, the acid constituent of hydrofluoric acid, has only quite recently been isolated for the first time by Moissan⁴ (in spite of a great many previous attempts⁵), by the electrolysis of hydrofluoric acid under suitable conditions, and, as was to be expected, has been found to be a substance of the most violent chemical energy.

The atomic weights, those all-important constants, have been determined for the metalloids already spoken of with great accuracy, and by various different methods in each

¹ Cf. p. 225.

² *Ann. de Chimie*, vol. xci. p. 5 (1813).

³ *Ann. Chim. Phys.* (2), vol. xxxii. p. 337.

⁴ *Ibid.* (6), vol. xii. p. 472 (1887); also *Comptes Rendus*, vol. cix. p. 861.

⁵ Cf. Gore, *Phil. Trans.* for 1869, p. 173.

case. For oxygen, nitrogen, chlorine, bromine, and iodine, the classical researches of Marignac¹ and Stas² have yielded the most reliable values; for fluorine the determination by Christensen³ may be regarded as the most exact.

Tellurium (chemically analogous to sulphur, which had been known for so long, but had first been characterised as an element by Lavoisier) was discovered by Muller von Reichenstein in 1782, and investigated by Klaproth⁴ in 1798; an intimate knowledge of it was, however, first arrived at through the investigations of Berzelius.⁵ Selenium was discovered by Berzelius⁶ in 1817, and, along with its most important compounds, examined by him in the most thorough manner. The atomic weights of the two last elements have only recently been settled, after great fluctuations, that of selenium⁷ being now taken as 79.07, and that of tellurium⁸ as 125, former determinations having for a long time caused the wrong value 127-128 to be ascribed to the latter. Since the work of Stas on the subject,⁹ the number 31.98 for sulphur has been accepted as being firmly established.

The discovery of the analogues of nitrogen,—phosphorus, arsenic, and antimony, to which bismuth may be added, took place a long time ago; but it is only of late years that they, and more especially their compounds, have been accurately investigated.¹⁰ For phosphorus, the correct atomic weight arrived at by Berzelius was confirmed by Dumas

¹ Cf. *Ann. Chem.*, vol. xlv. p. 1; vol. lix. p. 284; vol. lx. p. 180.

² *Untersuchungen über die Gesetze der chemischen Proportionen* (Leipzig, 1887), ("Researches upon the Laws of Chemical Proportions").

³ *Journ. pr. Chem.* (2), vol. xxxv. p. 541.

⁴ *Crell's Ann.*, vol. i. p. 91.

⁵ *Pogg. Ann.*, vol. xxxii. p. 28.

⁶ *Schweigger's Journ.*, vol. xxiii. pp. 309, 430.

⁷ Eckmann u. Pettersson, *Ber.*, vol. ix. p. 1210.

⁸ Brauner, *Ber.*, vol. xvi. p. 3055; Brauner has still more recently found a higher value than this, but he concludes from his experiments that, in those cases where the value obtained is greater than 125, this is due to the presence of some foreign substance (which has not yet been isolated) in the tellurium (cf. *Ztschr. Phys. Chem.*, vol. iv. p. 344).

⁹ Cf. note 2, above.

¹⁰ Cf. Thorpe and Tutton, *Journ. Chem. Soc.*, vol. lvii. p. 545.

(who found the value 31.02); similarly his atomic weight for arsenic (75) was corroborated by Pélouze and Dumas. But on the other hand R. Schneider and Cooke have proved, by their researches, that the value assumed by Berzelius for antimony was much too high.

Boron was discovered simultaneously and independently by Gay-Lussac¹ and Davy,² both of whom isolated it from boracic acid, which already Lavoisier had regarded as the oxide of an unknown element. Guided by a similar view, Berzelius succeeded in 1810 in discovering the element combined with oxygen in silica, although he was only able to prepare silicium pure for the first time in 1823 by the action of potassium on potassium silico-fluoride;³ herewith he devised an important method for the isolation of various elements.

The definite knowledge that diamond and graphite are modifications of the element carbon belongs to the beginning of the new period; in addition to the researches of Lavoisier in 1773, and those of Tennant in 1796, the proof furnished by Mackenzie—that equal parts by weight of graphite, charcoal, and diamond yield equal amounts of carbonic acid on combustion—was of especial importance for the recognition of the similar chemical nature of the three substances.

The phenomenon of allotropy, the term applied by Berzelius to the existence of one and the same substance in different modifications, has been observed with especial frequency among the metalloids. The oldest example of it was that offered by carbon, whose allotropic forms show the greatest conceivable differences among each other; the most remarkable case of it, however, is afforded by the conversion of ordinary oxygen into the chemically active ozone, which was discovered by Schönbein,⁴ although van Marum had a long time previously (in 1785) called attention to

¹ *Recherches Phys. Chim.*, vol. i. p. 276.

² *Phil. Trans.* for 1809, p. 75.

³ *Pogg. Ann.*, vol. i. p. 165.

⁴ *Ibid.*, vol. l. p. 616 (1840).

the peculiar change produced in oxygen by the electric spark. The beautiful investigations of Schönbein, Marignac, and de la Rive established the substantial identity of ozone and oxygen, while those of Andrews¹ and, more especially, of Soret² proved that the molecule of ozone was made up of three atoms of oxygen.

The allotropic modifications of sulphur were investigated by Mitscherlich, and those of selenium by Berzelius and, later, by Hittorff. The transformation of ordinary yellow phosphorus into red was also observed by Berzelius, but was first discovered with certainty by Schrotter³ in 1845, and its conversion into the metallic modification by Hittorff. The proof that boron and silicon, already long known in the amorphous state, also exist in the crystalline form, is due to Wohler. Lastly, reference may be made here to the discovery of allotropic modifications of chemical compounds, *e.g.* mercuric sulphide and iodide, arsenic trioxide, etc.

To the metals which were regarded as elements by Lavoisier many new ones were subsequently added, so a short account of the isolation of these must be given here. The memorable discovery of potassium and sodium, together with the conjoined discussion upon the nature of chlorine, had such a deep influence on the development of important chemical doctrines, that it has already been treated of in detail in the general section of this book. The relative atomic weights of these two alkali metals were determined by Berzelius with fair accuracy, allowing for the fact that he assumed their values to be four times greater than those now assigned to them. Marignac, Dumas, and Stas afterwards arrived at much the same figures in their investigations already referred to.

Lithium was discovered by Arfvedson,⁴ a pupil of Berzelius, in 1817; he found it to be a constituent of

¹ *Phil. Trans.* for 1856, p. 1; or *Ann. Chem.*, vol. xcvii. p. 371. Andrews and Tait, *Phil. Trans.* for 1861, p. 113; or *Pogg. Ann.*, vol. cxii. p. 241.

² *Compt. Rend.*, vol. lxiv. p. 904; or *Ann. Chem.*, Suppl., vol. v. p. 148.

³ *Pogg. Ann.*, vol. lxxxi. p. 276.

⁴ *Schweigger's Journ.*, vol. xxii. p. 93.

various minerals, *e.g.* petalite, and recognised its analogy to potassium and sodium, but was unable to isolate the metal itself. The latter was first properly investigated in 1855 by Bunsen and Matthiessen,¹ who obtained it by electrolysis. The red coloration which its salts impart to the spirit-of-wine flame was noticed by C. G. Gmelin in 1818.

The discovery of rubidium and caesium² in lepidolite and in the Dürkheimer mineral water by Bunsen and Kirchhoff, by the aid of spectrum analysis, was the first great gain which accrued to chemistry from this new method. Since the chemical reactions of the salts of these two alkali metals are very similar to those of the salts of potassium, their presence would undoubtedly have been overlooked but for the spectroscope. Indeed, several years before the discovery of caesium, the careful analyst Plattner³ had examined the mineral pollux, which is rich in that element, and was unable to explain the deficiency in the results of his analyses, this being really due to his taking the caesium sulphate present for a mixture of the sulphates of potassium and sodium. The atomic weights of caesium and rubidium were correctly estimated by Bunsen, although too low a value was at first assigned to the former, in consequence of the supply of material at disposal being insufficient. The atomic weight of lithium was definitely determined by Stas as 7.01.

The metals barium, strontium, calcium and magnesium were isolated by Davy from their amalgams, which Seebeck had been the first to prepare; but for a long time previous to this baryta and lime had been regarded as the oxides of unknown metals. Strontia had been discovered by Klaproth and Hope, independently of one another, and had been characterised as being similar to lime. Berzelius, Marignac, and Dumas determined the atomic weights of those four metals.

Beryllium, whose oxide Vauquelin had discovered in 1798 in the mineral beryl, was first obtained by Wohler⁴

¹ *Ann. Chem.*, vol. xciv. p. 107.

² *Pogg. Ann.*, vol. cx. p. 167; vol. cxiii. p. 337; vol. cxviii. p. 94.

³ *Ibid.*, vol. lxix. p. 443.

⁴ *Ibid.*, vol. xiii. p. 577.

in 1828, by acting upon its chloride with potassium. Its atomic weight gave rise to important discussions, since it remained for a long time uncertain whether this amounted to twice or three times its equivalent number. The point was only decided by the recent researches of Nilson and Pettersson¹ on the subject, which proved that beryllium, as a diatomic element, possesses the atomic weight 9.1.

Cadmium was first observed by Stromeyer in 1817, then subsequently rediscovered by others, and recognised as being similar to zinc in character. Thallium, isolated by Crookes² in 1861 from the selenious mud of the sulphuric acid manufacture, owes its discovery to the characteristic spectrum given by its salts. The chemical nature of this metal, which approximates on the one hand to lead, and on the other to the alkalies, was mainly established by Lamy, while Crookes determined its atomic weight.

Aluminium was isolated for the first time by Wöhler³ in 1827, by the action of potassium upon its chloride, and therewith the conjecture which had long been entertained, that alumina was the oxide of a metal, was confirmed. The elements indium and gallium, which, together with aluminium, constitute a family, were only discovered comparatively recently, the former of the two in 1863 by Reich and Richter,⁴ as a constituent of the Freiberg zinc blende, and the latter in 1875, also in zinc ores, by Lecoq de Boisbaudran.⁵ Here again it was the characteristic spectra of these two metals which led to their discovery. Their atomic weights were determined by the discoverers, and that of indium with especial accuracy also by Cl. Winkler⁶ and by Bunsen;⁷ while the atomic weight of aluminium has been worked out with the utmost care by Mallet.⁸

The isolation of the metals which constitute the cerium

¹ *Journ. pr. Chem.* (2), vol. xxxiii. p. 15.

² *Chem. News*, vol. iii. p. 193.

³ *Pogg. Ann.*, vol. xi. p. 146.

⁴ *Journ. pr. Chem.*, vol. lxxxix. p. 444; vol. xc. p. 172; vol. xcii. p. 480.

⁵ *Comptes Rendus*, vol. lxxxi. pp. 493, 1100.

⁶ *Journ. pr. Chem.*, vol. cii. p. 282.

⁷ *Pogg. Ann.*, vol. cxli. p. 28.

⁸ *Phil. Trans.* for 1880, p. 1003.

group has presented unusual difficulties. Although the discovery of yttria—impure, it is true, from admixture with other earths—was accomplished by Gadolin nearly a hundred years ago, the chemistry of the cerium metals is not even yet completely cleared up. After Klaproth and, simultaneously, Berzelius had prepared cerium sesquioxide from cerite, and the latter had recognised this as the oxide of a metal, Mosander discovered two new oxides in crude yttria, the metals of which—lanthanum and didymium—he isolated. A few years later (in 1843) he added to these two others, erbium and terbium, whose existence and nature is not yet, however, definitely settled, in spite of the admirable work which has been done on the subject. This has given us a better knowledge of yttrium, while yttria, which was formerly held to be a homogeneous substance, has proved itself a mixture of the oxides of various metals, of which, however, only one or two have as yet been isolated; witness the discovery of scandium by Nilson, and of ytterbium by Marignac.

The elements molybdenum, tungsten, and uranium, which belong to the same group as chromium, were discovered, like the latter itself, in the first decades of the modern period; their investigation is still being proceeded with, thanks to the extraordinary diversity of the compounds which they form with other elements (see below). Vauquelin discovered chromium in 1797 as a constituent of red lead spar, and he also contributed materially to a knowledge of its compounds; Klaproth pointed out independently at the same time that there was probably a new metal contained in that mineral. The presence of molybdenum and tungsten in their oxygen compounds was foreseen by Scheele and Bergman, the former being isolated in 1783 by Hjelm, and the latter by d'Elhujar. Uranium, lastly, or rather an oxide of it which was looked upon as the element, was detected by Klaproth in 1798 as a principal constituent of pitchblende; Péligot¹ was the first to correct this error by proving that the supposed element contained oxygen, and

¹ *Ann. Chim. Phys.* (3), vol. v. p. 5.

also by preparing metallic uranium itself. The atomic weights of chromium and uranium, as determined by Pélignot, have been corroborated by the recent estimations of Cl. Zimmermann and Berlin. For molybdenum, a somewhat higher value than that obtained by Berzelius has been arrived at from the work of Dumas, Rammelsberg, and others. The atomic weight got by Schneider, Marchand, and others for tungsten, viz. 183.5, has maintained its ground.

The elements which resemble tin in character, viz. titanium, zirconium, and thorium (to which germanium has within the last few years been added), belong practically to the chemical history of this century; for, although the oxides of titanium and zirconium were discovered at the end of last century, the isolation of the elements themselves was first accomplished by Berzelius, by means of the method already mentioned, viz. the decomposition of the double fluorides with potassium. Berzelius¹ also discovered thoria (ThO_2) and thorium in 1828; the atomic weight of this element was, however, only definitely established later on by Nilson,² the value then obtained by him having been quite recently corroborated by the determination of the vapour density of thorium chloride.³ Germanium, the youngest at present of all the elements, was discovered a few years ago by Cl. Winkler, and led in his hands to some admirable experimental work,⁴ which threw the clearest light upon its nature and that of its compounds. The impulse to look for a new element was given him by the analysis of a Freiberg silver ore, which invariably showed a deficit of about 7 per cent. This led to the surmise that some substance was present for which the analytical methods in use did not suffice, just as in the case of cesium, already mentioned. The atomic weight of germanium, as determined by Winkler, agrees with the position which falls to this element in the periodic system.

Vanadium, tantalum, and niobium—elements nearly

¹ *Pogg. Ann.*, vol. xvi. p. 385.

² *Ber.*, vol. xv. p. 2527.

³ Nilson u. Krüss, *Ber.*, vol. xx. p. 1671.

⁴ *Journ. pr. Chem.* (2), vol. xxxiv. p. 177; vol. xxxvi. p. 177.

related to antimony and bismuth—have only become well known through comparatively recent researches. Vanadium, recognised as a constituent of certain lead ores by del Rio so early as 1801, but more definitely by Sefström in 1830, was isolated in the metallic state by Roscoe¹ in 1867, who proved that the substance hitherto taken to be an element contained oxygen and nitrogen. The chemical relations of this element and its compounds were admirably worked out by him, and the atomic weight determined with certainty.

The investigations of Hatchett, Ekeberg, Wollaston, and Berzelius on the minerals columbite and tantalite, in the two first decades of our century, had already pointed to the presence of the elements which afterwards received the names of tantalum and niobium, without, however, these being got at themselves. Nor did the work of H. Rose² lead either to their isolation or to a correct knowledge of their compounds, for in this case, too, niobium dioxide (Nb_2O_5) was regarded as the element itself. It was the researches of Blomstrand³ and of Marignac⁴ which first furnished definite standpoints for a review of the chemical behaviour of the two elements and their compounds.

The metals of the platinum group, with the exception of platinum itself,⁵ have all been discovered during this century, as constituents of platinum ores. Platinum itself was only obtained perfectly pure after suitable methods had been worked out for separating it from the accompanying metals. Its employment for making certain kinds of apparatus, so important for the development both of scientific and of technical chemistry, also belongs to the present period.

Palladium came in 1803 under its present name into commerce, as a new metal, without its discoverer being

¹ *Phil. Trans.* for 1868, p. 1; or *Ann. Chem.*, Suppl., vol. vi. p. 86.

² *Pogg. Ann.*, vol. xcix. p. 80; vol. civ. p. 432.

³ *Journ. pr. Chem.*, vol. xcvii. p. 37.

⁴ *Ann. Chim. Phys.* (4), vol. viii. p. 5.

⁵ Cf. p. 141.

known; it was only at a later date that it was learnt to have been isolated by Wollaston from platinum ore.¹ The remarkable property which it possesses of combining with hydrogen was first observed by Graham.² The discovery of palladium led Wollaston³ on to that of another of the platinum metals, rhodium, which he thus named because of the rose-red colour of its solutions. It was investigated carefully by Berzelius,⁴ who made a minute study of the platinum metals generally, and by Claus;⁵ the separation of rhodium from other metals is due primarily to Bunsen,⁶ and to Deville and Debray. Tennant⁷ was the first to direct the attention of chemists to iridium and osmium, as two new metals which were contained in the residues left from the solution of platinum ores; while to Deville and Debray⁸ we mainly owe the mode of preparing both elements (the heaviest substances as yet known) pure. Ruthenium, lastly, was likewise discovered in platinum ores, as well as in osmiridium, by Claus,⁹ who has further given us most of our knowledge of this element and its compounds, together with its atomic weight. Debray¹⁰ has also recently investigated ruthenium and its oxygen compounds.

The atomic weights of the platinum metals have as yet only in part been determined with the requisite definiteness. For platinum itself the most reliable value was supposed to be that obtained by Berzelius, viz. 196.7, until Seubert¹¹ showed (in 1880) that this figure was too high by at least two units. The atomic weights of palladium, rhodium, and osmium were also determined by Berzelius, but require further corrobora-

¹ *Phil. Trans.* for 1804, p. 428.

² *Phil. Mag.* (4), vol. xxxii. p. 516.

³ *Phil. Trans.* for 1804, p. 419.

⁴ *Pogg. Ann.*, vol. xiii. p. 437.

⁵ *Beiträge zur Chemie der Platinmetalle* (Dorpat, 1854), ("Contributions to the Chemistry of the Platinum Metals").

⁶ *Ann. Chem.*, vol. cxlvi. p. 265.

⁷ *Phil. Trans.* for 1804, p. 411.

⁸ *Comptes Rendus*, vol. lxxxi. p. 839; vol. lxxxii. p. 1076.

⁹ *Ann. Chem.*, vol. lvi. p. 257; vol. lix. p. 284.

¹⁰ *Comptes Rendus*, vol. cvi. pp. 100, 328.

¹¹ *Ann. Chem.*, vol. ccvii. p. 29; *Ber.*, vol. xxi. p. 2179; also Dittmar and Arthur, *ibid.*, vol. xxi. ref. 412.

tion; this applies more especially to that of osmium,¹ which, according to Seubert's investigations, possesses a much lower atomic weight than that which Berzelius gave to it, and the same thing holds good for iridium.²

The above short survey of the discovery of elements during the present chemical period is sufficient to allow of our properly appreciating the extent of the achievements in this branch of the science. Since chemists have had before their eyes the task of assigning a definite place in the periodic system to each element, the discovery of a new one has possessed quite another charm and also a far higher significance than was formerly the case. What is now aimed at is to determine the atomic weight of each with such accuracy and to examine its chemical behaviour with such completeness as to permit of its being classified in this system. In the case of none among the recently discovered elements have those efforts been followed with such signal success as in that of the youngest of all, germanium.

We find in chemical literature many accounts of supposed new elements, which afterwards turned out either to have been prepared before, or to be mixtures of substances partly already known and partly unknown. A passing reference may be made here to the fantastic attempts of Winterl³ at the end of last and beginning of this century, who imagined that he had decomposed several metals into different elements. But even investigators of eminence fell into errors which could only be explained by defects in the analytical methods of their day; thus Bergman (in 1781) looked upon iron phosphide, prepared from "cold-short" iron by means of hydrochloric acid, as a new metal to which he gave the name of *siderum*, and Richter claimed impure nickel as an element, terming it *nickolanum*. Even Berzelius thought that he had discovered (in 1815) a hitherto unknown earth in some

¹ Seubert's latest results go to prove that the atomic weight of osmium is much lower (about 191) than that hitherto assumed (198.6).

² *Ber.*, vol. xi. p. 1770.

³ Kopp, *Gesch. d. Chemie*, vol. ii. p. 282.

Swedish minerals, but he corrected the error himself by showing that the supposed new body was phosphate of yttria. The history of the cerium metals, to which yttrium belongs, and also of tantalum and niobium, shows more than that of any others a great many such errors, while even at the present day a number of new elements are being brought forward whose homogeneous nature is in the highest degree doubtful, *e.g.* *decipium*, *mosandrium*, and *philippium*.¹

Extension of the Knowledge of Inorganic Compounds.

The general standpoints arrived at during the present chemical period for the comprehension of inorganic chemical compounds, more especially the opinions with regard to the constitution of acids, bases, and salts, have been entered into in detail in the first section of this book. It remains now to give an account of the development of special knowledge in this branch of the science. An exhaustive treatment of the subject is of course impossible here: only those researches of particular moment, which have materially aided in extending the knowledge of chemistry, can be mentioned.

Hydrogen Compounds of the Halogens.

The remarkable behaviour of hydrogen with respect to chlorine,—the readiness with which those two gases combine, was first investigated by Davy and Gay-Lussac, and afterwards made the subject of important physico-chemical work² by Roscoe and Bunsen.³ The researches of Davy and Faraday⁴ contributed greatly to a more intimate knowledge of hydrochloric acid, showing as they did how to condense the gas, while those of Roscoe and Dittmar⁵ established the chemical relations existing between hydrochloric acid

¹ *Comptes Rendus*, vol. lxxxvii. pp. 148, 559, 632, etc.

² Cf. *History of Physical Chemistry*.

³ *Phil. Trans.* for 1857, p. 355; or *Pogg. Ann.*, vol. c. p. 43; *Ann. Chem.*, vol. xevi. p. 357.

⁴ *Phil. Trans.* for 1823, p. 164.

⁵ *Ann. Chem.*, vol. cxii. p. 337.

and water. Gay-Lussac and Balard investigated hydriodic and hydrobromic acids, while the fundamental researches of Gay-Lussac, Thénard, and Berzelius contributed a knowledge of hydrofluoric acid in aqueous solution, and those of Gore¹ and Frémy² of this acid in the gaseous state, these latter thus establishing its composition. Nicklès fell a victim to the frightful action of anhydrous hydrofluoric acid in 1869. Ampère was the first to point out the analogy between fluorine and chlorine.

Oxygen Compounds of Hydrogen and of the Halogens.

The investigations which led to a knowledge of the composition of water have already been described; the first quantitative determination of its constituents, to which no exception could be taken, was made by Berzelius and Dulong.³ The discovery of peroxide of hydrogen⁴ by Thénard in 1818 showed that water was not the only oxide of that element, while the chemical behaviour of this peroxide, which was examined by Thénard, Schönbein, etc., and of recent years by Schöne⁵ and Traube,⁶ stamps it as one of the most remarkable of inorganic compounds. It also appears to play an important part in many of the processes of nature, and the interest in it is heightened still further by the significance which it promises to have for technical chemistry.

The various degrees of oxidation of which chlorine, iodine, and bromine are capable have been the cause of much valuable work since the beginning of our century, *e.g.* that of Gay-Lussac on chloric acid, of Stadion on perchloric acid, of Davy and Stadion on chlorine peroxide, of Millon⁷ on chlorous acid, and of Balard⁸ on hypochlorous acid. The

¹ *Phil. Trans.* for 1869, p. 173.

² *Ann. Chim. Phys.* (3), vol. xlvii. p. 5.

³ *Ibid.*, vol. xv. p. 386.

⁴ *Ibid.*, vol. viii. p. 306 (1818).

⁵ *Ann. Chem.*, vol. xcii. p. 258 (Schöne gives here a review of the previous literature on the subject).

⁶ Cf. *Ber.*, vol. xx. p. 3345.

⁷ *Ann. Chim. Phys.* (3), vol. vii. p. 298.

⁸ *Ibid.*, vol. lvii. p. 225.

knowledge of some of these compounds was much enlarged by Pebal's latest researches,¹ which established the nature of the so-called euchlorine and of chlorine peroxide. The oxygen compounds of iodine became known through the investigations of Davy and Magnus; periodic acid, which was discovered by the latter,² and iodic acid led later on to a knowledge of several series of salts, from whose composition important conclusions as to the saturation-capacity of iodine, and therefore of the halogens generally, were drawn.

Sulphur, Selenium, and Tellurium Compounds.

To the early known compounds of sulphur and oxygen, sulphurous and sulphuric acids (the anhydride of the latter of which was discovered by Vogel and Dobereiner), others came to be added, viz. "hyposulphurous acid" by Gay-Lussac,³ and dithionic acid by Welter and Gay-Lussac (in 1819). The constitution of the first of these, which is really thiosulphuric acid, was only made out at a much later date.⁴ The thio-acids containing more sulphur, and nearly related to sulphuric acid, were discovered at the beginning of the forties by Langlois, Fordos and Gélis, and Wackenroder; the question as to whether the pentathionic acid of the latter really exists has quite lately been vigorously discussed.⁵

To the above sulphur acids there has of late years been added Schützenberger's hyposulphurous acid (H_2SO_2) the chemical behaviour of which is of great interest.⁶ The two well-known oxides of sulphur have received an addition in R. Weber's sesquioxide, S_2O_3 .⁷ Lastly, mention may be made here of per-sulphuric acid, whose existence Berthelot

¹ *Ann. Chem.*, vol. clxxvii. p. 1; vol. cexiii. p. 113.

² *Pogg. Ann.*, vol. xxviii. p. 514.

³ *Ann. Chim.*, vol. lxxxv. p. 199; sodium hyposulphite (thiosulphate) was first prepared by Chaussier in 1799, and afterwards more carefully examined by Vauquelin.

⁴ Cf. Schorlemmer, *Journ. Chem. Soc.* (2), vol. vii. p. 256.

⁵ Cf. Curtius u. Henkel, *Journ. pr. Chem.* (2), vol. xxxvii. p. 37; Debus, *Journ. Chem. Soc.*, vol. liii. p. 278; or *Ann. Chem.*, vol. cexliv. p. 76.

⁶ *Comptes Rendus*, vol. lxix. p. 169.

⁷ *Pogg. Ann.*, vol. clvi. p. 53.

has shown to be probable, and for the anhydride of which he assumes the formula S_2O_7 . The enormous impetus given to chemical industries generally by the development of the sulphuric acid manufacture must also be referred to.

The compounds of selenium with hydrogen and oxygen were investigated by Berzelius, and an account of them given in his memorable treatise. After him there came Mitscherlich, who discovered selenic acid, and therewith furnished a beautiful confirmation of the analogy between selenium and sulphur. This chemical similarity has not however been maintained in all respects, Michaelis¹ having recently shown that the salts of selenious acid probably possess a constitution different from that of the sulphites.

The chlorine compounds of sulphur, selenium, and tellurium, whose study has aided in the characterisation of these elements, have been worked at at various times; by his investigation of tellurium tetrachloride Michaelis has lately furnished an excellent proof of tellurium being tetratomic.

Even if we desired to mention only the more important of the investigations which have aided in the discovery and elucidation of the hydrogen, oxygen, and haloid compounds of nitrogen, phosphorus, arsenic, and antimony, it would be necessary to record a long series. Among them were the researches of Davy, Berthollet, and Henry, which made clear the composition of ammonia,—so long looked upon as containing oxygen. The discovery of phosphuretted hydrogen (PH_3) by Gengembre² in 1783, and the examination of it by Pelletier (who was the first to prepare it pure), only became fruitful after Davy's investigations; the last-named elucidated the composition of this gas, and pointed out its analogy to ammonia, this being emphasised still more sharply by H. Rose later on. Thénard³ discovered liquid phosphuretted hydrogen, and recognised in it the cause of the spontaneous inflammability of the not completely pure gaseous compound. Arseniuretted and antimoniuiretted hydrogens,

¹ *Ann. Chem.*, vol. ccli. p. 150.

² *Crell's Ann.*, vol. i. p. 450.

³ *Ann. Chim. Phys.* (3), vol. xiv. p. 5.

which are analogous to ammonia in composition, were first obtained in the pure state by Soubeiran¹ and Pfaff.² The former compound cost Gehlen his life in 1815, from his not suspecting its extreme poisonousness. The great importance of the formation of arseniuretted hydrogen for the detection of minute quantities of arsenic in judicial-chemical analyses is well known.

The oxygen compounds of nitrogen played, as already described, a significant part in the history of the atomic theory, even although the true composition of all these oxides was not at that time made out. The number of the oxides of nitrogen known in Dalton's time was supplemented by nitrogen peroxide, whose relation to the others was arrived at through the researches of Berzelius, Gay-Lussac, and Dulong, and by nitric anhydride, discovered by St. Claire Deville. The various obscure points with respect to nitrous acid and nitrogen peroxide have been for the most part explained by the recent investigations of Hasenbach³ and Lunge.⁴ The discovery of hyponitrous acid,⁵ the acid corresponding to nitrous oxide, enlarged still further the series of the oxy-acids of nitrogen. Reference must also be made here to the important discovery of hydroxylamine,⁶ which, from its value as a reagent, has led to a knowledge of many remarkable compounds, especially in organic chemistry. Frémy's *acides sulfazotés* have only of late years been recognised as being really sulphoxyl-derivatives of ammonia and hydroxylamine (e.g. $\text{HO.N.}(\text{SO}_2\text{OH})_2$ and $\text{HO.NH.SO}_2\text{OH}$).⁷ The discovery of amido-amine⁸ (diamidogen or hydrazine, $\text{H}_2\text{N.NH}_2$), which fills up a long-felt gap, and that of the very striking compound, azo-imide (N_3H),⁹ are also worthy of note.

¹ *Ann. Chim. Phys.* (2), vol. xxiii. p. 307.

² *Pogg. Ann.*, vol. xl. p. 135.

³ *Journ. pr. Chem.* (2), vol. iv. p. 1.

⁴ Cf. *Ber.*, vol. xviii. p. 1376; vol. xxi. p. 67.

⁵ Divers, *Proc. R. S.*, vol. xix. p. 425; Zorn, *Ber.*, vol. x. p. 1306.

⁶ Lossen, *Ann. Chem.*, Suppl., vol. vi. p. 220.

⁷ Cf. Raschig's admirable paper (which also gives a review of the previous literature on the subject), *Ann. Chem.*, vol. cxxli. p. 161.

⁸ Curtius u. Fay, *Journ. pr. Chem.* (2), vol. xxxix. p. 27.

⁹ Curtius, *Ber.*, vol. xxiii. p. 3023.

Of the oxygen compounds of phosphorus, phosphorous and phosphoric acids were known, although very imperfectly, in Lavoisier's time; the former of the two was first prepared pure by Davy, by treating phosphorus trichloride with water, but its chemical constitution was only cleared up by later investigations. The recent admirable paper of Thorpe and Tutton¹ upon phosphorous oxide, P_4O_6 , shows that the real properties of this substance are very different from those formerly attributed to it. The labours of Clarke, Gay-Lussac, and Stromeyer prepared the way for the recognition of the mutual relations existing between ortho-, pyro-, and metaphosphoric acids, these being subsequently worked out by Graham;² and upon them Liebig established his far-reaching theory of polybasic acids.³ Hypophosphorous acid, whose salts were discovered by Dulong in 1816, has been the subject of important investigations and discussions.⁴ Hypophosphoric acid,⁵ $H_4P_2O_6$, has also lately been added to the above oxygen compounds.

The discovery of the halogen compounds of nitrogen and phosphorus has proved of particular interest, the latter of the two being largely employed for the preparation of many other substances, because of the readiness with which they enter into reaction. Chloride of nitrogen was discovered by Dulong,⁶ who suffered serious injury in consequence of some of it exploding unexpectedly; this dangerous substance has of late been made the subject of important investigations by Gattermann,⁷ who has succeeded in preparing the pure chloride, NCl_3 . The analogously formed iodide of nitrogen was first prepared by Serullas,⁸ while Bunsen, Stahlschmidt, and, quite recently, Raschig⁹

¹ *Journ. Chem. Soc.*, vol. lvii. p. 545.

² *Phil. Trans.* for 1833, vol. ii. p. 253.

³ Cf. p. 230.

⁴ Cf. Wurtz, *Ann. Chem.*, vol. xliii. p. 318; vol. lxviii. p. 41.

⁵ Salzer, *Ann. Chem.*, vol. clxxvii. p. 322; vol. xciv. p. 28; vol. ccxi. p. 1; vol. ccxxxii. p. 114; Sanger, *ibid.*, vol. ccxxxii. p. 1.

⁶ *Schweigger's Journ.*, vol. viii. p. 302.

⁷ *Ber.*, vol. xxi. p. 751.

⁸ *Ann. Chim. Phys.*, vol. xlii. p. 200.

⁹ *Ann. Chem.*, vol. ccxxx. p. 212.

have aided towards a knowledge of its composition. The chlorine compounds of phosphorus were prepared in the first decade of our century, the trichloride by Gay-Lussac and Thénard, and the pentachloride by Davy. The pentafluoride, isolated by Thorpe,¹ is of especial interest from its not decomposing even at high temperatures, unlike the other penta-haloid compounds of phosphorus. Wurtz discovered phosphorus oxychloride, which is of great value as a reagent in organic work, and H. Rose antimony pentachloride.

The haloid compounds of boron and silicon were mainly investigated by Berzelius and, later, by Wöhler and Deville,² and they constituted the material from which those elements themselves and others of their compounds were prepared; the above researches, in fact, greatly extended the knowledge of these substances generally. Among other points, the discovery of boron nitride and silicon hydride may be mentioned here.³ To the careful investigation of volatile silicon compounds is due the definite establishment of the atomic weight of that element, and, with this, of the composition of silica, to which another formula than the present was previously given.

Of the simple compounds of carbon, which from long custom are assigned to inorganic chemistry, the greater number were discovered and examined at the beginning of this century. Details have already been given with respect to carbonic acid and carbonic oxide. The study of the phenomena of combustion, in particular of the processes which go on in the flame of burning carbon compounds, in which the two gases just mentioned play a prominent part, was first taken up by Davy, who advanced the subject immensely by his beautiful researches. We must also refer here to the more recent investigations of Frankland, Blochmann, and Heumann on the theory of luminous flames.

Carbon oxychloride or phosgene, which has proved of

¹ *Ann. Chem.*, vol. clxxxii. p. 201.

² *Ibid.*, vol. cv. p. 67 *et seq.*

³ Wöhler u. Buff, *Ann. Chem.*, vol. cii. p. 120.

exceptional value as a reagent in organic chemistry, was first prepared by Davy in 1811, but carbon oxysulphide only comparatively recently by von Than.¹ Carbon disulphide, on the other hand, was obtained by Lampadius so early as 1796, and more minutely investigated by Clément and Desormes in 1802; its composition was arrived at correctly by Vauquelin and Berzelius, after the most confused opinions had previously been expressed as to its containing hydrogen and nitrogen. The profound influence which the classical researches of Gay-Lussac on cyanogen and its compounds exercised upon the development of chemistry has already been referred to.

Extension of the Knowledge of Metallic Compounds.

From the endless number of investigations which have contributed materially towards a knowledge of the metallic compounds, and therewith of the metals themselves, the most important must now be mentioned, if they have not already been so in the general section of this book.

The discoverers of the alkali metals also aided largely in their investigation; thus to Davy is due our knowledge of the oxides of potassium and sodium, to Gay-Lussac that of the corresponding peroxides, and to Bunsen that of rubidium and cesium compounds. The great fertilising influence which these labours on the compounds of the alkalies exercised upon the development of chemical industries will be detailed under the history of technical chemistry.

The peroxides of barium and calcium were discovered by Gay-Lussac and Thénard. The knowledge of the nature of chloride of lime was advanced by the researches of Balard, who was the first to express the opinion—still held by many—that this substance was a double compound of calcium chloride and hypochlorite. Since that time numerous further experiments have led many to regard it as an oxy-

¹ *Ann. Chem., Suppl.*, vol. v. p. 236. The properties of the pure compound were first established by Klason (*Journ. pr. Chem.* (2), vol. xxxv. p. 64).

chloride of calcium, which has given rise to a large amount of discussion.¹

The investigations which led to a knowledge of the compounds of beryllium and thallium have been cited above.² New oxygen compounds of copper, in addition to the oxides already known, were obtained by Rose³ and Thénard, while Wöhler discovered silver suboxide and peroxide; it must however be mentioned here that the existence of the former of these has of late been vigorously contested.⁴ The application of silver salts for the fixation of light-impressions (*i.e.* in photography), so pregnant in its results, will be treated of under the history of physical chemistry. Those chemists who shared in the discovery and investigation of aluminium, indium, and gallium, also contributed at the same time to a knowledge of their compounds. With respect to the compounds of alumina, pure chemistry has frequently been called upon to answer difficult questions pertaining to the technical manufacture of ultramarine, porcelain, glass, etc.

The compounds of the metals which form the iron group have been the object of a very large number of investigations, among which we may mention those on the different stages of oxidation of manganese by Liebig and Wohler,⁵ Mitscherlich,⁶ and, quite recently, Franke.⁷ The chlorine and fluorine compounds of manganese were studied by Christensen. To the two oxides of iron (FeO and Fe_2O_3), a knowledge of which we owe to Proust, and whose composition was established by Berzelius, Frémy added ferric acid, which he also carefully investigated. Light was thrown upon the nature of the cyanogen compounds

¹ Cf. the work of Göpner, Wolters, Kraut, Lunge, and others.

² Cf. pp. 379, 380.

³ *Pogg. Ann.*, vol. cxx. p. 1.

⁴ Wöhler, *Ann. Chem.*, vol. xxx. p. 1. Friedheim, *Ber.*, vol. xxi. p. 316. On the other hand, von der Pfordten, who at first believed that he had proved the existence of silver suboxide, has recently expressed himself in favour of a "hydrate of silver" (*Ber.*, vol. xxi. p. 2288).

⁵ *Pogg. Ann.*, vol. xxi. p. 584.

⁶ *Ibid.*, vol. xxv. p. 287.

⁷ *Journ. pr. Chem.* (2), vol. xxxvi. pp. 31, 166, 451.

of iron by the beautiful researches of Gay-Lussac, Berzelius, Gmelin (who discovered potassic ferricyanide), and Liebig, out of which the present views held with regard to these substances have developed themselves. The nitroprussides, so nearly allied to the ferrocyanides, were first obtained by Playfair, but their constitution has yet to be cleared up.

The chemistry of the cobalt salts was enriched by the discovery of the remarkable and highly varied ammonio-cobaltic compounds, which, observed for the first time in 1851, were afterwards investigated by Fr. Rose, Gibbs, Frémy, and especially Jörgensen.¹ The last-named has brought the extraordinarily difficult question of the chemical constitution of these bodies materially nearer to its solution, by systematically examining the ammonia compounds of those other metals analogous to cobalt in this respect—chromium and rhodium.²

The various combining relations which the different members of the group of elements comprising molybdenum, tungsten, and uranium, show towards other elements, have only been fully understood of recent years. The admirable work of Berzelius on molybdenum compounds has been supplemented by that of Krüss³ on the sulphides, and of Muthmann⁴ on the oxides, as well as by the earlier investigations of Blomstrand, Debray, Liechti, and Kempe on the halogen compounds of molybdenum. The chlorides of tungsten were examined in detail by Roscoe, who thereby advanced the knowledge of the saturation-capacity of this element. The complicated salts of tungstic acid were first studied by Margueritte, Scheibler, Marignac, and v. Knorre, but their ultimate constitution, as well as that of the phospho-molybdic and phospho-tungstic acids, has still to be unravelled. The chemical nature of uranium and its compounds has been worked out with most success by Cl.

¹ Cf. *Journ. pr. Chem.* (2), vol. xxiii. p. 227; vol. xxxi. pp. 49, 262; vol. xxxix. p. 1; vol. xli. p. 429.

² *Ibid.* (2), vol. xxv. pp. 83, 321; vol. xxx. p. 1; vol. xxxiv. p. 394.

³ *Ann. Chem.*, vol. ccxxv. p. 1.

⁴ *Ibid.*, vol. ccxxxviii. p. 109.

Zimmermann,¹ whose able researches largely supplemented the earlier ones of Péligot, Roscoe, and others.

Of the compounds of tin and its chemical analogues, the isomorphous double fluorides² aroused especial interest, from their proving the connection which exists between silicon, titanium, zirconium, and germanium. The peculiarity of titanium was set forth in a striking manner by the discovery of its nitrogen compounds,³ and more recently by the preparation of its various sulphides.⁴

To Roscoe's admirable work⁵ is due most of our knowledge of vanadium, he having correctly worked out the different stages of combination of this element with oxygen, chlorine, etc., and having put right the former erroneous assumptions with regard to their composition. Gerland's investigations⁶ on vanadyl salts and vanadic acids, and those of v. Hauer on the salts of the latter, have also aided here.

Similarly niobium, and tantalum, whose chemical nature had been completely misjudged, were given their proper position among the other elements by the recent investigations already cited, more particularly by the determination of the true composition of the chlorides of both and of niobium oxychloride,⁷ and by the examination of niobium fluoride and hydride.⁸

Valuable work has also lately been done on the compounds of gold,⁹ which has materially amplified the earlier researches of Proust, Berzelius, Figuier, etc., and has served to establish the chemical character of this element.

The literature on platinum and its compounds is very

¹ *Ann. Chem.*, vol. cexiii. p. 285 (contains a historical review); vol. ccxxxii. p. 274; also Alibegoff, *ibid.*, vol. ccxxxiii. p. 117.

² Marignac, *Ann. des Mines* (5), vol. xv. p. 221.

³ Wöhler, *Ann. Chem.*, vol. lxxiii. p. 43.

⁴ *Ann. Chem.*, vol. ccxxxiv. p. 257.

⁵ *Phil. Trans.* for 1869, p. 679; or *Ann. Chem.*, Suppl., vol. vii. p. 70.

⁶ *Ber.*, vol. ix. p. 874; vol. x. p. 2109; vol. xi. p. 98.

⁷ Deville and Troost, *Comptes Rendus*, vol. lx. p. 1221.

⁸ Krüss and Nilson, *Ber.*, vol. xx. p. 1676.

⁹ Cf. Krüss, *Ann. Chem.*, vol. ccxxxvii. p. 274 (contains a historical review); vol. ccxxxviii. pp. 30, 241; *Ber.*, vol. xxi. p. 126; Thorpe and Laurie, *Journ. Chem. Soc.*, vol. li. pp. 565, and 866.

voluminous, and gives evidence of most excellent experimental work. Reference may be made here to the discovery of the peculiar reactions to which platinum can give rise in virtue of its condensation of oxygen; and to the numerous investigations on the platinum-ammonium compounds, the first of which were prepared by Magnus, and whose peculiarities were studied by Gros, Reiset, Cléve, Thomsen, and Blomstrand. The recently published work of Jörgensen:¹ *Zur Konstitution der Platinbasen* ("On the Constitution of the Platinum Bases") marks an important step in the recognition of the constitution of these bodies.

The researches which have materially assisted towards a knowledge of the platinum metals have already been mentioned under the history of the individual elements.

If we throw a glance over the wide field of inorganic chemistry, with its seventy elements approximately and their endless compounds, we recognise the fact that in their classification the atomic theory primarily has rendered the most important service. The endeavour, too, to establish periodic relations between the properties of the elements and their atomic weights has introduced order among the motley array of the elements and their compounds. The question of the constitution of the latter allows in most cases of a simple and satisfactory answer; as soon, however, as the composition of inorganic compounds becomes complicated, the usual aids to the solution of such points no longer suffice. The consequence of this is that the rational composition of a large number of compounds, whose empirical composition has long been known, has not yet been cleared up; as examples of such we may refer to the metallo-ammonia compounds (*v.g.* those of cobalt and chromium), the poly-silicic acids, the tungstic acids, etc.

¹ *Journ. pr. Chem.* (2), vol. xxxiii. p. 489.

SPECIAL HISTORY OF ORGANIC CHEMISTRY IN THE NINETEENTH CENTURY.

The development of organic chemistry during the first few years of this century has already been described under the general history of the period (cf. p. 233); there, also, much of the pioneering work accomplished in this branch of the subject has been treated of, in so far, that is, as it had a determining influence on the origin and growth of important theoretical investigations. In this section the attempt will be made to pick out from the superabundance of work done in organic chemistry that which has proved of greatest significance, and to arrange it according to its nature (not according to its sequence in point of time),—more especially such investigations as have contributed to solving the question of the chemical constitution of whole classes of bodies. The general points of view by which experimenters have been guided in those researches have already been examined at various times in the first section of this book.

Before organic chemistry could be in a position to develop itself independently, the following two conditions had to be fulfilled:—In the first place, the determination of the empirical composition of organic substances was necessary (how this question was solved is described under the history of analytical chemistry):¹ in the second, it had to be proved that organic compounds were subject to the same atomic laws as inorganic, and that they were not, as many formerly assumed, to be classed as totally distinct from the latter. To Berzelius, more than to any other man, is due the setting aside of this dividing barrier between the two.

The most important methods, which have ever since remained standard ones in organic chemistry, were created by the fundamental researches of Gay-Lussac on cyanogen and

¹ Cf. p. 367.

its compounds, of Liebig and Wöhler on benzoyl, of Bunsen on the compounds of cacodyl, of Dumas and Péligot on wood-spirit, and by the investigations of Kolbe, Frankland, A. W. Hofmann, Williamson, Gerhardt, Wurtz, Kekulé, and others during the fifties and sixties. Many of those researches have already been referred to in the general section, because of the influence which they exercised on the development of views with regard to the chemical constitution of organic compounds; but it will not be altogether possible to avoid recurring to some of them in this portion of the book.

The recognition of the totally different behaviour of the so-called saturated, unsaturated, and aromatic substances was of the first importance for the systematising of organic compounds. A precise distinction between and definition of the above three classes, more especially of the two latter, has gradually been brought about in the course of the last few decades, as the knowledge of them has been extended.

Hydrocarbons and their Derivatives.

The hydrocarbons, from which as the simplest organic compounds all the others are derivable, have been, as befits their "typical" importance, the object of a great number of investigations, which have led to the development of doctrines of the utmost weight. We have only to think of the determination of the composition of marsh gas and of ethylene, which led to the recognition of multiple proportions, and therewith to the setting up of the atomic theory; of the significance of Faraday's researches on butylene for the evolution of what became known as polymerism; of the important labours of Regnault and others on ethylene and its haloid compounds, which afforded such rich food for the theories of substitution; and lastly, of the work of Kekulé and his pupils on benzene and its derivatives.

Mitscherlich's researches on benzene (which he then termed *Benzin*) taught new methods of preparing hydro-

carbons; the formation of this substance from benzoic acid, in consequence of the separation of carbon dioxide, became typical for a large number of similar reactions, *e.g.* the production of cumene from cumic acid, of methane from acetic acid, of chloroform from trichloroacetic acid, etc. Of great theoretical importance was Kolbe's mode of formation of hydrocarbons by the electrolysis of the alkaline salts of the fatty acids, and also that of Frankland by the action of zinc upon alkyl iodides; the latter pregnant investigations led to the discovery of the zinc alkyls, and opened up this especially fruitful field in the synthesis of organic compounds.¹ The researches of Wurtz,² which showed how the combination of different alkyls from hydrocarbons might be effected by the action of sodium upon two alkyl iodides, bore much fruit later on among the aromatic compounds; for, with this reaction as a model, the homologues of benzene were prepared synthetically, while at the same time this simple mode of formation allowed of their chemical constitution being deduced.³

Another synthesis⁴ of homologues of benzene, depending upon the peculiar interaction of aluminic chloride with mixtures of benzene and chlorine compounds (such as methyl chloride), has also proved itself of general application, as well as serviceable for the artificial production of other bodies, *e.g.* ketones, acids, etc. Notwithstanding the care with which these reactions have been studied, a conclusive explanation of the mode of action of the aluminium chloride has still to be given; this much, however, has been proved,—that their cause has to be sought for in the formation of peculiar intermediate compounds of the chloride with aromatic hydrocarbons (Gustavson).⁵

Berthelot's method⁶ of forming hydrocarbons out of different organic compounds by the action of hydriodic acid upon them at rather high temperatures, must also be men-

¹ Cf. p. 338.

² *Ann. Chim. Phys.* (3), vol. xlv. p. 275.

³ Cf. Fittig, *Ann. Chem.*, vol. cxxxi. p. 301.

⁴ Friedel and Crafts, *Comptes Rendus*, vols. lxxxiv., lxxxv., etc.

⁵ *Ber.*, vol. xi. p. 2751.

⁶ *Ann. Chim. Phys.* (4), vol. xx. p. 392.

tioned here, since it has led to important results in many cases. Reference must also be made to the method, so frequently employed, of reducing oxygen compounds to hydrocarbons by heating them with zinc dust.¹ The work of Berthelot on acetylene, of Butlerow and others on the butylenes and amylenes, of Freund on trimethylene, of W. H. Perkin jun. on the derivatives of tri- and tetramethylene, of Liebermann on allylene, etc., has materially enlarged our knowledge of the unsaturated hydrocarbons. The remarkable processes of the *isomerisation* of such compounds have quite recently been cleared up by the valuable researches of Faworsky.²

Out of the extraordinarily large number of investigations on aromatic hydrocarbons, whose constitution has given rise to important discussions, there may be mentioned here (in addition to the above) those of Fittig³ and Baeyer⁴ on mesitylene, which was found to be "symmetrical" trimethylbenzene, and also those of Graebe⁵ upon naphthalene, and of Graebe and Liebermann⁶ upon anthracene. Important conclusions were drawn from the two last with respect to the chemical constitution of these already long-known hydrocarbons, which from thenceforth were regarded as standing in a simple relation to benzene.

Other coal-tar hydrocarbons of complicated composition have likewise been satisfactorily investigated; thus phenanthrene, the isomer of anthracene, has been shown by Fittig and Graebe⁷ to be a diphenylene derivative of ethylene, fluorene by Fittig⁸ to be diphenylene-methane, and chrysene by Graebe⁹ to be phenylene-naphthalene-ethylene. To Bamberger¹⁰ is due the clearing up of the chemical nature of retene and pyrene. We must further refer to the admirable work of E. and O. Fischer, Zincke, and others on the

¹ Baeyer, *Ann. Chem.*, vol. cxl. p. 295.

² *Journ. pr. Chem.* (2), vol. xxxvii. pp. 382, 417, 532.

³ *Ztschr. Chem.* for 1866, p. 518.

⁴ *Ann. Chem.*, vol. cxl. p. 306.

⁵ *Ann. Chem.*, vol. cxlix. p. 22.

⁶ *Ibid.*, Suppl., vol. vii. p. 257.

⁷ *Ann. Chem.*, vol. clxvi. p. 361; vol. clxvii. p. 131.

⁸ *Ibid.*, vol. cxciii. p. 134.

⁹ *Ber.*, vol. xii. p. 1078.

¹⁰ *Ann. Chem.*, vol. cexxix. p. 102; *Ber.*, vol. xx. p. 365.

phenyl derivatives of methane, more especially triphenylmethane; this last was proved by E. and O. Fischer to be the mother-substance of exceptionally valuable aniline dyes, whose constitution was thus explained. The recent and striking researches of O. Wallach¹ on the hydrocarbons which occur in many plants, the *essential oils*, have begun to throw light on the nature of these.

The Alcohols and Analogous Compounds.

The close connection existing between the alcohols and the hydrocarbons was recognised as soon as methyl alcohol (the first member of a long series of compounds of this nature) had been successfully prepared from methane, by converting the latter into methyl chloride, and then transforming this into the alcohol. Formerly regarded as the hydrated oxides of hypothetical radicals, the alcohols were after this characterised as hydroxyl derivatives of the hydrocarbons. What an influence Williamson's researches on the formation of ether and Kolbe's views on the constitution of alcohol had upon the development of the opinions now held with regard to this point, has been already described.

Among the most important of the investigations which helped to establish our knowledge of the alcohols were those of Dumas and Péligot² on wood-spirit, whose analogy to ethyl alcohol they clearly recognised. The true composition of the latter was worked out by Saussure, who thus did away with the fundamentally erroneous ideas regarding it which had prevailed since the time of Lavoisier, although the latter had arrived at a correct knowledge of its constituents. Equally important were the fact that aethyl ($C_{16}H_{33}OH$), discovered by Chevreul, was characterised as an analogue of alcohol by Dumas and Péligot in spite of its unlikeness to the latter, and the corresponding proof by Cahours³ for the amyl

¹ Cf. *Ann. Chem.*, vols. cccxv. cccxvii. cccxx. cccxxviii. cccxxix. and ccxii.

² *Ann. Chim. Phys.*, vol. lviii. p. 5; vol. lxi. p. 93.

³ *Ibid.*, vol. lxx. p. 81; vol. lxxv. p. 193.

alcohol obtained from fusel oil, to which isobutyl alcohol¹ was afterwards added. The discovery of the secondary and tertiary alcohols, so memorable for the history of this class of compounds, was, as already stated, prognosticated by Kolbe. The series of the secondary carbinols was begun with isopropyl alcohol, isolated by Friedel, and that of the tertiary with Butlerow's trimethyl-carbinol. The modes of formation of these substances (that of isopropyl alcohol from acetone by the addition of hydrogen, and that of trimethyl-carbinol from acetyl chloride and zinc methyl) have since been extensively made use of for the preparation of analogous compounds.

Carbinols of other series were investigated by Cannizzaro, who discovered benzyl alcohol,² the simplest carbinol of the aromatic series, and by Cahours and Hofmann, who isolated allyl alcohol;³ while an accurate acquaintance with various new primary carbinols of the fatty series was arrived at by the valuable systematic researches of Lieben and Rossi.⁴ The above-mentioned investigations were also of great significance for the development of the views upon chemical constitution, and more especially upon the isomerism of organic compounds.

The knowledge of the polyatomic alcohols had its beginning in the already-mentioned important researches of Berthelot on glycerine, as representing the triatomic carbinols, and especially in those of Wurtz on the diatomic glycols. In connection with these we would call attention here to the notable discovery of the poly-ethylene alcohols, and of ethylene oxide (distinguished by the readiness with which it enters into reaction).⁵

The derivatives of the alcohols known as the simple ethers have, with common ethyl ether at their head, frequently been the subjects of important investigations. The discussions upon the constitution of ether and its mode

¹ Wurtz, *Ann. Chem.*, vol. xciii. p. 107.

² *Ibid.*, vol. cxxiv. p. 324.

³ *Ibid.*, vol. c. p. 356.

⁴ Cf. *Ibid.*, vol. clviii. p. 137.

⁵ *Comptes Rendus*, vol. xlviii. p. 101; vol. xlix. p. 813.

of formation—discussions which lasted for many years—were brought to an end by the work of Williamson and Chancel, which led to the discovery of mixed ethers.¹

The knowledge of the compound ethers, sometimes now also called *Esters*, has been greatly extended within the last sixty years. To the neutral ethers of the acids, the number of which has gone on continuously increasing (but regarding which it is impossible to mention here even the more important researches), there have been added the so-called ether-acids, whose chemical nature has been cleared up by the work of Hempel, Serullas, Magnus, and Regnault on ethyl-sulphuric and ethionic acids, of Pérouze on the ethyl-phosphoric acids, of Mitscherlich on ethyl-oxalic acid, and other more recent labours, *eg.* that upon phenyl-ethyl-sulphuric acid by Baumann, and upon ethyl-oxalic acid by Anschütz.

Certain of the compounds prepared from ethyl alcohol and other carbinols have played an important part in the synthesis of organic substances, thanks to their capability of reaction; we have but to recall here the discovery of sodium ethylate by Liebig, that of chloro-carbonic ether by Dumas, and Debus' investigations of the important products which result from the oxidation of ethyl alcohol by nitric acid.

The first step towards a knowledge of those compounds so nearly allied to the alcohols, which have received the generic name of *phenols*, was Laurent's investigation of carbolic acid and its derivatives.² Gerhardt was the first to point out the analogy between alcohol and phenol. Of great importance for the opening up of this class of compounds, and more especially for their technical production, was that mode of formation of phenol itself which was first observed by Kekulé³ and Wurtz,⁴ viz. by fusing benzenesulphonic acid with potash.

¹ Cf. p. 281.

² *Ann. Chim. Phys.* (3), vol. iii. p. 195.

³ *Lehrb. der. organ. Chemie*, vol. iii. p. 13.

⁴ *Ann. Chem.*, vol. cxliv. p. 121.

Carboxylic Acids.

A field of immense size and fruitfulness became opened up to chemical research with the systematic investigation of the acids contained in animal and vegetable fats, as well as in other natural products. The important work on the fatty acids, suggested in the first instance by Liebig, which was accomplished by his pupils Varrentrapp, Rochleder, Bromeis, Fehling, Redtenbacher, and others, and that of Heintz¹ upon palmitic and stearic acids, not only materially supplemented the earlier investigations of Chevreul on the fats, but led to the discovery of new and wider domains. Important methods for the separation of the fatty acids resulted from these labours. The common link which unites the compounds of this class was only discovered when their chemical constitution came to be understood. The successful efforts of Kolbe, who was the first to recognise acetic as methyl-carboxylic acid, and who established this view by direct experiment, have been already described in the general section. It has indeed been from acetic, as the most fully investigated of all the carboxylic acids, that our present ideas upon the constitution of the whole class of compounds have developed themselves. The recognition of the correct atomic composition of acetic acid by Berzelius in 1814, and of its relation to alcohol by Döbereiner in 1821, was of great importance for the solution of this problem.

After the constitution of the carboxylic acids had once been grasped, it became possible for Kolbe to predict the existence of other members of this class, as he had done in the case of the alcohols, and thus existing blanks could be filled up. Of special importance here was the discovery of isobutyric acid,² of the isomers of valeric acid — itself already long known, and of other acids richer in carbon, in the systematic investigation of which Lieben and

¹ *Ann. Chem.*, vol. lxxxiv. p. 297; vol. lxxxviii. p. 297; *Journ. pr. Chem.*, vol. lxvi. p. 1.

² Erlenmeyer, *Ztschr. Chem.* for 1865, p. 651.

Rossi¹ and Krafft, among others, have rendered great service.

The knowledge of the polybasic saturated carboxylic acids, whose chemical constitution was likewise only thoroughly cleared up through Kolbe's view, was greatly advanced by the work of Berzelius, Fehling, and others on succinic acid (synthetised from ethylene cyanide by Maxwell Simpson²), by that of Arppe on adipic acid and homologous compounds,³ and by the discovery and investigation of malonic acid,⁴ etc. The ethers of this last acid have served for the synthesis of homologues of malonic and other polycarboxylic acids,⁵ thanks to the facility with which they exchange hydrogen for sodium; while from aceto-acetic ether, which so closely resembles malonic, there have been prepared numerous compounds belonging to this class, to be afterwards systematically investigated. Drechsel's memorable synthesis of the simplest dibasic acid, oxalic, from carbon dioxide and sodium,⁶ also deserves mention here.

The wide field of unsaturated carboxylic acids, some of which (*e.g.* acrylic, angelic, fumaric, and maleic) were discovered at an early date, first became cultivated with success after a clear idea of the constitution of these compounds had been arrived at through Kekulé's admirable investigations⁷ on the two last-named and on the pyro-citric acids, which explained the behaviour of these bodies to nascent hydrogen, and after Frankland and Duppa⁸ had made their beautiful syntheses, which resulted in the conversion of oxalic ether into unsaturated carboxylic acids. In fact this last investigation led Frankland to express the view that acrylic acid and its homologues were

¹ Cf. *Ann. Chem.*, vol. clix. p. 75; vol. clxv. p. 116.

² *Proc. R. S.*, vol. x. p. 574; or *Ann. Chem.*, vol. cxviii. p. 373.

³ *Ann. Chem.*, vol. cxv. p. 143; vol. cxv. p. 288.

⁴ *Ibid.*, vol. cxxxi. p. 348.

⁵ Cf. Conrad, Bischoff, Guthzeit, *Ann. Chem.*, vol. cciv. p. 121; vol. ccix. p. 211; vol. ccxiv. p. 31.

⁶ *Ztschr. Chem.*, for 1868, p. 120.

⁷ *Ann. Chem.*, vol. cxxx. p. 21; vol. cxxxi. p. 81; Suppl., vol. i. p. 129; vol. ii. p. 108.

⁸ *Journ. Chem. Soc.*, vol. xviii. p. 133; or *Ann. Chem.*, vol. cxxxvi. p. 1.

derivatives of acetic acid, and a simple explanation was given of their transformation into the latter. The more recent systematic researches of Fittig¹ and his pupils on the unsaturated carboxylic acids have contributed in great degree to the rounding off and deepening of our knowledge of this class of compounds. The discovery of tetrolic and propiolic acids² prepared the way for an acquaintance with the carboxylic acids derived from acetylene.

The class of the aromatic carboxylic acids, with benzoic acid at their head, has been the subject of innumerable and fruitful researches. We have but to recall here the discovery of the peculiar mode of formation of these compounds from hydrocarbons by oxidation, as well as by the direct introduction of the elements of carbonic acid by means of aluminic chloride;³ and the splendid investigations on the di-, tri-, and poly-carboxylic acids of benzene,⁴ to the last class of which the already long-known mellitic acid was found to belong. The aromatic carboxylic acids of unsaturated character, like cinnamic acid, etc., proved particularly easy of examination after Perkin⁵ had worked out the reaction now known by his name—a reaction which can be generally applied to their formation. Lastly, the isolation of phenyl-propiolic acid⁶ and its derivatives has led to results of importance.

The discovery of the chlorides, anhydrides, and amides of the carboxylic acids deserves particular mention here, since these classes of compounds fill an important place in the history of organic chemistry. The first organic acid chloride was benzoyl chloride, obtained by Liebig and Wöhler by the action of chlorine on oil of bitter almonds, in their classical research already so frequently referred to.

¹ *Ann. Chem.*, vol. clxxxviii. p. 87; vol. cxev. p. 50; vol. cc. p. 21; vol. ccvi. p. 1; vol. ccviii. p. 37.

² Geuther, *Journ. pr. Chem.* (2), vol. iii. p. 448; Bandrowski, *Ber.*, vol. xiii. p. 2340.

³ Friedel and Crafts, *Comptes Rendus*, vol. lxxxvi. p. 1368.

⁴ Baeyer, *Ann. Chem.*, Suppl., vol. vii. p. 1; vol. clxvi. p. 325; Fittig, *ibid.*, vol. cxlviii. p. 11; Graebe, *ibid.*, vol. cxlix. p. 18, etc.

⁵ *Journ. Chem. Soc.*, vol. xxi. p. 53; or *Ann. Chem.*, vol. cxlvii. p. 229.

⁶ Glaser, *Ann. Chem.*, vol. cliv. p. 140; Baeyer, *Ber.*, vol. xiii. p. 2258.

The general mode for the preparation of such compounds, *i.e.* by acting upon organic acids with phosphorus pentachloride, is due to Cahours;¹ since then this reagent has been a standard one in organic chemistry, and has proved its value in the most various circumstances, but more especially for the replacement of oxygen or hydroxyl by chlorine. Phosphorus oxychloride was applied by Gerhardt,² and phosphorus trichloride by Béchamp³ for the same purpose.

The great capability of reaction which the acid chlorides possess had already been shown by Liebig and Wöhler in the case of benzoyl chloride, from which they prepared the amide of benzoic acid with ammonia, the ether with alcohol, and the sulphide with sulphide of lead, thus introducing at the same time general modes of formation for those classes of compounds. The acid chlorides afterwards led Gerhardt⁴ on to the important discovery of the acid anhydrides, which have likewise proved of great value for the synthesis of organic compounds. Brodie⁵ then prepared from those anhydrides the peroxides of the acid radicals, so remarkable in their behaviour, which have since been ranked alongside of peroxide of hydrogen. To the acid amides, a class which had been opened up by Dumas' discovery of oxamide, Gerhardt added the anilides, and thus gave the impulse to the subdivision of the former into primary, secondary, and tertiary amides. The discovery of the aminic acids and the imides of polybasic acids must also be mentioned here,—compounds which are closely related to the amides; oxamic acid was isolated by Balard, and succinimide by Fehling. And reference must be made, too, to the connection between the acid nitriles and the primary amides of the acids.

The investigation of certain derivatives of the carboxylic acids has led to results of very great moment, in that a thorough grasp has been gained of the relations of two

¹ *Ann. Chem.*, vol. lx. p. 254.

² *Ann. Chim. Phys.* (3), vol. xxxvii. p. 285.

³ *Comptes Rendus*, vol. xl. p. 944.

⁴ *Ann. Chem.*, vol. lxxxii. p. 131; vol. lxxxvii. p. 151.

⁵ *Proc. R. S.*, vol. xii. p. 655; or *Ann. Chem.*, vol. cxxix. p. 282.

great classes of compounds—the oxy- and amido-acids—to them. The distinct idea which is now connected with the terms “oxy-carboxylic acid” and “amido-carboxylic acid” has developed itself from lactic acid and alanin as oxy- and amido-propionic acids, and from those other compounds already known for such a long time before their constitution had been deciphered,—glycollic acid and glycocoll. The work of Wurtz,¹ and of R. Hofmann and Kekulé,² among others, upon those substances, and especially the decisive investigations of Kolbe, which furnished the key to a thorough explanation of the facts, laid the foundation of our present knowledge of these classes of compounds.³

Of great importance for the true recognition of the relations of the substances just named to one another, and to the carboxylic acids from which they are derived, was the transformation of the amido- into oxy-acids by means of nitrous acid (Piria, Strecker, etc.), and that of the latter into the corresponding carboxylic acids by means of hydriodic acid. In this way the constitution of malic, tartaric, aspartic, lactic, and many other acids was definitely arrived at,⁴ so that the method may be considered as a peculiarly valuable aid to the clearing up of the rational composition of many organic compounds. Wislicenus⁵ has contributed in a very marked degree to a knowledge of the various lactic acids, his work on the subject having helped much to extend the doctrine of isomerism. The idea of “physical isomerism,” which was started by the different behaviour of substances of the same composition towards polarised light, has since developed itself more and more. Pasteur’s admirable researches⁶ on levo- and dextro-tartaric acids, and on the inactive racemic acid produced by their com-

¹ *Ann. Chim. Phys.* (3), vol. lix. p. 171.

² *Ann. Chem.*, vol. cii. p. 11; vol. cv. p. 288.

³ Cf. p. 300.

⁴ Cf. Schmitt, *Ann. Chem.*, vol. cxiv. p. 106; Kolbe, *ibid.*, vol. cxxi. p. 232; Lautemann, *ibid.*, vol. cix. p. 268.

⁵ *Ann. Chem.*, vol. cxxviii. p. 11; vol. elxvi. p. 3; vol. elxvii. p. 302.

⁶ *Ann. Chim. Phys.* (3), vol. xxiv. p. 442; vol. xxviii. p. 56; vol. xxxviii. p. 437.

bination, having previous to this thrown much light upon the subject.

Once the constitution of many of the naturally occurring oxy- and amido-acids became known, the synthetic preparation of such compounds was merely a question of time; thus lactic acid was prepared artificially from propionic acid as well as from aldehyde,¹ inactive tartaric acid from dibromo-succinic,² citric acid from acetone,³ hippuric acid (first recognised as a definite compound by Liebig) from glycocoll,⁴ and salicylic acid from phenol.

This last leads us to the aromatic oxy-acids, and to the important method of their formation from phenates and carbonic acid, discovered by Kolbe.⁵ A complete explanation of this general reaction has only of late been given by Schmitt,⁶ who has proved that the production of an isomer (sodium phenyl-carbonate, $C_6H_5O.CO_2Na$) precedes that of the sodium salicylate. The observation that the phenates behave very differently according to the nature of their alkali,—that, for instance, phenol-potassium and carbonic acid yield the isomeric para-oxy-benzoic acid instead of salicylic,—deserves to be noted here as especially important. Ost's discovery of the phenol-di- and tri-carboxylic acids,⁷ which result from the same reaction at a higher temperature, must also be recalled.

Of late years a special group has been formed of a peculiar class of oxy-acids which readily change into the so-called *lactones*, with separation of water. Fittig,⁸ in conjunction with his pupils, has investigated this remarkable class of compounds systematically, and has largely contributed towards a knowledge of the relations between lactones

¹ Wislicenus, *Ann. Chem.*, vol. cxxviii. p. 11.

² Kekulé, *ibid.*, vol. cxvii. p. 124.

³ Grimaux and Adam, *Comptes Rendus*, vol. xc. p. 1252.

⁴ Dessaigne, *Jahresber. d. Chem.* for 1857, p. 367.

⁵ Cf. *Ann. Chem.*, vol. cxiii. p. 125; vol. cxv. p. 201; *Journ. pr. Chem.* (2), vol. x. p. 95.

⁶ *Journ. pr. Chem.* (2), vol. xxxi. p. 397.

⁷ *Ibid.* (2), vol. xiv. p. 95.

⁸ Cf. *Ann. Chem.*, vol. ccxxvi. p. 322; vol. ccxxvii. p. 1; vol. ccxvi. p. 27; vol. ccviii. p. 111; vol. ccl. p. 166; vol. cclvi. p. 50.

and the corresponding acids, and also of their constitution, which latter was formerly interpreted differently; thus, the simplest member of the series, butyro-lactone, was previously held to be the aldehyde of succinic acid. Many lactonic acids have also been examined, and found to be carboxylic derivatives of the lactones.

Aldehydes.

The knowledge of the aldehydes, so important from many different points of view, has gone on steadily increasing ever since bitter almond oil or benzoic aldehyde was first investigated by Liebig and Wöhler, and ordinary aldehyde also by the former; the latter compound, first obtained by Fourcroy and Döbereiner, was thoroughly examined by Liebig. The chemical constitution of the aldehydes and of the nearly allied ketones was first definitely grasped and given expression to by Kolbe. Both classes of compounds acquired an especial significance after their capacity for combining with other organic bodies became known, and was hence utilised for the synthesis of compounds richer in carbon.

Liebig¹ was the first to explain the relation of the aldehyde of acetic acid to alcohol on the one hand, and to acetic acid on the other, whereupon Berzelius pointed out clearly the analogy existing between aldehyde and acetic acid, and bitter almond oil and benzoic acid respectively. The mode of formation of the aldehydes, by oxidation of the alcohols, has since then remained the general one. It was only discovered at a much later date that members of this class of compounds could be prepared from the salts of the acids, by heating these with sodium formate.² Still more recent is the discovery of the method of preparing aromatic aldehydes from phenols and chloroform (*i.e.* nascent formic acid), a reaction which has led to the finding out of some remarkable compounds.³ The aldehyde of formic acid, the first member of

¹ *Ann. Chem.*, vol. xiv. p. 133; vol. xxii. p. 273.

² Piria, *Ann. Chem.*, vol. c. p. 114; Limpricht, *ibid.*, vol. ci. p. 291.

³ Reimer, *Ber.*, vol. ix. p. 423; Tiemann, *ibid.*, vol. ix. p. 824; vol. x. p. 63.

its series, was prepared by A. W. Hofmann,¹ the simplest representative of the di-aldehydes, glyoxal, having already been obtained long before by Debus (in 1856) as one of the products of the oxidation of alcohol. With regard to aldehydes of complicated composition, many of these were long ago isolated from various ethereal oils, *e.g.* oil of cinnamon, oil of cumin, etc., and recognised as analogues of ordinary aldehyde. This last compound has been ever and anew the subject of important investigations, more especially since Liebig and Fehling observed its tendency to polymerise (into para- and meta-aldehydes).² And those researches gained an increased interest through the discovery of aldol³ (a compound resulting from aldehyde, and of the same percentage composition with it), and of its nearly allied compound, crotonic aldehyde;⁴ the perception of the constitution of the last-named substance was of importance in that it led to an explanation of this "condensation," and therefore also of other similar processes.

The numerous investigations which have been made with the object of explaining such reactions of aldehydes with other compounds, under elimination of water, cannot be given in detail here. Reference can only be made to those of W. H. Perkin sen., who showed how the condensation of aromatic aldehydes with fatty acids might be effected, — a reaction which still continues to yield rich fruit;⁵ and to the researches of L. Claisen, who has systematically examined the manifold condensation processes of which the aldehydes and ketones are capable.⁶

Ketones and Ketonic Acids.

The work done upon the ketones, compounds so closely allied to the aldehydes, has also been most fruitful. The

¹ *Proc. R. S.*, vol. xvi. p. 156.

² *Ann. Chem.*, vol. xxv. p. 17; vol. xxvii. p. 319.

³ Wurtz, *Comptes Rendus*, vol. lxxiv. p. 1361.

⁴ Kekulé, *Ann. Chem.*, vol. clxii. pp. 92, 309.

⁵ Cf. *Ann. Chem.*, vol. cexvi. p. 115; vol. cccxxvii. p. 48, etc.

⁶ Cf. *Ibid.*, vol. clxxx. p. 1; vol. cccviii. p. 121; vol. cccxiii. p. 137; vol. cccxxvii. p. 261; *Ber.*, vol. xxi. p. 1135.

simplest member of this class of bodies, acetone, had already been known for a long time and had been the subject of frequent investigation when Liebig¹ definitely established its composition. Important points in the earlier history of the ketones were (1) the discovery of their mode of formation from acid chlorides and zinc alkyls,² and (2) the preparation of mixed ketones.³ The formation of those peculiar compounds, mesityl oxide, phorone, and mesitylene, from acetone was observed a long time ago, but it was only completely explained after similar processes depending upon the condensation of aldehyde had been correctly interpreted.

The transformation of ketones into secondary carbinols by the addition of hydrogen has been already spoken of.⁴ Equally worthy of notice was the conversion of acetone into pinacone,⁵ a diatomic alcohol, and that of the latter into pinacolone; those reactions, extended to other—especially to aromatic—ketones, have led to important results.⁶

Entirely new fields have been opened up by the investigation of the di-ketones, to which acetonyl-acetone, naphthoquinone, anthraquinone, and, as recent researches have shown, benzoquinone and similar compounds belong,—substances whose nature has been cleared up by the comprehensive labours of Graebe, Liebermann, Fittig, Zincke, Paal, Claisen, Combes, and others.

The acids known as croconic acid, carboxylic acid ($C_{10}H_4O_{10}$), etc., prepared from potassium carboxide, were obtained a long time ago by Will and Lerch; the recent beautiful researches of Nietzki⁷ have shown that some of them are related to benzoquinone, while others are derived from a compound (not yet isolated) containing five atoms of carbon in the molecule. The obscurity hitherto surrounding the constitution of these remarkable bodies has thus been dispersed.

¹ *Ann. Chem.*, vol. i. p. 223. ² Freund, *Ann. Chem.*, vol. cxviii. p. 1.

³ Williamson, *Journ. Chem. Soc.*, vol. iv. p. 238; or *Ann. Chem.*, vol. lxxx. p. 86.

⁴ Cf. p. 403.

⁵ Fittig, *Ann. Chem.*, vol. cx. p. 25; vol. cxiv. p. 54.

⁶ Cf. Zincke, *Ber.*, vols. x. and xi.

⁷ *Ber.*, vol. xviii. pp. 499 and 1833; vol. xix. pp. 293 and 772.

The so-called ketonic acids, certain of which (*eg.* pyro-racemic) have been known for a long time, have of late years awakened the interest of a large number of investigators, and rightly so; we have but to think of the splendid results, more especially from the synthetic point of view, which have been achieved with aceto-acetic ether,¹ levulinic acid,² acetone-dicarboxylic acid,³ benzoyl-carboxylic acid⁴ (which has become of importance through its relation to isatin), and other similar compounds. These ketonic acids acquire a still greater interest from the circumstance that they show a double chemical behaviour, their constitution, as judged from certain reactions, being that of hydroxyl compounds, and as judged from certain others, that of carbonyl ones.⁵ Attention will be called later on to the remarkable compounds which the ketones and aldehydes yield with hydroxylamine and phenyl-hydrazine.

Carbohydrates—Glucosides.

The sugar varieties, which are so widely distributed in nature, and many of which have been known from an early age, belong partly to the alcohols and partly to the aldehydes. Just as the practical interest in many of these bodies has increased in an extraordinary degree, so has also the purely scientific augmented with an advancing knowledge of the close relations which exist between the sugar varieties and compounds whose constitution has already been worked out. Thus, glucose has been transformed into mannite, which is now known to be primary hexyl alcohol containing six hydroxyl groups in place of five hydrogen atoms; the rational composition of saccharic, mucic, and levulinic acids, which are more or less intimately related to the sugars, has

¹ Cf. Wislicenus, *Ann. Chem.*, vol. clxxxvi. p. 161 (contains a historical review).

² Conrad's investigations showed this to be β -aceto-propionic acid (*Ann. Chem.*, vol. clxxxviii. p. 223).

³ v. Pechmann, *Ber.*, vol. xvii. p. 2542.

⁴ Claisen, *Ber.*, vol. x. p. 430.

⁵ See General Section, p. 334, note 1.

been arrived at ; and the acid ethers of the latter have been obtained, etc. Such observations as these give support to the assumption that those carbohydrates which are comprised under the term glucoses are to be regarded as hexatomic alcohols, from which two atoms of hydrogen have been withdrawn in such a manner that they contain the formyl of the aldehydes or the carbonyl of the ketones (Baeyer, Fittig, V. Meyer). And this view is further rendered more probable by E. Fischer's recent observation of the formation of peculiar products, which he terms *osazones*, by the action upon carbohydrates of phenyl-hydrazine.¹

The investigation of the individual sugars—of their chemical behaviour and the products of their decomposition—has been taken part in by a great number of chemists ; among those who have actively busied themselves with the subject we may mention Brouhardat, Brown and Heron, Kiliani, v. Lippmann, O'Sullivan, Salomon, Scheibler, Soxhlet, and Tollens.² With respect to the relations of the other carbohydrates (especially starch, dextrine, etc.) to the glucoses, a large amount of work has also been done, among others, by Brown and Heron,³ but much still remains to be accomplished.

The glucosides,⁴ which stand in the most intimate relation to the glucoses, and whose occurrence in the vegetable and animal kingdoms awakened the interest of chemists of the highest eminence at a very early date, have been the subjects of important work ever since the memorable investigation of Liebig and Wöhler on amygdalin, and that of Piria on salicin. Among other researches we would refer here to those of Will on myronic acid, of Tiemann and Haarmann on coniferin, of Will on aesculin, etc. ; researches which resulted in the elucidation of the decomposition-products of the glucosides named, and which laid the foundation

¹ Cf. *Ber.*, vol. xvii. p. 579 ; vol. xx. pp. 825, 2566 ; vol. xxi. p. 988.

² Cf. Tollens' *Kurzes Handbuch der Kohlenhydrate* ("Short Text-Book of the Carbohydrates," 1888).

³ *Journ. Chem. Soc.*, vol. xxxv. p. 596 ; or *Ann. Chem.*, vol. cxcix.

⁴ Cf. the article *Glycoside*, by O. Jacobsen, in Ladenburg's *Handwörterbuch der Chemie*.

for a knowledge of their constitution. The expectation that these natural products will ultimately be prepared artificially, will doubtless be fulfilled before long.

*Haloid Derivatives of the Hydrocarbons and other
Compounds.*

As an appendix to the results of the investigations named above, investigations which have largely increased our knowledge of the hydrocarbons, alcohols, carboxylic acids, aldehydes, and ketones, some others must be mentioned here which bear upon the haloid and other similar derivatives of those compounds.

Hand in hand with the examination of the hydrocarbons went that of their haloid- and nitro-derivatives; for in some cases those were easily obtained from the hydrocarbons, while in others they often served for the preparation of the latter. The formation of chlorine and bromine compounds from hydrocarbons was the subject of important discussions, arising from the experiments upon substitution-reactions made and suggested by Dumas and Laurent, and for the explanation of which special theories were brought forward; take, for example, the first investigations made in this direction, — those upon the action of chlorine on naphthalene, ethylene, and ethylene chloride.

Other views began to prevail when, with the setting up of a new theory of the aromatic compounds, the difference between the hydrogen atoms of the benzene molecule and those belonging to the substituting radicals which had entered it came to be recognised. This difference was markedly apparent in the case of the halogens, and was clearly demonstrated by the work of Kekulé, Fittig, Beilstein, and others.¹ Further, the study of the remarkable isomeric relations, predicted on theoretical grounds by Kekulé for the derivatives of benzene, led to the thorough examination of the haloid substitution-products of the aromatic hydrocarbons.

¹ Cf. *Ann. Chem.*, vol. cxxxvi. p. 301; vol. cxxxvii. p. 192; vol. cxxxix. p. 331.

Most instructive and successful were the researches on the so-called "halogen carriers," which include a large number of the elements,—those, namely, whose compounds with the halogens are capable of partially yielding up the latter again; this explains their action as halogen conveyers. The above action has been examined more especially in the case of the aromatic hydrocarbons; without entering into details, we would refer here to the investigations¹ on the subject carried out at L. Meyer's suggestion by Aronheim, Page, Scheufelen, Schwalb, and others, and to those of Willgerodt.² The earliest observations on this point were made by H. Müller in 1862, when he noticed how chlorine was conveyed by iodine in the action of the former upon benzene and its homologues.

The numerous researches on the combination of halogens with unsaturated hydrocarbons were of very great moment, the first example of such an addition being afforded by ethylene. It would be out of place here even to mention only the more important investigations bearing upon reactions of this nature; but it may be stated generally that our present views with respect to the constitution of unsaturated compounds have resulted in great degree from the behaviour of such hydrocarbons to the haloids and halogen hydrides.

The modes of formation of haloid derivatives of the hydrocarbons are typical, *i.e.* are also applicable to other classes of compounds, *e.g.* acids, ketones, etc. And the same holds good for the chemical behaviour of such compounds, this having been in most cases first established for the haloid derivatives. To mention only one or two of the researches which have advanced our knowledge of the subject,—take the discovery and investigation of trichloroacetic acid by Dumas,³ that of chloral by Liebig and Dumas,⁴ and that of monochloro-acetic and monochloro-propionic acids, from whose chemical behaviour the constitution of the corresponding oxy-

¹ Cf. *Ann. Chem.*, vol. ccxxxi. p. 152 (contains a historical review).

² *Journ. pr. Chem.* (2), vol. xxxiv. p. 264; cf. also Neumann, *Ann. Chem.*, vol. ccxli. p. 33 ("Sulphuric Acid as a Carrier of Iodine").

³ *Ann. Chem.*, vol. xxxii. p. 101.

⁴ *Ibid.*, vol. i. p. 189; *Ann. Chim. Phys.*, vol. lvi. p. 123.

and amido-acids was established by Kolbe. The method of obtaining haloid derivatives of the alcohols, acids, etc., from unsaturated compounds of the same nature, by the combination of the latter either with the halogens themselves or with their hydrogen acids, proved itself of great significance. We have but to recall to mind the compounds thus prepared from fumaric and maleic acids, the pyro-citric acids, allyl alcohol, and so on, in order to appreciate how they led to the recognition of the connection existing between the unsaturated compounds just named and saturated ones.

Nitro- and Nitroso-Compounds.

Mitscherlich's discovery and investigation of nitro-benzene¹ paved the way for a knowledge of the nitro-compounds; the formation of this substance from benzene and its relation to the latter were, however, only clearly understood after Dumas and Gerhardt's view of nitro-benzene being a substitution-product of benzene came to be adopted. Since then the group nitroxyl (NO_2) has been ranked as a substituent alongside of the halogens. There is scarcely any reaction which has been more frequently applied among the aromatic compounds than the action of nitric acid upon them; take, for instance, the discovery of nitro-naphthalene, of di- and tri-nitrobenzenes, and of the nitro-derivatives of benzoic acid, benzoic aldehyde, etc. Picric acid, which was so much earlier known than nitro-benzene, was first characterised as trinitro-phenol by Gerhardt. Attention will be called later on to the history of some of the classes of compounds proceeding from these nitro-derivatives, *e.g.* the amines and azo-compounds, which have been destined to play such a prominent part in industrial chemistry.

The first nitro-derivatives of saturated compounds date from the year 1872, when Kolbe discovered nitro-methane² and Victor Meyer nitro-ethane. The modes of formation of

¹ *Ann. Chem.*, vol. xii. p. 305.

² *Journ. pr. Chem.* (2), vol. v. p. 427.

these substances were particularly calculated to arouse the reflection of chemists, since it was to have been expected here that compounds of quite other constitution—ethers of nitrous acid—would have been obtained instead. The thorough investigation and explanation of the chemical nature of nitro-ethane is due to V. Meyer. Those splendid researches¹ of his further resulted in the discovery (by himself) of other remarkable compounds, which include the nitrolic acids and nitrols.

These last have been proved to be representatives of the two classes of isonitroso- and nitroso-compounds, which have repeatedly, and more especially of late years, awakened the interest of chemists. It was those investigations of Victor Meyer and his pupils which established the constitution of the isonitroso-compounds, and showed how they were formed by the action of hydroxylamine upon substances containing the radical carbonyl. Thanks to this perfect² reaction, so universally applicable, many substances which were formerly numbered among the nitroso-compounds have since been recognised as really belonging to the class of their isomers. On the other hand, the above reaction has proved itself a convenient means of testing whether or not compounds contain the radical carbonyl.³ There have thus been drawn from those simple researches valuable conclusions with respect to the constitution of whole classes of compounds, *e.g.* of the quinones. The investigations of V. Meyer and Auwers⁴ upon benzil-oxime, and those of Beckmann⁵ upon benzaldoxime have thrown an unlooked-for light upon the constitution of certain of the oximes, as most of the isonitroso-compounds are now called. The chemical structure of the "hydroxamic acids" has been cleared up by Lossen.⁶

¹ *Ann. Chem.*, vol. clxxi. p. 1; vol. clxxv. p. 88; vol. clxxx. p. 111.

² German, *glatt*.

³ In phenyl-hydrazine E. Fischer discovered an analogous and equally serviceable reagent for carbonyl compounds, which has proved of the utmost value in establishing the constitution of a very large number of substances.

⁴ *Ber.*, vol. xxi. pp. 784, 3510; vol. xxii. pp. 537, 705, and 1985.

⁵ *Ibid.*, vol. xx. p. 2766; vol. xxii. pp. 429, 514, and 1531.

⁶ *Ann. Chem.*, vol. cclii. p. 174.

Development of the Knowledge of Sulphur Compounds.

The examination of organic sulphur compounds has proved of great importance for the development of our views upon the constitution of organic compounds generally, and more especially upon the saturation-capacity of the sulphur group of elements. Their investigation has led to the doing away once for all with the one-sided opinion that sulphur, selenium, and tellurium can only act as divalent, by furnishing proofs that they may also be tetra- or hexa-valent.

The earliest known of those compounds, which contain sulphur combined in the same manner as the alcohols, carb-oxylic acids, ethers, etc., contain oxygen, was mercaptan, discovered by Zeise; its true constitution, as a hydrosulphide corresponding to alcohol, was recognised by Liebig.¹ To this there were soon added ethyl sulphide and its polysulphides, whose analogy to the sulphides of the metals was obvious. The similarly constituted selenium and tellurium compounds were to a great extent worked out by Lowig² and Wöhler.³

Of organic acids which contain sulphur in place of oxygen, thiacetic acid,⁴ discovered by Kekulé, was the first known, although benzoyl sulphide had previous to this been regarded as the "thio-anhydride" of such an acid. Since then the number of those acids and their corresponding aldehydes has been greatly extended. Thio-glycollic acid (analogous to glycollic) and its analogues have been investigated mainly by Klason.⁵

By the action of powerful reagents on many of the compounds containing divalent sulphur, which have just been spoken of, it has been found possible to prepare others in which the sulphur present possesses a higher valency,—compounds which are comparable with sulphurous and

¹ *Ann. Chem.*, vol. xi. pp. 2, 11.

² *Pogg. Ann.*, vol. xxxvii. p. 552.

³ *Ann. Chem.*, vol. xxxv. p. 111; vol. lxxxiv. p. 69.

⁴ *Ibid.*, vol. xc. p. 311.

⁵ Cf. *Ibid.*, vol. clxxxvii. p. 113.

sulphuric acids, and which can be derived and in part prepared from the latter. The earliest known of these were the sulphonic acids and sulphones, whose first representatives—phenyl-sulphonic acid and diphenyl-sulphone (*Sulphobenzid*)—were obtained by Mitscherlich,¹ by acting upon benzene with sulphuric acid. These compounds, however, only came to be fully understood after Kolbe had shown them to be derivatives of sulphuric acid and its anhydride. Previous to this (in 1844) he had enlarged the then existing knowledge of the sulphonic acids by his work upon methyl-sulphonic acid and its chlorine derivatives. The important discovery² of the transformation of hydrosulphides, disulphides, and sulphocyanides into sulphonic acids furnished a general method for the preparation of the latter.

In a similar manner the conversion of the alkyl sulphides into sulphones, which contain two atoms of oxygen more in the molecule, was effected.³ Kolbe was again the first here to point out definitely the analogy between sulphones and ketones, and sulphonic and carboxylic acids. There have been added quite lately to the di-ketones the di-sulphones and the sulphone-ketones (products intermediate between the two), in whose investigation R. Otto⁴ has done more than any one else. The di- and tri-sulphonic acids, which correspond to the poly-carboxylic, have been known for a long time, Hofmann and Buckton⁵ having been the first to investigate them.

Von Oefele's discovery of the sulphines⁶ was especially pregnant in its results, because the existence of these compounds stood in contradiction to the assumption that the sulphur atom was invariably divalent. And the same applies to the investigation of the sulph-oxides by Saytzeff.⁷

¹ *Pogg. Ann.*, vol. xxix. p. 231 ; vol. xxxi. p. 628.

² Löwig, *Pogg. Ann.*, vol. xlvii. p. 153 ; Muspratt, *Ann. Chem.*, vol. lxxv. p. 251.

³ Von Oefele, *Ann. Chem.*, vol. cxxxii. p. 80.

⁴ *Journ. pr. Chem.* (2), vol. xxx. pp. 171, 321 ; vol. xxxvi. p. 401.

⁵ *Ann. Chem.*, vol. c. p. 133.

⁶ *Ibid.*, vol. cxxvii. p. 370 ; vol. cxxxii. p. 82.

⁷ *Ibid.*, vol. cxliv. p. 148.

and to that of the sulphinic acids, whose formation and chemical behaviour was cleared up by the work of Kalle, Otto, Klason, and others. Mention must also be made here of the remarkable conversion of sulphimates into sulphones¹ and of sulphites into sulphonic acids² by means of alkyl iodides, those reactions having led to conclusions respecting the constitution both of the sulphinic acids and the sulphites.

Organic Nitrogen Compounds.

An exceptionally wide field in organic chemistry was opened up by the discovery of the nitrogenous bases corresponding to ammonia. With the finding out of their connection with the latter, the question of their chemical constitution in general was solved. A. W. Hofmann's classical researches³ on the substituted ammonias and ammonium bases, which result from the action of alkyl iodides upon ammonia, deserve the first mention here, since they led to the true perception of the constitution of these bodies, and established a basis upon which they might be systematised. His splendid work upon aniline and its numerous derivatives, begun in 1843,⁴ and on the addition products of this base (*cf.* cyan-aniline) immensely enriched organic chemistry. These investigations resulted in the discovery of a wealth of new and striking facts, *cf.* the observation of the influence exerted by halogens entering the aniline molecule upon the chemical character of the resulting compounds.⁵ Upon the basis of those labours, which prepared the way for a knowledge of the aromatic bases, the aniline colour industry has developed itself in the most brilliant manner. From a theoretical point of view, also, these researches on the di- and tri amines and on the corresponding ammonium bases

¹ Otto, *Ber.*, vol. xiii. p. 1274.

² Strecker, *Ann. Chem.*, vol. cxlviii. p. 90.

³ *Ann. Chem.*, vol. lxxiv. p. 117; vol. lxxv. p. 356; *cf.* also pp. 279-280 of this book.

⁴ *Ibid.*, vol. xlvii. p. 37, and numerous later papers.

⁵ *Ibid.*, vol. liii. p. 1; *cf.* also p. 268 of this book.

(obtained from ethylene bromide and ammonia) were of especial importance; Hofmann, in fact, has worked out and explained organic nitrogen compounds generally as no other man has done. His investigations on the formation of substitution-products of ammonia contributed more than anything else to the establishment of the "typical" theory towards the end of the forties. The important work also effected by Hofmann on the mustard oils brought out clearly the relation existing between this class of compounds and the amines, and furnished a firm basis for arriving at their constitution.

The observation that the organic ammonias result from the nitro-compounds by reduction was a point of especial significance in their history, this having been first effected by Zinin¹ in the conversion of nitro- into amido-benzene. The above reaction has proved itself of the greatest use as a general method, and has since been applied with success in innumerable instances, besides having been extended to the later discovered nitro-compounds of the fatty series. The mode of formation of the primary amines from the cyanic ethers, discovered by Wurtz,² must also be referred to here as of historical importance, since the simplest organic ammonia, methylamine, was first prepared in this way.

From the vast number of observations on the chemical behaviour of the classes of compounds in question, we can but pick out a very few, such, namely, as have led to the clearing up of their constitution and to the discovery of new and important groups. To what an unlooked-for significance the action of nitrous acid upon amines and similar bodies (a reaction which had already been studied by Hofmann and others) attained in the hands of P. Griess, who demonstrated the conditions under which diazo-compounds were formed, and examined these with the greatest success! To the latter there were afterwards added the azo-compounds and hydrazines, classes which are of such importance as also

¹ *Journ. pr. Chem.*, vol. xxvii. p. 149.

² *Ann. Chim. Phys.* (3), vol. xxx. p. 443.

to merit a detailed description (see below). The transformation of aromatic amines into valuable dyes by oxidation, observed by W. H. Perkin sen., A. W. Hofmann, and others, marked the commencement of a new era in chemical industry.

Only a passing reference need be made here to the conversion of the organic ammonias into quinoline, acridine, quinoxaline, and other basic substances by similar processes of condensation, since these reactions will be considered later on, especially in their connection with the pyridine and quinoline bases, and the relations of the latter to the alkaloids.

Great advances have been made in the artificial production of naturally-occurring nitrogenous substances, by suitable transformations of ammonia or amines. As examples of these we may take the synthesis of trimethylamine, and that of neurine¹ from the latter substance; the production of sarcosine from methylamine and chloroacetic acid; the conversion of sarcosine into creatine by means of cyanamide;² the formation of uric acid derivatives from urea (syntheses of parabanic, oxaluric, and barbituric acids); and, lastly, the synthesis of uric acid itself.³

Hand in hand with this advancing knowledge of the amine bases there has gone that of the amides, which have already been spoken of. Here we can only refer to the important conversion of these into the cyanides, and their re-formation from the latter; and to the interesting behaviour of the substituted amides with phosphorus pentachloride—a reaction which has been examined particularly by Wallach,⁴ and which has led to a knowledge of certain peculiar bases, the oxalines. The corresponding thiamides, which have been investigated by Cahours, Hofmann, and many others, have on their part been converted into other

¹ Wurtz, *Ann. Chem.*, Suppl., vol. vi. pp. 116 and 197.

² Volhard, *Ann. Chem.*, vol. cxxiii. p. 261; *Jahresber. d. Chemie* for 1868, p. 685.

³ Behrend and Roosen, *Ann. Chem.*, vol. celi. p. 235.

⁴ *Ann. Chem.*, vol. clxxxiv. p. 1; vol. ccxiv. p. 193.

nitrogenous compounds, *e.g.* the amidines,¹ whose investigation has likewise yielded many useful results.

Through the discovery and investigation of the organic compounds of phosphorus, antimony, and arsenic, the connection existing between those three elements and also their relation to nitrogen were proved in the clearest manner, so that here, as well as in other cases, the study of organic compounds has thrown a brilliant light upon particular branches of inorganic chemistry. The phosphines and phosphonium bases first became known through the classical and comprehensive researches of A. W. Hofmann,² and the corresponding compounds of the aromatic series through those of Michaelis.³ The organic compounds of phosphorus were thenceforth recognised as derivatives of the well-known inorganic ones,—phosphuretted hydrogen (PH_3) and phosphonium iodide, and phosphorus tri- and penta-chlorides. The study of the organic compounds of arsenic and antimony, the former of which were admirably worked out by Bunsen, and at a later date by Cahours, Baeyer, and Michaelis,⁴ and the latter by Löwig, Landolt, Michaelis,⁵ and others, likewise led to the conclusion that these substances were derivable from the inorganic compounds of the elements. The influence exercised by some of those researches upon the development of the doctrine of valency has already been sufficiently referred to in the general section.

The field comprising the organic compounds of nitrogen is by no means exhausted with the description of those classes which have been shortly mentioned above. A number of others must be referred to here, with regard to the chemical constitution of which much has also been accomplished.

¹ Wallach, *Ann. Chem.*, vol. clxxxiv. pp. 5 and 91; Bernthsen, *ibid.*, vol. clxxxiv. p. 321; vol. cxcii. p. 1. Pinner, in especial, has extended this branch by his beautiful researches on imido-ethers and amidines (cf. *Ber.*, vol. xvi. p. 1654; vol. xvii. p. 2520; vol. xviii. p. 759).

² *Ber.*, vol. iv. p. 605; vol. v. p. 104; vol. vi. p. 306.

³ Cf. *Ann. Chem.*, vol. clxxxviii. p. 275.

⁴ For the literature on the subject, cf. *Ann. Chem.*, vol. cci. p. 184.

⁵ Cf. *ibid.*, vol. ccxxxiii. p. 39.

Of the azo-compounds, azo-benzene was the first to be discovered (by Mitscherlich),¹ while much later there came azoxy-benzene by Zinin² and hydrazo-benzene by A. W. Hofmann.³ The now universally accepted views held with regard to these three kinds of azo-compounds are due to Erlenmeyer,⁴ and still more to Kekulé,⁵ who assumed in azo-benzene two doubly-linked nitrogen atoms, and in oxyazo- and hydrazo-benzene two singly-linked ones. The ready production of these and similar substances from diazo-compounds has tended greatly to advance our knowledge of them. The investigations of Griess, Kekulé, V. Meyer, Witt, and others, which showed how diazo- could be converted into azo-compounds, have led to the establishment of a flourishing industry—the manufacture of azo-dyes. The doctrine of isomerism has also been enriched by a wealth of observations arising out of these labours.

The diazo-compounds, so remarkable for their reactions, were discovered by Griess and investigated by him in a long series of admirable researches, which disclosed their most important characteristics. Griess showed how they were formed by the action of nitrous acid on aromatic amido-compounds,—a reaction which had previously been studied under other conditions, and had not led then to the discovery of those bodies. In a number of papers,⁶ dating from the year 1859, which followed one another with great rapidity, the above-named investigator made the chemical world acquainted with the diazo-derivatives of phenol, aniline, and benzoic acid. The view accepted by most chemists with respect to the constitution of these bodies, according to which two atoms of nitrogen are linked together as in the azo-compounds, originated with Kekulé.⁷ Another view, in which one of the nitrogen atoms is assumed to be pentavalent and

¹ *Pogg. Ann.*, vol. xxxii. p. 324.

² *Ann. Chem.*, vol. lxxxv. p. 328.

³ *Jahresber. d. Chemie* for 1863, p. 424.

⁴ *Ztschr. Chem.* for 1863, p. 678.

⁵ *Lehrb. d. Chem.*, vol. ii. p. 703.

⁶ *Ann. Chem.*, vol. cxiii. p. 201; vol. cxvii. p. 1; vol. cxxi. p. 257; vol. cxxvii. p. 39.

⁷ *Ztschr. Chem.* for 1866, p. 689.

the other trivalent, has been expressed by Blomstrand,¹ who has brought forward cogent arguments in its favour.

The existence of diazo-compounds in the fatty series has only been proved comparatively recently by the excellent researches of Curtius² on diazo-acetic and diazo-succinic ethers.

Another class of bodies, the hydrazines, which stand in a near relation to the diazo-compounds, was discovered in 1875 by E. Fischer³ and carefully investigated by him.⁴ Hydrazine itself has only quite lately become known (cf. p. 390), but phenyl-hydrazine has proved of the greatest value both as a specific reagent and as an aid in the synthesis of complicated compounds. Its relation to diazo-compounds was definitely proved by Fischer, through its formation from diazo-amido-benzene and conversion into diazo-benzene-imide.

Since Scheele's discovery of hydrocyanic acid, the cyanogen compounds have been the subject of frequent investigation by the most able chemists, so that the knowledge of them has been immensely increased. The development of this branch of organic chemistry is in a great degree due to the marked property possessed by those compounds of changing into isomers, and also of combining with other substances to yield new compounds.

The composition of prussic acid and of many of the cyanides was worked out by Berthollet and Ittner, and, especially, by Gay-Lussac in his classical researches, in which he discovered cyanogen and recognised its analogy to the halogens. He it was, too, who assumed in yellow prussiate of potash (a substance already known for a long time) the presence of the radical *ferrocyanogen*, while Berzelius explained it as being a double salt of iron protocyanide and cyanide of potash. The discovery of potassium ferricyanide by L. Gmelin in 1822, and that of

¹ In his *Chemie der Jetztzeit*, p. 272; cf. also *Ber.*, vol. viii. p. 51; and Strecker, *ibid.*, vol. v. p. 786.

² *Ber.*, vol. xvi. p. 2230; vol. xvii. p. 953; vol. xviii. pp. 1283 and 2371.

³ *Ibid.*, vol. viii. p. 589.

⁴ *Ann. Chem.*, vol. cxc. p. 67; vol. cxcix. p. 281; vol. ccxii. p. 316.

the so-called nitro-prussides by Playfair¹ extended the knowledge of cyanogen compounds of complex composition, in which, at Graham's suggestion, the radical tri-cyanogen was assumed.

Sulphocyanic acid, together with its salts, was discovered by Porret, and subsequently investigated by Berzelius, who established its composition: Liebig succeeded in isolating cyanogen sulphide in 1829, and he also showed what remarkable products were obtained from the decomposition of ammonium sulphocyanide.² Of recent years Reynolds, Volhard, Delitzsch and, more especially, Klason,³ among others, have advanced our knowledge of this class of compounds.

Cyanic acid, whose chemical behaviour and relation to its own isomers gave rise to important discussions respecting the constitution of all of them, was first isolated by Wohler,⁴ who was led during the investigation of its salts to his memorable discovery of the artificial formation of urea. Cyanuric acid, obtained by Serullas from the solid cyanogen chloride which he discovered, was recognised by Liebig and Wohler as being of the same percentage composition as cyanic acid. The influence which this observation, taken in conjunction with that of the isomerism of both of these compounds with fulminic acid, had upon the doctrine of isomeric substances, has already been discussed in the general section of this book.⁵ The haloid compounds of cyanogen have been known for a long time, cyanogen chloride having been obtained by Berthollet, and the iodide by Davy; but cyanamide, which was destined to acquire so great a significance for the synthesis of organic compounds,⁷ was first prepared in 1851 by Cloëz and Cannizzaro.⁸

¹ *Phil. Trans.* for 1849, vol. ii. p. 477.

² *Ann. Chem.*, vol. x. p. 11.

³ Cf. more particularly *Journ. pr. Chem.* (2), vol. xxxvi. p. 57; vol. xxxviii. p. 366.

⁴ *Pogg. Ann.*, vol. xv. p. 619; vol. xx. p. 369.

⁵ Cf. p. 338.

⁶ Cf. p. 236.

⁷ Cf. Volhard's, Strecker's, and Drechsel's researches, more especially *Journ. pr. Chem.* (2), vol. xi. p. 284.

⁸ *Comptes Rendus*, vol. xxxi. p. 62.

Owing to the readiness with which they unite with other substances, the cyanogen compounds as a whole have been of great service for the opening up of new branches of the science, and for advancing our knowledge of these; take, for example, the formation of guanidine and its derivatives from cyanamide or cyanogen chloride and ammonia, and also the formation of derivatives of the last-named compound.¹ The tendency shown by hydrocyanic acid to combine with aldehydes and ketones has already been mentioned.

The compounds of cyanogen as well as of thiocyanogen with organic radicals have, thanks to their diversity and capability of transformation, yielded an almost inexhaustible material for investigation. The alkyl cyanides or nitriles, with methyl cyanide at their head, were first prepared by Dumas² from the ammonium salts of the fatty acids, by acting upon these with phosphoric anhydride. The exceptionally important connection which exists between those nitriles and the fatty acids was demonstrated by Frankland and Kolbe³ when they converted the former into the latter by treatment with caustic potash. Another passing reference may be made here to the generalisation of this reaction, and the consequent production of an immense number of carboxylic acids and their derivatives from simpler compounds, even although it was spoken of when those compounds themselves were being described. The investigation of mandelic acid,⁴ resulting from oil of bitter almonds and hydrocyanic acid in presence of hydrochloric, gave the first impetus to the study of the compounds obtained under similar conditions from other aldehydes and ketones. The simplest nitrile of the aromatic series, phenyl cyanide, was first observed by Fehling.⁵

The isocyanides, isonitriles, or carbamines, which are isomeric with the nitriles, were discovered simultaneously

¹ Cf. Erlenmeyer, *Ann. Chem.*, vol. cxlvi. p. 253; A. W. Hofmann, *ibid.*, vol. lxvii. p. 129, etc.

² *Comptes Rendus*, vol. xxv. pp. 383 and 442.

³ *Ann. Chem.*, vol. lxx. p. 269.

⁴ Liebig, *Ann. Chem.*, vol. xviii. p. 319.

⁵ *Ann. Chem.*, vol. xlix. p. 91.

by A. W. Hofmann¹ and Gautier,² by different procedures, their existence having previously been foreseen by Kolbe. The perception of the cause of the isomerism existing between these two classes of compounds marked an important advance in theoretical chemistry. The conclusive explanation of the similar isomerism between the alkyl thiocyanates and the mustard oils, of which mustard oil proper (allyl iso-thiocyanate) was the earliest known, is due to Hofmann: the latter succeeded both in preparing the iso-thiocyanates artificially, and in proving at the same time their chemical constitution from their various decompositions.³ Hand in hand with the acquirement of this knowledge went the gradual establishment of the views upon the analogously constituted cyanic and isocyanic ethers: and here again Hofmann acted as the pioneer with his researches, after the simplest compounds of this nature had been obtained by Wurtz and Cloëz. The ease with which isocyanic ether and the corresponding mustard oils assimilate the elements of ammonia and the amines led to the discovery of the extensive class of the substituted ureas;⁴ the simplicity of the reaction, upon which the formation of these substances was based, allowed of the explanation of the numerous cases of isomerism which occur here.

The question of the chemical constitution of the polymeric cyanogen compounds presented far greater difficulties, the number of these having increased to an extraordinary extent since it was proved that cyanuric, fulminic and cyanic acids had all the same percentage composition. It is only quite recently (since 1884) that a certain amount of clearness has been arrived at with regard to the constitution of the cyanuric and isocyanuric compounds, and this has been due more particularly to the admirable investigations of A. W. Hofmann and of Klason, and also to those of Rathke, Weddige, and others. These researches have proved that isocyanuric acid

¹ *Ann. Chem.*, vol. cxliv. p. 144; vol. cxlv. p. 107.

² *Comptes Rendus*, vol. lxv. pp. 468 and 862.

³ *Ber.*, vol. i. pp. 26 and 169; vol. ii. pp. 116 and 452.

⁴ Cf. Wurtz, *Ann. Chem.*, vol. lxxx. p. 346; A. W. Hofmann, *ibid.*, vol. xxxiii. p. 57.

and isomelamine are not in themselves capable of existence, although derivatives of both are. The doctrine of stable and unstable modifications, already referred to,¹ developed itself mainly from observations made upon those polymeric compounds. The obscurity surrounding the compounds of this nature, as well as those decomposition-products of ammonium sulphocyanide known under the names of mellone and melam, the bases resulting from the nitriles by polymerisation (cyan-ethine, etc.), is now beginning to vanish, and a knowledge of their constitution is gradually being acquired. Thus E. v. Meyer has recently established the constitution of the cyan-alkines, his investigations² having shown them to be amido-miazines; and the discovery of the dimolecular nitriles³ also deserves mention here. The rational composition of fulminic acid and allied compounds, *e.g.* fulminuric acid and other isomers, is now becoming much better understood, thanks to the pioneering researches of Liebig,⁴ and the investigations of Kekulé,⁵ Schischkoff,⁶ and, more recently, of Steiner, Carstanjen, Ehrenberg,⁷ and others, although it has not yet been made absolutely clear.

*Historical Notes on Pyridine and Quinoline.*⁸

An extensive group of nitrogen compounds—the pyridine and quinoline bases—has only been worked at with success of quite recent years, although these substances were in part discovered during the earliest decades of the century; their investigation has been carried on with the utmost zeal ever since it came to be recognised that the vegetable alkaloids were among their derivatives. The

¹ Cf. p. 334.

² *Journ. pr. Chem.* (2), vol. xxxix. p. 262.

³ *Ibid.*, vol. xxxviii. p. 336; vol. xxxix. p. 188.

⁴ *Ann. Chem.*, vol. xxvi. p. 146.

⁵ *Ibid.*, vol. cv. p. 279.

⁶ *Ibid.*, vol. ci. p. 213.

⁷ *Journ. pr. Chem.* (2), vol. xxv. p. 232; vol. xxx. p. 38.

⁸ With regard to the sources of the following notes, cf. the pamphlets of Metzger and of Heseckiel on these bases, and Calm-Buehka's work, *Die Chemie des Pyridins und seiner Derivate*.

researches of Anderson¹ on the volatile bases of bone oil, those of Williams² on the similar bodies contained in coal tar, and Gerhardt's observation on the production of quinoline from quinine³ were the first beginnings in the cultivation of this field, which has since been worked with such wonderful success. The investigation of these substances received an especial impetus with the first recognition of the similarity between the pyridine bases and quinoline, and of the distinct analogy between these substances and the aromatic compounds. The earliest attempt to explain the constitution of pyridine and quinoline was due to Körner,⁴ and it bore the richest fruit; he assumed these bodies to be benzene and naphthalene respectively, in which a methine group (CH)^{'''} was replaced by the trivalent nitrogen atom. This hypothesis was applied to the facts already known, to which a large number of new ones were being continually added, with the result that they were without difficulty made to accord with it. The theory of the aromatic compounds, which had by this time become strongly developed, gave those endeavours a more or less secure basis to go upon, more especially when it came to criticising and sifting the rapidly augmenting number of isomers among the pyridine derivatives.

The connection of pyridine and quinoline with benzene and naphthalene, assumed in the above hypothesis, was clearly proved by a succession of beautiful researches. We may refer here to the analogous behaviour with regard to oxidising agents shown by the alkylated pyridines and the alkyl derivatives of benzene. The investigation of these relations, more especially those of the isomeric methyl- and ethyl-pyridines and the pyridine mono-carboxylic acids, we owe to the admirable work of Weidel, Skraup, Ladenburg, and Wischnegradsky. Just as the admissibility of the hypothesis

¹ *Phil. Trans. E.*, vol. xvi. p. 4, and vol. xx. (2), p. 247; *Phil. Mag.* (4), vol. ii. p. 257; *Ann. Chem.*, vols. lx., lxx., lxxv., lxxx., lxxxiv.

² *Phil. Mag.* (4), vol. viii. p. 24; *Phil. Trans. E.*, vol. xxi. (2), p. 315, etc.

³ *Ann. Chem.*, vol. xlii. p. 310.

⁴ Cf. p. 330.

respecting the constitution of benzene was arrived at from the number of its substitution-products which could actually be prepared, so in like manner a similar deduction was drawn for pyridine, viz. that only the theoretically possible methyl-pyridines and pyridine-carboxylic acids were capable of preparation, and no more.

Among the experimental researches which have furnished further support for the above view must be mentioned those of Königs, Ladenburg, and A. W. Hofmann, which distinctly proved the connection between pyridine and piperidine (the latter containing six atoms of hydrogen more in the molecule than the former). The analogy between this compound and pyridine on the one hand, and hexahydro-benzene and benzene on the other, thus became at once apparent.

The different modes of formation of pyridine bases from substances of simpler composition likewise assisted towards a knowledge of their constitution. We may refer here to the synthesis of one of the collidines from aldehyde-ammonia, as well as from ethylidene chloride and ammonia; to that of a chloro-pyridine from pyrrol-potassium and chloroform; and to the researches of Hantzsch, which resulted in the artificial production of lutidine.

The synthetic investigations of quinoline and its derivatives have proved themselves extraordinarily fruitful; they served more particularly to confirm the constitution ascribed to those bodies, this being also deducible from the products of decomposition of the latter. Out of the great amount of work done in this branch, only one or two researches can be mentioned here, viz. those of Skraup, who (doubtless stimulated by the previous investigations of Königs and Graebe) discovered the general method of preparing quinoline and its derivatives, by the action of glycerine on the aromatic amines; Baeyer's beautiful investigations on the formation of quinoline, oxy-quinoline, etc., by the condensation of *o*-amido-phenyl-compounds; the synthesis of quinoline and its homologues from a mixture of *o*-amido-benzaldehyde and other aldehydes by Friedländer; and that from aniline and aldehyde by v. Miller and Döbner.

While the above syntheses have made clear the constitution of quinoline, other investigations have established its connection with pyridine; thus it was seen that the product obtained by oxidising quinoline was a pyridine-dicarboxylic acid, the formation of which was in every respect analogous to that of benzene-dicarboxylic acid from naphthalene.

The minute study of the quinoline derivatives, and the increasing certainty with which their constitution has come to be criticised, have led to the discovery of numerous other similar bodies, which are derivable from quinoline itself. Among these are iso-quinoline, quinoxaline, and cinnoline, with their derivatives, all of which may, like quinoline itself, be looked upon as derivatives of naphthalene.

A still greater interest than that aroused by the discovery of the compounds just named was awakened by the proof (gradually arrived at from a long series of admirable researches) of the intimate connection existing between pyridine and quinoline and various vegetable alkaloids, whose constitution was thereby explained. Wischnegradsky and then Königs were the first to express the opinion that the alkaloids were derivatives of pyridine or quinoline. They grounded this view upon the conversion of pyridine into piperidine, which is a decomposition-product of the alkaloid piperin contained in pepper, and on the retransformation of piperidine into pyridine; to this was added later on the precisely analogous conversion of conine into conyine, a propyl-pyridine.¹ Quickly following the recognition of this last important fact came the further one² that this alkaloid of hemlock was the dextro-rotatory modification of α -propyl-piperidine.³

¹ A. W. Hofmann, *Ber.*, vol. xvii. p. 825.

² Cf. Ladenburg, *Ann. Chem.*, vol. cexlvii. p. 80 (1888).

³ The memorable synthesis of conine was effected by the aid of α -allyl-pyridine; the reducing mixture (sodium and alcohol), which served for the conversion of this pyridine derivative into conine, has since been found useful in other similar cases. Ladenburg in this way transformed trimethylene cyanide into piperidine, which was thus completely synthesised; at the same time he discovered, as the final product of the reaction, pentamethylene-diamine, which proved to be identical with one of the ptomanes, —cadaverine.

The investigation of the decomposition-products of many of the organic bases has likewise resulted in proving that the latter are most intimately related to pyridine or quinoline. Thus Ladenburg has obtained dibromo-pyridine from atropine, Weidel pyridine-tricarboxylic acid from berberine, and v. Gerichten a pyridine-dicarboxylic acid from narcotine.

The chemistry of the alkaloids has been greatly advanced within the last few years by much admirable work; great light has, for instance, been thrown upon the constitution of atropine by Ladenburg, of narcotine by Roser, and of papaverine, hydrastine, and cocaine by Goldschmiedt, Freund, Merck, Einhorn, and others. In every case those vegetable alkaloids have been found to be related to pyridine, quinoline, or isoquinoline.

The above very short summary of but a few of the many investigations which have been carried out in this branch is of itself sufficient to show how necessary is a knowledge of the chemical nature and constitution of the pyridine and quinoline bases for the proper understanding of the alkaloids, and what a rich harvest may still be expected here.

Certain non-nitrogenous compounds also, viz. meconic, comenic, pyromeconic, and chelidonic acids, whose constitution remained quite obscure although the substances themselves had long been known, have been shown, more particularly by the recent researches¹ of Ost and of Lieben and Haitinger, to be naturally connected with pyridine. A light has been thrown upon their constitution, as also upon that of the similarly constituted compounds obtained from citric and malic acids,² by the important observation that they are converted by ammonia into oxypyridine-carboxylic acid.

Pyrrrol and Analogous Compounds.

Another group of compounds, of which pyrrrol, furfurane, and thiophene are the representatives, has been the subject

¹ *Journ. pr. Chem.* (2), vol. xxvii. p. 257; vol. xxix. p. 57; *Ber.*, vol. xvi. p. 1259.

² A. W. Hofmann, *Ber.*, vol. xvii. p. 2687; v. Pechmann, *ibid.*, vol. xvii. p. 936; vol. xix. p. 2694.

of the most ardent investigation during recent years, with the result that the constitution of these substances and also that of many of their derivatives has been cleared up. The analogy which exists between those compounds gradually came to be recognised; they all contain the same nucleus, consisting of four atoms of carbon and four of hydrogen, this being combined in pyrrol with the imido-group (NH), in furfuran with one atom of oxygen, and in thiophene with one atom of sulphur. Their similarity to benzene became more apparent the better they came to be known, and was shown in a particularly striking manner in the investigation of thiophene (discovered by Victor Meyer) and its derivatives. The work¹ which has been done upon this class of bodies is amongst the most brilliant of our time.

The artificial formation of thiophene from succinic acid and phosphorus trisulphide,² that of pyrrol from succinimide by means of zinc dust, and the conversion of pyrrol into the compounds richer in hydrogen—pyrroline and pyrrolidine (Ciamician)—are reactions of special importance, which greatly aided in elucidating the constitution of these bodies. Pyrrol, which was discovered by Runge in coal tar and first isolated by Anderson, has with its rapidly-augmenting host of derivatives been closely and comprehensively examined by Ciamician, Demstedt, Paal, and others of late years. Schwanert³ a long time ago having made the fundamental observation that pyrrol could be produced from ammonium mucate.

The work done upon furfuran (which was discovered by Limpricht⁴) is to be taken in conjunction with that upon pyromucic acid (first observed by Scheele, and recognised as a distinct compound by Labillardière) and its aldehyde furfural (discovered by Dobereiner and examined by Stenhouse, Fownes, and others). The analogy in behaviour of the latter to benzoic aldehyde was proved more especially by

¹ Cf. p. 331.

² *Ber.*, vol. xviii. p. 454.

³ *Ann. Chem.*, vol. cxvi. p. 273.

⁴ *Ibid.*, vol. clxv. p. 281.

Baeyer and E. Fischer,¹ and the close connection between pyromucic and maleic acids by Hill.²

Among the aromatic compounds proper, to which the substances just named show a great similarity, indole (discovered by Baeyer) was recognised by him as being an analogue of pyrrol, and was made the basis of important researches which resulted in showing its relation to the compounds of the indigo group, more particularly to isatin, oxindole, and dioxindole. Various derivatives of indole have lately been prepared by a method discovered by E. Fischer,—*i.e.* from the condensation of phenyl-hydrazine with aldehydes and ketones.³ Cumarone, obtained by Fittig and Ebert from cumarine, has been designated by Hantzsch⁴ the “furfurane of the naphthalene series,” and he has confirmed this view by some ingenious syntheses of its derivatives. The analogy existing between the three compounds furfurane, thiophene and pyrrol, and diphenylene oxide, sulphide and imide (carbazole) respectively, was perceived about the year 1885.

Organo-metallic Compounds.

After it had come to be seen that not only hydrogen, oxygen, nitrogen, sulphur, and the halogens could combine directly with carbon, but also arsenic as well—a point which Kolbe was the first to indicate in his interpretation of cacodyl,⁵—new fields in organic chemistry became opened up in rapid succession. Frankland's discovery⁶ of the action of zinc on methyl and ethyl iodides, in which the metal breaks up the iodide in order to combine with the alkyl, led to a knowledge of the organo-metallic compounds. Thanks to the readiness with which these enter into reaction, they have been destined to aid in the development of organic chemistry to an unlooked-for extent, more especially as regards syn-

¹ *Ber.*, vol. x. p. 13.

² *Ibid.*, vol. xiii. p. 734; *Journ. Chem. Soc.*, vol. xl. p. 36.

³ *Ann. Chem.*, vol. cexxxvi. p. 116.

⁴ *Ber.*, vol. xix. p. 1290.

⁵ *Cf.* p. 296.

⁶ *Journ. Chem. Soc.*, vol. ii. p. 263; or *Ann. Chem.*, vol. lxxi. p. 171 (1849).

thetic methods. With the aid of the zinc-alkyls many other organo-metallic compounds were prepared and minutely investigated in due course, *e.g.* the ethyl compounds of tin, mercury, lead, sodium, aluminium, and other elements.¹ Among the last were those non-metals of which organic compounds had not previously been known; boric methide and other similar substances were prepared by Frankland,² and the important alkyl compounds of silicon by Friedel and Crafts, the composition of these latter proving the complete analogy between that element and carbon. To the organo-metallic compounds of the fatty series, various others belonging to the aromatic have since been added, the first of these having been mercury di-phenyl.³

The short description which has just been given of the development of organic chemistry is sufficient, notwithstanding its incompleteness, to allow of our recognising the main currents which have prevailed, and which still do so, in this branch of the science. The review of the numberless organic substances, which have been investigated during the last fifty or sixty years, is materially facilitated by the general points of view which have become gradually established from the classification of those compounds, and from the deduction of their chemical constitution. A prominent place in this respect is to be given to the gradually growing perception that organic compounds might be looked upon as derivatives of inorganic, and to the increasing certainty with which their constitution could be defined on the basis of the saturation-capacities peculiar to the atoms of the various elements.

¹ Cf. the papers of Buckton, Odling, Frankland, Cahours, Ladenburg, etc., in the *Philosophical Transactions*, *Journal of the Chemical Society*, and *Annalen der Chemie*.

² *Proc. R. S.*, vol. xii. p. 123; or *Ann. Chem.*, vol. cxxiv. p. 129.

³ R. Otto, *Ann. Chem.*, vol. cliv. p. 93.

HISTORY OF PHYSICAL CHEMISTRY IN RECENT TIMES.¹

The influence which certain branches of physics have exercised on the development of chemical doctrines cannot be estimated too highly. Through the introduction of physical methods, more particularly through the application of weighing, measuring, and calculating to chemical problems, chemistry first became an exact science. The importance of those methods, in so far as they had a determining influence on the chemical tendency of the present period, has already been gone into in the general section of this book. From the time of Lavoisier on, it came to be more and more clearly seen that an intimate connection exists between the chemical and physical properties of substances. Definite relations were found to hold good both between the proportions by weight of substances which enter into chemical combination and between the volumes of combining gases (Avogadro, Gay-Lussac). Investigators sought to determine the more important physical constants of compounds in their various states of aggregation, *e.g.* the specific gravity, specific heat, etc., as well as the changes in physical properties which were brought about by chemical reactions, and thus to arrive at general relations from which the chemical constitution and physical behaviour of different substances could be worked out. To the efforts at solving such problems as these, physical chemistry owes its origin and gradual development.

Although Lavoisier, in conjunction with eminent physicists (Laplace, in particular), entered into some of the above problems, and Gay-Lussac at a later period established the relations which exist between the volumes of different gases and their chemical composition, while Dulong and

¹ With regard to the sources of information on which this and the following section are based, the reader is referred to W. Ostwald's admirable work, *Lehrbuch der allgemeinen Chemie* (1885-87), ("Text-Book of General Chemistry").

Petit pointed out the connection between the specific heat and atomic weight of the elements, the boundary land between physics and chemistry was first systematically worked at by Hermann Kopp; with the investigations of the last-named chemist on the relations between atomic weight and specific gravity, on the laws which regulate the boiling temperatures of liquids, and so on, the history of physical chemistry is intimately bound up. The attention paid to physico-chemical questions has gone on steadily increasing during the last three decades, and this applies in a special degree to such as bear upon the relations between the thermo-chemical and optical behaviour of substances and their chemical constitution.

But there is another branch also, viz. that of chemical affinity (*Verwandtschaft*), which has been greatly enriched by the investigations just referred to. With the aid of physico-chemical methods, and the calculations requisite for these, a beginning is being made towards the solution of the old problem respecting the cause and nature of chemical affinity. It will therefore be appropriate to speak of the history of the doctrine of affinity while describing the development of physico-chemical researches. Through both of these branches there runs the continuous endeavour to make chemical reactions capable of mathematical treatment.

The behaviour of gases and vapours has been, more than anything else, the subject of fruitful physico-chemical investigations, doubtless because the physical properties of a substance in the gaseous state are observable with fewer complications than in any other, and hence definite relations between these and the chemical constitution of the compound are most readily apparent.

Determination of Vapour Density.

The laws of Boyle and Mariotte and of Gay-Lussac, which expressed the connection between the volume of a gas and its temperature and pressure, prepared the ground for a

knowledge of other relations. Gay-Lussac's law of volumes, which has already been treated of,¹ was the first result in this branch which benefited chemistry in an exceptional degree. The recognition of the intimate connection between the specific gravity of a gas and its molecular weight we owe to Avogadro,² although it was a long time of taking root in the science; this "law of Avogadro," which expresses the above relation, still governs chemical research, and is an indispensable aid in the determination of the molecular weights of chemical compounds.

The due appreciation of its value has led to continuous endeavours towards simplifying and refining the methods for determining the specific gravity of gases and vapours. Dumas, as already mentioned, was the first to devise a generally applicable method for vapour density determinations,³ and by this he achieved great results. Another plan, according to which the volume of vapour produced from a given weight of substance is accurately estimated, was worked out by Gay-Lussac and afterwards modified by Hofmann.⁴ And to the above methods there was added in 1878 that of Victor Meyer,⁵ which depends upon the measurement of the air (or any other indifferent gas) which is expelled from the apparatus by the vapour resulting from a given weight of the substance in question. The improvements which those methods have undergone since their introduction cannot be entered into here, but emphasis must be laid upon the point that through their means the all-important knowledge of the relative weights of the atoms and molecules of elements and compounds has been immensely advanced.

The determination of the specific gravity of vapours has proved in certain cases the most reliable means of deciding between the values arrived at by different methods, stöchiometric or otherwise, and so getting at the correct atomic weights. To give only some more or less recent

¹ Cf. p. 202.

² Cf. pp. 202 and 277.

³ *Ann. Chim. Phys.*, vol. xxxiii. p. 341.

⁴ *Ber.*, vol. i. p. 198.

⁵ *Ibid.*, vol. xi. pp. 1867 and 2253.

instances of this, we would refer to the deduction of the atomic weights of silicon, beryllium, thorium, and germanium from the vapour densities of their chlorides. Starting with Avogadro's hypothesis—that the vapour density is proportional to the molecular weight—chemists have been able to deduce from the specific gravities of gasified elements most striking conclusions with respect to the number of atoms in their molecules at different temperatures. One has but to think of the results of Dumas' and Mitscherlich's investigations¹ on the vapour densities of sulphur, arsenic, phosphorus, and mercury, the molecules of which contain different numbers of atoms, as was deduced at a later date from the specific gravity of their vapours after the revivification of Avogadro's law. The reader is further referred to the important work of V. Meyer and of Nilson and Pettersson on the vapour densities of compounds, more especially of such as show a varying composition with changing temperature. Aluminium chloride, for instance, has the simplest molecular weight which is possible (that expressed by the formula AlCl_3) at a temperature sufficiently high, but one double as great (Al_2Cl_6) at lower temperatures; and the same applies to stannous chloride (SnCl_2 or Sn_2Cl_4), etc.

These few examples are sufficient to illustrate what has just been said above. The significance which is attached to the results of vapour density determinations is most strikingly shown in the fact that such estimations are held to be the most reliable means of getting at the valency of an element. The amount of care, however, which is requisite here, is proved by the different results obtained by different experimenters, and is particularly well shown by the behaviour of aluminic chloride, from whose vapour density the conclusion was drawn (and held to until quite recently) that aluminium was tetravalent, although the whole behaviour of the element pointed to its tri-valency; this has now been confirmed by the determination of the normal density of vapour of the chloride.

¹ Cf. p. 212.

Dissociation.

From the observations made upon what are known as anomalous vapour densities, the cause of which has been recognised in a gradually increasing decomposition of the compound with rise of temperature, the doctrine of dissociation—so important for physical chemistry—has developed itself; the name “dissociation” was first made use of by H. de St. Claire Deville to express decompositions of this nature. He was the earliest (from the year 1857)¹ to work systematically at this branch of the science, which has also been made the subject of important investigations by others since, *e.g.* Debray, Cahours, Wurtz, Horstmann, Isambert, and A. Naumann. Most of these experimenters did not confine themselves to cases of abnormal vapour density alone, but studied generally the gradual increase in decomposition of chemical compounds under an increasing temperature.

The Liquefaction of Gases.

The investigation of the transition of gases and vapours into the liquid and solid states has given rise to work of exceeding importance. We have but to recall the comprehensive researches of Faraday² on the liquefaction of gases which were at that date held to be uncondensable, and especially the late investigations of R. Pictet,³ Cailletet,⁴ Wroblevsky and Olzevsky,⁵ which have proved that there is no gas that can withstand the combined effect of sufficiently high pressure and low temperature. Nitrogen, hydrogen, and oxygen have thus all been reduced to the liquid and solid states, and their boiling temperatures determined,—observations of very great moment.

Previous to the date of these last investigations

¹ Cf. *Comptes Rendus*, vol. xlv. p. 857.

² *Phil. Trans.* for 1823, p. 160; and for 1845, p. 1.

³ *Comptes Rendus*, vol. lxxxv. p. 1214; also in subsequent volumes of the *Archives des Sciences Naturelles*.

⁴ *Comptes Rendus*, vol. lxxxv. p. 1213 (1877).

⁵ *Ann. Phys.*, N. F., vol. xx. p. 243, etc.

Andrews¹ had made a thorough study of the conditions under which a gas can be liquefied, and had established the important conceptions of "critical temperature" and "critical pressure," Mendelejeff² having a few years before this made some valuable observations on the subject.

Light was thrown upon the behaviour of gases to liquids in the first decade of the century by the investigations of Henry and Dalton, which established the fact that the amount of absorption of a gas or of a mixture of gases by a liquid is dependent upon the pressure, and this law was afterwards confirmed by Bunsen's classical researches.³

The Kinetic Theory of Gases.

The thorough investigation of gases, of their physical behaviour in particular, led to the setting up of a theory by means of which the various phenomena exhibited by them—specific heat, diffusion, and friction—have been brought together under one common standpoint and explained in a satisfactory manner. The fundamental idea that a gas was an assemblage of moving particles had previously been put forward by D. Bernoulli and by Herepath, and Joule had in 1851 made a great step in advance by calculating the mean translational velocity of these particles. This idea, in the hands of Krönig and Clausius, gave birth to the modern kinetic theory of gases, which has been so splendidly worked out by Clausius and Maxwell, and since then perfected in detail by Boltzmann, O. E. Meyer, van der Waals, and many others. It may be regarded as springing from the mechanical theory of heat.⁴

Spectrum Analysis.

The examination of the optical behaviour of glowing gases and vapours has exercised a most profound influence

¹ *Phil. Trans.* for 1869, p. 575; or *Pogg. Ann.*, Suppl., vol. v. p. 64 (1871).

² *Ann. Chem.*, vol. cxix. p. 11.

³ *Ibid.*, vol. xciii. p. 1 (1855).

⁴ For an account of the development of the above theory, see O. E. Meyer's work, *Die Kinetische Theorie der Gase* (Breslau, 1877); also Watson's *Kinetic Theory of Gases*.

upon physical chemistry. Spectrum analysis has grown out of some apparently insignificant and disconnected observations made by Marggraf, Herschel, and others upon the light emitted by flames coloured by certain salts. The spectra of such flames were investigated by various physicists, among whom Talbot, Miller, and Swan deserve first mention ; but it was only after Kirchhoff¹ (in 1860) had made and proved the definite statement—that every glowing vapour emits rays of the same degree of refrangibility that it absorbs,—that spectrum analysis became developed by Bunsen and himself into one of the great branches of our science. Its importance for analytical chemistry has already been touched upon.

The application of the spectroscope to the determination of the composition of the heavenly bodies, and with this the firm establishment of stellar-physics, must be mentioned here. With respect to general chemistry, the efforts to arrive at harmonic relations between the lines of the spectrum themselves, and at a connection between those lines and the atomic weights of the elements which give rise to them, appear to be well founded, as is seen from the work of Maxwell, Stoney, Soret, and Lecoq de Boisbaudran. A complete theory of the spectral phenomena peculiar to gases remains still a problem for the future, although much admirable preparatory work has been done on the subject.

Atomic Volumes of Solids and Liquids.

The endeavour to establish relations between the physical properties of solid and liquid bodies and their chemical composition has given rise to a large amount of investigation, of which the most important must be mentioned here. H. Kopp was the first to work out in a thorough manner the connection between the specific gravity of elements and compounds and their atomic composition. After establishing the atomic or specific volumes of these latter, this chemist succeeded in discovering a number of relations, and, more particularly, in working out the specific volumes of

¹ *Pogg. Ann.*, vol. cix. p. 275.

the elementary atoms in compounds; it thus became possible to calculate the atomic volumes of complicated compounds.¹

The work done of recent years in this branch, among which that of Thorpe, Lossen, Staedel, and R. Schiff may be mentioned, has for the most part been carried out upon the principles laid down by Kopp; it has resulted in bringing out many new points of view, and has led to a number of modifications in the values arrived at by him. The formerly accepted opinion—that the atomic volumes of the elements in their compounds are mostly invariable—has been greatly shaken by this later work. Among the numerous researches (in addition to H. Kopp's) which have been made with the object of discovering a connection between the volumes of solid compounds and their atomic composition, those of Schroeder are especially worthy of note. He assumes volume units² of chemically analogous elements, and believes that he has therewith found the key to the solution of the above problem. But here again we are still far from a knowledge of any law governing the atomic volumes of solid or liquid compounds, whereas, in the case of gases, the simple relations existing between specific gravity and composition were worked out a long time ago.³

Laws regulating the Boiling Temperature.

Kopp was likewise the first, in his classical researches,⁴ to point out a connection between the boiling temperature and the composition of compounds (more especially of organic ones), in so far that he drew from his results the deduction that approximately equal differences in boiling-point correspond to equal differences in the composition of organic substances. And even although this supposed regularity turned out to be only applicable to certain compounds, and

¹ Cf. Kopp's pioneering researches, *Ann. Chem.*, vol. xli. p. 79; vol. xevi. pp. 153 and 303.

² These units he terms *Steren*.

³ Cf. p. 441.

⁴ *Ann. Chem.*, vol. xli. pp. 86 and 169; vol. lv. p. 166, etc.

could not be relied upon for other series, still Kopp's work gave a powerful impetus to the search after actual relations between boiling-point and chemical composition.

The question arose,—in what manner does the different chemical constitution of isomeric and chemically analogous compounds exercise an influence on their boiling temperatures—to be subjected to examination by Kopp.¹ Other more recent and more extended investigations, *e.g.* those² of Linnemann, Schorlemmer, Zincke, Naumann, and others, have resulted in showing that there are a number of definite relations here also, without, however, having rendered it possible to formulate a precise law setting forth the dependence of boiling-point upon chemical constitution; but it has been clearly established that there is a distinct connection between them. It is possible that a closer knowledge of the intimate relation sought for may be arrived at rather from the occasionally observed anomalies (*e.g.* the lowering of boiling temperature with increasing molecular weight, as in the case of the glycols and certain chlorine compounds, etc.) than from regularities.

There have not been wanting zealous endeavours also to discover regular relations between the temperatures at which solid substances become liquid and their composition, but no definite results have been arrived at in this way. Of more importance, however, have been the researches made with the object of determining melting-point and heat of solidification, *e.g.* those of Pettersson and Nilson, and those on the influence of pressure upon melting-point (James Thomson, Bunsen).

Raoult's Law of Solidification.

The work of Raoult, de Coppet, and others in this branch has been of quite exceptional significance, having led to the striking result that the temperature of solidification

¹ *Ann. Chem.*, vol. l. p. 142; vol. xvi. p. 1.

² Cf. A. Naumann, *Allgemeine und Physikalische Chemie* ("General and Physical Chemistry"), (1877), p. 553 *et seq.*

of a solvent is depressed in a regular manner by any substances dissolved in it. It has been proved, in fact, that if molecular quantities of different substances are dissolved in equal amounts of a solvent, they reduce its point of solidification to the same extent.¹ This simple observation, which has been abundantly verified, can thus be applied (conversely) to the determination of molecular weights;² it has therefore led to conclusions of the utmost value. Special mention must be made of the interesting results deduced therefrom by van 't Hoff, Arrhenius, Ostwald, and Planck with respect to the constitution of salts contained in dilute solutions.³ These have led to the view that the latter become dissociated under such conditions, that, for instance, chloride of sodium is not present as such in a dilute aqueous solution, but is there separated into its sodium and chlorine atoms. And however improbable such conceptions may appear to chemists, cogent reasons have been adduced in favour of this "dissociation hypothesis." The phenomena of electrolysis and of electrical conduction are readily explained by its aid, as can be seen from the investigations and discussions carried on by Ostwald and others.⁴

A similar relation (parallel to the observation of Raoult and others upon the dependence of the temperature of solidification on the molecular weight of the compound dissolved) has been discovered between the vapour pressure of a solvent and the molecular weight of the dissolved substance, the diminution in vapour-pressure being proportional to the latter (*i.e.* to the molecular weight of the substance dissolved).⁵

¹ Raoult designates this regularity as the *Loi générale de la congélation*, and states it as follows: "If 1 mol. of a substance be dissolved in 100 mols. of any given solvent, the temperature of solidification of the latter will be reduced by 0.63°."

² Chemists have not neglected to carry out determinations of molecular weight by the aid of Raoult's method (cf. V. Meyer, *Ber.*, vol. xxi. p. 1068; Auwers, *ibid.*, p. 701; Beckmann, *ibid.*, vol. xxi. pp. 766, 1163, etc.)

³ Cf. *Ztschr. physik. Chem.*, vol. ii. (1888).

⁴ *Ibid.*, vol. ii. p. 901; and especially vol. iii. p. 588; Arrhenius, *ibid.*, vol. iv. p. 96.

⁵ Cf. Raoult, *Comptes Rendus*, vol. lxxxvii. p. 167.

Specific Heat of Solid Bodies.

The work which has been done upon the specific heat of elements and compounds is among the most important in the whole field of physical chemistry, the dependence of this property on the atomic composition having been definitely established. We would recall here the Dulong-Petit law of the approximate equality in the specific heats of solid elements, the significance of which for the development of the atomic theory has already been detailed in the general section;¹ the extension of this law by Neumann; and its enlargement by Regnault's classical researches, as well as by those of H. Kopp, Weber, and others, which proved that the specific heat varies with the temperature at which it is determined. And even if the confidence felt in the applicability of the Dulong-Petit law was shaken by the marked deviation from it shown by certain elements, still its usefulness in a very large number of cases and the great value of its principle remained; as Berzelius had predicted, it formed "the foundation of one of the most beautiful pages in chemical theory." The investigation of the specific heat of liquids has not led to conclusions of such a general nature as have resulted in the case of solids.

Optical Behaviour of Solids and Liquids.

A long series of excellent experimental researches has been induced by the endeavour to discover definite relations between the optical behaviour of solid and liquid substances and their chemical composition. The earlier labours of Becquerel, Cahours, and Deville, and the later ones of Gladstone and Dale, Landolt, Brühl, Kanonikoff, and others have led to conclusions of importance respecting the connection between the constitution of a substance and its power of refracting a ray of light.²

¹ Cf. p. 207.

² For the literature on this subject, cf. Landolt and Börnstein's *Physikalische-Chemische Tabellen*, p. 220; and Ostwald's *Lehrbuch*, vol. i. p. 437 *et seq.*

The working out of the refraction-equivalents pertaining to the individual elementary atoms within their compounds has led to the discovery of stöchiometric regularities with respect to refraction. Of especial interest is the proof that the varying function or mode of combination of the elements, carbon in particular, has a determining influence on the molecular refraction. If the latter is accurately known, then conclusions may be drawn from the refractive power as to the constitution.

Only a passing reference can be made here to the importance to crystallography of the observed relations between light refraction and crystalline form, and to the pioneering work of Brewster and Fresnel on the subject.

Another optical property of many substances, more especially organic, has greatly excited the interest of chemists in quite recent years, viz. circular polarisation, which it has been attempted, and with success, to connect closely with the chemical constitution of the compounds in question. After the first memorable investigations of Arago, Biot, and Seebeck had been made, the observation—that certain substances, whether in the solid or liquid state, are capable of turning the plane of polarisation of light—was held to be of importance for physics alone. It has only been since Pasteur's beautiful researches¹ on the optically active tartaric acids, and the inactive racemic acid produced by their combination, that relations between optical activity and crystalline form have been discovered, and deductions drawn from these as to chemical constitution.

The desire to gain light upon this point produced in 1874 a theory, which was given out at the same time and independently by Lebel² and van 't Hoff,³ and which is based upon the hypothesis that the cause of this optical activity is to be sought for in the presence of an asymmetric carbon atom, *i.e.* one which is linked to four other different

¹ *Comptes Rendus*, vol. xxiii. p. 535 (1848); vol. xxix. p. 297; vol. xxxi. p. 480.

² *Bull. Soc. Chim.* (2), vol. xxii. p. 337.

³ *Ibid.* (2), vol. xxiii. p. 295.

atoms or radicals. Should this assumption become fully demonstrated (and it has this in its favour,—that an asymmetric carbon atom has been found in every optically active substance whose constitution has been determined with the necessary accuracy), then it may with confidence be stated that there is an intimate connection between this physical property and chemical constitution.

We may again refer shortly here to van 't Hoff's spacial conception of the distribution of the four valencies of the carbon atom (represented as in the middle of a tetrahedron, with its four affinities at the four corners), and to the extension of this hypothesis by Wislicenus, who has explained by its means the constitution and formation of geometrical isomers, *e.g.* fumaric and maleic acids, and the crotonic acids, with their derivatives.¹ Even already speculations of this nature have proved themselves fruitful, in that they have led to the perception of relations which had been hitherto overlooked.

In conclusion to what has just been said with regard to circular polarisation, mention must be made here of the work done upon the rotation of the plane of polarisation by a magnet, since stöchiometric regularities, *i.e.* relations between magnetic polarisation and chemical constitution, have been brought to light in this case also by the careful investigations of W. H. Perkin sen.²

Diffusion, etc.

The properties of liquids which are comprised under the designation "capillarity," together with the friction and diffusion of liquids and of solutions of solids in liquids, have given rise to numerous and valuable researches. It is to be hoped that, through the accurate measurement of the constants appertaining to those properties, material will be gained for the determination of the molecular weights of liquids—a problem which is as yet but imperfectly solved.

¹ Cf. p. 336.

² *Journ. pr. Chem.* (2), vol. xxxi. p. 481; vol. xxxii. p. 523.

Reference may be made here to the work upon capillarity-phenomena of Quincke, Mendelejeff, Wilhelmy, Volkmann, R. Schiff, and Traube, which has placed beyond all doubt the existence of a connection between capillarity and chemical constitution.

Graham's memorable researches¹ gave a powerful impulse to the investigation of fluid friction and diffusion: here, too, relations have been found between these phenomena and chemical composition. Mention must be made, in conjunction with this, of his division of substances into crystalloids and colloids, according to their behaviour on diffusion. The reader is also referred to the work upon osmose (so nearly connected with diffusion, and of such great importance for physiology) by Jolly, C. Ludwig, Pfeffer, and Brücke.

Electrolysis of liquid or of dissolved Substances.

The importance of the first work which was done upon this subject for the development of the electro-chemical theory has already been shortly touched upon in the general section.² The connection, so early assumed between electricity and chemical action, received the most brilliant confirmation from Faraday's electrolytic law, according to which equal amounts of electricity, when passed through different electrolytes, set free equivalent quantities of analogous substances at the two poles.³ This law was vigorously contested by Berzelius, because it appeared to him to imply that all the components of the substances decomposed by the current were held together in these by equal affinities. Later experimental researches have corroborated the validity of this law in its full extent, and permit of our hoping for a definite solution of the important problem of chemical equivalents, and, therewith, of the true saturation-capacities of the elements; the

¹ *Phil. Trans.* for 1850, 1851, and 1861; or *Ann. Chem.*, vols. lxxvii., lxxx., and cxxiii.

² Cf. p. 216.

³ Cf. p. 215.

reader is here reminded of Renault's investigations¹ on the various "electrolytic equivalents" of one and the same element, according to the nature of the compounds in which it is contained.

These and other observations have helped to make clearer the process of electrolysis itself, in so far that they have shown the intimate mutual relations existing between chemical and electrical energy. In the light of this Faraday's law appears as the expression of the fact that equal quantities of electricity require equivalent amounts of *ions* in their passage through different electrolytes. Electric conductivity and its relations both to physical properties and chemical composition have frequently been made the subject of investigations, among others by G. Wiedemann, Lenz, Long, and W. Ostwald. The recent work of the last-named chemist, more especially, has proved that a close connection exists between the conductivity of acids and their affinity for bases.

It has been attempted, too, to connect magnetism with chemical properties. The researches of Plücker, and especially those of G. Wiedemann,² have, in fact, resulted in showing that there are certain definite relations between the intensity of the magnetism of compounds and their chemical nature.

Isomorphism, etc.

The investigation of the connection between the forms of solid bodies and their composition has been of great significance for the development of chemical doctrines. The growth of crystallography benefited mineralogy in the first instance, but it also led to the discovery of isomorphism, which—as already stated in the general section³—exercised great influence upon the atomic theory. The services rendered here by E. Mitscherlich, to whom even G. Rose owed much, may again be recalled at this point. Mitscher-

¹ *Ann. Chim. Phys.* (4), vol. xi. p. 137.

² *Pogg. Ann.*, vol. cxxvii. p. 1 ; vol. cxxxv. p. 177.

³ Cf. p. 208.

lich did away with the erroneous conceptions which ascribed the crystalline form of a substance to the presence of minute quantities of other bodies, and proved irrefutably the connection existing between crystalline form and chemical composition. The deduction drawn both by himself and by Berzelius, viz. that in true cases of isomorphism of several substances, the chemical constitution of all became known as soon as that of any one of them was made out, because similarity of crystalline form is "a mechanical consequence of similarity in atomic constitution,"—this deduction was soon overthrown by observations of a contrary nature. It was found that dissimilarly constituted substances might be isomorphous, and analogously constituted ones heteromorphous; Mitscherlich himself added to his brilliant discovery of isomorphism that of *dimorphism* and *polymorphism*, while Scherer pointed out cases of the so-called *polymeric* isomorphism, which proved that elementary atoms might be replaced by atomic groups without change of crystalline form.

These and other similar observations have resulted in the view that isomorphism is only to be applied with great caution as a means for determining chemical constitution, as otherwise false conclusions are unavoidable. A passing reference may be made here to the later researches of H. Kopp upon the relations between isomorphism and atomic volume, and to those of Schrauf, Pasteur, and others upon the phenomena of *isogonism*. The problem—what changes of crystalline form are produced through the substitution of particular atoms by other atoms or radicals—has been systematically worked at by P. Groth¹ in the case of certain groups of organic compounds; the phenomenon of the partial alteration of crystalline form, in consequence of such substitution, he terms *morphotropism*. But much study is still required for the investigation of this newly opened out branch of the science.

The so-called *allotropism* of elements and compounds is probably closely connected with polymorphism, i.e. with the

¹ *Pogg. Ann.*, vol. cxli. p. 31.

fact that the same chemical substance can exist in different forms. A most important distinction between the two kinds of phenomena consists, however, in this,—that we have in the former case chemical as well as physical differences. Reference has been already made, under the history of the elements, to the discovery of certain of the more striking “allotropic modifications” of these.¹ But it may be mentioned at this point that material progress has recently been made in this branch through the investigation of the physical constants of such allotropic bodies, *e.g.* their specific heat, heat of combustion, atomic volume, etc.²

Speaking generally, chemists lean to the idea that the same cause underlies both allotropism and polymerism, and that therefore the former is to be explained by assuming that different numbers of atoms (of one and the same element) are grouped together into dissimilar molecules; as has been stated already, the molecular weights of oxygen and ozone have been established, and therewith the difference between them explained.

Thermo-Chemistry.

It is now a long time since the first attempts were made to determine the amounts of heat liberated during and in consequence of chemical reactions, with the object of thereby arriving at a measure of the affinities active in those processes. But the efforts of Laplace and Lavoisier, Davy, Rumford, and others in this direction remained incomplete, their methods for the estimation of heat quantities being too inexact.

Thermo-chemistry only became firmly established with the exact measurement of the thermal changes accompanying chemical reactions. Of the earlier investigations, those of Favre and Silbermann on heat of combustion deserve especial mention, because the calorimeter was materially improved by these chemists. Emphasis must also be laid here upon

¹ Cf. p. 377.

² Cf. the work of Hittorff, Lemoine, and others.

the almost forgotten labours of G. H. Hess,¹ who deduced from his own observations the all-important principle of the *Constanz der Wärmesummen* (i.e. that the heat evolved in the formation of a given compound is always the same), and thus taught in 1840 the application of the first law of the mechanical equivalent of heat to chemical reactions, before ever the latter had been brought forward.

From this principle Hess² established the point that the amount of heat evolved in any chemical reaction was always the same, whether the reaction was consummated at once or by degrees in separate instalments. This law, taken in conjunction with the principle at which Lavoisier and Laplace had arrived fifty years before—viz. that the decomposition of a compound into its constituents requires exactly the same amount of heat as is evolved during its formation from the latter—constitutes the basis of thermo-chemistry.

Since the conception of heat as energy of motion found perfect expression in the mechanical theory, and especially since the development of the term *energy*, the above principles appear as self-evident deductions from that theory. The earliest application of the mechanical theory of heat to thermo-chemical processes was made by Julius Thomsen,³ who has devoted himself to investi-

¹ To Ostwald belongs the merit of having referred with emphasis, in his *Lehrbuch der allgemeinen Chemie*, to the services of the St. Petersburg chemist, Hess, as the founder of thermo-chemistry. On p. 12 of vol. ii. of his book (among other passages) Ostwald expresses himself as follows: "In his fate we find a repetition of that which befel Richter, the importance of whose work for stöchiometry was for so long overlooked. Hess himself (*Journ. pr. Chem.*, vol. xviv. p. 120) assigned to the latter his proper position by correcting the mistake of confounding Richter with Wenzel, which was due to Berzelius. It is now again needful that the same loving service should be rendered to him, who on his part did justice to an investigator wrongly criticised and insufficiently esteemed in his own day."

² *Pogg. Ann.*, vol. l. p. 385 (1840).

³ Julius Thomsen, born at Copenhagen in 1826, where he continues to work as Professor at the University, has since 1852 applied himself with the utmost ardour to the building up and development of thermo-chemistry. The large number of scattered papers, which contain the records of his comprehensive researches, were a few years ago collected together and published by him in four volumes under the title *Thermochemische Untersuchungen* ("Thermo-chemical Researches").

gating thermo-chemically the more important chemical reactions, *e.g.* the formation of salts, oxidation and reduction, and the combustion of organic compounds. This branch of the science has been enriched by him in an extraordinary degree, both by the working out of new methods and by the systematic investigation of numerous chemical processes. In addition to Thomsen, Berthelot¹ and (since 1879) F. Stohmann² have contributed in conjunction with their pupils a large number of important observations in thermo-chemistry, and have materially assisted in the refinement of calorimetric methods.

The efforts of these investigators were mainly directed to the discovery of relations between the thermo-chemical values (which, calculated upon the molecular weights of the reacting substances, were termed *molecular heats*) and the chemical constitution of compounds. The heats of combustion, in particular, furnished much food for speculations of this nature. But although regularities of various kinds became apparent, *e.g.* with respect to the heats of combustion and heats of formation in homologous and other series, very great caution requires to be exercised in forming deductions as to constitution from calorific values; this has lately been clearly shown by Brühl,³ in a critique upon such attempts. A salutary limit has thus been placed upon the too great extension and over-valuation of the conclusions drawn from thermo-chemical work, a temperate criticism (on the part of Lothar Meyer and others) having previously done away

¹ M. P. E. Berthelot, born in Paris in 1827, became Professor in the Collège de France there, and recently held for a short period the post of Minister of Education; he first made himself known by the beautiful researches, already spoken of, entitled, *Sur les Combinaisons de la Glycerine avec les Acides*. He soon directed his attention to the synthesis of organic compounds, which at that time had been but little studied, and in his comprehensive work, *Chimie Organique fondée sur la Synthèse* (1860), gave a detailed account of the observations and discussions in this branch of the science. Later on he turned with all his energy to the experimental solution of thermo-chemical problems, which he collated together in the two-volume book, *Mécanique Chimique fondée sur la Thermochimie* (1879). To him we also owe the recent historical work *Les Origines de l'Alchimie* (cf. p. 23).

² Cf. his papers, published in the *Journ. pr. Chimie* since 1879.

³ *Journ. pr. Chimie* (2), vol. xxxv. pp. 181, 209.

with the erroneous view that an absolute measure of affinity was furnished by the heat evolved or absorbed in the formation or decomposition of chemical compounds. In spite, however, of this failure, thermo-chemical investigations will certainly prove to be indispensable for the perfected doctrine of affinity of the future.

Photo-Chemistry.

This short account of the growth of physical chemistry would be incomplete if nothing were said respecting the chemical action of light. The latter, a particular form of radiant energy, gives rise—as is well known—to various chemical reactions, of which the great process of assimilation in plants was the earliest to attract the attention of chemists. The detailed treatment of this process, first observed towards the end of last century, belongs to the recently developed science of vegetable physiology.

The earliest superficial observations on the action of light upon compounds of silver were made by Schultze so long ago as at the beginning of the eighteenth century: indeed, Boyle had noticed the blackening of chloride of silver, but had ascribed it to the influence of the air. The fundamental experiment which called photo-chemistry into life was made by Scheele, who thus proved himself a pioneer in this as in other branches of the science; he studied the action of the solar spectrum upon paper covered with silver chloride, and established the point that the effect begins first and is strongest in the violet portion. We must recall here too the experiments of Ritter, who observed the action of the ultra-violet rays; and, especially, the epoch-making discoveries of Daguerre and Talbot, which gave birth about the year 1839 to the art of photography, so enormously developed of recent years, they having succeeded, after many attempts, in permanently fixing light-pictures.¹

¹ The following notes may be added here upon the history of photography: Niépce had associated himself with Daguerre in his work, but did not live to see the perfecting of the Daguerreotype process. Talbot replaced Daguerre's silver plates by paper rendered sensitive to light. Among the further advances made in photography may be mentioned the production of negatives

The foundation of comparative photo-chemistry, which is termed actinometry, was laid by the memorable researches of Bunsen and Roscoe,¹ Draper² having previously made important experiments in a similar direction. These investigators, along with others, *e.g.* B. H. W. Vogel, made clear the laws to which the actinic rays are subject. Especially remarkable were the results of the observations on the absorption of chemically active rays, and upon photo-chemical induction, a term employed by Bunsen and Roscoe to designate the process by means of which the substance sensitive to light was brought into such a condition that it underwent decomposition proportional in amount to the intensity of the light. In addition to the above, mention must be made here of the remarkable researches of Tyndall upon vapours and gases sensitive to light, in whose decomposition the action of the light is shown; thus he proved that the vapour of amyl nitrite (to give an instance) was decomposed by the actinic rays.

The phenomena, whose investigation has just been treated of, come properly speaking under the doctrine of affinity, whose task it is to show that chemical reactions, *i.e.* the formation and decomposition of chemical compounds, are the results of definite measurable forces. True, this important branch of the science is still far from attaining to such a goal; but the development of the doctrine of affinity, a short sketch of which now falls to be given, shows that much zealous work is being done with the view of solving the difficult problems here involved.

upon glass and the application of collodion (Niépce de St. Victor—the nephew of the Niépce mentioned above—and Legray, 1847); the multiplication of photographic pictures through pressure by means of the so-called photolithography, heliography, and the phototype method, which in time became superseded by the splendid autotype process (Meisenbach) and the heliotype one (Obernetter); and, lastly, the preparation of plates particularly sensitive to light (bromo-gelatine, etc.)

¹ *Phil. Trans.* for 1857, p. 355, and for 1863, p. 139; or *Pogg. Ann.*, vol. c. p. 43 (1857); vol. cxvii. p. 531.

² *Phil. Mag.* for 1843.

*Development of the Doctrine of Affinity since the
Time of Bergman.*

In a previous section of this book an account has been given of the earlier efforts to arrive at a knowledge of the phenomena of affinity. Through most of the speculations upon this question, ever since the time of Boyle, there runs the assumption that the so-called force of chemical affinity is in the main identical with that of gravity; only in that the former is exerted within very small distances, whereby the form of the material particles has to be taken into account, are differences between the two forces apparent. The attempts to estimate the affinity of substances for one another remained at that time (*i.e.* previous to Berthollet) very imperfect, because it was sought to determine qualitatively the relative intensities of the affinities under arbitrary conditions, without taking physical considerations into account. This period, from about the time of Geoffroy (1718) to that of Berthollet (1800), is characterised by the bringing out of "Tables of Affinity" (*Verwandtschaftstafeln*).¹

Bergman's doctrine of chemical affinity and his determinations of the latter belong in part to this evolutionary stage, although he paid more attention to the influence of temperature upon the phenomena investigated by him than his predecessors had done. The reaction proper against the merely empirical conception of these latter is, however, to be found in Berthollet, whose *Essai de Statique Chimique* was a protest against the neglecting of physical conditions during chemical processes.

*Bergman's Doctrine of Affinity.*²

Although the work of this investigator belongs to the phlogistic period, his doctrine of affinity can only be conveniently treated of here, in order that it may be compared or rather contrasted with that of Berthollet.

¹ Cf. p. 131.

² Cf. Bergman's *Opuscula phys. et chem.*, vol. iii. p. 291 (1783).

Bergman's conception of the phenomena of affinity came into such general adoption, that it is to be found even now, at least portions of it are, in many text-books.

The chief law of his doctrine states that the value of the affinity between two substances which act chemically upon one another is constant under similar conditions, and therefore that it is independent of the masses of those substances. Bergman assumed the universal force of gravity as the cause of affinity, this being, however, greatly modified by the form and position of the small particles of the reacting bodies. Partly from his own speculations with regard to affinity, and partly from the incorrectly determined composition of neutral salts, he drew erroneous conclusions with respect to the magnitudes of the affinities of bases to acids, and *vice versa*: he thus set up the tenet that an acid has the strongest affinity for that base of which it saturates the largest quantity, in order to form a neutral salt. Berthollet, as will presently be shown, deduced precisely the opposite from his own assumption,—that mass-action comes into play in chemical processes. It is noteworthy that Bergman recognised the impossibility of carrying out absolute affinity-determinations, and that he devoted his entire energies to making relative ones (by decomposing one compound by another), and then collating these in "affinity tables."

Berthollet's Doctrine of Affinity.

Against Bergman's ideas, and especially against the assumption that affinity is independent of the masses of the interacting substances, Berthollet raised a lively opposition. Setting out, like Bergman, with the hypothesis that affinity is identical with gravity, he went on to emphasise the undeniable conclusion that the forces of chemical affinity, like those of general attraction, must be proportional to the masses of the acting substances. The further deductions from this principle he worked out with a masterly clearness in his *Essai de Statique Chimique*.

These views of Berthollet did not at the time receive

the recognition which they merited, mainly, no doubt, because their author came into collision with established facts by carrying his deductions from them too far. His fundamental law of the dependence of chemical action upon the masses of the substances concerned in it led him to regard the "chemical effect" of any body as the product of its affinity and mass. From this he drew the further conclusion that the formation and composition of a chemical compound depended substantially upon the masses of the acting constituents which went to produce it. According to this view, any two substances must combine with one another in constantly varying proportions; with this deduction, however, Berthollet found himself in a serious dilemma.

But, if he went too far here, he so immensely advanced the doctrine of affinity and followed up its true aims by a more discreet application of his fundamental principle, that the errors into which he fell may well be forgotten. He pointed out with perfect clearness that it was impossible to determine the absolute values of chemical affinities, seeing that these were necessarily materially dependent upon the physical properties of the substances which were formed or decomposed by the chemical reactions in question. According to him, such determining (and opposite) properties were *cohesion*, i.e. the mutual attraction of the small particles of any substance for one another, and *elasticity*, i.e. the tendency of those particles to occupy the greatest possible space. He saw in the greater or lesser insolubility of substances a measure of cohesion, and in their volatility a measure of elasticity, and by means of such conceptions conclusively explained chemical changes in which the separation of a precipitate or the escape of a gas or vapour had a determining influence on the course of the reaction. In fact, he stated distinctly that a complete rearrangement (*Umsetzung*) of substances can only take place if cohesion or elasticity comes into play, and never by the mere action of affinity alone. He thus brought forward entirely new points of view, which have borne much rich fruit.

*The Supplanting of Berthollet's Opinions by
other Doctrines.*

The first good which resulted from Berthollet's conception consisted in the recognition of the uselessness of tables of affinity, in so far as these were supposed to give the relative affinities of different substances. The important fundamental idea of his doctrine of affinity, viz. that the chemical action of a body is proportional to its mass, and is therefore to be expressed by the product of this into the affinity (*i.e.* by a factor still to be determined), led Berthollet to conclusions which were directly opposed to many known facts, and to numerous other data worked out at that time by Proust. The controversy between these two men, which turned upon the question whether chemical compounds are built up of elements in proportions which only alter in amount by certain definite increments, or in proportions which continually vary, has already been treated of in the general section (*cf.* p. 175).

In bringing forward his theory Berthollet either neglected to pay sufficient heed to the stöchiometric relations known at that time, or else his knowledge of these was incomplete. It is precisely to the circumstance that he carried his theory of mass-action too far, and made it the starting-point for the most far-reaching deductions, that we have to ascribe the miscredit into which his principles—notwithstanding their clearness—fell, in fact they were held to be totally erroneous. It was thus that Bergman's doctrine, although based upon wrong assumptions and therefore leading its author to false conclusions, kept for so long a time the upper hand, and this all the more readily since it could be better made to accord with the atomic theory. The revival of Berthollet's principles was reserved for quite recent times, after various isolated experimental researches had furnished proof of their admissibility.

After Berthollet's temporary overthrow, the rapidly developing atomic theory formed the main subject of in-

terest for chemists; and hand in hand with the building of it up there went the development of electro-chemical doctrines, whose object it was to show that the closest connection existed between electricity and the force termed affinity.

The doctrine of affinity now sought to perfect itself through the development of electro-chemistry: Berzelius' theory caused Berthollet's to be neglected. The successful work which has since been accomplished, with the object of getting at the actual relations between electrolysis and affinity, enables us to perceive now that in those efforts the investigators of that time were carried too far.

Those endeavours could only result in showing the qualitative differences in the affinities of different substances; in fact, the electro-chemical theories reached their culminating point in the proof of an analogy between the electrical and chemical properties of substances. Faraday's electrolytic law, which threw light upon the quantitative side of electrolytic processes, did not give any information as to the relative magnitudes of the affinities of the substances in question.

The fortunes of the most important of the electro-chemical theories, that of Berzelius, have already been described. Blomstrand's ingenious attempt¹ to bring it back to life again has indeed shown how valuable it is for the explanation both of chemical processes and of the constitution of compounds; but it has been unable to aid materially in penetrating the obscure domain of the phenomena of affinity.

New prospects were opened out for the doctrine of chemical affinity by the thorough investigation of thermo-chemical processes, whose importance for physical chemistry has already been referred to. But in this case also, as in the application of electro-chemical conceptions to the problems of affinity, the worth of thermo-chemical determinations very soon became greatly over-estimated. Thus, even Julius Thomsen, who was for a long period the most

¹ Cf. his work, *Die Chemie der Jetztzeit* (1869).

eminent worker in this field, regarded the heat evolved or absorbed in chemical reactions (more especially in the formation and decomposition of compounds) as an absolute measure of the affinity; in his view the work of affinity was transformed into measurable heat.

But although the insufficiency of thermo-chemistry for the solution of the problems of affinity has now been made manifest, its present and future significance must not be depreciated. On the contrary, by the careful application of the mechanical theory of heat to the interpretation of chemical processes, great benefits have already accrued to the doctrine of affinity.

The Revival of Berthollet's Doctrines.

The most powerful impulse to a further healthy development was given to the doctrine of affinity by the revivification of Berthollet's theory. This was accomplished in its fullest extent by the publication in 1867 of the work of two Scandinavian investigators, Guldberg and Waage.

Several years previous to this H. Rose had proved with absolute clearness the mass-action of water in many reactions, *e.g.* in the decomposition of alkaline sulphides and of potassium bisulphate, and in the formation of basic salts. The attention of such distinguished workers as Rose, Malaguti, Gladstone, and others had further been directed to the study of the mutual decomposition of two salts, whether those were soluble or one of them insoluble. In fact, attempts were made to work out in various ways the relative affinities of particular substances, and thus to solve a problem which Berthollet had sketched out theoretically.

The ideas of the latter received, lastly, valuable experimental confirmation from the extremely important researches of Berthelot and Péan de St. Gilles¹ on the formation of compound ethers and ether-acids from an alcohol and an acid. In subsequent theoretical discussions, these and the more recent valuable experiments of Mensch-

¹ *Ann. Chim. Phys.* (3), vols. lxxv. lxxvi. and lxxviii.

kutkin¹ (which furnished information with regard to the chemical equilibrium existing between different substances and to the time-rate of reaction, were applied with success to proving and confirming the correctness of Berthollet's axioms.

The observations on chemical equilibrium in *reciprocal* processes especially contributed to the general adoption of those doctrines of Berthollet: it was thought then (and still is) that the values thus obtained offered the surest data for arriving at the relative affinities of substances taking part in a reaction. With regard to the ideas held respecting such states of equilibrium, the opinion prevailed for a long time that a statical equilibrium must be assumed. A reversal of this was prepared for by the view originated and propounded by Williamson² in 1850, which was also worked out independently by Clausius several years afterwards, viz. that the atoms of substances are in a state of continual motion, not merely during chemical reactions but also when the substances are apparently at rest. A dynamical equilibrium thus took the place of a statical, *i.e.* an equilibrium of the opposing reactions. Pfaunder has of late ingeniously applied such speculations to the explanation of the phenomena of dissociation and of reciprocal reactions generally.

But although Williamson emphasised the point that his speculations were in accord with Berthollet's principles, a sufficiently secure and broad basis was still wanting, upon which they could at that time be further developed. Such a foundation for the building up of the doctrine of affinity was furnished by the work of Guldberg and Waage,³ who took Berthollet's axioms as their immediate starting-point, reanimated these anew, and proved their agreement with facts.

Like Berthollet, the investigators just named stated the chemical action of a substance as being proportional to its *active amount*,⁴ the latter being given by the quantity con-

¹ Cf. *Ann. Chem.*, vol. xciv.; *Journ. pr. Chemie* (2), vols. xxv. xxvi. xxix.

² In a paper read before the British Association at Edinburgh; *Ann. Chem.*, vol. lxxvii. p. 37.

³ *Études sur les Affinités Chimiques* (1867); this was published in German in the *Journ. pr. Chemie* (2), vol. xix. p. 69.

⁴ " . . . seiner wirksamen Menge proportional."

tained in unit of space. The intensity of the interaction of two substances is expressed, according to them, by the product of the active amounts; but a coefficient¹ still remains to be determined which shall express the dependence of the reaction upon the nature of the substances taking part in it, the temperature, and other factors. By the aid of such hypotheses the relations existing between the amounts of the reacting substances and their actions² can be deduced mathematically. Important conclusions have also been drawn from them with respect to time-rate of reaction and chemical equilibrium, and these have been found to agree sufficiently well with the results of actual experiment.

The latest Development of the Doctrine of Affinity.

Guldberg and Waage's theory, based as it was upon Berthollet's principles, has had an extraordinarily stimulating effect. In particular, it has led to the successful determination of the specific affinity-coefficients of different substances, especially of bases and acids; and these experimentally-determined constants have been made use of to test the correctness of the theory itself. Among the work done with this aim in view, that of Ostwald³ deserves especial mention; he has determined by different methods, volumetric and optical, the manner in which a base is distributed among different acids present in excess, and has deduced from this the specific affinity-coefficients of the latter. J. Thomsen⁴ had previously attempted to solve the same problem by thermo-chemical methods.

Ostwald⁵ further sought, a few years ago, to deduce the affinity-coefficients of acids from reactions which go on with a measurable velocity under the influence of those acids, *e.g.* the decomposition of acetamide and of methyl acetate, and

¹ Such affinity-coefficients have hitherto only been determined in particular cases, and then only approximately.

² *Wirkungen.*

³ Published in the *Journ. pr. Chemie* since the year 1877.

⁴ *Pogg. Ann.*, vol. cxxxviii. p. 575.

⁵ Cf. *Journ. pr. Chemie* for 1884 and 1885.

the inversion of cane sugar; in this case, too, the results obtained have shown a sufficiently near agreement with calculation. The reader is referred, lastly, to the remarkable relations which have been discovered by Arrhenius, and also by Ostwald, between the affinity-coefficients and the capacity for (chemical) reaction of acids and bases on the one hand, and their electrical conductivity in dilute solution on the other. Ostwald's researches¹ have thrown a surprisingly new light upon the chemical relations—especially upon the constitution—of the compounds investigated, since they showed that the affinity-coefficients of substances alter definitely according to the constitution of the latter. At the same time it has turned out that the position or function of the atoms has a determining influence upon those coefficients, this important fact being most apparent in the case of isomeric compounds, *e.g.* the oxy-benzoic and chloro-propionic acids, etc.²

The limits, within which this short account of the development of the doctrine of affinity is necessarily confined, would be widely overstepped were the results of other investigations—even taking only those of importance—to be described. Only a passing reference can be made to the work of Wilhelmy, which has led to a better knowledge regarding time-rate of reaction, and to that of Menschutkin, van 't Hoff, Horstmann, and others, which has resulted in making clear the conditions of chemical equilibrium in various reactions.

The hypothesis that the small particles of substances are in continual motion, not merely during chemical reactions, but also when the whole system is in a state of equilibrium, is now held to be indispensable for the new doctrine of affinity. The great aim of this latter is to convert chemistry into a branch of applied mechanics—an aim which Berthollet and Laplace, notwithstanding the imperfection of the appliances at command in their day, had the prescience to designate as the highest possible one.

¹ Cf. *Journ. pr. Chemie* (2), vol. xxxii. p. 300; and especially *Ztschr. phys. Chemie*, vol. iii. pp. 170, 241, and 369.

² Cf. also Raoult's work bearing on affinity-coefficients, as developed by Planck and others.

A SKETCH OF THE HISTORY OF MINERALOGICAL CHEMISTRY
DURING THE LAST HUNDRED YEARS.¹

Mineralogy only attained to the rank of a science after it had recognised the fact that chemistry was indispensable to it for ascertaining the composition of minerals. It is true that even in this century Mohs,² who did so much for mineral physics, almost denied that the chemical characters of minerals had any signification; but the system which he set up was only temporarily adopted by a few scientists. The benefits which accrued to mineralogy from the application of chemical aids were so obvious that the latter could never again be dispensed with. Mineralogy has been brought to its present high position by the joint assiduous work of mineralogists and chemists. The beautiful aim—of making clear the connection which exists between the physical and chemical properties of individual minerals—has firmly retained its place for the mineral chemist ever since the labours of Berzelius, Mitscherlich, G. Rose, and others were consummated.

The first modest attempts to gain a knowledge of the chemical composition of minerals were made in the seventeenth and first half of the eighteenth centuries, but these did not extend beyond mere superficial observations of a few qualitative reactions. In the second half of last century, however, there was much important preparatory work done, which helped materially to found the science of mineralogy.

¹ Cf. Kopp, *Geschichte der Chemie*, vol. ii. p. 84 *et seq.*; v. Kobell, *Geschichte der Mineralogie* (1650-1860), more especially p. 303 *et seq.*

² Mohs set up the axiom that a mineralogist had merely to consider the natural-historical properties of minerals, *i.e.* crystalline form, specific gravity, hardness, and so on. If their chemical behaviour is taken into account, then, he expressly states, mineralogy oversteps its legitimate bounds and entangles itself in difficulties. This renunciation of the most important aid to mineralogical research is certainly characteristic. Berzelius was fully justified in comparing such a mineralogist to a man who objects to use a light in the dark, on the ground that he would thereby see more than he requires to.

Mineral chemistry had its distinguished exponents in Bergman, and, a little later, in Klaproth and Vauquelin, whose services in devising methods for the analysis of inorganic substances have already been referred to.¹ The chemical investigation of minerals was carried on at that time, upon the principles which they laid down, by numerous other workers, among whom we may name Lampadius, Bucholz, Wiegleb, Westrumb, Valentin Rose the younger, Kirwan, Gadolin, and Ekeberg.

The extraordinary benefit which accrued to mineralogy from the introduction of the blowpipe by Cronstedt, and its subsequent use by Galin, Bergman, Rinman, and, particularly, Berzelius, may again be emphasised at this point.²

Even before the gradual development of a mineral chemistry, and also simultaneously with it, Romé de l'Isle, Werner, Haüy, and Bergman had recognised crystallography as being essential to the study of mineralogy, and had applied themselves to it. Haüy, in particular, achieved wonderful results in this branch; he referred back the various crystalline forms to a few primary ones, and took account of chemical as well as of physical properties in classifying minerals. That he carried his deductions too far here is seen from his well-known axiom that difference in crystalline form signifies also difference in chemical composition.

The endeavours made to classify minerals during that period are for the most part characterised by the desire to recognise their chemical as well as physical properties. If this had only a subordinate signification in Cronstedt's, Haüy's, and, especially, Werner's systems, it was on the other hand put prominently forward by Bergman³ as an essential aid to the classification of minerals, so far as this was possible with the then existing chemical knowledge. But few of the mineralogists of that day, however, subscribed to Bergman's principles, most of them giving in their

¹ Cf. p. 358 *et seq.*

² Cf. p. 359.

³ In his *Sciagraphia Regni Mineralis*, etc. (1782).

adhesion to Werner's system, in which only a very modest place was assigned to mineral chemistry.

A new life began for mineralogical chemistry when Berzelius turned himself to its study. Basing his arguments upon his own comprehensive labours, which had for their aim the exact determination of the composition of minerals and artificial inorganic compounds, he was enabled to show that the doctrine of chemical proportions (and therefore the atomic theory) was applicable in its fullest extent to minerals also.¹ He was the first to characterise these latter as being in every respect "chemical compounds." At the same time this gave him occasion to classify them similarly to substances prepared artificially, and thus arose his Chemical System,² in which he gave definite expression to the view that mineralogy should only form a part of, or appendage to, chemistry. The order of the minerals in his system was determined by the position of their electro-positive constituents in the so-called "tension series." Ten years later³ Berzelius altered his principle of classification, in so far that he came to look upon the electro-negative constituents as primarily determining this, and he arranged the minerals accordingly. For his two main classes he took non-oxidised and oxidised substances, and between those two he divided minerals with a marvellous perspicacity. All previous attempts at classifying minerals according to chemical principles were thrown into oblivion by Berzelius' system.

The development of this latter, whose main features were subsequently reproduced in later classifications, was influenced in the highest degree by an observation made by N. Fuchs, viz. that certain substances can replace each other in minerals, and still more by the extension of this doctrine through Mitscherlich's discovery of isomorphism.⁴ The results of the analyses of minerals hitherto obtained were

¹ Cf. p. 193.

² *Schweigger's Journ.*, vols. xi. and xii. (1814).

³ *Leonhard's Zeitschrift für Mineralogie*, vol. i.

⁴ Cf. p. 208.

henceforth regarded from entirely new points of view. A high, perhaps too high, significance was now attributed to crystalline form in its connection with chemical composition. This over-estimate quickly became manifest after Mitscherlich discovered the first cases of dimorphism,—to be extended later on to tri- and polymorphism. Haüy's principle—that a difference in crystalline form also means a difference in chemical composition—was thereby overthrown; and, in spite of the opposition of this distinguished investigator, the doctrine of isomorphism took its place triumphantly in mineralogy.

The various mineralogical systems which were brought forward after that of Berzelius, *i.e.* after the year 1824, are almost all characterised by the endeavour to classify minerals according to their chemical composition, a greater or lesser signification being at the same time attached to their physical properties. In addition to G. Rose's classification of mineral bodies, which rested upon a purely chemical basis, the mixed systems of Beudant, C. F. Naumann, and Hausmann may be named here as having become best known.

The nomenclature of minerals has by no means kept equal pace with their strictly scientific classification. The empirical principle still prevails here, this being apparent from the way in which minerals are named after their discoverers, or after localities in which they are found, or according to their physical properties, etc., instead of the name expressing or at least indicating their chemical composition.

Mineralogy owes its present flourishing condition to the immense development of mineral chemistry. Berzelius and his pupils, among whom Chr. Gmelin, E. Mitscherlich, Wöhler, H. and G. Rose, Svanberg, and Mosander may be mentioned, were the first to really open up the ground which Bergman, Klaproth, Vauquelin, and others had prepared. It is impossible to give a detailed account here of the wealth of new methods which have been devised for the analysis of minerals, and for the separation of their

individual constituents. The almost inexhaustible field of minerals has ever since then been investigated chemically by numberless workers. To the problem which naturally comes first, viz. the establishing of their empirical composition, the further and higher one was added of getting at their chemical constitution. The silicates in particular, on account of their extraordinary variety, have given rise to continually renewed investigations.¹

The limits of this short account of the development of mineralogical chemistry do not permit the citing of even a few examples of the services rendered to this branch of the science by such men as Stromeyer, Th. Scheerer, Rammelsberg,² Bunsen, and others. Among other chemists who have done good work for mineralogical chemistry the following may be named:—v. Bonsdorff, O. L. Erdmann, Marignac, Th. Thomson, Blomstrand, Deville, v. Hauer, Hermann, Th. Richter, Sandberger, Smith and Brush, Streng, Cl. Winkler, P. Jannasch, Lemberg, Th. Petersen; to these many more names might be added.

The Artificial Production of Minerals³—Beginnings of Geological Chemistry.

To the older analytical method, which was the one naturally first followed in the investigation of minerals, the synthetic method has in recent times been added, with the

¹ Efforts have not been wanting to apply specially to minerals the more recent chemical views which have been arrived at with respect to the constitution of organic compounds. Wurtz was the first to do this, by comparing the poly-ethylene alcohols (discovered by himself) with the poly-silicic acids.

² C. Rammelsberg, born in Berlin in 1813, has worked from the year 1840 partly at the Technical College (*Gewerbeschule*) and partly at the University there, and has been since 1874 head of the second chemical laboratory of the latter. His researches, which have greatly enriched inorganic and especially mineralogical chemistry, have appeared for the most part in *Poggendorff's Annalen*. He has rendered very great service by the publication of his *Handbuch der Mineralchemie* (second edition, 1875), and of his *Krystallographisch-physikalische Chemie* (1881-82).

³ Cf. *Die Künstlich dargestellten Mineralien*, etc. ("Artificially-prepared Minerals"), by C. W. C. Fuchs (Haarlem, 1872); and the *Synthèse des Minéraux et des Roches*, by Fouqué and Michel Lévy (Paris, 1882).

result that mineralogical chemistry has been enriched by an extraordinary number of new facts and has led to the development of geological chemistry. The endeavour to imitate and to explain the natural production of minerals, by preparing them artificially under various conditions, has been the cause of many memorable researches, of which a short account must be given here.

After Berzelius had defined minerals as chemical compounds whose composition was dependent upon the same laws as that of compounds artificially produced, the problem at once arose of preparing mineral substances from their components. But several decades passed by, during which mineral chemistry was developed by improved analytical methods, before the synthesis of minerals was definitely taken in hand with this conscious aim in view. Only isolated observations on the artificial formation of such, *e.g.* of calc-spar and arragonite by G. Rose, and some experiments made by Gay-Lussac, Berthier, and Mitscherlich, fall to be recorded during the first half of this century;¹ the brilliant development of this branch of mineralogical or geological chemistry only began in 1851 with the memorable labours of Ebelmen, Durocher, Daubrée, and Sénarmont. These investigators elaborated a series of methods which led to the production of minerals under conditions similar in part to those found in nature. It was justifiable to draw careful deductions with respect to naturally occurring processes from these methods of formation; at any rate, hypotheses which were brought forward to explain the formation of minerals and rocks could be put to the test in this way. Geology thus gained a firmer foothold, and found in chemistry an indispensable helpmeet.²

¹ The earliest observation of this nature was doubtless that made by James Hall upon the transformation of chalk into marble in 1801.

² Sénarmont expressed himself in the following significant words with regard to the necessity of chemistry for geology: "*C'est à la chimie minéralogique, que la géologie doit l'utile contrôle expérimental de ces conceptions rationnelles. Les minéraux cristallisés ont, en effet, une origine toute chimique, et c'est l'expérience chimique qui doit servir d'appui à la géologie, si elle veut faire un pas de plus dans l'étude des roches, qui en sont composées.*"

Reference may be made here to Bunsen's beautiful investigations¹ upon the geological conditions of Iceland, and especially upon the geysers, and to those on the formation of volcanic rocks, all of which were productive of new views; and also to the labours of G. Bischof,² who was indefatigable in advancing chemical geology.

Among the distinguished array of investigators who made further advances in this direction, and, in particular, who discovered new modes of formation of minerals, H. St. Claire Deville and Troost, Becquerel, Debray, Hautefeuille, Wöhler, Rammelsberg, R. Schneider, and especially Fouqué and Michel Lévy, stand out pre-eminent. Of recent years Friedel and Sarasin have carried out important syntheses of minerals.

The chief founders of the synthetic method in mineralogical-geological investigations have been Frenchmen, and so reference is with perfect justice made to a *French School* in this branch, the four gentlemen last named being now its chief exponents.³

The modes of formation of minerals observed by them vary greatly, the processes being partly wet and partly fusion ones. To mention only one or two of the more important, take the production of many natural minerals by the slow mutual decomposition of two salts in solution, *e.g.* the formation of quartz and calc-spar from gypsum and silicate of potash in presence of carbonic acid; the deposition of artificial minerals from solution (formation of gypsum); the decomposition of various substances by water under increased pressure (formation of quartz, wollastonite, apophyllite, etc.); and, lastly, the production of numerous

¹ *Ann. Chem.*, vol. lxii. p. 1; vol. lxv. p. 70.

² Cf. his *Lehrbuch der chemischen Geologie*.

³ Fouqué and Michel Lévy consider that the cause of this pre-eminence in the above field is to be found in the "nature of the French national character." The argument with which they support this assumption (see p. 5 of the work, *Synthèse des Minéraux*, etc.) is so characteristic, that it may find a place here: "*Notre génie national répugne à l'idée d'accumuler un trop grand nombre de faits scientifiques, sans les coordonner, et si cette tendance nous entraîne quelquefois à des hypothèses hasardées, elle a, d'autre part, le mérite, de nous induire aux expériences synthétiques.*"

minerals by processes requiring fusion and a white heat—processes similar to those which go on in volcanoes (formation of tridymite, olivine, and other silicates).

Since nature but seldom allows her workshops to be spied into, the numerous experiments on the production of minerals, made in imitation of natural processes, and which have been carried to a successful issue, possess the highest significance for the explanation of those processes. The repeated proofs that one and the same mineral can be artificially prepared in the most diverse ways, by wet as well as by fusion methods, has rendered the former one-sided conception of geological processes (*i.e.* the view that rock-masses have been produced *either* in the wet way *or* by igneous action) almost impossible now. The synthesis of minerals has riveted still more firmly than before the already long-established link between mineralogy and chemistry.

DEVELOPMENT OF AGRICULTURAL AND OF PHYSIOLOGICAL CHEMISTRY.

The history of these branches of chemistry is primarily associated with the work done by Liebig, of which a short description has already been given in the General Section. It is true that this gifted investigator had many predecessors, who found out various isolated chemical facts of great importance for vegetable and animal physiology; but it was he who first with far-seeing glance collected such facts (and, especially, new observations of his own) together under general points of view. The ideas of a Palissy upon the necessity of mineral substances for plant life;¹ the investigations which towards the end of the seventeenth century led Malpighi and Mariotte to definite conclusions with respect to the nutrition of plants through their leaves and roots; the bold and comprehensive speculations of Lavoisier² regarding metabolism in plants and animals,—his conviction that the life processes are made up of a series of chemical reactions; lastly, the work of Fourcroy, Vauquelin, Proust, Berzelius, and Chevreul upon products of the animal body;—all these, together with other labours, served to prepare the ground upon which Liebig afterwards raised the edifice of chemistry in its relation to agriculture, physiology, and pathology.

Those branches of chemistry are most closely interlaced with organic, for one of their main problems consists in isolating compounds of an organic nature and establishing the composition of these. To this is added the further task of finding out the *rôle* which such substances fill in the organism. Vegetable and animal physiology are especially indebted to chemistry in questions of nutrition.

¹ Cf. p. 87.

² These are set forth in a paper written in 1792, but only published in 1860 (in vol. iv. of the *Œuvres de Lavoisier*).

*Agricultural Chemistry and Vegetable Physiology.*¹

The work done in physiological chemistry towards the end of last century and the beginning of this by Priestley, Ingen-Houss, Senebier, and Th. de Saussure had led to many important results with respect to the nourishment of plants. One might now suppose that, from the analysis of the ashes of plants, a distinct connection between the plants and the soil would have been apparent. The decomposition of carbonic acid by the leaves, which was observed by those workers, ought, one might further suppose, to have pointed to carbonic acid as the main source of the organic matter of plants. In like manner the early made observation that salts of ammonia were highly conducive to the growth of vegetables,² might have found an explanation in the recognition of ammonia as the source of their nitrogenous constituents.

These deductions, however, which now appear to us self-evident, were not drawn, and it was sought to credit humus as being the universal nutrient of plants, without paying any heed to those older fundamental observations which have just been mentioned. The processes of nutrition of plants were thus entirely misunderstood, for, according to this doctrine, they fed like animals upon organic matter.

This assumption, which dominated agricultural chemistry for many decades, found its chief advocates in Germany and France in Albrecht Thaer³ and Mathieu de Dombasle respectively. In their opinion inorganic salts, the significance of which could not be absolutely denied, acted merely as

¹ For the literature consulted on this subject (in addition to the books and papers cited below), see the *Geschichte der Botanik*, by J. Sachs; *Lehrbuch der Pflanzenphysiologie*, by Pfeffer; *Lehrbuch der Agrikulturchemie*, by W. Knop; *Chimie et Physiologie appliquées à l'Agriculture*, etc., by L. Grandeau; and *Neues Handwörterbuch der Chemie*, vol. ii. pp. 119 and 1012.

² Nicolas Leblanc pointed out the importance of salts of ammonia in this respect so long ago as at the end of last century.

³ Cf. his work, *Grundsätze der rationellen Landwirthschaft* ("Principles of Rational Husbandry"). Even Saussure, the originator of the doctrine of plant nutrition, fell into the humus theory error.

stimulants, and not as if they were essential to the growth of the plant.¹ Indeed, Thaer held that the formation (*i.e.* creation) of earths in plants through their vital forces was possible. In this assumption he followed the opinion of Schrader, who so early as the year 1800 imagined that he had proved by actual experiments the generation of the ash-constituents of plants by the vital forces.²

Liebig put an abrupt end to this period of unscientific attempts at explaining the process of plant nutrition by his critical demolition of the humus doctrine. Taking his stand upon a large number of investigations carried out by himself and his pupils, in conjunction with earlier work done by others, he brought out in 1840 his book, *Die Chemie in ihrer Anwendung auf Agrikultur und Physiologie*³ ("Chemistry in its Application to Agriculture and Physiology"); in this he did battle with the arbitrary axioms of the humus theory, and completely undermined the foundations of the latter, hitherto looked upon as so secure. The following sentences by Liebig constitute the quintessence of his doctrine; they already contain the complete programme of the agricultural chemistry which has been created since that time. "The nutritive materials of all green plants are inorganic substances." . . . "Plants live upon carbonic acid, ammonia (nitric acid), water, phosphoric acid, sulphuric acid, silicic acid, lime, magnesia, potash, and iron; many of them also require common salt." . . . "Dung, the excrements of the lower animals and of man, does not act upon plant life through (the direct assimilation of) its organic elements, but indirectly through the products of its decomposition- and putrefaction-processes, *i.e.* by the transformation of its carbon into carbonic acid, and of its nitrogen into ammonia or nitric acid. Organic manure, which consists of portions or *débris* of plants and animals, may be

¹ Several writers have ascribed to Sprengel, who achieved so much for botany, the merit of having proved the indispensability of the ash-constituents for plants, but this is incorrect.

² This erroneous view was first combated upon good grounds by Saussure, and then by Davy.

³ The incitement to this work came from the British Association for the Advancement of Science.

replaced by the inorganic compounds into which it breaks up in the ground." From these axioms Liebig drew the all-important conclusion that the soil must be replenished with whatever constituents have been withdrawn from it by the culture of plants, if its exhaustion is to be provided against.

In the further development of this pregnant doctrine, whose victory over the old system was soon complete, distinguished pupils of Liebig as well as he himself took part. Indeed, almost all agricultural chemists since that time have come either indirectly or directly from Liebig's school. Boussingault¹ strove independently after similar goals; the services which he rendered in carrying out researches on the nutrition of plants by new methods must be emphasised here.

Definite researches were first made in order to explain the chemical conditions existing in the soil, from which plants are supplied with their purely mineral constituents. These included the investigation of the processes involved in the weathering of rocks, through which soil is produced. Liebig, Boussingault, Déherain, Dietrich, and others showed by their investigations what were the parts played by the active agents here,—water, carbonic acid, and oxygen; they also came to the conclusion that free nitrogen as such was not directly assimilated by plants, but this view has recently been overthrown by the work of Hellriegel and others. It is only after rocks have been "weathered" that the inorganic substances necessary for the nutrition of plants are brought into such a condition that they can be assimilated by these. The valuable experimental work done by E. Wolff, Henneberg, W. Knop, F. Stohmann, Zoller, Lehmann, and Nobbe, among others, upon the composition

¹ J. B. Boussingault, who was born in 1802 and who died in 1886, first became known through his adventurous journeys in South America, where he turned his catholic knowledge to brilliant account. After returning to France he devoted himself more and more to agricultural-chemical questions, which he treated partly in experimental researches, and partly in his detailed works, *Economie Rurale*; *Agronomie*; and *Chimie Agricole et Physiologie* (1864).

of different soils must be mentioned here, and also the closely allied experiments by them on the nutrition of plants in sterile soils and in solutions of salts,—*dry culture* and *water culture*. These methods have served to solve the most important questions regarding plant nutrition.

Those researches all went to prove that the same substances as are found in the ashes of plants are the true nutrients of the latter, and are absolutely indispensable to them. But they did more than this, in showing the significance—indeed, the determining influence—as regards nutrition, not merely of the nature of the nutritive materials contained in the soil, but also of the form in which these are present, and of their action upon the other constituents.

The earliest series of experiments on the absorption by different soils of the mineral constituents which serve as food for plants was due to Liebig, while similar work by Henneberg and Stohmann, Peters, Knop, Zöller, etc., must also be recorded; these observations were likewise of great importance for the explanation of the action of manures.

Notwithstanding, however, that an immense number of new facts have been brought to light through these and other labours, the fundamental principles of Liebig's doctrine have undergone no alteration since he first gave them to the world in his pioneering work of 1840. He clearly recognised in all its broad features how plants draw their nutriment from the constituents of the air and the soil. Upon this he based his doctrines of rational husbandry, which have already borne the richest fruits, and in the elaboration of which scientific and practical men are still engaged.

Development of Phyto-Chemistry.

After the importance of various inorganic substances for the life of plants had come to be recognised, the pressing question arose for physiologico-chemical investigation—How and in what phases is the formation of organic substances from carbonic acid, ammonia, and water consummated?

The problem to be solved here consists in isolating the chemical compounds present in the various organs of plants, and in establishing their physiologico-chemical relations to one another,—a magnificent task, and one which has already occupied many eminent investigators.

The conversion of carbonic acid into organic compounds under the influence of water and light, the process of the assimilation of carbon, which was already correctly apprehended in its main outlines by Saussure,¹ has naturally formed the subject of numerous investigations. Thus recent researches by Lommel, Pfeffer, N. J. C. Müller, Engelmann, and others have elucidated the nature of the light rays which are active here. Much valuable work too has been done upon chlorophyll, although the opinions of men like Sachs, Pringsheim, etc., differ as to the part which this substance plays in the assimilation of carbon. Speculation has still, however, free play in the answering of the question—What is the organic compound which is in the first instance produced from the carbonic acid; for, according to the experiments of J. Sachs, the first visible and palpable product is starch, a complicated compound whose chemical constitution is as yet unknown.

The multifarious substances produced by plants have been the objects of ardent investigation, more especially since the stimulus which was given to the subject by Liebig's work; the chemistry of plant life has been developed alongside of that of animal life, particularly since the close of the forties. Reference must be made here, in passing, to

¹ Cf. his *Recherches Chimiques sur la Végétation* (1804). Previous to this Ingen-Houss had observed the assimilation of carbonic acid and water by the leaves of plants, but, being enchained by the phlogistic theory, had not perceived that the oxygen thereby liberated came from this carbonic acid. The above relation was first made clear by Senebier, and became a certainty after Saussure's masterly researches, through which the balance between the substances absorbed and eliminated was approximately ascertained. Ingen-Houss, too, and Saussure still more definitely, recognised that the converse of this assimilation process (*i.e.* a breathing in of oxygen and giving out of carbonic acid) goes on in various parts of plants. Saussure and, after him, Dutrochet and others further observed the evolution of heat which accompanies respiration in plants, and thus established a noteworthy analogy between the processes in the vegetable and animal organisms.

Rochleder's researches (so important from the chemical point of view) in this field, upon caffeine, various glucosides, tannic acids, and other vegetable products. The attention of phyto-chemists has been directed in a special degree to the nitrogenous compounds which are formed in plants, *i.e.* to the albumens in the first instance, and then to the compounds produced by the breaking up of these. After Mulder had pointed out the similarity of the former to animal albumen, they were investigated by Liebig and his pupils, and have, more particularly, formed the subject of excellent work by Ritt-hausen during recent years. The hope that conclusions might be arrived at with regard to the constitution of the albumens from the nature of their decomposition-products, more especially from the amido-acids like leucine, asparagine, glutamic acid, etc., has not indeed been realised; but, from the point of view of vegetable physiology, the researches on the nitrogenous compounds which are formed during the germination of seeds and other processes have furnished much valuable preparatory work for the future development of that branch of the science.¹

There are, besides, many other vegetable products containing nitrogen which have occupied the attention of chemists as well as of physiologists, *e.g.* various glucosides such as myronic acid and amygdalin, and, in particular, the great class of the alkaloids,—compounds whose importance for chemistry has already been treated of.

The carbohydrates, in their signification for the life of plants have likewise been much investigated, with regard both to the conversion of some of them into others by chemical means, and to their physiological modes of formation; but here again the necessary link is often wanting between particular products. The reader is referred to the pioneering investigations of Brücke, Nägeli, Sachs, and others upon starch, and upon the connection which exists between the formation of the latter and the activity of chlorophyll; to the numberless researches on the sugar varieties, especially dextrose and cane sugar, the occurrence

¹ Cf. the investigations of E. Schulze and others.

of the latter in beetroot and its technical production from the same having created a chemistry of its own: and to the laborious work which has been and still is being done with the object of clearing up the chemical nature of the glucosides and their peculiar behaviour to ferments. The most important of the investigations upon vegetable fats, ethereal oils, and various other (vegetable) compounds belong in the main to organic chemistry proper, and have been referred to under the history of this.

Development of Zoo-Chemistry.¹

The physiological chemistry of the animal body, zoo-chemistry, has made extraordinary progress since the early investigations of Fourcroy and Vanquelin, Chevreul, Berzelius, and others were made. From the examination of the chemical constituents of animal organs, secretions, etc., an advance was made to the infinitely more difficult problem—Under what conditions are those substances formed in the organism, and what are their relations to one another? From the chemical investigations which arose from this, animal physiology was first constituted into the science as we now know it. And this applies in a special degree to the important question of nutrition, and, speaking generally, to the modern views of the metabolic processes of the animal body. Chemical investigation has thus been the means of dispelling the obscurity in which so many erroneous views grew and flourished.

Since the publication of the above-mentioned researches, the most distinguished physiologists and chemists have co-operated in the development of zoo-chemistry, in so far as this has aimed at a knowledge of the substances of which the animal body is composed. From the large number of excellent investigations of this kind, only one or two can be touched upon here. Reference must first be made to

¹ The numerous sources of physiologico-chemical investigations are to be found in Hoppe-Seyler's *Lehrbuch der physiologischen Chemie*. Only in a few instances have direct references been given here.

the work of v. Bibra, Mulder, Frémy, and Heintz upon the constituents of bones, through which the true composition of these was established. The question as to the nature of the albumens has given rise to many important researches, especially since Mulder first proved the presence of compounds of this kind in plants, and Liebig and his pupils strove to get at their composition; but they have not as yet led to a knowledge of the true constitution of these bodies. Among those who have worked at this subject may be mentioned A. Schmidt, Graham, Brücke, Johnson, Hoppe-Seyler, Kühne, Hammarsten, Lehmann, Schützenberger, Nencki, Drechsel, and Harnack. To the physiologist the question of the behaviour of albumen in the animal body (in particular, the changes which it undergoes during digestion, etc.) is of more importance than its rational composition. Some investigations will be referred to later on, in which an answer to such physiological questions is attempted.

The most important of the researches which led gradually but ultimately to a true explanation of the composition of fats have already been spoken of.¹ The part played by fats in metabolism has only been satisfactorily worked out of recent years, and the same remark applies to the carbohydrates. The pathological occurrence of those substances has also given much occupation to chemists, who, by furnishing definite tests for sugar, albumen, etc., have in many cases lightened, and even rendered possible, the diagnosis of a disease by the physician.

As in all the other branches of chemistry, so too in physiological and pathological, have special methods of a zoo-chemical analysis gradually developed themselves and become indispensable.

The investigations which have been made with the object of elucidating the chemical processes which go on in the animal organism, and therewith the processes which condition or accompany life, are almost innumerable. Our present knowledge of the various animal fluids which take part in such processes has only been attained to by the most

¹ Cf. p. 405.

arduous labours. To mention but one or two of these, reference may be made in the first instance to the more important of the researches on the secretions which promote digestion. The classical investigations of C. Ludwig, Brücke, and Cl. Bernard proved that the secretions from the glands were to be looked upon as resulting from essentially chemical processes. The importance of the saliva for digestion was also shown by its chemical investigation; Leuchs, in 1831, discovered the ferment *ptyalin* which saliva contains, and which has the power of transforming starch into sugar, and the chemistry of the saliva has since been materially advanced by the later work of O. Nasse, C. Ludwig, Brücke, Herter, and others.

Many scientists of repute have occupied themselves with the investigation of the gastric juice: thus the work of C. Schmidt, Bidder, Beaumont, Frerichs, Lehmann, v. Wittich, and others has resulted in establishing the composition of this secretion, and also the peculiar nature of *pepsin*, the ferment which it contains. The excessively important part played by the latter in the digestion of the albumens, which are thereby converted into *peptones*, has been mainly arrived at through the labours of Lehmann, Hofmeister, Hemminger, and, quite recently, Kühne and Chittenden.

Our knowledge of the pancreatic fluid and of its powerful influence on the digestive process, which is due to the presence in it of particular ferments, we owe to W. Kühne, Hufner, and others.

The chemistry of the bile, lastly, which originated with Strecker's memorable work¹ on the bile-acids and their decomposition-products, has been subsequently extended by Städeler, Frerichs, Gorup-Besanez, Maly, etc.

The present knowledge of the chemical composition of the blood and of its various constituents (so difficult to separate from one another), together with the chemical behaviour of these, is the outcome of an infinite number of laborious investigations; and it is still very far from being complete. Reference must be made here to the pioneering work of Al.

¹ *Ann. Chem.*, vols. lxi. lxv. lxvii. and lxx.

Schmidt upon the causes of the coagulation of blood; to that of C. Schmidt, Hoppe-Seyler, Hüfner, Preyer, and others on hæmoglobin and oxy-hæmoglobin, and the behaviour of these to gases, and also to the successful application of the spectroscope here; further, to the memorable researches which finally established the composition of the blood-gases, and especially the difference existing between arterial and venous blood in this respect. The services rendered by C. Ludwig deserve to be particularly emphasised, the investigations which he has carried out along with his pupils since 1858 far surpassing the earlier ones of Magnus and of L. Meyer in accuracy.

The numerous researches, by means of which the quantitative relations between the air inhaled and exhaled by animals were exactly determined, have been of the utmost value for a knowledge of the metabolic processes of the animal body. We have only to recall here the experiments carried out on a large scale by Pettenkofer and by Regnault and Reiset since the year 1862, and the important observations by C. Ludwig, and by Pettenkofer and Voit, on the effect of muscular exertion upon the consumption of oxygen and the production of carbonic acid.

The exceedingly numerous researches on the substances which occur in blood serum, on the inorganic constituents of blood, and on the pathological changes which the latter undergoes, cannot be entered upon here.

Milk has been the subject of frequent investigation ever since Chevreul, Lerch, Heintz, and others established its principal constituents. Much attention has been paid in more recent work to the process of coagulation, to the changes which milk undergoes in the organism, to the nature of the albuminous compounds which it contains, and so on; witness the important researches on the subject by Soxhlet, Hammarsten, Lehmann, Hoppe-Seyler, Struve, etc.

Much excellent chemical and physiological work has been done upon urine—the secretion of the kidneys. Take, for instance, the observations on the artificial production of urea, of such moment from a chemical point of view.

and those upon uric acid and its manifold transformation-products, the synthesis of some of which has already been achieved.¹ Then there are, too, the important physiological and pathological investigations by Liebig, Voit, Bischoff, and others on the separation of urea in its bearing upon metabolism; the researches on the formation of hippuric acid, by Wöhler, Liebig, Dessaignes, and Meissner; on that of the phenol-sulphuric acids by Baumann; on the formation of sugar, albumen, glycuronic acid, cynurenic acid (an oxy-quinoline-carboxylic acid), and indole; and on the separation of all of those substances just named in the urine.

The explanation of the manner of origin of these and other substances, which are partly found under normal conditions and partly under pathological, has long been recognised as constituting an important problem of physiological chemistry. From the results of a large number of observations, a systematic method of analysing urine has gradually become developed, and this daily stands the practising physician in good stead; for, from the occurrence or accumulation of certain substances in the urine, the latter can recognise particular diseases with greater precision than by any other sign.

The work which has been done upon the chemical composition of flesh,² a subject to which peculiar difficulties are attached, can only be briefly referred to. Liebig's classical researches on "the constituents of the fluids of flesh,"³ and the nearly allied ones of his pupils Schlossberger, Scherer, Strecker, and Städele, prepared the way for later and even more ambitious labours; we would refer here to the observations of Helmholtz, Ranke, Brücke, and others on the effect of muscular action upon the chemical processes which go on in muscle-substance,—observations to which the first incitement may have been given by Liebig's ingenious and far-reaching speculations. The important part which glycogen plays in these, as well as in other processes (*e.g.* the processes

¹ Cf. *The History of Organic Chemistry*, p. 424.

² Cf. (*e.g.*) Falk's book, *Das Fleisch* (1880).

³ *Ann. Chem.*, vol. lxii. p. 257.

of the liver), was arrived at through the admirable work of Brücke, Cl. Bernard, Külz, v. Mering, etc.

From the rich material of facts relating to the chemical composition and physiological importance of particular parts of the animal organism, which, as shortly stated above, have thus been accumulated, the views regarding the metabolic processes of the animal body have been developed, and indeed completed, in certain of their details. The establishing of the laws which govern the nutrition of animals was long ago felt to be of the first importance. And here again Liebig gave the powerful impulse to the first, even if incomplete, solution of this question from the chemical standpoint.

The service which he rendered with regard to the development of the doctrine of metabolism appears especially great when one recalls to mind how erroneous were the opinions of physiologists respecting the chemical processes going on in the animal body, before he set forth his views on nutrition and other physiological processes in his standard work, *Die Thierchemie oder die Organische Chemie in ihrer Anwendung auf Physiologie und Pathologie* (1842), ("Animal Chemistry, or Organic Chemistry in its Application to Physiology and Pathology"). The most eminent physiologists of that time, Tiedemann, Burdach, and others, were by no means fully convinced of the necessity of chemistry for their science; to explain the processes in the organism they had recourse to "vital forces," many of them indeed flatly refusing the aid of chemistry. It was left to Liebig to form a truer estimate of the problems of physiology and of the means to be used in solving these; the opinion which he expressed—that it must adopt the methods of physics and chemistry—coming as this did with the full weight of his authority, was quickly taken to heart. And what a change came over physiology in consequence!

The powerful influence exercised by Liebig on the development of the doctrine of metabolism has already been frequently referred to. But a short *résumé* may be given here of the main conclusions of his comprehensive work and ingenious speculations. He recognised the various importance

of different nutritives for the animal body, in so far that he defined the albumenoids as *plastic* compounds, which served mainly for the building up of the tissues and as the source of muscular power, and the fats and carbohydrates as *respiratory* compounds, which went for the most part to produce the animal heat. It was in fact he who first drew sharp distinctions between nutritive substances among themselves, and between these and other substances which, while not directly nutrient, bring about metabolic changes in the organism.¹ And he also successfully determined the relative values of the former by direct experiment.

The potent effect of Liebig's ideas respecting nutrition and metabolism showed itself during the succeeding years in the splendid work which was done by Bidder and Schmidt, Bischoff, Voit, Pettenkofer, Frerichs, and others, as the result of his stimulus. By the aid of improved methods and, especially, by the use of larger respiration-apparatus, Liebig's views were subjected to a sharper scrutiny, and thus underwent many corrections, more particularly with respect to the *role* of albumen and to the formation of fat. But in all essential points he was right. To the elucidation of the functions of particular nutritives in the animal body, the classical researches² of Voit and Pettenkofer, together with those of their pupils (among whom were Ranke, Forster, Rubner, Falek, and Fr. Hofmann) upon nutrition, and therefore upon metabolism, have contributed in an especial degree.

The aims of the above branch of physiological chemistry are so intimately connected with those of hygiene that the two overlap at this point. Hygiene may indeed be looked upon as a branch of chemistry, having found in the latter science the most powerful of all aids to her development. Reference has already been made in the history of analytical chemistry³ to the continuous improvement in the methods of analysis of foods and drinks, a point of such immense importance to the community in general.

¹ *Genussmittel*.

² Most of these were published in the *Zeitschrift für Biologie*.

³ Cf. p. 371.

*Fermentation ; Putrefaction.*¹

The various processes by which ferments are set in action, and by which their action is conditioned, have now attained to such a supreme importance for hygiene, and for physiology as a whole, that a few words must be said here with regard to the development of our knowledge of the processes of fermentation and putrefaction during recent years.

It is a long time since the vinous fermentation first attracted the attention of chemists, but Lavoisier was the earliest to recognise that the two main products resulting from it—alcohol and carbonic acid—came from the sugar present; at the same time he attempted to work out the quantitative relations between the latter and the two former compounds. As to the reason for the breaking up of sugar in the presence of yeast, no views were expressed at that time which were at all tenable. Before it was known that yeast consisted of living cells, Liebig's mechanical-chemical theory of fermentation² gained many adherents. This theory, which was propounded in the year 1839, attempted to explain alcoholic fermentation and other similar processes from one common point of view. Liebig here regarded ferments in general as easily decomposable bodies, from which the stimulus to the decomposition of fermentable substances proceeded. This view recalls that which Stahl and Willis had brought forward long before, for they also assumed a transference of the motion of fermenting particles to a large number of other ones. Some investigators had contented themselves with attributing to yeast a "catalytic" action, but this simply meant the employment of a word to cover their ignorance of the subject.

Then came in 1836 the important discovery, made simultaneously and independently by Cagniard de Latour,

¹ For the literature consulted here, see the articles "*Fermente*" and "*Gärung*" in the *Handwörterbuch der Chemie*; A. Mayer, *Lehrbuch der Gährungschemie*; and Schützenberger, *Gährungserscheinungen*.

² Cf. *Ann. Chem.*, vol. xxx. pp. 250 and 363.

Schwann, and Kützing, that yeast consists of low organisms which are self-propagating. The subsequent comprehensive researches of Pasteur¹ entirely confirmed the correctness of these observations. From all this the vitalistic theory of fermentation followed as a necessary consequence; according to this theory the decomposition of the sugar is dependent upon the vitality and consequent activity of the yeast fungus.

Other processes of fermentation were now investigated from the standpoint thus obtained, with the result that low organisms were found to be the cause of the action in their cases also. We would refer here to the splendid researches of Pasteur upon the acetic and lactic fermentations, of equal importance physiologically and chemically; to the discovery of the particular fission fungi which give rise to various fermentations; and to the work of Rees, de Bary, Brefeld, A. Mayer, Fitz, and others, the object of which was to elucidate the conditions of the life and especially of the nutrition of organised ferments (more particularly yeast and its connection with fermentation), and also the products of these latter.²

Much vigorous discussion has ultimately led to agreement upon the most important of the disputed points, with regard to which the views of different workers were formerly far apart. A twofold growth of the yeast cells is now established, viz. (1) a growth in presence of oxygen, which is not followed by fermentation, and (2) one in absence of oxygen, through which fermentation is produced.³

The difference between organised and unorganised ferments, the latter of which are termed *enzymes*, came to be

¹ Cf. his large works, *Études sur la Bière*,—*sur le Vin*,—*sur le Vinaigre*.

² C. Schmidt found succinic acid, and Pasteur glycerine, among the products of the vinous fermentation.

³ Liebig maintained an antagonistic attitude to the vitalistic theory of fermentation; he did not indeed contest the organised nature of yeast, but would not acknowledge that the latter itself gave rise to fermentation through its life processes. Instead of this he assumed in yeast the presence of an albuminous ferment, which, on the death of the former, he imagined to bring about the decomposition of the sugar into alcohol and carbonic acid.

clearly recognised, this being mainly due to Pasteur's work. The extraordinarily important functions of these unorganised ferments in the animal and vegetable organisms has led physiologists and chemists of the highest eminence to devote their close attention to the subject, but as yet no satisfactory theory of the action of such ferments has been brought forward; in conjunction with this, reference must be made here to the work of Nasse, Hüfner, Traube, Hoppe-Seyler, Nencki, Al. Schmidt, and Wurtz.

The phenomena of putrefaction, which were classed by Liebig in the same category with the processes of fermentation (both being brought about, in his view, by similar mechanical-chemical causes), acquired a heightened physiological interest after it was perceived that they were connected with the presence of certain peculiar organisms. Here again the researches of Pasteur and also of Nencki, Hoppe-Seyler, etc., stand out pre-eminent. The chemical examination of the products of putrefaction has led to remarkable results, which have also a high importance for the chemist. Most interest has been centred in the nitrogenous compounds which originate from the decomposition of animal albuminous substances by putrefaction; thus we would recall here the discovery of various amido-acids, of indole and its homologues, and, particularly, of the so-called *ptomaines*.¹ The formation of these latter, which have also been called *corpses alkaloids*, because of their likeness to the alkaloids from plants, is of the first importance to the forensic chemist,² seeing that cases have occurred in which the ptomaines have been confounded with the true alkaloids, on account of similarity in reaction. The Italian toxicologist Selmi was the first to clearly recognise the important rôle, from a forensic point of view, of these putrefaction bases, and he it was who gave them the generic name by which they are now known,—the ptomaines.

¹ For a historical notice of these peculiar compounds, cf. Beckurts' *Ausmittlung giftiger alkalioide* ("Detection of Poisonous Alkaloids"), (*Archiv Pharm.* for 1886, p. 1041); also Armstrong, *Journ. Chem. Ind.*, vol. vi. p. 482.

² Cf. p. 371.

In addition to Schmi—Otto, Husemann, Dragendorff, Kobert, Brieger, and others have rendered good service in extending our knowledge of these substances. Brieger, in especial, and also Nencki, Étard, Gautier, Guareschi, and Mosso have succeeded in characterising certain ptomaines chemically. The constitution of some of them has been quite recently established, witness the beautiful synthesis of cadaverine.¹

The Relation of Chemistry to Pathology and Therapeutics.

The phenomena of putrefaction possess the highest interest for pathologists, because such processes lie at the root of many diseases. An increasing knowledge of the causes of these processes has thus resulted in the establishment of a close connection between chemistry and pathology, the former having now become indispensable to the latter. And this necessity for chemistry has shown itself not merely in the investigation of the products of putrefaction: through its means the more delicate tests for the recognition and distinction of bacteria have been elaborated, and it has thus been instrumental in helping to found the new science of *bacteriology*.

Above all, it has been reserved for chemistry to direct the attention of physicians to remedies for counteracting the pathological processes induced by micro-organisms. Only a passing reference can be made here to the wonderful results which have been achieved in medicine and surgery, and also on the large scale in the preservation of food and drink, by the use of antiseptics. One is probably not wrong in assuming that the old practices of smoking flesh and of dipping wood into tar drew attention to the carbolic acid which the latter contains, and the antiseptic action of which has now found such world-wide application in Lister's method of treating wounds. The discovery of the anti-fermentation and anti-putrefaction powers of salicylic acid by H. Kolbe originated in the idea that this compound

¹ Cf. p. 434.

tended to break up into carbolic and carbonic acids in its passage through the organism. The assumption made by various investigators—that antiferments and antiseptics act by precipitating or chemically altering the readily decomposable albuminous substances—explains the rôle of these in a sufficiently satisfactory manner: for, when those bodies are got rid of, the ferments are deprived of their necessary nutriment.

The nearly allied question of the great benefit which chemistry has conferred upon medicine by enlarging its stock of remedies can only be touched upon very briefly, as any detailed treatment of the subject here would overstep the limits of this work. With the history of medicine in the earlier ages the conditions were quite otherwise; for, in the iatro-chemical as well as in the phlogistic periods the latter was in the main conjoined with the history of chemistry, whereas now chemical investigation pursues totally distinct aims.

To mention only one or two of the specially important services which chemistry has rendered to medical science, take the introduction of narcotics and anæsthetics—chloroform, ether, nitrous oxide, chloral, bromide of potassium, sulfonal, etc. Recall, too, the success with which naturally occurring medicaments have been replaced by others artificially prepared, *e.g.* quinine by antipyretic remedies like salicylic acid, acetanilide, antipyrine, phenacetine, etc. It has already been shown¹ how, with the acquisition of the knowledge that the alkaloids are derivatives of pyridine or quinoline, a firmer foothold was gained for the artificial formation of those natural products—an object which has been striven after for so long.

The Relation of Chemistry to Pharmacy.

With the rapid augmentation of the medical treasury, the problems which confront the pharmacist have likewise grown in a very high degree. If the latter is to do jus-

¹ Cf. p. 434.

tice to the demands which are made upon him, he must be equipped with a catholic and thorough knowledge of chemistry. The development of pharmaceutical chemistry in recent years is for the most part concurrent with that of particular branches of the pure and applied science. The discoveries of inorganic and organic compounds which have proved of importance for pharmacy have likewise been of great value for chemistry itself.¹

In the domain of analytical chemistry we see the assiduous and scientifically educated pharmacist striving after similar aims with the chemist. The former ought to have a thorough knowledge and be master of the approved analytical methods which are required for the testing and examining of official drugs as well as of food and drink, and also of those employed in legal cases where chemistry comes into play.²

Pharmaceutical chemistry is in fact connected in the closest manner with pure chemistry, for both have the same foundations. If we would convince ourselves of this, we have but to look through the numerous recent text-books of the former, to perceive that in contents and arrangement they are much the same as those of the pure science. So long ago as 1844 H. Kopp³ expressed himself pertinently on the subject as follows: "Since the end of last century pharmaceutical chemistry has deviated more and more from the direction which it still followed during the earlier decades of the latter, when it merely borrowed from the investigations of scientific chemistry those results which had a bearing upon the preparation of medicines. It became more and more nearly allied to purely scientific chemistry: pharmaceutical text-books, which formerly were mere collections of empirical recipes, came to have a genuine scientific character, while the journals originally brought out for pharmacy became important miscellanies for pure chemistry."

At the close of last century and beginning of this one

¹ Cf. *The History of Pure Chemistry*, p. 373 et seq.

² Cf. p. 371.

³ *Geschichte der Chemie*, vol. ii. p. 119.

the relation of chemistry to pharmacy was, however, different from what it is now. Then the latter was an *Alma Mater* for the former, whereas now these positions are exactly reversed: pharmacy enjoys to-day the fruits of a highly developed chemistry. In earlier times the study of pharmacy was in truth the only road to that of pure chemistry, and this is why the most eminent chemists from the end of last century until well on in this one came from the pharmaceutical school. We have but to recall here the names of Scheele, Rouelle, Klaproth, Vauquelin, Liebig, H. Rose, and many others.

The pharmaceutical institutes which began to spring into life at the close of last century were of great value for the education of chemists who wished at the same time to become pharmacists, for in these any young man who was anxious to learn received a course of systematic instruction. The Trommsdorff Institute in Erfurt, founded in 1795, deserves especial mention in this connection. And good text-books of pharmacy were not wanting then either, *e.g.* Hagen's *Apothekerkunst* ("The Art of Pharmacy," 1778), Götting's *Handbuch der Pharmazie* (1800), Hermbstädt's, Trommsdorff's, Westrumb's, and Buchholz's text-books, etc.

A historical account of how pharmacy proper has developed along with chemistry during the present century is unnecessary here, for the reasons already given.

HISTORY OF TECHNICAL CHEMISTRY DURING THE LAST HUNDRED YEARS.¹

The immense development of large chemical industries and, in fact, of all the branches of chemical technology during the present century is the natural consequence of the great advances in chemical knowledge, and the rational application of these to technical processes. The light of scientific research has thus been shed upon the latter, and new branches of industry have been grounded upon exact investigations. The history of technical chemistry offers a continuous series of examples of this beneficial action of theory upon practice. On the other hand, numerous questions have arisen in the course of technical working which have given rise to investigations of the highest value for pure chemistry.

The great advances which have been made in chemical technology only became possible with the development of analytical chemistry, which allowed of a clear insight into the composition of the original, intermediate, and final products of technical processes. Since the beginning of this century methods of research have gradually become more perfect, methods which more and more meet the requirements of the technical chemist, and which have constituted and still constitute the most important aids to the development of chemical industry. Many of these methods have already been referred to in the history of analytical chemistry,² but the reader may also be reminded at this point of their use with respect to the wants of everyday life. The testing and examination of articles of food and drink are now carried on

¹ For the literature on the subject, see *Wagner's Jahresberichte* and his *Lehrbuch der Technologie* ("Annual Reports" and "Text-Book of Technology"); A. W. Hofmann's *Bericht über die Entwicklung der Chemischen Industrie* ("Report on the Development of Chemical Industries," etc., 1875-77); Karmarsch, *Geschichte der Technologie*, etc. ("History of Technology," etc.)

² Cf. pp. 364, 366, and 371.

in a very large number of laboratories, the methods employed here having been elaborated from purely chemical investigations. This applies in a special degree to the analysis of water, which is of such enormous importance alike from a hygienic and an industrial point of view. We have only to think how necessary it is to establish the chemical composition of a water before employing it for any manufactures; and the various processes of purification, too, to which it has to be subjected, before it can be used for many purposes, are based upon rational chemical researches and observations. Another benefit which water analysis has conferred upon the community at large consists in its having rendered possible the artificial production of mineral waters, and thus called a flourishing industry into life; the great services rendered in respect to this by F. A. Struve (since 1820) deserve to be recalled here.

In the following pages mention will be chiefly made of such work as has either led to the introduction of important novelties into chemical technology or to the opening up of new branches of the latter.

It is hardly possible to estimate the benefit to the national wellbeing which has accrued, more especially in England, Germany, France, Switzerland, and Belgium, from the growth of chemical industries. Take, for example, the coal-tar colour manufacture in Germany, which has arisen upon foundations of purely scientific work, and the alkali and sulphuric acid manufactures in Great Britain. The former illustrates in the most perfect manner the principle of the refinement of matter, a troublesome and almost worthless waste product—tar—being now worked up by chemical processes into a vast number of valuable substances. And the same applies in greater or less degree to the great chemical industries of all the countries mentioned above; in every case men are striving to bring individual chemical processes to the highest state of perfection by utilising all waste products. The soda industry of to-day offers a specially good instance of this, for in it we find two competing processes both flourishing, simply because they have called to their aid every means of rational

chemical investigation. There is indeed hardly any branch of chemical manufacture of which the same may not more or less be said.

Reference may also be made here to the development of technical instruction, which has of course contributed immensely to the advancement of chemical industries. Technical schools and colleges belong for the most part to the present century. The earliest of those on the continent of Europe were the *École Polytechnique* of France, founded in 1795, the Vienna Polytechnic Institute (1815), and the Berlin Technical College (1820). As every one knows, Great Britain is by no means so well equipped with technical schools and colleges as many of its neighbours on the continent, but public opinion is now becoming thoroughly awakened on the subject, and the want is being rapidly supplied. The chemical laboratories of the above and other similar institutions have gone on acquiring more and more importance as aids to the furtherance of chemical manufactures.

The literature on technical chemistry has sprung from insignificant beginnings. Hermstadt's works on *Dyeing*, *Bleaching*, *Distilling*, etc., which were published in and after the year 1820, deserve mention on account of their value at that time. During the last fifty years immense strides have been made in this respect, as is witnessed (*e.g.*) by the excellent encyclopedias of Precht and Karmarsch, Muspratt-Stohmann-Kerl, Bolley, Ure, and Watts, and also by the text-books upon chemical technology, among others those of Dumas, Payen, Knapp, Wagner, and Ost, in which the results of theory and practice are given together. In addition to these, the weekly and monthly journals, among which Dingler's *Polytechnisches Journal*, Wagner's *Jahresberichte*, and the recently established *Journal of the Society of Chemical Industry* may be named, supply us with information upon the results of current chemico-technical investigation. By such means the closest connection between chemical industry and the pure science is permanently maintained.

The Progress of Metallurgy.

Although the production of iron and steel, as carried on in the phlogistic period, gave rise to chemical work through which the mutual relations of cast-iron, wrought-iron, and steel were in some measure explained, there still remained a variety of problems in connection with these to be solved at a later date. The improvement of analytical methods rendered it possible to detect and estimate the various impurities in iron,—silicon, phosphorus, sulphur, arsenic, etc.,—and at the same time to recognise their influence in modifying the properties of the latter. The blast furnace process was explained by the excellent investigations of Gruner, Tunner, L. Rinman, and others, the analyses of the furnace gases by Bunsen¹ and Playfair² aiding in an especial degree towards the elucidation of the reactions which go on in it. The determination of the composition of pig-iron—the proof that a chemical compound of iron and carbon exists—was also conducive to the establishment of a theory of the blast furnace process. The Bessemer process for the production of steel (1856) was the result of the clear perception of the connection existing between iron and steel, while the chemical investigation of the products which are formed during its various stages greatly aided its development.

The Thomas-Gilchrist process for dephosphorising iron, which has been such a wonderful success, must be recalled here. Light was shed upon the theory of it by various analytical researches, *cf.* those of Finkener;³ while, on the other hand, scientific experiments by A. Frank, P. Wagner, and others have led to the utilisation of the phosphoric acid which accumulates in the slag produced in the process—the *Thomas* slag,—so that this latter has now become an artificial manure of the first importance. The ingenious application of the spectroscope to the examination of the Bessemer flame, whereby the end point of the

¹ Cf. *Pogg. Ann.*, vol. xlvi. p. 193.

² *Brit. Assoc. Reports* for 1845, etc.

³ Cf. *Wagner's Jahresber.* for 1883, p. 136.

reaction can be clearly distinguished, must also be referred to.¹

As an example of the utilisation of by-products, we may take the successful working up into iron of iron pyrites from which all the sulphur possible has been driven off.² The desire to waste no material of any value is also shown in the process of manufacturing copper from pyrites whose sulphur has already been utilised,—a process elaborated from chemical researches.

The metallurgy of nickel has developed rapidly since German silver (already long known to the Chinese, having been used by them for making a number of articles, began to be prepared upon a rational system, and especially since its employment as an ingredient of coins: the German nickel coinage dates from 1873. The recent experiments made by the United States naval authorities point to the use of an alloy of nickel and iron for armour-plating ships of war.

Numerous improvements have been made in respect to the production of silver, among others the Augustin and Ziervogel extraction processes, and the Pattinson and Parkes processes for the desilverisation of lead; while the metallurgy of gold has also been facilitated by the introduction of good methods for separating the latter from other metals, *e.g.* by that of d'Arcet (1802). The most important additions to the technology of platinum were made by Deville in 1852 and Debray in 1857, in the fusion of large quantities of the metal.

The galvano-plastic process, *i.e.* the precipitation upon one metal of a thin layer of another one by means of electricity, has proved itself of great importance. The original observation in this direction was made by de la Rive in 1836, and this was followed by the publication in 1839 by Jacobi, and a little later by Spencer, of the process from which the more perfect electro-metallurgy of to-day has developed itself.

Among the metals which have been isolated during the present century, aluminium has been made available for

¹ Roscoe, *Chem. News* for 1871.

² Gossage, *Chem. Centr.* for 1860, p. 783.

technical purposes by the assiduous and successful labours of H. St. Claire Deville,¹ while the Stassfurt mineral carnallite has proved itself a convenient source from which to prepare magnesium. The methods by which those metals are actually produced have grown out of the work of their discoverers.²

Numerous improvements have also been made in the course of the century in the manufacture of alloys of every kind. Thus, from zinc and copper there have been prepared malleable brass, similor, etc., and from aluminium and copper aluminium bronze, besides a great many alloys of tin, including type metal; this last used to be made from antimony and lead only, but to these tin is now added.

This century has also witnessed the production of all sorts of metallic compounds, among which mineral pigments take a prominent place. The most important improvement in the manufacture of white lead was due to Thénard (1801). Zinc white, which was made on an experimental scale by Courtois so long ago as at the end of last century, was first brought into general repute by Leclaire in 1840, after which it came to be produced on the large scale. The introduction of chrome colours, especially of chrome green and chrome red, both of which are so highly valued for enamelling, belongs to the present century. Schweinfurt green, a double compound of cupric arsenite and acetate, was discovered by Sattler in 1814; it was greatly in vogue for a long time, but has now become superseded by other colours on account of its poisonous nature. The extended application of many metallic salts, formerly prepared in small quantities only, to new purposes (*e.g.* of nitrate of silver in photography, and of the yellow and red prussiates of potash in dyeing) has led to the rise of entirely new branches of manufacture. There are now but few salts of any of the more plentifully occurring metals which have not some use on the large scale; for instance, stannous chloride and various salts of aluminium, iron, and manganese in

¹ *Comptes Rendus*, vols. xxxviii. xxxix. and xl.

² Cf. *The History of Pure Chemistry*.

dyeing, and compounds of mercury, bismuth, antimony, zinc, etc., in pharmacy.

Development of the Great Chemical Industries.

The great chemical industries are a product of our own time, their growth having gone hand in hand with the growth of pure chemistry. The manufactures of sulphuric acid and soda, which may be looked upon as the basis of all the others, and which are naturally followed by those of hydrochloric acid, bleaching powder, chlorate of potash, nitric acid, etc., only attained to their full vigour after the various processes involved had been explained by chemical investigation, and after the most favourable conditions for those processes had been worked out. The introduction of easy methods of analysis into technical industries has also been of the utmost service to these.

Important practical improvements were made in the manufacture of sulphuric acid¹ so early as the beginning of the present century, *e.g.* the amount of steam required was regulated, and the process was made continuous (the latter by Holker). The first attempt to explain this remarkable chemical process of the formation of sulphuric acid from sulphurous acid, air, water, and nitrous gas was made by Clément and Desormes,² who recognised the important part played by the nitric oxide. Later researches by Péligot, and more especially by Cl. Winkler,³ R. Weber,⁴ Lunge, and others, have served to elucidate the reactions which go on between the above mentioned substances, and have therefore been of the utmost value in respect to the manufacture of the acid: they have led, for example, to an exact knowledge of disturbing conditions, which can therefore now be provided against. To Reich is due the merit of having brought the technical process under due control, by his analysis of the chamber gases; and, ever since Cl. Winkler called

¹ Cf. Lunge's *Manufacture of Sulphuric Acid and Alkali*.

² *Ann. de Chimie*, vol. lix. p. 329.

³ Cf. *Hofmann's Bericht*, etc., vol. i. p. 282.

⁴ *Journ. pr. Chemie*, vol. lxxxv. p. 423; *Pogg. Ann.*, vol. cxxvii. p. 54 :

technical gas analysis into life, this has been a regular part of the operation. How essential for the manufacture the observations on the chemical behaviour of nitrous acid to sulphurous and sulphuric have been, is sufficiently evidenced by the introduction of the Gay-Lussac and Glover towers (so called after the inventors) to which they gave rise, and which have made the process into one complete whole.

But if scientific chemistry has thus proved itself so necessary for technical, the latter has likewise done a great deal to advance the former; for many important discoveries, *e.g.* those of selenium and thallium, have been rendered possible by its aid, and researches of high value, such as those of Lunge upon the various stages of the oxidation of nitrogen, have been given rise to by technical questions.

The preparation of sulphuric anhydride from sulphur dioxide and oxygen, which was formerly merely a lecture-room experiment, has been converted into a technical process through the admirable researches of Cl. Winkler,¹ and thus an important reagent has been made available for many branches of chemical industry. Sulphurous acid, whose sole technical application (practically speaking) for a long time was in the manufacture of sulphuric acid, is now used on the large scale for the bleaching of wool and silk, and as a refrigerant, and it has also quite recently found an extensive employment in the production of cellulose. The utilisation of sulphurous acid for those purposes is all the more striking when we remember that in the roasting of sulphides it used often to be allowed to escape into the air, to the great detriment both of human beings and vegetation.

The Soda Industry.—The transformation of common salt, which occurs so abundantly in nature, forms the foundation of this immense industry, whose history commences with the beginning of the present chemical period. Nicolas Leblanc² was the first to succeed in converting salt into

¹ *Wagner's Jahresber.* for 1879 and 1884.

² This remarkable man, who was born at Issoudun (Indre) in 1742 (and not, as usually stated, in 1753), derived no pecuniary benefit from his great

soda, with sodic sulphate as an intermediate product, Malherbe and De la Metherie having some time previously attempted to utilise the latter substance in the same way, but without success. In the year 1823 Muspratt began the erection of his alkali works at Liverpool; his name deserves a foremost place in connection with the development of the soda industry. The advantages which have accrued to the manufacture of soda from chemical investigation are incalculable. The simple analytical methods which supplied the necessary information as to the composition of the raw, intermediate, and final products were and still are of the first importance for the regulation of the technical process. The formation of soda from the sulphate, by fusing the latter with coal and limestone, was ultimately so far explained by exact chemical experiments¹ (after various unsuccessful speculations on the subject by Dumas and others), that a tenable theory of this fusion process could be set up.

Scientific researches have also given rise to numerous important improvements in the soda manufacture, *e.g.* to the beautiful process of Hargreaves and Robinson (by which sulphate of soda is prepared directly without the previous production of sulphuric acid), to the introduction of revolving soda furnaces, and to many processes for utilising and rendering harmless the unpleasant alkali waste. With respect to the last, we would refer here to the work of Guckelberger, Mond, and Schaffner and Helbig, who succeeded in making various laboratory reactions practicable on the large scale. But the greatest advance of all in this direction is the recently published and exceedingly simple process of Chance,² by which nearly all the sulphur in alkali waste can be recovered, and at a very cheap rate; the result of this will probably be to give the Leblanc process a new lease of life, since it will

labours. He died in the utmost poverty in 1806, his death being due to despair. A monument has quite recently been erected at his birthplace to his memory.

¹ Cf. Dubrunfaut in *Wagner's Jahresber.* for 1864, p. 177; Scheurer-Kestner, *ibid.* 1864, p. 173; and, especially, Kolb, *ibid.* 1866, p. 136.

² *Journ. Chem. Ind.*, vol. vii. p. 162.

enable it to compete on more equal terms with the younger ammonia-soda process (see below).

Purely chemical observations, lastly, have led to what is unquestionably the most important of all the innovations in the soda industry, viz. the conversion of common salt into carbonate of soda, without the intermediate formation of sulphate at all, by the ammonia-soda process.¹ Although the reaction upon which this method is based is extremely simple, it took a very long time before the most favourable conditions for it were worked out, and before it was made into a practical success; but this was ultimately achieved by E. Solvay. The production of "ammonia soda" has now attained to such a height that the manufacture of "Leblanc soda" has been greatly prejudiced. For many years back chemists have been striving to solve the problem—how to obtain hydrochloric acid or chlorine from the waste products of the ammonia-soda process;² should this be ultimately accomplished on the practical scale, then it is hardly conceivable that the Leblanc process can continue to exist.³

Chemical labours have exercised a less profound influence upon the manufacture of hydrochloric acid, which is necessarily produced in such quantity in the Leblanc process, although laboratory researches have led to important improvements with regard to its condensation by water, and to its purification from admixed substances. It may be mentioned here, as a curious point in chemical history, that this acid which is at present so cheap, and which has at times been almost worthless, was in Glauber's time the most costly of the mineral acids.

The manufacture of chloride of lime, which uses up large quantities of hydrochloric acid, has also derived great benefit

¹ For the history of this, cf. *Hofmann's Bericht*, vol. i. p. 445.

² The quite recent process of Weldon and Pechiney seems to come near to the solution of this problem (cf. *Journ. Chem. Ind.*, vol. vi. p. 775; or *Mon. Scient.* for April 1888).

³ The production of "ammonia soda" and of artificial manures has grown so enormously of late years that the demand for ammonia salts has immensely increased; but this requirement has in its turn been met by the introduction of improved apparatus for the working up of gas liquor. Here, again, the advances in this branch of manufacture are due to chemical investigation.

from chemical research, in fact it may be said to have arisen from the latter. Berthollet's experiments upon the bleaching action of chlorine and the chlorides (*i.e.* hypochlorites) of the alkalis led to the manufacture of the bleach liquor known under the name of *Eau de Javelle*. Chloride of lime was first produced by Messrs. Tennant and Co. in Glasgow in the year 1799. Weldon's beautiful process¹ for the recovery of the manganese dioxide, required in the preparation of chlorine, from the otherwise worthless chlorine waste—a process which has been in practical working since 1867, grew out of exact laboratory experiments: at the same time its development gave rise to a rich harvest of scientific results. Deacon's method of producing chlorine² directly from hydrochloric acid likewise originated in apparently trivial observations; a strictly scientific explanation of the action of the copper salt on the mixture of hydrochloric acid and air in this process has, however, still to be given.

Bleaching powder itself has been the subject of numberless investigations, made with the object of getting at its constitution. It may, in fact, be said that there is no other substance of equally simple composition regarding the nature of which so much doubt still prevails, notwithstanding all the efforts which have been made to clear it up.³

The two other halogens, bromine and iodine, also became in due course important from a technical point of view, although their much lesser abundance in nature, and consequent less extended practical application, cause them to be produced in small quantities as compared with chlorine. The manufacture of these is based upon the original work of Gay-Lussac and Balard. Laboratory experiments have also led to the production of iodine from mother liquors which were formerly looked upon as valueless, *e.g.* those from Chili saltpetre and from phosphorite after its treatment with acid. To A. Frank⁴ is due the merit of having made bromine

¹ Cf. *Chem. News* for September 1870.

² Cf. *Journ. Chem. Soc.* for 1872, p. 725.

³ Cf. *The History of Inorganic Chemistry*, p. 393.

⁴ Cf. *Hofmann's Bericht*, etc., vol. i. p. 127.

available for technical purposes, by preparing it from the Stassfurt waste salts. Large quantities of both of those halogens (in combination with silver) are now employed in photography.

Nitric acid also plays an important part in chemical industries, especially since the development of the manufacture of explosives on a large scale. Potassium nitrate, which has been known and valued for so long, is still an indispensable ingredient of gunpowder. Since the introduction of the nitrate of soda from the Chili deposits, nitric acid has been prepared from it (instead of from the more expensive nitrate of potash) by the old process of distillation with sulphuric acid. At the same time nitrate of soda is now largely converted into the potash salt by double decomposition with chloride of potassium. This process, so simple from a chemical point of view, could however only be carried out on an extensive scale after the rich deposits of potash salts at Stassfurt had been discovered; and it required careful chemical investigation to make those salts available,¹ for their composition had to be worked out, and proper methods for separating them from one another had to be devised. The large quantities of potassium chloride which occur in the Stassfurt mines have led in certain instances to the carrying out of the Leblanc process with it instead of with common salt, and to the consequent production of carbonate of potash, or *mineral potash*, as it was called (H. Gruneberg, 1861). The extensive use of the Stassfurt potash (and other) salts in the manufacture of artificial manures may also be referred to here; an immense new industry has thus been developed concurrently with the increased production of superphosphate of lime and salts of ammonia.

A reference to the history of gunpowder, and of explosives generally, must not be omitted here, and this all the more because the discovery and use of the latter are connected in the most intimate manner with the development of the chemistry of the time. It is known that the Chinese and

¹ Cf. A. Frank, *Hofmann's Bericht*, etc., vol. i. p. 351; also Pfeiffer's *Kaliindustrie* ("The Potash Industry," 1877).

Saracens made use long ago of mixtures similar to gunpowder for fireworks, while in Europe it has been employed for the propulsion of projectiles since the beginning of the fourteenth century. But five hundred years passed before the chemical reactions, which go on during the combustion of powder, were in some degree understood. That its effect was due to the production of gas was stated by van Helmont; but it was only through the exact experiments of Bunsen and Schischkoff¹ upon the composition of powder gases and residues that the foundation was laid for a theory of its combustion, this being further developed by the later work of Linck, Karolyi, Abel and Nobel, and Debus.

The explosives (with the exception of gunpowder), whose preparation now forms such a great industry, have all been made available for practical use by chemical investigations. The epoch-making discovery of gun-cotton by Schonbein and Böttger in 1846 must be recalled here: its chemical nature and reaction upon ignition were cleared up by the laborious work of Lenk, Karolyi, Heeren, Abel, and others. Nitro-glycerine had been known as a chemical preparation for fifteen years before it began to find extended application in 1862, as the result of Nobel's researches. The careful investigations of Abel, E. Kopp, and Champion upon its modes of formation and chemical behaviour immensely facilitated both its own manufacture and that of its various preparations,—dynamite, etc. The "smokeless powder," of which we hear so much at the present moment, is also to be placed in the same category as the explosives just mentioned, since it contains nitro-cellulose. Reference must also be again made at this point to the famous researches, already mentioned, of Liebig and other chemists upon the fulminates, which rendered the manufacture of fulminate of mercury and its use in the preparation of fuses possible.

The whole match industry likewise owes its enormous development to the increased knowledge of chemical preparations and processes. What a contrast there is between the "chemical tinder" of 1807—*i.e.* matches containing a

¹ *Pogg. Ann.*, vol. cii. p. 53.

mixture of chlorate of potash and sulphur, which were ignited by dipping them into sulphuric acid—and our present friction matches! Those prepared with phosphorus were introduced in 1833 by Romer of Vienna and Moldenhauer of Darmstadt; they have since then undergone many improvements, the most important of these being subsequent to the discovery of amorphous (non-poisonous) phosphorus, which has been used since the year 1848, although for a long time only in small quantity, either in the match itself or in the material of the surface upon which the match is rubbed. Phosphorus, which last century was still a chemical curiosity, has been manufactured on the large scale for about fifty years. Scheele's process for its preparation was improved upon by Nicolas so far back as 1778, and has been materially modified in recent years, *e.g.* by Fleck.

Hand in hand with the development of the soda industry went the expansion of other branches of chemical manufacture, prominent among which was that of soap. In order to appreciate the influence of chemical investigation upon this, we have to recall to mind the pioneering labours of Chevreul on the subject. The knowledge of the chemical nature of fats to which they led was perfected by later work, particularly by that of Heintz and of Berthelot, which finally proved that the fats were neutral glycerine ethers of various fatty acids.¹ The manufacture of stearine candles and of glycerine, which are important both as commercial and household products, may be regarded as the fruits of the labours just spoken of, in addition to which those of A. de Milly (the originator of the stearine industry), Melsens, and Frémy deserve especial mention.

Closely connected also with the soda industry stand the manufactures of ultramarine and of glass. The former substance, which is in a special degree a product of chemical research, was discovered in 1828 by Chr. Gmelin, and also at about the same time by Guimet. It has given rise to a large amount of scientific investigation,² which has led

¹ Cf. p. 405.

² The work of Leykauf, Buchner, R. Hoffmann, Knapp, and Guckelberger may be referred to here.

to material improvements in the manufacture of the various kinds of ultramarine, and has also cleared up particular parts of the firing process, but from which no final opinion has yet been formed as to the chemical nature of this curious product. The two hypotheses still oppose one another, viz. (1) that ultramarine is a definite chemical compound, and (2) that it is a mixture similar to glass. The recent work of F. Knapp¹ has, however, begun to throw some light upon the cause of the colour of ultramarine.

Although the production of glass reached a high state of development in olden times through pure empiricism, it too has greatly benefited by chemical research. The manufacture of glass with sulphate of soda and the improvements in flint and crystal glasses belong to the present century, while progress has also been made in silvering (by Liebig), and in glass painting, through the discovery of new mineral colours. The investigations of Wöhler, Knapp, Ebell, M. Müller, and others resulted in elucidating the chemical reasons for the different colours of different glasses. Lastly, laboratory work has greatly advanced the art of imitating the precious stones, and, generally, of producing new varieties of glass. The chemical reactions which go on during the formation of glass have given rise to much experimental work,² but the conclusions drawn from this—as to whether glass is a true chemical compound or not—have been very various.

Water glass, which was known to Agricola, Glauber, etc., was made available for technical purposes by Fuchs in 1818, and has since then been used for a great number of different purposes, *e.g.* for impregnating wood, preparing cements, protecting frescoes, etc.

Earthenware and Pottery.—Important practical improvements in this old field of industry are associated with the

¹ *Journ. pr. Chemie* (2), vol. xxxviii. p. 48.

² Pérouze, *Ann. Chim. Phys.* (4), vol. x. p. 184; R. Weber, *Wagner's Jahresbericht* for 1863, p. 391; Benrath, *ibid.* 1871, p. 398; also Benrath's book, *Die Glasfabrikation* ("The Manufacture of Glass," 1875).

names of Wedgwood, Littler, Sadler, and others. C. Bischof,¹ Richters,² and, quite recently, Seger³ have rendered good service in their chemical investigations upon the nature of fireclay, and on the connection between its composition and its behaviour at high temperatures. The labours just cited have also done much to improve the manufacture of pottery, by enabling the proper mixtures of the ingredients to be made. The ceramic art is further greatly indebted to chemistry as regards glazing and the burning-in of colours.

The preparation and application of mortar, especially of hydraulic cement, have likewise been greatly advanced by purely chemical work, whereby a nearer approach has been made to the solution of the much-discussed problem,—how its hardening is to be explained from a chemical point of view. Many investigations have been made with a view of arriving at the explanation of this, the chief property of cements, among others by Winkler, Feichtinger, Michaëlis,⁴ F. Schott,⁵ Fr. Knapp,⁶ and Michel.⁷ The old view of the hardening process, viz. that it consists entirely in the gradual formation of a calcium silicate, had to be abandoned as insufficient ; but a complete theory of it still remains to be given.

The advances made in the manufacture of paper can be but partially touched upon here, the more especially since they belong chiefly to the domain of mechanics. The attempts to utilise raw vegetable products, particularly wood and straw, for the production of paper, were first successfully carried out in the year 1846. In caustic soda a reagent was found by means of which cellulose could be prepared from these ; while of late years a solution of calcium sulphite in sulphurous acid has shown itself especially well adapted for this purpose. The above process for the production of *sulphite cellulose* resulted from the chemical investigations of Tilghman and Al. Mitscherlich. The conversion of cellulose

¹ *Dingl. Journ.*, vols. clix. xciv. cxviii. and cc.

² *Ibid.*, vol. xc. p. 150.

³ *Ibid.*, vol. cxxviii. p. 70.

⁴ Cf. his pamphlet, *Die hydraulischen Mörtel*, etc. (Leipzig, 1869).

⁵ *Dingl. Journ.*, vol. ccii. p. 434 ; vol. ccix. p. 130.

⁶ *Ibid.*, vol. ccii. p. 513.

⁷ *Journ. pr. Chemie* (2), vol. xxxiii. p. 548.

into cane-sugar or alcohol is another problem which has been often attacked, and from many different sides, but it still remains to be solved. Should this ultimately be successfully carried out on the large scale, a complete revolution would be effected in agriculture and husbandry generally.

The manufacture and working up of starch has also derived great advantages from chemical investigations. The transformation which starch undergoes upon treatment with acids has only been quite recently cleared up in some degree by the work of Märker, Musculus, O'Sullivan, Payen, Brown and Heron, Salomon, Allihn, etc. The earliest observation on the production of starch-sugar was made by Kirchhoff in 1811, and from this an important branch of industry has now arisen.

The beet-sugar industry has developed into something enormous from experiments instituted by chemists on a small scale.¹ Marggraf's discovery, in 1747, that sugar was present in the juice of beet, was not at that time capable of being applied commercially. Achard, a pupil of Marggraf, and, in a lesser degree, Hermstadt, Lampadius, and others, again took up at the end of last century the problem of obtaining sugar from beet on the large scale, and they did succeed in devising a process which was carried out in numerous factories during the years of the Napoleonic wars, when the trade of the continent was driven in upon itself. But this process was unable to live long, being but a very imperfect one, and giving but a small yield of sugar. It is from the year 1825 that the real rise of the beet-sugar industry dates, various factors entering into its growth, not the least of which was the practical application of chemical knowledge. We have but to think, for example, of the development of saccharimetric methods, whose aim was the determination—either by chemical or by physical means—of the percentage of sugar in beet juice; of the improvements in the refining process;² of the recovery of the crystallisable

¹ Cf. Stohmann's *Zuckerfabrikation* (1885).

² The decomposition of saccharate of lime by carbonic acid was introduced by Barruel and Kuhlmann.

sugar in molasses, and so on. The filtration of the refined juice through bone charcoal was first recommended by Figuier in 1811, and then by Derosne in 1812, and has since become an essential part of the process. The use of vacuum pans for evaporating the syrup was introduced by Howard in 1813, since which time many improvements have been made in them. The extremely convenient diffusion process, for obtaining the juice of the beet, was discovered by Roberts in 1866, and soon came into general use, at first in Austria.

A passing reference may be made here to the good done to this branch of industry by agricultural chemistry, in the determination of the most favourable conditions for the growth of beet, and the investigation of the composition of the soils and manures employed, etc. Indeed, there is hardly any other branch of technical chemistry so intimately connected with agriculture as the beet-sugar manufacture. The production of artificial manures has received a powerful impulse from the immense quantity of beet now under cultivation. Lastly, pure chemistry itself has benefited in many respects from the careful investigation of beet juice.

The so-called *saccharine*, a compound containing sulphur which is now manufactured from the toluene of coal-tar, and which is used to a certain extent in lieu of sugar, offers an example of the assiduity with which every branch of chemical industry is being exploited with the object of imitating natural products by artificial ones, and even of replacing the former by other more active substances.

*The Aniline Colours and other similar Dyes.*¹

There is no industry which better illustrates the practical good that accrues from scientific chemical researches than that of coal-tar, the working up of this substance and perfecting of the numerous methods involved in so doing

¹ Cf. especially Nietzki's *Chemie der Organischen Farbstoffe* ("The Chemistry of the Organic Colouring Matters," 1889); and G. Schultz's *Chemie des Steinkohlentheers*, etc. ("The Chemistry of Coal-Tar," etc., 1886-90).

having set in motion and continued to occupy permanently the energies of a large army of chemists. It was clearly proved here that pure chemical work was the necessary preliminary to the development of each and every branch of the whole coal-tar industry. Out of the large number of important investigations by which the latter has been advanced, only the most striking can be mentioned here,¹—those which have had an undoubted influence in shaping this branch of chemical manufacture. This applies to A. W. Hofmann's classical researches upon aniline and its derivatives, and upon rosaniline, the base of fuchsine (magenta), and its derivatives; and also to the notable work done by E. and O. Fischer upon para-rosaniline and rosaniline, which established the constitution of these compounds. What a deep significance for technical industry the investigations of Couper and Rosenstiehl on the toluidines possessed, is sufficiently well known, while important results also accrued from these to the pure science. The beautiful discovery of green dyes from oil of bitter almonds and benzo-trichloride by O. Fischer and Döbner (working separately) in 1877 may likewise be recalled, as also the proof that these substances were, like rosaniline and aurine, derivatives of triphenyl-methane. It must not be forgotten that Mansfield's work of more than forty years ago laid the necessary foundation for the development of the aniline industry,² for it rendered possible the production of benzene and its homologues from coal-tar on the large scale, and also of nitro-benzene.

The first aniline dye which was produced upon a technical scale was the violet prepared by Perkin in 1856, by acting upon aniline with bichromate of potash and sulphuric acid. A. W. Hofmann observed in 1858 the formation of aniline red, which was shortly afterwards manufactured by another method by Verguin of Lyons, and

¹ For the references to special papers, see *The History of Organic Chemistry*, p. 422 *et seq.*

² *Journ. Chem. Soc.*, vol. i. p. 244; vol. viii. p. 110. Mansfield fell a victim to his work, dying of the severe burns which he received as the result of an explosion.

introduced into commerce under the name of fuchsine. This was quickly followed by Hofmann's discovery of aniline blue, aniline violet, and aniline green, which were further proved by that chemist to be derivatives of fuchsine. The discovery of methyl violet by Lauth in 1861¹ and that of aniline black by Lightfoot in 1863 were of great practical importance. While the constitution of this last compound is still enveloped in mystery, that of the other aniline dyes is now for the most part known, thanks especially to the investigations of E. and O. Fischer, mentioned above.

A. Baeyer's successful conversion of phthalic acid into colouring matters (the phthaleins) was of practical importance, since it led to Caro's discovery of the beautiful eosin dyes, while it also proved itself fruitful from a purely scientific point of view, as the elucidation of the constitution of these phthaleins threw light upon other branches of the subject. From the memorable researches of P. Griess, supplemented by those of Caro, Nietzki, Witt, and others, the manufacture of azo-dyes has arisen; the modes of formation and constitution of these were so clearly made out by the above experimenters that an endless series of valuable colouring matters can now be produced by the aid of certain typical reactions.

The first azo-dye was brought into commerce under the name of aniline yellow so long ago as 1864, without, however, its true constitution being known. It is only since 1876 that the enormous development of this industry dates; quickly following upon one another came chrysoidin, the tropaeolines (most of which are yellow and orange dyes), the *Ponceaux* and "Fast Red" of commerce (red dyes distinguished by their purity), together with Biebrich scarlet and crocein scarlet. The most important discovery of the last few years in this direction was that of the "substantive cotton dyes," obtained from benzidine and similar compounds as examples of which we may cite Congo red and chrysamine. The fact that there are more than 150 azo-colours in the

¹ This dye was not, however, prepared on the large scale until 1867.

market is sufficient evidence of the immense number of such compounds.

Chemical research has also borne rich fruit in respect to the alizarine industry. This valuable dye was formerly prepared entirely from the madder root, but is now, practically speaking, obtained only from coal-tar, this revolution having been brought about by Graebe and Liebermann's successful synthesis (in 1869) of alizarine from anthracene, a constituent of coal-tar. In fact, the madder plantations of Alsace, the south of France, and Algiers, which were in a flourishing condition twenty years ago, have now almost ceased to exist. In addition to this great practical triumph, the purely scientific results, which consisted in the determination of the chemical constitution of alizarine and similar compounds, must also be borne in mind.

The chemical investigation of methylene blue and the safranines, new dyes of great value, has been of much importance both practically and theoretically, the rational composition of the former having been arrived at by Bernthsen, and that of the latter by Nietzki and Witt. The great aim of so many of the researches upon such organic colouring matters, viz. the elucidation of their relations to other compounds from which they are readily derivable, has in the above cases been attained: methylene blue is derived from thio-diphenylamine, and the safranines from phenazine. Similarly rosaniline, aurine, and numerous allied substances have been proved to be derivatives of triphenyl-methane: the azo-dyes to be derivatives of azo-benzene and azo-naphthalene, and alizarine, purpurine, etc., to be derivatives of anthraquinone. Various attempts have lately been made, by Witt and Nietzki among others, to discover definite relations between the chemical constitution of dyes and their colouring properties, but these speculations have as yet no claim to be looked upon as constituting a theory.

The chemical investigation of indigo blue, the most valuable of all blue dyes, has also been ardently prosecuted, with a view of arriving at its constitution, which, however, is not yet definitely settled. Most of our knowledge on the

subject is due to Baeyer. He has succeeded in preparing indigo artificially from simpler compounds contained in coal-tar, but as yet no one has been able to convert any one of the known syntheses into a practical commercial process.

Dyeing.

The processes by which colours are fixed upon vegetable or animal fibres have been greatly improved since the chemical nature of dyes came to be known, although there are some cases in which a true explanation is still wanting of the mode in which certain mordants act. The earliest attempt, even if it was an imperfect one, to get clear ideas upon this subject, was made by Macquer in 1795. The empiricism which prevailed for so long in the dyeing industry has gradually been done away with, thanks to the efforts of chemists to obtain a truer insight into the reactions which dyeing involves. With respect to the application of the more important dyes, previous to the discovery of the coal-tar colours, it may be mentioned that indigo was used in Europe from the first half of last century, and madder red from the second half, while picric acid came into vogue at the beginning of the present one. The use of extract of Campeachy-wood (which is still very considerable) dates from about the year 1840, and that of the dye from the yellow berries of the Chinese plant *Sophora japonica* from about 1848. Reference must also be made to the improvements in the application of metallic colours in dyeing, *e.g.* Prussian blue, chrome yellow, chrome orange, etc.

Tanning, whose processes up to 1860 or so were almost purely empirical, has been made susceptible of scientific treatment through the investigations of Knapp, Eittner, Böttinger, and others. This subject ought to have a great interest for chemists, seeing that, according to Knapp, it constitutes a special case of dyeing. The researches on the various tannic acids have been of value from a theoretical point of view. Among the important practical innovations, for which this branch of manufacture has to thank chemistry,

the mineral tanning introduced by Knapp, Heinzerling, and others deserves notice.

Fermentation Processes.

The development of the various processes of manufacture involving fermentation has been immensely advanced by chemical investigation, while at the same time the nature of the processes themselves has been brought into clear relief. In place of the contact theory of Berzelius and Mitscherlich, which was merely a statement of the facts in other words and no explanation, we now have Pasteur's vital theory of fermentation. To this also the "mechanical" theory of Liebig had to give way in its main point, although Pasteur's opinion with respect to the physiological functions of yeast became subsequently modified to a material extent through the researches of others.¹

The labours which were undertaken with the object of testing or establishing theoretical views have also had a determining influence upon the practical working of fermentation processes, since the knowledge thus gained has rendered it possible to subject those processes to a better control than was formerly the case.

The good which has been done by the application of analysis to fermented liquors is evident at a glance, since any defects in their mode of preparation thus become apparent. A knowledge of the normal composition of wine and beer has led to rational suggestions for the improvement of those drinks. It would be out of place here to attempt even a bare enumeration of the more important innovations in this branch, many of which are due to Pasteur.

The manufacture of spirits may be cited as one of the great branches of industry which has been helped to its present high state of development by chemical work. Take, for example, the enormous production of alcohol preparations²

¹ Cf. *The History of Physiological Chemistry*, p. 491.

² *E.g.* Ethyl iodide, bromide and nitrite; propyl and isobutyl compounds, etc.

from spirit itself, as well as from the first and last runnings of the still; the manufacture of ordinary and of compound ethers, the latter of which are so largely used in perfumery and for making artificial *liqueurs*; and that of chloroform, iodoform, and chloral, whose importance in a medicinal sense is sufficiently well known.

The knowledge that the formation of acetic acid from alcohol depended upon the oxidation of the latter, formed the basis of the *Quick Vinegar Process*,¹ the development of which was the direct consequence of Döbereiner's work; while, on the other hand, the technical production of pyro-ligneous acid, methyl alcohol, etc., arose from the chemical investigation of the products of the distillation of wood.

The manufacture of many of the other organic acids is likewise based upon a sound knowledge of chemical reactions, *e.g.* that of salicylic acid from carbolic, of benzoic acid from toluene, of phthalic acid from naphthalene, of oxalic acid from wood by treating the latter with alkali, etc. This last process was discovered by Gay-Lussac in 1829, and its practical application now constitutes an important industry.

Various Products from Coal-tar; Illuminating Materials.

Coal-tar is the raw material from which most of these bodies are prepared—it is, in fact, a rich mine for numberless useful substances. Formerly a troublesome waste material, it is now of at least equal value with the other products from the distillation of coal. The manufacture of ammonia and salts of ammonia from gas liquor is now a thoroughly rational one, thanks to the careful chemical examination of the latter, and it forms a large and important branch of industry. In consequence of the rapidly increasing consumption of ammonia salts, more and more attention is being paid to the problem of utilising the ammonia which escapes into the air when coal is either converted into coke or completely burnt. L.

¹ This process was first carried out by Schützenbach in Freiburg in 1823, and then by Wagenmann in Berlin in 1824.

Mond¹ has quite lately set up an ingeniously constructed apparatus on a large scale at Northwich in Cheshire, which serves not merely for heating purposes, but at the same time allows of the condensation of the ammonia produced. The manufacture of coal-gas has developed in a more empirical manner, and has thus been less influenced by recent chemical researches bearing upon the subject, than many other branches of industry; but here, too, much good has been done by the application of the methods of gas analysis, and chemical experiments have also borne fruit in the introduction of improved modes of purifying gas.

The great influence which chemical investigations have exercised upon the production of other illuminants has already been touched upon, and is shown in the manufacture of stearic acid from animal fats.² Attempts have been made upon the large scale to employ (the liquid) oleic acid, which occurs so plentifully in nature, for the production of candles, by making use of the well-known reaction with caustic potash which converts it into (solid) palmitic acid. The flourishing paraffin manufacture³ of Scotland and Germany, with its various valuable by-products, also owes much to chemistry. But the latter has still many problems to solve both in this field and in that of the petroleum industry, as is evident from the recent work of Markownikoff, Beilstein, and Engler upon the chemical nature of petroleum. A short reference has already been made (p. 392) to the theoretical points which bear upon illumination, and to the causes of the luminosity or non-luminosity of different flames.

Heating Materials.

That the knowledge gained through chemical analysis of the composition of different kinds of fuel, of their products of combustion, and of their chemical behaviour generally, is

¹ *Journ. Chem. Ind.* for 1889, p. 505.

² Cf. p. 511.

³ Paraffin, which was discovered in wood-tar by Reichenbach in the year 1830, is obtained practically from lignite or bituminous shale.

of the first consequence, requires no demonstration. It is of course impossible to refer here to the large number of important investigations in this field, but reference must be made to the fundamental work of E. Richters and F. Muck;¹ to the improvements in the methods of analysis of furnace gases,² which permit of conclusions being drawn with regard to the course of any particular combustion; and to the improvements in heating apparatus which have been brought about by chemical work,—the construction of generators and regenerators, whose history is inseparably connected with the names of Aubertot, Thomas, Laurens, and, above all others, Siemens.

Speculations regarding the origin of coal deposits, and the metamorphoses which these undergo, have received much support from the work which has been done upon the composition of coal and of the gases which are found enclosed in it. And it is mainly to chemical research that we owe the means of averting or at least diminishing the great dangers to which coal miners are exposed from explosions of fire-damp,—witness the Davy safety lamp. The subject is still being assiduously worked at from time to time by eminent chemists and practical engineers. The zeal which the various recent "Fire-Damp Commissions" of different countries showed in their investigations is still fresh in the public memory.

The above short sketch is sufficient to indicate how enormous have been the benefits which laboratory research has conferred upon every branch of technical chemistry, and how the latter has been raised to a higher level by a continuous infusion of the scientific spirit. Nowhere can we find a better illustration of Bacon's maxim: *Scientia est potentia*.

¹ Cf. Muck, *Grundzüge und Ziele der Steinkohlenchemie* ("The Outlines and Aims of the Chemistry of Coal," 1881).

² Cf. Winkler's *Anleitung zur technischen Gasanalyse* ("Methods of Technical Gas Analysis"); also p. 366 of this book.

THE GROWTH OF CHEMICAL INSTRUCTION IN THE NINETEENTH CENTURY, MORE ESPECIALLY IN GERMANY.

At the beginning of this century there was a marked want of those facilities which, during the last few decades, have been at the command of any one desirous of devoting himself to the study of chemistry. At that time there were practically no laboratories for general instruction. In lectures upon physics, mineralogy, and anatomy, chemistry was relegated to a very inferior place. It is true that there were chairs of chemistry in various universities and colleges, but the lectures on this subject were usually conjoined with those upon one of the others, just named, in such a manner that chemistry was forced into the background. Chemical literature, lastly, was still poor in works which either gave a review of the state of the science at the time, or furnished regular reports of the latest discoveries in it.

In France, where towards the end of the eighteenth century it began to be perceived that instruction in natural science must be aided by every means at command, a start was made far ahead of other countries in respect to the development of chemical study. Up till then apothecaries' shops were the only places where work in practical chemistry could be carried on, and there merely after certain prescriptions and not according to scientific methods. Vauquelin was the first to organise a course of instruction in his small laboratory for students anxious to learn, while Gay-Lussac and Thénard also taught in their laboratories, which however were exceedingly cramped. It was only after Liebig had taken up the subject with his accustomed energy, that chemistry came to be taught in the higher schools in essentially the same manner as that to which we are now accustomed.¹

¹ Cf. below; also O. L. Erdmann's valuable pamphlet, *Ueber das Studium der Chemie*.

The importance of lectures on chemistry, illustrated by experiments, for the proper understanding of chemical reactions, was recognised a long time ago, more especially in France.¹ But during the early decades of the present century this aid to study hardly existed in the higher teaching institutions of Germany, and the so-called natural philosophy of that day was such that it sorely handicapped the development of exact scientific research. Chemistry, in particular, was not looked upon by the natural philosophers as a science at all, but was degraded by them into a mere experimental art.

The efforts made by Davy, however, backed as these were by an exceptional talent for devising and carrying out experiments, resulted from the beginning of this century in an increasing demand for lectures with appropriate experimental illustrations. We know that it was the lectures given by Marcet in London which induced Berzelius in 1812 to abandon the old method of instruction, and to make use of experiments in introducing students to chemical science; and the result of this was conclusive. The subsequent good achieved by Faraday, Liebig, Wöhler, Bunsen, Wurtz, Kolbe, and, especially, A. W. Hofmann, from the new lecture experiments which they devised, requires but to be mentioned. Those experiments and many others have since taken a permanent place in the teaching of chemistry.

Practical instruction in chemical laboratories, as commonly carried out at the present day, was developed by Liebig. The gradual introduction into laboratories, through his example, of teaching methods based upon a strictly scientific foundation, created a wholesome reaction against the still prevailing tendency of the natural philosophy of the day, which was combated by Liebig all the more energetically from his having himself suffered under its pernicious influence.² He first emphasised with all the force at his command that the true centre-point of chemical study lay not in lectures but in practical work. With what energy

¹ Cf. The work of Rouelle, p. 115, note 1; also p. 164.

² Cf. p. 248.

and under what sacrifices he gave personal proof of this is well known.¹ True, Berzelius had already given instruction in his laboratory to a limited number of pupils, mostly elder ones, who in their turn propagated their master's doctrines, but the real development of chemical teaching is due to Liebig. He it was who laid down the order, now classical, in which the various branches of the subject should succeed one another, viz. (1) the systematic study of qualitative and then of quantitative analysis,² (2) exercises in the making of preparations, and (3) attempts at independent research.

Liebig's laboratory was the centre from which, after about the end of the twenties, the brightest light radiated. He was the first to enunciate and apply the principle that his pupils, be they students of pharmacy, technical chemistry, mineralogy or physiology, should learn to treat chemical questions practically. Thanks to the wonderful stimulus which he was able to exert, there was founded in his modest laboratory a school which left its stamp upon the chemistry of the succeeding decades, and whose beneficial influence is still felt all over the world at the present day. The peculiarity of Liebig as a great teacher consisted, according to Kolbe,³ in his "being able to stimulate his pupils to original thought, and to inculcate them with the scientific spirit while they were working out his own ideas."

The most eminent among the teachers of chemistry since the time of Liebig, of whom Wohler, Bunsen, Erdmann, Kolbe, Wurtz, and A. W. Hofmann may be named here, made the essential principles of his method of teaching their own, while each added of course much that was new, with the most beneficial results.

Numerous teaching laboratories were in due course founded in the other German universities and colleges on

¹ Cf. the Memoir of him by Kolbe in the *Journ. pr. Chemie* (2), vol. viii. p. 435.

² The co-operation here of R. Fresenius, who was at one time assistant to Liebig, and the stimulus given by him towards the creation of a systematic course of analytical work will remain in lasting remembrance (cf. p. 363).

³ In his work, *Das chemische Laboratorium der Universität Marburg*, etc., p. 26.

the model of the Giessen one, and about these a few notes may fitly find a place here. How badly off Austria and Prussia were in this respect, even so recently as the year 1840, was vividly depicted by Liebig in his two pamphlets entitled *Ueber den Zustand der Chemie in Österreich*,¹—und in *Preussen*² ("On the State of Chemistry in Austria, and in Prussia"). Even in Berlin there were up to that time no facilities for the study of practical chemistry. H. Rose and Mitscherlich were not in a position to give regular laboratory instruction; and the same thing applied to the other "high schools" of Prussia.

In the meantime laboratories began to be established elsewhere in Germany, *e.g.* at Göttingen, where Wöhler set up one in the course of the thirties, to be rebuilt and enlarged in 1888; and at Marburg, where Bunsen began a regular practical course in 1840. The chemical laboratory which Erdmann³ instituted in Leipzig in 1843 remained for a long time the pattern of what a well-organised place of the kind should be. It was only in the course of the fifties that Heidelberg, Karlsruhe, Breslau, Greifswald and Königsberg followed suit with laboratories properly equipped for the purposes in view.

A new era in the history of chemical institutions began about the middle of the sixties, the famous laboratories at Bonn and Berlin,⁴ both built according to A. W. Hofmann's

¹ *Ann. Chem.*, vol. xxv. p. 339.

² *Ibid.*, vol. xxxiv. pp. 97 and 355.

³ Otto Linné Erdmann was born at Dresden in 1804, and died in 1867 while holding the post of Professor of Chemistry at Leipzig, where, since 1827, and especially since the organising of the laboratory which he had himself founded, he laboured with wonderful energy and with great success. His rich experiences, and the views arising from them to which he was led, were set forth in the weighty, if short, pamphlet entitled, *Ueber das Studium der Chemie* (1861). That he was also active in a literary sense, his *Lehrbuch der Chemie* and *Grundriss der Waarenkunde* ("Outlines of a Knowledge of Technical Products"), etc., prove. In 1828 he started the *Journal für technische und ökonomische Chemie*, which developed in 1834 into the *Journal für praktische Chemie*. His numerous experimental researches have helped to enrich mineral chemistry, the chemistry of the carbon compounds, and also chemical technology.

⁴ Up to that date Berlin was without any large laboratory for general instruction.

plans, being completed in 1867, while the equally well-known Leipzig laboratory, designed by Kolbe, was finished in 1868. The experience gained, both during the erection of these and by their subsequent use, has been applied with good results in the planning of later and even in some respects finer institutes. Of the other new German laboratories, those of Aachen¹ (1870), Munich (1877), the Berlin Technical College (1879), Kiel (1880), Strassburg (1885), and Göttingen (1888) may be especially named. In Austria, too, various admirable laboratories have been built during the last two decades, among which those of Graz and Vienna stand out prominent.

The other countries of Europe have not kept pace with Germany in the establishment of institutes for the teaching of chemistry. There were, it is true, laboratories in France at the beginning of the century in which such men as Gay-Lussac, Thénard, Dulong, Chevreul, and others carried out their work, but the opportunities for general chemical instruction were extremely few, the above institutes receiving but trifling support from the State. And the fees which a laboratory student had to pay were exorbitant, being 1500 francs for an eight-months' course. Even the efforts made to establish teaching laboratories during the thirties by Dumas and Pérouze, and later on by Wurtz, Gerhardt, and others, were followed by but scant success.

Those conditions were only improved upon after Wurtz in 1869 sent in his report² upon the German laboratories to the French Minister of Education, in which he insisted upon the necessity of establishing properly equipped laboratories for practical instruction in chemistry. He stated that at that date there was in France only one chemical institute with the necessary means at command,—that of the *École Normale Supérieure*, under the direction of H. St. Claire Deville.

In Great Britain, too, it is only within the last twenty

¹ *i.e.* Aix la Chapelle.

² *Les hautes Études pratiques dans les Universités Allemandes* (1870).

or twenty-five years that the lack of roomy and well-fitted-up laboratories has been remedied; and to this, especially of late years, the recognition of the fact that the industries of the country would be enormously benefited thereby has very greatly contributed. The first laboratory in Britain, small though it was, in which a young man had the opportunity of working practically at the subject, was that of Thomas Thomson¹ in Glasgow, established in 1817. After the founding of the College of Chemistry² in London in 1845 (which quickly rose into a flourishing condition under the leadership of A. W. Hofmann), the country became by degrees well supplied with suitably equipped laboratories, in which instruction upon the lines of the German school was given. In addition to the Universities and a few of the older institutions for higher education in London, etc., each of the University Colleges now scattered over the country possesses its own chemical laboratory, and the same thing applies in greater or less degree to the colleges and schools for technical instruction which continue to be founded with considerable rapidity. In fact, the mind of the country is now becoming pretty well awakened to the importance of the subject. Among the chemical laboratories, more or less recently erected, those at Manchester, Leeds, Edinburgh, the City and Guilds Institute (South Kensington, London), and Cambridge may be specially named.

In Switzerland, Holland, Belgium, Italy, Russia, Scandinavia, and America are now to be found numerous chemical teaching institutes, arranged and fitted up in accordance with the requirements of the age.

The increasing necessity for specialisation in chemistry, and the consequent resulting division of labour, has made

¹ Cf. p. 183.

² The College of Chemistry was taken over by Government in 1853, and was made a part of the Royal School of Mines, while at the same time retaining a *quasi*-separate existence under its own name. In 1872 it was moved from its old premises in Oxford Street to South Kensington. The name College of Chemistry was finally merged into that of the Normal School of Science and Royal School of Mines in 1881. In 1890 the N. S. S. and R. S. M. were rechristened the Royal College of Science.

itself evident in the establishment of laboratories for certain definite purposes only. Thus we now find institutions existing solely for researches in chemical physics, technological chemistry, physiological chemistry, pharmaceutical chemistry, and hygiene. What a contrast between the present facilities for chemical study and the opportunities of only a few decades back!

Among the more important improvements which have been aimed at and achieved in the construction of laboratories during these last decades, are those which have reference to arrangements for supplying plentiful ventilation and good light. Then the means for carrying out chemical operations have also been both greatly increased and improved upon, *e.g.* coal and charcoal fires have been superseded by gas, the *Bunsen* burner having played an important part here. The apparatus, too, employed by chemists has undergone many refinements, as is readily seen in the delicate balances and the appliances for filtering, distilling, heating, etc., which are now in common use.¹ The making of preparations is at present an easy matter compared with what it used to be, this being in part due to better methods of procedure; by far the greater number of these substances can now in fact be bought quite pure. Chemists are thus freed from the difficulty which was ever present with them sixty years ago,—of having laboriously to prepare even their most simple reagents. Berzelius had to make his own yellow prussiate of potash, the pure mineral acids, spirits of wine

¹ The following points may be referred to with advantage here:—Water suction pumps were introduced by Bunsen in 1868, and injector pumps a little later by Arzberger, Zulkowsky, etc., to be used for filtering and producing a vacuum. Simple distillation was immensely facilitated by the introduction of the Liebig condenser, while a reflux condenser appears to have been first made use of by Kolbe and Frankland in 1847. Dittmar and Anschütz (independently of one another) were the first to distil under diminished pressure. The water-bath, for which Berzelius devised a convenient form, has since been improved by arrangements, devised by Fresenius, Bunsen, Kekulé, and others, for keeping the water in it at a constant level. The use of gas regulators for the maintenance of a uniform temperature may also be mentioned, and this again in conjunction with Bunsen's name. Caoutchouc tubing appears to have been first brought into general employment by Berzelius.

for burning, etc. And how simple were the arrangements generally in his laboratory!¹ Many of the aids to practical work which are now accepted as a matter of course had in his day no existence.

Chemical Literature.

The manuals and text-books of chemistry and also the journals have increased to a very great extent of late years, thus much facilitating the study of the science. For a long time Lavoisier's *Traité de Chimie* remained the pattern of what such a book should be, and upon it numerous others were modelled, e.g. those of Girtanner, Gren, and Thomson. Berzelius' large book on chemistry exercised an extraordinary influence, more especially after it had been translated into other languages, and contributed in an exceptional degree to the spread of chemical knowledge.

This great work, great both in its conception and in the manner in which it was carried out, was afterwards taken in many cases as the standard for the arrangement of chemical matter in text-books which appeared later. Of these a few may be mentioned here:—Thénard's *Traité de Chimie Élémentaire*; Mitscherlich's *Lehrbuch der Chemie*; Liebig's *Organische Chemie*; Wöhler's *Grundriss der Chemie* ("Outlines of Chemistry"), from which sprang the well-known and widely-read work of the same title by Fittig; Regnault's *Cours Élémentaire de Chimie*, which formed the basis of Strecker's *Kurzes Lehrbuch der Chemie*; Graham's *Elements of Chemistry*, from which arose Otto's large work, the organic portion of which was written by Kolbe, while H. Kopp wrote the theoretical part, and Buff and Zamminer the physico-chemical. Gerhardt's *Traité de Chimie Organique* (1853-56), known as the text-book of the type theory, greatly contributed to the propagation of the latter, while Kekulé's book, which began to appear shortly after the last volume of Gerhardt's *Traité* had been published, served to develop the "typical" view, and (in its

¹ Cf. Wöhler's description, *Ber.*, vol. xv. p. 3139.

second volume) strengthened his own assumption as to the mode in which atoms are combined with one another. It is unnecessary to mention here even a few of the numerous text-books of chemistry which have been written since then, for, belonging as they do to the present era, they are already sufficiently well known. A palpable want has recently been supplied by the publication of W. Ostwald's and Horstmann's admirable text-books of general theoretical and physical chemistry.

There has likewise been no lack of chemical encyclopedias since the great success of Liebig, in conjunction with Wöhler and Poggendorff, in the *Handwörterbuch der reinen und angewandten Chemie*, which began to appear in 1837. Wurtz's *Dictionnaire de Chimie pure et appliquée*; Watts' *Dictionary of Chemistry*, and Laudenburg's *Handwörterbuch der Chemie* have been written upon a similar plan. The publication of Frémy's *Encyclopédie de Chimie* must also be recalled.

Among the larger treatises of chemistry, which are intermediate between the text-books proper and the dictionaries, that of L. Gmelin justly excited the admiration of his contemporaries by its consistent thoroughness. In Beilstein's *Handbuch der Organischen Chemie*, the second edition of which is only just completed, the present huge mass of material on the subject has been sifted and arranged in a masterly manner.

The periodical journals, whose number has gone on steadily increasing, have exercised the greatest influence upon the enlargement and spread of chemical knowledge, more especially since the beginning of this century. A short account has already been given¹ of the condition of this class of literature towards the end of last century. In Germany, after the third decade of the present one, all the most important chemical researches were for long published either in Poggendorff's *Annalen der Physik und Chemie* or in the *Annalen der Chemie und Pharmazie*,² which was

¹ Cf. pp. 165 and 169.

² Until the year 1839 this journal bore the simpler title, *Annalen der Pharmazie*.

at first edited by Liebig alone, but afterwards in conjunction with Wöhler. The latter, more particularly, soon became the medium in which were discussed the experimental and speculative chemical questions of the day. And no one was better qualified to deal with those in a thorough manner than Liebig himself.

In France the *Annales de Chimie*, founded in 1789 by Lavoisier, Fourcroy, and Berthollet, has always been appreciated and loyally supported. Since 1816 it has appeared as the *Annales de Chimie et de Physique*, its first editors under this new title having been Gay-Lussac and Arago, and it has all along contained the records of pretty nearly all the more important French chemical researches. The *Comptes Rendus*, which has been published weekly by the *Académie Française* since the year 1835, includes among its numerous papers only comparatively few and short accounts of chemical investigations.

In Great Britain, up to the year 1841, papers on chemical subjects were published either in the *Philosophical Transactions*, the *Transactions* of the Royal Society of Edinburgh, etc., or in other more recent journals, which have since been superseded, such as Nicholson's *Philosophical Journal*, and Thomson's (later Phillips') *Annals of Philosophy*. Since 1841, or at least since 1848, the *Journal of the Chemical Society* has been the main organ of scientific chemistry in this country. Apart from the original memoirs which it contains, this journal has since 1871 greatly extended its usefulness by giving copious abstracts of papers which have appeared in the chemical journals of other countries.

And the other European countries have not been behind-hand in the publication of chemical journals; according to the degree in which chemistry has found in them a permanent home, so have journals of every shade and variety sprung up. Most of these were and are still connected with learned corporations—academies and chemical societies—in Austria, Italy, Holland, Belgium, Switzerland, Russia, and Scandinavia, and the same applies also to North America.

In Germany more particularly, which has now for long been the chief centre for scientific chemical interests, thanks to the favourable conditions for scientific instruction there, a number of new journals for the publication of papers on purely chemical subjects have been added to those older ones just mentioned. Among these are the *Journal für Praktische Chemie*, begun by Erdmann in 1834, and, especially, the *Berichte der Deutschen Chemischen Gesellschaft*, which was brought into life with the founding of the German Chemical Society at Berlin in 1868, and in which one finds a record of pretty nearly all that is being done in scientific chemistry, either in the form of original papers or of abstracts from other journals. Mention must also be made here of the *Kritische Zeitschrift*, known later on as the *Zeitschrift für Chemie*, which was supported by such men as Kekulé, Erlenmeyer, Fittig, and others, and the critical utterances in which have often helped to shed light upon disputed points in chemistry.

Mention still remains to be made of the *Jahresberichte* ("Yearly Reports") on the progress of chemistry and allied branches of science. The reports which were edited by Berzelius (from 1821 to 1847) are unique, and are absolutely indispensable to any one who desires to make a detailed study of the progress of chemistry during those years. The continuation of them, which was undertaken by Liebig in conjunction with other chemists, cannot be compared with these earlier volumes, the new *Jahresberichte* having been restricted into mere epitomes of reference with regard to current chemical work.

The critic, whose use as a fermentive and corrective agent will be denied by no one, seems either to have disappeared from the chemical literature of recent years, or at all events to be at present dormant. It is well to remember that the critical acumen which was brought to bear upon the occasional errors of chemical investigation by Berzelius and Liebig, and at a later date by Kolbe, had a clarifying and not a disintegrating effect, even in those cases where the critic's argument had a strongly polemical flavour.

The value of a minute study of good original papers has time and again been insisted upon by the great teachers of chemistry. The records of such experimental labours offer to the student the best means of following out the author's train of thought; they thus strengthen the historical sense, and at the same time strongly incite to criticism and to emulation. They are therefore to be looked upon as among the best literary aids to the study of chemistry. At the same time they possess a high educational value from their style and form alone. As Erdmann well says in his short treatise, already cited, p. 60: "By making use of such sources of information the student learns at one and the same time from a master of the science how and in what form scientific results should be stated, how to distinguish between what is and is not essential, and how to condense the subject-matter, while at the same time omitting from it nothing of importance, so that no necessary element is wanting for its critical examination."

INDEX OF AUTHORS' NAMES

The figures in thick type refer for the most part to those pages upon which biographical notices occur, although they are also employed in some cases for others on which points of special importance are recorded.

- ABEL, 510
 Abukases, 29
 Achar, 514
 Afzelius, 192
 Agricola, 58, 80, 84, 85, 90
 Albertus Magnus, 31, 40, 130
 Allihn, 514
 Ampère, 387
 Anaximenes, 8
 Anderson, 432, 436
 Andrews, 378, 444
 Anschütz, 404
 d'Arcet, 502
 Archimedes, 14
 Arfvedson, 378
 Aristotle, 2, 6, 8, 16, 20, 38
 Armstrong, 493
 Arnaldus Villanovanus, 32, 40, 43
 Arppe, 406
 Arrhenius, 448, 468
 Augustin, 502
 Avenzoar, 29
 Avenrhoes, 29
 Avicenna, 29
 Avogadro, 202, 275, 441
- BACON, Francis, 98
 Bacon, Roger, 31, 43
 Baeyer, 330, 333, 334, 337, 340, 401,
 415, 425, 433, 437, 517, 519
 Balard, 278, 375, 387, 393, 408, 508
 Bamberger, 401
 de Bary, 492
 Basilius Valentinus, 35, 41, 46, 47, 50,
 51, 52, 53, 54, 57
 Baumann, 404, 488
 Baumé, 144
 Bayen, 124, 166
- Beaumont, 486
 Béchamp, 408
 Becher, 105, 106, 127
 Beilstein, 416, 522, 532
 Bergman, 119, 122, 131, 132, 133,
 136, 137, 138, 233, 360, 460, 470
 Berlin, 382
 Bernard, Claude, 486, 489
 Bernoulli, 444
 Bernthsen, 518
 Berthelot, 310, 388, 401, 403, 457,
 465
 Berthier, 474
 Berthollet, 161, 162, 163, 175, 389,
 427, 460 *et seq.*
 Berzelius, 170, 191-201, 203-206,
 209, 210-212, 214, 216-224, 226,
 227 *et seq.*, 234-238, 241 *et seq.*,
 262, 265-269, 279, 358, 359, 361-
 363, 369, 374 *et seq.*, 385 *et seq.*,
 428, 452, 469 *et seq.*, 524 *et seq.*
- Bendant, 208, 472
 v. Bibra, 485
 Bidder, 486, 490
 Biringuccio, 82
 Bischof, C., 513
 Bischof, G., 475
 Bischoff, 488, 490
 Black, 116, 121, 122, 136, 169
 Blaise de Vigenère, 94
 Blochmann, 392
 Blomstrand, 223, 311, 322, 383, 395,
 397, 427, 464, 473
 Boerhave, 59, 110, 128, 129
 Boltzmann, 444
 v. Bonsdorff, 473
 Böttger, 510
 Bötticher, 62

- Böttinger, 519
 Boullay, 239
 Boussingault, 258, 480
 Boyle, Robert, 3, 58, 97, 100-104,
 121, 123, 126, 128, 130, 132, 133,
 134, 138, 141, 458
 Brand, 141
 Brandt, 141
 Brefeld, 492
 Brieger, 494
 Brisson, 144
 Brodie, 408
 Bromeis, 405
 Brown, 415, 514
 Brücke, 452, 483, 485 *et seq.*
 Brühl, 449, 457
 Brush, 473
 Buchholz, 167, 470
 Buckton, 421
 Buff, 310
 Buffon, 132
 Bunsen, 246, 293, 359, 363, 366,
 379 *et seq.*, 391, 393, 425, 444, 445,
 447, 459, 473, 475, 501, 510, 525
 et seq.
 Butlerow, 300, 316 *et seq.*, 325 *et seq.*,
 401, 403

 CÆSALPIN, 85
 Cæetano, 61
 Cagliostro, 61
 Cagniard de la Tour, 491
 Cahours, 244, 403, 408, 424, 425
 Cailletet, 443
 Cannizzaro, 317, 403, 428
 Carlisle, 216
 Caro, 517
 Carstanjen, 431
 Cavendish, 117-118, 122, 137, 142,
 157, 169
 Chance, 506
 Chancel, 281, 404
 Chevreul, 402, 405, 511
 Christensen, 376, 394
 Ciamician, 436
 Claisen, 412
 Clarke, 391
 Classen, 364
 Claus, 311, 326, 330, 384
 Clausius, 444
 Clément, 393
 Cléve, 397
 Cloëz, 428, 430
 Combes, 413
 Cooke, 377
 de Coppet, 351, 447
 Couper, 310, 315-316, 319, 326
 Coupier, 516

 Courtois, 375, 503
 Cousin, 158
 Crell, 169
 Croll, 71
 Cronstedt, 136, 141
 Crookes, 349, 350, 380
 Curtius, 388, 390, 427

 DAGUERRE, 458
 Dale, 449
 Dalton, 170, 177-182, 183, 187, 444
 Daniell, 229
 Daubrée, 474
 Davy, Humphry, 184-187, 224, 225,
 226, 228, 374 *et seq.*, 386 *et seq.*,
 428, 455, 525
 Deacon, 508
 Debray, 384, 395, 443, 475, 502
 Debus, 412, 510
 Déherain, 480
 Delitzsch, 428
 Democritus, 8
 Deunstedt, 436
 Derosne, 515
 Descroizille, 365
 Desormes, 393, 504
 Dessaigues, 488
 Deville, H. St. Claire, 384, 390, 392,
 443, 475, 502, 528
 Dewar, 331
 Dietrich, 480
 Diodorus Siculus, 13
 Dioscorides, 7, 13, 16, 18, 19
 Dittmar, 386, 530
 Dobereiner, 235, 388, 405, 436, 521
 Döhner, 433, 516
 Dombasle, 478
 Dragendorff, 371, 494
 Draper, 459
 Drebbel, 92
 Drechsel, 406, 485
 Duhamel de Monceau, 114, 129, 139,
 141
 Dulong, 206, 207-208, 228, 229, 387,
 390, 391
 Dulong and Petit, 207-208, 449
 Dumas, 191, 212-214, 239, 243, 244
 et seq., 256-259, 260 *et seq.*, 269,
 270, 346, 362, 376 *et seq.*, 402,
 404, 408, 417, 429, 441, 528
 Duppa, 406
 Durocher, 474

 EBELL, 512
 Ebelmen, 474
 Ehrenberg, 431
 Eittner, 519
 Ekeberg, 170, 192, 383, 470

- Elhujar, 381
 Eller, 112
 Empedocles, 8
 Engelmann, 482
 Engler, 522
 Erasmus of Rotterdam, 61
 Erastus, 70
 Erdmann, 191, 346, 362, 370, 473,
 526 *et seq.*
 Erlenmeyer, 316 *et seq.*, 320, 326, 426
 Étard, 494

 FAGGOT, 147
 Falck, 490
 Faraday, 214-215, 237, 386, 443, 452,
 464
 Faworsky, 401
 Fehling, 365, 405, 406, 408, 412, 429
 Feichtinger, 513
 Figuier, 396, 515
 Finkener, 501
 Fischer, E., 340, 341, 402, 415, 427,
 437, 516
 Fischer, G. E., 173
 Fischer, O., 340, 402, 516
 Fittig, 327, 340, 401, 407, 410, 413,
 415, 416
 Fitz, 492
 Flamel, Nicolas, 34
 Fleck, 511
 Flügge, 372
 Fordos, 388
 Forster, 490
 Fouqué, 475
 Fourcroy, 153, 161, 164-165
 Fownes, 436
 Frank, A., 501, 508
 Franke, 394
 Frankland, E., 294, 295, 297, 304-
 307, 308, 311, 312, 316, 319, 338,
 343, 392, 399, 406, 429, 437, 438
 Frémy, 387, 390, 394, 485, 511
 Frerichs, 486, 490
 Fresenius, R., 359, 363, 370, 526
 Freund, 401
 Friedel, 300, 403, 475
 Friedländer, 433
 Frobenius, 144
 Fuchs, 512
 Fuchs, N., 471

 GADOLIN, 381
 Gahn, 136, 138, 141, 142, 170, 470
 Galen, 48
 Gattermann, 391
 Gautier, 430, 494
 Gay-Lussac, 173, 187-189, 201, 208,
 225, 226, 235, 237, 364, 369, 375,
 387 *et seq.*, 427, 439, 440, 441, 474,
 505, 503, 524
 Geber, 28, 38, 47, 49, 50, 51, 52, 56
 Gehlen, 390
 Gélis, 388
 Gengembre, 166, 389
 Geoffroy the elder, 113, 114, 129
 Geoffroy the younger, 114
 Gerhardt, 266, 269-275, 276 *et seq.*,
 283-289, 309, 404, 408, 418, 432,
 531
 v. Gerichten, 435
 Gerland, 396
 Geuther, 343
 Gibbs, 395
 Gilbert, 169
 Gilchrist, 501
 Girtanner, 531
 Gladstone, 449, 465
 Glaser, 105, 146
 Glauber, 83, 84, 88, 89, 91, 92 *et seq.*,
 130
 Glover, 505
 Gmelin, Chr., 192, 511
 Gmelin, L., 230, 274, 395, 427
 Gore, 387
 Gorup-Besanez, 486
 Göttling, 358
 Goulard, 146
 Graebe, 333, 401, 413, 433, 518
 Graham, 230, 384, 391, 360, 452
 Gren, 169
 Grew, 146
 Griess, 301, 423, 426, 517
 Grimaux, 153
 Gros, 397
 Groth, 454
 Grüneberg, 509
 Gruner, 501
 Guareschi, 494
 Guckelberger, 506
 Guimet, 511
 Guldberg, 465 *et seq.*
 Gustavson, 400
 Guyton de Morveau, 126, 153, 161,
 162-163

 HAARMANN, 415
 Hagen, 146
 Haitinger, 435
 Hales, 121, 122, 124
 Hall, 474
 Hammarsten, 485, 487
 Hantzsch, 337, 433, 437
 Hargreaves, 506
 Hasenbach, 390
 Hatchett, 169, 383
 v. Hauer, 396, 473

- Hausmann, 359, 472
 Hautefeuille, 475
 Haüy, 208, 470
 Hawksbee, 104
 Heeren, 510
 Heintz, 405, 485, 487
 Helbig, 506
 Hellot, 139
 van Helmont, 58, 60, 72-77
 Helvetius, 60
 Hempel, 366
 Henneberg, 480, 481
 Hennel, 404
 Henninger, 486
 Henry, 169, 389, 444
 Henry, L., 323
 Heracitus, 8
 Herepath, 444
 Hermann, 473
 Hermbstädt, 514
 Hermes Trismegistos, 24
 Heron, 415, 514
 Herschel, 445
 Herter, 486
 Hess, 370, 526
 Heumann, 392
 Hiärne, 138
 Higgins, 184
 Hill, 437
 Hisinger, 192, 216
 Hittorff, 378
 Hjelm, 381
 van 't Hoff, 336, 448, 450, 468
 Hoffmann, 110, 133, 135, 143, 144
 Hofmann, A. W., 268, 279-280, 403,
 412, 422 *et seq.*, 441, 516, 525, 527
 et seq.
 Hofmann, Fr., 490
 Hofmann, R., 409
 Hofmeister, 486
 Homberg, 104-105, 136
 Hooke, 104, 125
 Hope, 379
 Hoppe-Seyler, 485 *et seq.*
 Horstmann, 443, 468
 Howard, 515
 Häfner, 487, 493
 Humboldt, A. von, 201, 248, 257
 Husemann, 371

 INGEN-HOUSS, 478, 482
 Isaac Hollandus, 34
 Isambert, 443
 Ittner, 427

 JACOBI, 502
 Jolly, 452
 Jørgensen, 395, 397

 Joule, 444
 Julius Firmicus, 25

 KALLE, 422
 Kane, 242
 Kanonikoff, 449
 Karolyi, 510
 Kay, 309
 Kekulé, 289, 290, 291, 300, 310-
 312, 314 *et seq.*, 319 *et seq.*, 325
 et seq., 328 *et seq.*, 332 *et seq.*, 340,
 404, 406, 409, 416, 420, 426, 431
 Kempe, 395
 Kiliani, 415
 Kircher, 61
 Kirchhoff, 514
 Kirchhoff, G., 359, 379, 445
 Kirwan, 169, 470
 Klaproth, 167-169, 360, 362, 376, 379,
 381, 470
 Klason, 393, 420, 428, 430
 Klinger, 323
 Knapp, 512, 513, 520
 Knop, W., 480, 481
 Knorr, 340
 v. Knorre, 395
 Kobert, 494
 Kolb, 506
 Kolbe, 289, 292-303, 311, 312, 319,
 322, 326, 338, 399, 405, 409, 410
 et seq., 418 *et seq.*, 429, 437, 494,
 525
 König, 372
 Königs, 340, 433, 434
 Kopp, E., 310
 Kopp, H., 350, 440, 445, 446, 454
 Koppfer, 370
 Körner, 330, 432
 Krafft, 406
 Krönig, 444
 Krüger, 323
 Krüss, 395, 396
 Kuhne, 485, 486
 Külz, 489
 Kunkel, 105-106, 138, 139, 141,
 147
 Kützing, 492

 LAAR, 334
 Ladenburg, 330, 333, 340, 341, 432,
 433, 434
 Lampadius, 358, 393, 470
 Lamy, 380
 Landolt, 425, 449
 Langlois, 388
 Laplace, 153, 439, 456
 Laurent, 261-263, 269-270, 276-278,
 283, 404

- Laurie, 396
 Lauth, 517
 Lavoisier, 4, 128, 149, **151-162**, 258, 361, 367 *et seq.*, 374, 439, 455, 477, 491
 Lebel, 336, 450
 Leblanc, 140, 505, 506
 Leclaire, 503
 Lecoq de Boisbaudran, 380, 445
 Lefèvre, 105
 Lehmann, 480, 485, 486
 Lemberg, 473
 Lemery, **104-105**, 127, 132
 Lenk, 510
 Lerch, 413, 487
 Leuchs, 486
 Lévy, 475
 Libavius, 56, 58, **71-72**, 85, 88
 Lieben, 403, 405, 435
 Liebermann, 401, 413, 518
 Liebig, 231 *et seq.*, 237, 240, 241 *et seq.*, 244, **247-254**, 255, 258, 262, 265 *et seq.*, 279, 289, 305, 365, 370, 394, 405, 408, 411 *et seq.*, 417 *et seq.*, 477 *et seq.*, 512, 524 *et seq.*
 Liechti, 395
 Lightfoot, 517
 Lipprecht, 436
 Linck, 510
 Linnenmann, 447
 v. Lippmann, 415
 Lister, 494
 Littler, 513
 Lommel, 482
 Lossen, 323, 446
 Löwig, 375, 421, 425
 Loysel, 153
 Lubbock, 170
 Ludwig, C., 452, 486
 Lucretius, 9
 Lunge, 390, 504

MACKENZIE, 377
 Macquer, 113, 115, 138, 139, 519
 Magnus, 388, 397, 404, 487
 Malaguti, 465
 Mallet, 380
 Malpighi, 477
 Maly, 486
 Mansfield, 516
 Marat, 153
 Marcet, 525
 Marchand, 191, 346, 362, 370, 382
 Märcker, 514
 Marggraf, 112-113, 133, 135, 136, 138, 142, 514
 Margueritte, 365, 395
 Marignac, 191, 346, 362, 376 *et seq.*, 395, 473
 Mariotte, 103, 477
 Markownikoff, 522
 van Marum, 377
 Matthiessen, 379
 Maxwell, 444
 Mayer, A., 492
 Mayow, 104, 123, 125, 128, 156
 Meissner, 488
 Melsens, 267, 511
 Mendelejeff, **348-349**, 444, 452
 Menschkutkin, 465, 466, 468
 Merck, 435
 v. Mering, 489
 Meyer, E. von, 431
 Meyer, Lothar, **347**, 417, 487
 Meyer, O. E., 444
 Meyer, Victor, 331, 335, 337, 340, 415 418-419, 426, 436, 441, 442
 Michael, 336, 337
 Michaëlis, A., 389, 425
 Michaëlis, W., 513
 Michel, 513
 Miller, 360, 445
 v. Miller, 433
 Milly, A. de, 511
 Mitscherlich, Al., 513
 Mitscherlich, E., 209, 212, 378, 394, 399, 404, 418, 421, 426, **453-454**, 471 *et seq.*
 Mohr, 365, 371
 Mohs, 469
 Moissan, 375
 Moitrel d'Elément, 122
 Moldenhauer, 511
 Mond, 522
 Monge, 153
 Morveau. See Guyton de Morveau
 Mosander, 192, 381, 472
 Mosso, 494
 Muck, 523
 Mulder, 483, 485
 Muller, H., 417
 Müller, M., 512
 Müller, N. J. C., 482
 Müller, v. Reichenstein, 376
 Musculus, 514
 Muspratt, 506
 Muthmann, 395
 van Mynsicht, 71, 94

NÄGELI, 483
 Naquet, 320, 325
 Nasse, O., 486, 493
 Naumann, A., 443, 447
 Naumann, C. F., 472
 Nencki, 485, 493, 494

- Neri, 86
 Neumann, 449
 Neumann, Kaspar, 112
 Newlands, 347
 Newton, 132
 Nicholson, 216
 Nicklès, 387
 Nicolas, 511
 Niépce, 458
 Niépce de St. Victor, 459
 Nietzsche, 413, 517, 518
 Nilson, 380, 381, 442, 447
 Nobbe, 480
 ODLING, 289, 308, 309
 v. Oefele, 301, 421
 Olympiodor, 25, 26
 Olzevsky, 443
 Ortholph von Baierland, 48
 Ost, 410, 435
 Ostwald, 448, 456, 467, 468, 532
 Otto, J., 371
 Otto, R., 371, 421, 494
 PAAL, 413
 Palissy, 61, 82, 477
 Paracelsus, 58, 65-70, 89, 95
 Parkes, 502
 Parmentier, 166
 Pasteur, 409, 450, 454, 492, 520
 Pattinson, 502
 Payen, 514
 Péan de St. Gilles, 465
 Pebal, 388
 Pechiney, 507
 v. Pechmann, 340
 Pélégot, 244, 258, 381, 396, 402, 504
 Pelletier, 166, 389
 Pélouze, 363, 377
 Perkin, W. H., jun., 401
 Perkin, W. H., sen., 340, 407, 412, 424, 451, 516
 Peters, 481
 Petersen, 473
 Petit, 206, 207-208
 Pettenkofer, 487, 490
 Pettersson, 380, 442, 447
 Pfaff, 359, 390
 Pfaundler, 466
 Pfeffer, 452, 482
 Pietet, 443
 Piria, 409, 415
 Planck, 448
 Plato, 38
 Plattner, 364, 379
 Playfair, 293, 395, 428, 501
 Pliny, 7, 8, 12, 13, 14, 15, 17, 19, 20
 Plücker, 453
 Popoff, 323
 Porret, 428
 Porta, 86
 Pott, 112
 Preyer, 487
 Priestley, 118-119, 123, 124, 155, 366, 478
 Pringheim, 482
 Proust, 171, 174-177, 361, 362, 396, 463
 Prout, 190, 191
 Psellus, Michael, 30
 Pseudo-Democritus, 26
 QUINCKE, 452
 RAMMELSBERG, 363, 383, 473, 475
 Ranke, 490
 Raoult, 351, 447-448
 Raschig, 390, 391
 Rathke, 430
 Raymund Lully, 32, 33, 40, 43, 56
 Réaumur, 139, 144
 Redtenbacher, 405
 Rees, 492
 Regnault, 243, 264, 399, 404, 449, 487
 Reich, 330, 504
 Reiset, 397, 487
 Renault, 453
 Rey, 125
 Reynolds, 428
 Rhazes, 29
 Richter, J. B., 171-174, 361
 Richter, Th., 380, 473
 Richters, 513, 523
 Riecke, 337
 Rinman, 138, 470, 501
 Ripley, 34
 Ritter, 458
 Ritthausen, 483
 de la Rive, 378, 502
 Roberts, 515
 Robinson, 506
 Robiquet, 235
 Rochleder, 405, 483
 Romé de l'Isle, 208, 470
 Romer, 511
 Roscoe, 345, 383, 386, 395, 459
 Rose, 323
 Rose, Fr., 395
 Rose, G., 363, 453, 472, 474
 Rose, H., 359, 363, 383, 389, 392, 465, 472
 Rose, Valentin, the elder, 363
 Rose, Valentin, the younger, 144, 363, 470
 Rosenstiehl, 516
 Rossi, 403, 406
 Rouelle, 113, 115, 129, 140, 151
 Rubner, 490

- Rumford, 455
 Runge, 436
 Rutherford, 123

 SACHS, 482, 483
 Sadler, 513
 Sala, Angelus, 58, 77, 94
 Salomon, 415, 514
 Sandberger, 473
 Sarasin, 475
 Sattler, 503
 Saussure, 402, 478, 482
 Schaffner, 506
 Scheele, 119-120, 123, 124, 133, 135,
 142, 144, 145, 155, 458
 Scheerer, 363, 473
 Scheibler, 395, 415
 Scherer, 169, 454, 488
 Schiel, 285
 Schiff, R., 446
 Schischkoff, 431, 510
 Schlossberger, 488
 Schmidt, A., 485, 487, 493
 Schmidt, C., 486, 490, 492
 Schmieder, 61
 Schmitt, 410
 Schneider, 363, 377, 382
 Schönbein, 378, 387, 510
 Schöne, 387
 Schorlemmer, 323, 447
 Schott, 513
 Schrauf, 454
 Schroeder, 446
 Schrötter, 378
 Schürer, 86
 Schützenberger, 388, 485
 Schützenbach, 521
 Schultze, 458
 Schulze, E., 483
 Schwanert, 436
 Schwanhardt, 140
 Schwann, 492
 Seebeck, 379
 Sefström, 383
 Seger, 513
 Selmi, 493, 494
 Sénarmont, 474
 Sendivogius, 61
 Senebier, 478
 Sennert, 58, 77
 Serullas, 391, 404, 428
 Seubert, 385
 Siemens, 523
 Simpson, Maxwell, 406
 Skraup, 333, 432
 Smith, 473
 Solvay, 507
 Soret, 378, 445

 Soubeiran, 390
 Soxhlet, 415, 487
 Spencer, 502
 Stadion, 387
 Staedel, 446
 Städeler, 486, 488
 Stahl, 4, 107-110, 139
 Stahlschmidt, 391
 Stas, 191, 346, 362, 371, 376 *et seq.*
 Steiner, 431
 Stenhouse, 436
 Sterry Hunt, 283
 Stohmann, 457, 480, 481
 Stoney, 445
 Strecker, 409, 486, 488
 Streng, 473
 Stromeyer, 363, 380, 391, 473
 Struve, F. A., 499
 Struve, H., 487
 Suidas, 2
 Svanberg, 472
 Swan, 360, 445
 Sylvius, de le Boë, 58, 77-78, 89, 91
 Synesios, 26.

 TACHENIUS, 58, 77, 79, 89, 90, 95
 Talbot, 359, 445, 458
 Tennant, 377, 508
 Tertullian, 24
 Thaer, 478
 Thales, 8
 v. Than, 393
 Thénard, 144, 188, 225, 226, 369, 393,
 503, 524
 Theophilus Presbyter, 45
 Theophrastus, 7, 16
 Thomas and Gilchrist, 501
 Thomas Aquinas, 30
 Thomsen, J., 397, 456, 457, 464, 467
 Thomson, James, 447
 Thomson, Th., 178, 183, 190, 473,
 531
 Thorpe, 391, 392, 396, 446
 Thurneysser, 61, 70
 Tiemann, 415
 Tilghman, 513
 Traube, 387, 452, 493
 Trommsdorff, 167, 497
 Troost, 475
 Tunner, 501
 Turner, 191
 Turquet de Mayerne, 71, 94
 Tutton, 391
 Tyndall, 459

 VALERIUS CORDUS, 95
 Varrentrapp, 405
 Vauquelin, 165-166, 360, 362, 365,
 379, 381, 470, 524

- Verguin, 516
 Vinzenz of Beauvais, 30
 Vogel, 388
 Vogel, B. H. W., 459
 Voit, 487
 Volhard, 301, 365, 424, 428
 Volkmann, 452
 Volta, 137

 WAAGE, 465 *et seq.*
 van der Waals, 444
 Wagenmann, 521
 Wallach, 402, 424
 Watson, 444
 Watson, W., 141
 Watt, 158
 Weber, 449
 Weber, R., 388, 504
 Weddige, 430
 Wedgwood, 513
 Weidel, 333
 Weldon, 507, 508
 Welter, 388
 Wenzel, 360
 Werner, 470
 Westrumb, 167, 470
 Wiedemann, 453
 Wiegleb, 167, 470
 Wilhelmy, 452, 468
 Will, 413, 415
 Willgerodt, 417

 Williams, 432
 Williamson, 278, 280-282, 283 *et seq.*, 399, 319, 404, 466
 Willis, 104, 127
 Winkler, Cl., 366, 380, 382, 473, 504
 Winterl, 385
 Wischnegradsky, 432, 434
 Wislizenus, 335, 336, 337, 343, 409, 451
 Witt, 426, 518
 v. Wittich, 486
 Wöhler, 237, 240, 252, 254-256, 293, 363, 370, 378 *et seq.*, 392, 394, 420, 428, 472, 475, 525 *et seq.*
 Wolff, E., 480
 Wollaston, 183, 187, 383, 384
 Wray, 144
 Wren, 104
 Wroblevsky, 443
 Wurtz, 164, 278-279, 282, 283, 289, 300, 309, 310, 319 *et seq.*, 327, 392, 400, 403, 404, 412, 423, 430, 443, 493, 528, 532

 ZEISE, 420
 Ziervogel, 502
 Zimmermann, 382, 396
 Zincke, 401, 413, 447
 Zinin, 423
 Zoller, 480, 481
 Zosimos of Panopolis, 24, 25

INDEX OF SUBJECTS

The figures in thick type refer to those pages upon which subjects are treated in detail or points of special importance are recorded.

- ABSORPTION** of gases by water, 122
Academia Cæsarea Leopoldina, 99
del Cimento, 99
Académie Française, 165
Royale, 99
 Academies, Spanish, 28
 Acetaldehyde, 411
 Acetic acid, 19, 56, 93, 144, 521
 acid, constitution of, 266, 267, **298**, 405
 acid (glacial), 144
 acid, synthesis of, 338
 Aceto-acetic ether, 334, 406, 414
 Acetone, 413
Acetyl, 296
 theory (Liebig), 243-244
 Acetylene, 401
 Acid amides, 407
 anhydrides (Gerhardt), 282, 407, 408
 chlorides, organic, 407
 nitriles, 295, 429
 theory of Lavoisier, 157
Acides, 161
 Acids, 50, 88
 constitution of (Berzelius), 220
 constitution of (Davy), 228
 constitution of (Liebig), 230-231
 nomenclature of (Lavoisier), 161-162
 organic, 144, 405
 Acrylic acid, 406
 Actinometry, 459
Ætherium, 242
Ætherin, 239
 theory, the, 239
Affinitas, 130
 Affinity, chemical, 130
 -coefficients, specific, 467 *et seq.*
- Affinity, degrees of, 316
 determinations of, 460, 467 *et seq.*
 doctrine of, 460-468
 doctrine of (Bergman), 460
 doctrine of (Berthollet), 461 *et seq.*
 doctrine of, its latest development, 467
 simple elective, 130
 tables of, 114, 131
 units of, 323
 views as to its causes, 130-132
Affinivalenten, 320
 Agricultural chemistry, 478-481
 Air, composition of atmospheric, 117, 123-124, 125, 137
 Albumens, vegetable, 483
 animal, 485
 Alchemistic period, the, 21-62
 Alchemists, practical-chemical knowledge of the, 44 *et seq.*
 Alchemy among the Arabians, 27 *et seq.*
 at the European courts, 34, 63
 decay of, 57 *et seq.*
 during the last four centuries, 57-62
 general notes upon, 2
 in Egypt, 23, 37
 in the Christian countries of the West, 30 *et seq.*
 its relation to the Platonist philosophy, 22-24
 origin of, 21, 23 *et seq.*
 problems of, 31 *et seq.*
 relations of, to astrology, 25
 theories of, 37 *et seq.*
 Alcohol, constitution of, 241
 meaning of the word, 95
 preparations from, 520
 Alcoholometry, beginnings of, 143-144

- Alcohols, 402-404
 constitution of (Kolbe), 299
 secondary, 300
 tertiary, 300
- Aldehydes, 411-412
 constitution of (Kolbe), 299
- Aldol, 412
- Alexandrian Academy, the, 21, 26, 37
- Algaroth, powder of, 91
- Alizarine, 518
- Alkahest*, 51, 93
- Alkali*, 50
- Alkali metals, discovery of the, 185, 224
 waste, 506
- Alkalies, decomposition of, 224
- Alkalimetry, 365
- Alkaloids, derivatives of pyridine, etc., 434
 tests for, 371
- Alkarsine*, 246
- Alkyl-pyridines, 432
- Allo-isomerism, 336
- Allotropy, 238, 377, 454
- Alum, 18, 51, 87
 earth, confounding of this with lime, 90
- Aluminium, 380, 502
 bronze, 503
 chloride, syntheses with, 400
- Amalgamation processes for obtaining silver, 85
- Amides, 407
- Amidines, 425
- Amido-acids, 409
 constitution of, 300
- Amine bases (Wurtz, Hofmann), 278
et seq.
- Amines, 422
- Ammonia as a type, 280, 284
 gas, discovery of, 122
 salts of, 89, 90
- Ammonia soda*, 507
- Amygdalin, 252
- Analysis, development of, 358 *et seq.*
- Analysis*, introduction of the word by Boyle, 134
- Analysis, legal-chemical, 370, 371
 of articles of food and drink, 371
 of gases, 366
 of inorganic substances, 358 *et seq.*
 of organic substances, 367
 qualitative, 79, 92, 133 *et seq.*, 358, 367
 qualitative, system of, 359
 quantitative, beginnings of, 136
 quantitative, system of, 360 *et seq.*, 369
- Analysis, technical, 371
 volumetric, 364
- Ancients, practical-chemical knowledge of the, 10 *et seq.*
- Anhydrides of organic acids, 407
- Aniline, 422, 516
 black, 517
- Aniline blue, 517
 colours, 516 *et seq.*
 green, 517
 red, 516
 violet, 516
- Annalen, Chemische*, 169
der Chemie und Pharmazie, 250, 532
der Physik, 169
der Physik und Chemie (Poggendorff), 169, 532
- Annales de Chimie*, 166
de Chimie et de Physique, 533
- Anthracene, 401
- Antimony, 47, 53, 376
 compounds, organic, 425
 pentachloride, 392
 pills, 91
 preparations, 53, 90, 91
- Antiphlogistic system, the, 149 *et seq.*, 160, 167 *et seq.*, 170
- Antiseptics, 146, 495
- Apothecaries' shops, 48, 88
- Apparatus for collecting gases, 122
- Aqua fortis*, 50
regia, 51
vitæ, 56
- Arcana*, 68
- Archeus*, 68, 75
- Aromatic compounds, meaning of the name, 330
 compounds, theory of, 328 *et seq.*
- Arsenic, 54, 91, 377, 390
 compounds, 54, 91, 389
- Arsines, etc., 425
- Asparagine, constitution of, 301
- Assimilation in plants, 458, 481 *et seq.*
- Asymmetric carbon atom, 450-451
- Atomic heat, 207
 hypothesis, the, 170, 179
 theory (Dalton's), 177-182
 theory, further development of, 182 *et seq.*
 theory, further development of, by Berzelius, 198 *et seq.*
 volume, 445-446
 weight determinations, 362
 weight (Laurent), 276-278
 weight system of Berzelius and the opposition to it, 210-212, 214

- Atomic weight tables (Berzelius), 205, 211
weight tables (Dalton), 181
weights (Berzelius), 199 *et seq.*
weights, correction of, 349
weights, deduction of, by Cannizzaro, 317
weights (Dumas), 212 *et seq.*
weights (Gerhardt), 274
weights of the metalloids, 375 *et seq.*
weights of the metals, 378 *et seq.*
weights, periodic arrangement of, 346 *et seq.*
weights, relative (Berzelius), 198 *et seq.*, 205, 210
weights, relative (Dalton), 179 *et seq.*
weights, relative (Thomson), 183
weights, uncertainty as to them generally, 214
weights, uncertainty as to those of the metals, 205, 208
- Atomicity, 316, 320
- Atoms, conception of, 170
of various orders (Dalton), 180
spacial arrangement of, 238, 325, 335-337
- Atropine, 435
- Aurum potable*, 52
- Azo-compounds, 425-426
- Azo-dyes, 426, 517
- Azo-imide, 390
- BACKWARD substitution, 267
- Bacteriology, 494
- Balance, importance of the (Lavoisier), 152
- Barium, 379
- Bases, designation of, by Lavoisier, 162
- Basicity, law of (Gerhardt), 272
of acids, 230 *et seq.*
of acids, criterion of, 231
- Beet sugar industry, 514
- Beetroot sugar, 113, 141, 514
- Benzene, constitution of (Baeyer), 330
constitution of (Claus), 330
constitution of (Kekulé), 328
constitution of (Ladenburg), 330
derivatives, isomerism of, 329, 332 *et seq.*
hexagon formula of, 328
- Benzil-dioximes, isomerism of, 337
- Benzin*, 399
- Benzoic acid, 94, 407, 521
- Benzoyl the radical of benzoic acid, 240
- Berichte der Deutschen chemischen Gesellschaft*, 534
- Berlin Academy, the, 99
blue, 139
- Beryllium, 379, 394
- Biblical characters as alchemists, 24
- Bile, chemistry of the, 486
- Bismuth, 47
preparations, 91
- Bitter almond oil, 240
salt, 146
- Blast-furnace process, the, 501
- Blood, chemistry of the, 486, 487
gases, 487
- Blowpipe, 136, 359, 364, 470
- Boiling point, laws regulating the, 446
- Bone charcoal for sugar refining, 515
- Bones, constituents of, 485
- Boron, 377, 378, 392
methide, 438
- Brass, 16, 139
- Bromine, 375, 508
- Butylene, 399, 401
- CACODYL compounds, 246
compounds, constitution of, 305-306
- Cadaverine, 494
- Cadmium, 380
- Cæsium, 379, 393
- Calcination of the metals (Lavoisier), 154 *et seq.*
- Calcium, 379
- Capillarity, 451
- Carbamines, 429-430
- Carbinols, secondary and tertiary, 403
- Carbohydrates, 414-415, 485
- Carbon, 377, 392
as a constituent of organic compounds, 233
bisulphide, 393
compounds, saturated and unsaturated, 326 *et seq.*
determination of, 368
- Carbonate of ammonia, 52
- Carbonic oxide, composition of (Dalton), 179
acid, 76
acid, composition of, 361, 368
acid, composition of (Black), 116
acid, composition of (Dalton), 179
- Carboxylic acids, 405-411
acids, constitution of (Kolbe), 299
- Carbures*, 162
- Cellulose, 513
- Cement, 513
copper, 46
- Ceramic art, the, 17, 139, 512
- Cerium metals, the, 381
- Chelidonic acid, 435
- Chemia*, first use of the word, 25
- Chemical compound, different from a mixture, 128

- Chemical compound, meaning of, 10, 41-42, 128
 compound, old ideas regarding, 10, 42
 industries, the great, 504
 nomenclature (Lavoisier), 161
 notation (Berzelius), 222
 notation (Dalton), 182
- Chemistry, agricultural and physiological, 477-497
 analytical, in the modern period, 358-372
 analytical, in the phlogistic period, 133
 analytical, its development by Boyle, 103
 antiphlogistic, 151 *et seq.*
 applied, in the phlogistic period, 138
 geological, 399 *et seq.*
 in ancient times, 5
 inorganic, 373-397
 its meaning at different periods, 2 *et seq.*
 meaning and origin of the word, 2
 mineralogical, 469-476
 organic, 398-438
 pharmaceutical, 47, 145
 physical, 439-468
 pneumatic, founded by van Helmont, 75
 pneumatic, its further development, 116-118, 121-126
 tasks of, in the various ages, 1
 technical, in recent times, 498-523
 technical, in the iatro-chemical age, 84
 technical, in the phlogistic period, 138
- Chili saltpetre, 508
Chimie Française, 164
Chloraldehyde, 243
 Chloride of lime, 393, 507
 Chlorimetry, 365
 Chlorine, 375
 discovery of, 120, 141
 recognised as an element, 185, 225, 226
 supposed composition of, 226
 the name, 226
 Chlorophyll, 482
 Chromium, 381
 Chrysamine, 517
 Chrysoidine, 517
 Cinabar, 92
 Cinnamic acid, 407
 Circular polarisation, 450
 polarisation, magnetic, 451
 Citric acid, 145
- Classification of organic compounds, 262-263, 272, 284
 Coal-tar colour industry, 515 *et seq.*
 products from, 521
 Cobalt, 141
 -ammonia compounds, 395
 blue, 86
 Cocaine, 435
 Cohesion, 462
 Coins of alchemistic gold, 60
 Colcothar, 55
 Colloids, 452
 Combining proportions, proof that these are constant, 175
 weights (Gmelin), 214, 274, 278
 Combustion according to Stahl, 108
 according to Hoffmann, 110
 according to Mayow, 104
 correct explanation of, by Lavoisier, 155-157
 -ladder (Gerhardt), 273
 phenomena of, 392
 theory of (Lavoisier), 155 *et seq.*
 Composition of substances according to Becher, 106-107
 Compounds, atomic, 321
 classification of, 161
 molecular, 321
Comptes Rendus, 533
Condensation, 412
 meaning of, 339
 Congo red, 517
 Conine, 434
 Constitution, chemical (Berzelius), 219, 238
 chemical (Gerhardt), 286-288
 Constitutional formulæ (Kolbe), 299-300
 Copper, 14, 46, 85
 oxide for organic analysis, 369
 vitriol, 18
 Copulæ, 266, 267, 294 *et seq.*
 Copulated or conjugated compounds, 266, 271, 294 *et seq.*, 306
 Copulation, a consequence of saturation-capacity (Frankland), 307
 meaning of, 266, 304
 Corpse alkaloids, 493, 494
 Corpuscular theory (Berzelius), 204
 theory (Boyle), 103
 Creatine, 424
 Criticism, importance of (Kolbe), 534
 Crotonic aldehyde, 412
 Crystalline form, its connection with composition, 208-210
 Crystallography, 470
 Cumarone, 437
 Cyanic acid, 237, 430

- Cyanamide, 428
 Cyanogen, 235
 compounds of, 427-431
 iron compounds of, 394, 395
 polymers of, 430
 Cyanuric acid, 430
- DEDUCTION, significance of (Aristotle),
 7, 11
 Dephlogisticated air (oxygen), 124
 Dephosphorisation of iron, 501
 Desmotropism, 335
 Diagonal formula of benzene, 330
 Diazo-compounds, 423, 426
 Diffusion, 451
 Diffusion process for sugar, 515
 Di-ketones, 413
 Dissociation, 443
 Distillation, 5, 19, 48, 87, 530
 Distilleries, 87
 Di-sulphones, 421
 Docimacy, 364
 beginnings of, 46
 Dualism (Berzelius), 220, 230 *et seq.*
 fight against, 224 *et seq.*, 259 *et seq.*
 overthrow by Unitarism, 265 *et seq.*
 Dyeing, 18, 47, 87, 139, 519
- EARTHENWARE, 17, 47, 86, 512
 Earths (Becher's), 106
Eau de Javelle, 508
Effect, chemical, 462
 Elasticity, 462
 Electro-chemical theory (Berzelius), 217
 et seq., 464
 -chemical theory (Davy), 216
 Electrolysis, 216 *et seq.*, 229, 452
 Faraday's law of, 452, 464
 of salts of fatty acids, 296
 Electrolytic law (Faraday), 215
 Element, meaning of (Boyle), 102, 127
 meaning of (Lavoisier), 160
 Elements, Aristotle's four, 9
 classification of, 161
 discovery of new, 121, 374 *et seq.*
 discovery of supposed new, 385, 386
 old views regarding, 8
Elixir, 40
Enzymes, 492
 Eosin dyes, 517
 Equilibrium, dynamical, 466
 statical, 466
 Equivalents, 214
 electro-chemical, 215
 first table of, 173
 of the elements (Gerhardt), 273 *et seq.*
 of the elements (Laurent), 276 *et seq.*
 of the elements (Wollaston), 187
- Ester*, 404
 Ether as a fifth element, 10
 -acids, 404
 constitution of, 242
 from alcohol, 95, 144
 Ethers, compound, 404
 mixed, 281, 404
 simple, 403
 varieties of, 144
 -Williamson), 281
Ethyl, 241
 theory, the, 241, 242
 Ethylene, composition of (Dalton), 178
 oxide, 403
 Experimental lectures, 250, 524 *et seq.*
 methods, development of, by Boyle,
 100-101
 Explosives, 509-510
 Extraction of silver, 502
- FAMILIES, natural, of the elements,
 347
 Fats, 95, 145, 485
 Fatty acid series, structure of com-
 pounds of the, 326-327
 Fatty acids, 405
 acids, constitution of (Kolbe), 296
 et seq.
 Fermentation, 491 *et seq.*
 former views regarding, 144
 processes, 520
 significance of (v. Helmont), 75
 theories, 491 *et seq.*
 Ferments, organised and unorganised,
 492
 Ferrocyanogen, 427
 Filter papers, incineration of, 362
 Fixation of carbonic acid by alkalies
 (Black), 121, 122
 Fixed air (Black), 116
 Flame colorations (Marggraf), 135
 Flesh, chemistry of, 488
 Fluorine, 375
 Formic acid, 144
 Formulæ, graphical (Kekulé), 325
 Four-volume formulæ, 276
 Fulminate of mercury (Kekulé), 290
 Fulminic acid, 430, 431
 acid, isomerism with cyanic acid,
 237
 Fulminuric acid, 431
 Fumaric acid, 406
 Furfurane, 331, 437
 Furfurol, 436
 Furnace gases, 501
- GALL apples, juice of, 94
 Gallium, 380

- Galvanic current used in analysis, 364
 Galvano-plastic process, 502
 Gas analysis, 366
 analysis, beginnings of, 137
 analysis, technical, 366
Gas sylvestre, 76
 Gases, 75-76
 absorption of, 444
 Gases, discovery of many, by Priestley, 118, 122
 kinetic theory of, 444
 liquefaction of, 443-444
 Gastric juice, 486
 Geometrical isomerism, 335
 Germanium, 382
 German silver, 502
 Glass, manufacture of, 16, 47, 86, 139, 512
 Glauber's salt, 89
 Glucose, 414, 415
 Glucosides, 415
 Glycerine, 145, 310, 511
 Glyceryl, 310
 Glycogen, 488
 Glycol, 310
 Glycols, 403
 Glyoxal, 412
 Gold, 13, 45, 502
 amalgamation with silver, 14
 compounds of, 396
 determination of its atomic weight, 396
 separation from silver, 14, 85
Gradverwandschaft, 316
 Great chemical industries, the, 504
 Groups of elements, 344, 347-348
 Gun-cotton, 510
 Gunpowder, 509
 Gypsum, 90
 HALOGEN carriers, 417
 derivatives of hydrocarbons, 416 *et seq.*
 Halogens, the, 225, 226, 375
 compounds of the, 226, 386 *et seq.*
Handwörterbuch der Chemie, 251, 532
 Heat, latent (Black), 116
 latent (Lavoisier and Laplace), 153
 nature of (Lavoisier), 153, 159, 160
 of combustion, 455 *et seq.*
 of fusion, 448
 of solidification, 447, 448
 specific, its relation to atomic weight (Dulong-Petit), 207
 Heating materials, 522
 Hermetic, 24
 art, 24
 Heterologous compounds, 285
 Hexagon formula of benzene, 328
Historia naturalis of Pliny, 7
 History of chemistry, alchemistic period, 21-62
 of chemistry, from Lavoisier till now, 148-535
 of chemistry, iatro-chemical period, 63-96
 of chemistry in early times, 7-20
 of chemistry, phlogistic period, 97-148
 Homologous compounds, 285
Hydracides, 227
 Hydrates, discovery of those of the metallic oxides, 176
 Hydrazine, 390
 Hydrazines, 427
 Hydrocarbons, 399-402
 synthesis of, 338, 400
 Hydrochloric acid, 51, 88
 acid gas, 122
 acid, manufacture of, 507
 Hydrocyanic acid, 145, 427
 Hydrofluoric acid, 140, 387
 Hydrogen, 76, 117, 375
 acids, 228
 acids, theory of the, 228 *et seq.*, 231
 a constituent of organic compounds, 233
 as a type, 283 *et seq.*
 as the primary material, 190
 determination of, 367 *et seq.*
 Hydroxylamine as a reagent, 419
 Hygiene, relations of, to chemistry, 490 *et seq.*
 IATRO-CHEMICAL doctrines of Paracelsus, 67 *et seq.*
 doctrines of Sylvius, 78
 doctrines of van Helmont, 74-75
 period, the, 63 *et seq.*
 Iatro-chemistry, general notes upon, 3
 problems of, 64
 Iatro-chemists, practical-chemical knowledge of the, 84 *et seq.*
 Illuminants, 522
 Indigo blue, 518
 Indium, 380
 Indole, 437
 Induction, photo-chemical, 459
 Inductive methods, the gradual appreciation of, 32, 63, 98
 Industries, the great chemical, 504
Inflammable air, 117
 Inorganic compounds, structure of, 325, 343 *et seq.*

- Inorganic compounds, systematising of, 343, 348, 349
- Instruction, growth of chemical, 524-535
systematic chemical, 195, 249, 255, 524 *et seq.*
technico-chemical, 500
- Iodine, 375, 508
recognised as an element, 226
- Iridium, 384
- Iron, 15, 138, 501
industry, the, 501
ores used in olden times, 15
- Isocyanides, the, 429
- Isogonism, 454
- Isologous compounds, 285
- Isomerism, geometrical, 335-336, 451
of position, 333
- Isomers, 236-239
structural-chemical, interpretation of, 332 *et seq.*
- Isomorphism (Mitscherlich), 209, 453 *et seq.*, 471
appreciation of its value by Berzelius, 209
polymeric, 454
- Iso-nitroso compounds, 419
- Jahresberichte der Chemie* (Berzelius), 196
der Chemie (Liebig), 251, 534
- Journal, Allgemeine J. der Chemie*, 169
de Physique, 165
für praktische Chemie, 294, 534
- Journal of the Chemical Society, 533
of the Society of Chemical Industry, 500
- Journals, chemical, 532-534
- Juices, the animal, van Helmont's views upon, 75
- Kermes* mineral, the, 91, 146
- Ketones, 412-414
constitution of (Kolbe), 299
- Ketonic acids, 414
- LABORATORIES, establishment of, 500, 525 *et seq.*
for instruction, 525 *et seq.*
for students, 195, 249, 293, 524 *et seq.*
improvement of, 528-529
instruction in, 524 *et seq.*
recent, 528 *et seq.*
technical, 530
- Laboratory arrangements in old times, 5
- Lactic acid, 145
acid, constitution of, 300
- Lactic acids, the, 409
- Lactones, 410
- Lana philosophica*, 54
- Lapis infernalis*, 92
- Law of Boyle and Mariotte, 103
of solidification (Raoult), 447
- Lead, 15
acetates of, 56, 93
sugar of, 56, 93
- Lecture experiments, 525
- Light, chemical action of, 458, 482
refraction of, 450
- Lime, 19
- Linking bars (Couper), 316
of atoms, 313, 324
of atoms (mode of), 314 *et seq.*, 318
- Literature, chemical, 531 *et seq.*
technico-chemical, 500
- Lithium, 379
- MADDER red, 518
- Magisterium*, 40
- Magnesium, 379
- Magnetism of chemical compounds, 453
- Maleic acid, 406
- Malonic acid, 406
- Manganese, 141
black oxide of (Scheele), 120
compounds, 394
- Manuals of chemistry, 5, 72, 105, 111, 115, 146, 154, 186, 195, 255, 294, 531, 532
- Marsh gas, 76
gas as a type, 291
- Mass-action, 461-462, 465 *et seq.*
- Matches, 510, 511
- Materia Medica* (Dioscorides), 7
prima, 42, 62
- Matière de chaleur* (Lavoisier), 160
- Meconic acid, etc., 435
- Medicines of various orders, 29, 40
- Mellitic acid, 407
- Mercaptan, 420
- Mercury, 16, 46
as a constituent of metals 38, 40
salts of, 53, 92
- Mesitylene, 333, 401
- Metabolism, animal, 489 *et seq.*
vegetable, 481 *et seq.*
- Metalepsy, 260
- Metallic chlorides, 88
compounds, 393 *et seq.*
oxides, 54
salts, 52
- Metallo-organic compounds, 437
- Metallurgy, furtherance of, by Agricola, 81, 85

- Metallurgy in the alchemistic period, 45
 in the phlogistic period, 138
 of the Ancients, 12
 of recent times, 501-502
- Metals, ancient views as to their production, 13
 colouring of, 37
 derivation and meaning of the word, 12
 duplication of, 25
 ennobling of, 22, 34, 43
 increase in weight on calcination, 125-126
 nature of (Boyle), 127
 nature of (Stahl), 127
 old chemical theory of the, 38
 oldest knowledge of the, 12 *et seq.*
 supposed composition of, 29, 38, 41
 transmutation of, 25, 26, 43
- Metamerism, 238
- Methane, composition of (Dalton), 178
- Methods, analytical, 360 *et seq.*
 technico-chemical, 371
- Methylene blue, 518
- Microscope, its application to chemical researches, 113, 135
- Milk, chemistry of, 487
- Mineral colours, 503
- Mineral potash*, 509
- Mineral system, the chemical (Berzelius), 193
 tanning, 520
 waters, analysis of, 135
 waters, artificial, 499
- Mineralogical chemistry, 469-476
 systems, 472
- Minerals, analysis of, 359 *et seq.*
 artificial production of, 473-476
 classification of, 470 *et seq.*
 classification of (Bergman), 119
 Klaproth's researches on, 168
 nomenclature of, 472
 tests for, 136
- Mixture weights (Gmelin), 191
- Molecular heat, 457
 weight (Laurent), 277
 weight, determination of, 351
- Molybdenum, 381, 395
- Mordants, 87, 139, 519
- Morphotropism, 454
- Mortar, 513
- Mosaic gold, 139
- Multiple types, 284
- Mustard oils, the, 423
- NAPHTHALENE, 401
- Narcotics, 495
- Nestorians, the, 27
- Neurine, synthesis of, 424
- Neutralisation, law of (Richter), 172
- Nickel, 141, 502
- Niobium, 383, 396
- Nitric acid, 51, 88, 140
 acid, composition of, 117, 118, 142
- Nitriles, 429
- Nitro-benzene, 418
 -ethane, 418
 -glycerine, 510
 -methane, 418
 -prussides, 428
- Nitrogen, 374
 compounds, inorganic, 389 *et seq.*
 compounds, organic, 422, 435-436
 discovery of, 123, 370
 estimation of, 258, 370
- Nitroso-compounds, 418-419
- Nitrous oxide, 185
- Nitrum*, 17, 52
- Nomenclature, chemical (Lavoisier, etc.), 161-163
 chemical (Berzelius), 221
- Notation, chemical (Dalton), 182
 chemical (Berzelius), 222-224
- Nuclei (Laurent), 262
- Nucleus theory (Laurent), 261-263
- Nutrients, 490
- Oelsüss*, 145
- Oils, ethereal*, 402
- Organic chemistry, development of, up to 1811, 232 *et seq.*
 compounds, chemical behaviour of, 341
 compounds, constitution of (Kolbe), 298 *et seq.*
 compounds, modes of decomposing, 342
 compounds, structure of, 326 *et seq.*
 compounds, their arrangement in series, 284
 substances, qualitative composition of, 233, 367
 substances, quantitative composition of, 233, 234, 368 *et seq.*
- Organo-metallic compounds (Frankland), 260 *et seq.*, 437
 -metals (Frankland), 304 *et seq.*, 437
- Orthrin*, 240
- Osazones, 415
- Osmium, 384
- Osmose, 452
- Oxalic acid, 145, 521
 acid, synthesis of, 406
- Oxalines, 424
- Oxidation theory (Lavoisier), 157
- Oxy-acids, 409-410

- Oxy-acids, constitution of, 300
Oxydes, 162
 Oxygen, 374
 acids, theory of the, 157, 221
 as the centre-point of the chemical system, 160, 198, 204
 compounds of, 387 *et seq.*
 discovery of, 118, 120, 123-124, 152
 law, the (Berzelius), 200
 Ozone, 377
- PALLADIUM, 383, 384
 Paper manufacture, 513
 Paraffin industry, the, 522
 Para-rosaniline, 516
 Pathology, its relations to chemistry, 494 *et seq.*
 Pattinson process, the, 502
 Pepsine, 486
 Peptones, 486
Periods of the elements, 347
 Periods, the various chemical, 1 *et seq.*
 the various chemical, their characteristics, 1-4
 Petroleum industry, the, 522
 Pharmacy, early development of, 69
 in the iatro-chemical age, 87 *et seq.*
 its relations to chemistry, 495 *et seq.*
 Phenacetine, 495
 Phenanthrene, 401
 Phenols, 404
 Phenyl-hydrazine as a reagent, 419
 -propionic acid, 407
 Philosopher's stone, the, 26, 30, 40, 42 *et seq.*
Philosophical Transactions, the, 99, 533
 Phlogistic period, general history of the, 100-121
 period, merits of the, 147-148
 period, special history of the, 121-148
 period, the, 97-148
 system, fall of the, 158
 Phlogisticated air (nitrogen), 124
 Phlogiston an element, 127
 assumption of the hypothetical, 107
 its identification with hydrogen, 118, 122, 150
 theory, beginnings of (Becher), 107
 theory, development of (Stahl), 107 *et seq.*
 theory, general notes on the, 4
 theory, its value, 109
 Phlogistonists, practical-chemical knowledge of the, 133 *et seq.*
 Phloroglucin, 334
 Phosgene, 392
 Phosphines, etc., 425
- Phosphoric acid, 142
 acids, basicity of, 230
 Phosphorous oxide, 391
 Phosphorus, 141, 376, 378, 389 *et seq.*
 manufacture of, 511
 pentafluoride, 322
 Photo-chemistry, 458
 Photography, 458
 Phthaleins, 517
 Phthalic acid, 521
 Physical chemistry, history of, 439-468
 chemistry, its general significance, 350
 methods, application of, to chemistry, 147
 Phyto-chemistry, 481 *et seq.*
 Picric acid, 418, 519
 Pigments of antiquity, the, 18
 Piperidine, 434
 Plant-nutrients, 479 *et seq.*
 Platinum, 141, 383, 502
 bases, 397
 Poisons, methods for detecting, 370-371
 Polarity, electric, of atoms, 217-218
 Polybasic acids, doctrine of the, 230-232
 Poly-carboxylic acids, 406
 -sulphonic acids, 421
 Polymerism, 238
 Polymorphism, 454, 472
 Porcelain, 17, 62, 139, 513
 Position, determination of chemical, 333
Position-isomers, 333
 Potash, distinction of, from soda, 140
 salts, 89
 Potashes, 17, 18
 Potassium, 225, 378, 393
 carboxide, acids from, 413
 Potter's art, the, 17, 47, 86, 512
 Powder, smokeless, 510
 Precious stones, artificial, 14, 512
 Prediction of alcohols, 300
 of new elements, 349
 Preparations, chemical, 18
 medicinal, 18, 19
 officinal, 146
 technico-chemical, 139
 Pressure, critical, 444
 Primary material (Crookes'), 350
 Princes as patrons of alchemy, 60
Principe oxygine, 157
 Prism-formula of benzene, 330
 Progression, law of (Richter), 173
Proin, 240
 Proportional numbers of the elements (Davy), 187
 Proportions, doctrine of chemical, 170 *et seq.*

- Proportions, law of constant, 171, 175-176
 law of multiple, 178 *et seq.*, 183, 193, 198, 234
Protyle, 350
 Prout's hypothesis, 189-191, 346
 Prussian blue, 139
 Prussic acid, 145, 427
 Pseudo-compounds, 334
 Ptomaines, 371, 493, 494
 Ptyalin, 486
 Putrefaction, 491 *et seq.*
 bases, 371, 493, 494
 Pyridine, constitution of, 330, 331, 432
 bases, 431 *et seq.*
 Pyroligneous acid, 144
 Pyrrol, 331, 435
- QUALITATIVE tests for substances, 134-135
 Quantitative researches, period of, 149 *et seq.*
 Quick vinegar process, the, 521
 Quinoline, 330, 431 *et seq.*
 derivatives of, 431 *et seq.*
 Quinone, 333
Quinta essentia, 10
Quintessence, 93
- RACEMIC acid, 194, 238
 Radical theory, first steps towards the, 235, 236
 theory, the newer, 294-300
 theory, the older, 239-247
Radical vinegar, 93
 Radicals, assumption of oxygenated, 240-241
 chemistry of compound, 244
 compound, 233, 235
 polyatomic, 284
 the clearer definition of, 245
 variability of, 243
 Reaction, time-rate of, 466, 467
 Reactions, specific, of organic compounds, 342
 Reagents, 134-136
 Reciprocal reactions, 466
 Refraction-equivalents, 450
 Replaceable value of elements, 305, 308, 309
 Residues, theory of (Gerhardt), 270
Résidu, 270
 Rhodium, 384
 Rosaniline, 516
 Royal Society, the, 99
 Rubidium, 379, 393
 Ruby glass, 86, 139
 Ruthenium, 384
- SACCHARIMETRY, 514
Saccharine, 515
 Safety lamp, the Davy, 185
 Safranines, 518
Sels, 49
Ammoniacum, 502
mirabile, 89
 Salicylic acid, 410, 495, 521
 Saliva, chemistry of the, 486
 Salmiac, 51, 89, 141
 Salt as a constituent of the metals, 41
 Salt, its meaning in the alchemistic and iatro-chemical periods, 49, 79-80
 Rouelle's definition of a, 115, 129
 Salts, constitution of (Berzelius), 220
 constitution of (Liebig), 231
 nomenclature of (Lavoisier), 162
 Saltpetre, 51, 89
 Saturation-capacity, assumption of a constant (Kekulé), 320 *et seq.*
 -capacity of the elements (Frankland), 304 *et seq.*, 307, 318
 Scandium, 381
 Scarlet dyes, 517
 Schweinfurt green, 503
 Secretions, animal, 486
 Seignette salt, 94
 Selenium, 376
 compounds, inorganic, 389
 compounds, organic, 420
 Silicates, solution of, 362
 Silicon, 378
 alkyl compounds of, 438
 compounds of, 392
 Silver, 14, 46, 53, 85, 502
 mirrors, 253
 nitrate of, 53, 92, 503
 salts, 92, 503
 Smalt, 87
 Soap, 17, 140
 manufacture, 511
 Societies, learned, 98-99
 Soda, 17, 18
 artificial preparation of, 114, 140, 505
 industry, the, 505-507
 Sodium, 225, 378, 393
 Solar spectrum, chemical action of the, 458
 Solidification, Raoult's law of, 447
 Spagirie art, the, 24
 Special history of modern chemistry, 353 *et seq.*
 Specific heat, 449
 volume, 445
 Spectrum analysis, 359, 444
 Spirit-lamp (Berzelius), 362
 Spirit of wine, 48, 56, 95, 143

- Spirit of wine, sweetening of, 56
Spirits, manufacture of, 520
Spiritus, 49
 fumens Libanii, 92
 ignis-aëreus (Mayow), 123
 Mindereri, 90
 salis, 49, 51
Starch, 19, 415, 482, 514
Stassfurt salts, 509
Statique Chimique (Berthollet), 164, 175, 461
Stearine candles, 511
Steel, 86, 138, 501
Steren, 446
Stibines, etc., 425
Stoichiometry, 361
 founding of (Richter), 174
 meaning of the term (Richter), 174
Strontium, 185, 379
Structural formulæ, 316, 318, 325 *et seq.*
Structure, chemical, 318, 325
 theory, beginnings of the, 314 *et seq.*
 theory, development of the, 324 *et seq.*
Sublimate, 53
Substitution, first observations upon, 260
 -form (Gerhardt), 271
 laws of (Dumas), 260
 theory of, 261-265
Succinic acid, 94, 406
Sugar, 19
 from beet juice, 113, 141, 514
 estimation of, 365
 refining of, 514-515
Sulphines, 421
Sulphite cellulose, 513
Sulphone-ketones, 421
Sulphones, 421
Sulphonic acids, 421
Sulphonal, 493
Sulphoxides, 421
Sulphur, 378
 as a constituent of the metals, 38, 40
Sulphur auratum, 91
Sulphur, compounds of, 388
 compounds of the alchemistic period, 54-55
 compounds, organic, 420-422
Sulphur ether, 144
Sulphur, milk of, 55
Sulphur philosophorum, 51
Sulphur, recovery of, from alkali waste, 506
Sulphures, 162
Sulphuric acid, 50, 88, 140
 acid, anhydrous, 505
 acid, fuming, 140
 acid, manufacture of, 504-505
Sulphurous acid gas, 122
 acid, practical utilisation of, 505
Symbols, chemical, 182, 222
 with bar across them, 222-223
Synthesis of organic compounds, 338 *et seq.*
System, chemical, of the minerals, 471
 natural, of the elements, 344
 periodic, 346 *et seq.*
Système unitaire, 286

Tables des Rapports, 114
Tables of affinity, 114
Tanning, 519
Tantalum, 383, 396
Tar, products from, 521
Tartar, 93, 94
 emetic, 71, 94
Tartaric acid, 145
 acid, isomeric with racemic, 238
Tartarus, 68, 94
Taurine, constitution of, 301
Tautomerism, 334-335
Tellurium, 376
 compounds, inorganic, 388
 compounds, organic, 420
Temperature, critical, 444
Tension-series of the elements (Berzelius), 218-219
Tetramethylene, 401
Text-books of chemistry. See Manuals
Thallium, 380, 394
Thermo-chemistry, 455-458, 464
Thiacetic acid, 420
Thiamides, 424
Thiophene, 331, 436, 437
Thomas slag, 501
Thorium, 382
Tin, 15, 396
 perchloride of, 91, 92
Tincture, 38
Tinder, chemical, 511
Titanium, 382, 396
Titrimetry, 189, 364
Toluidines, 516
Trichloroacetic acid, 259, 263, 266, 267
Trimethylene, 401
Triphenyl-methane, 401, 402
 -phosphine oxides, isomeric, 322
Tropæolines, the, 517
Tungsten, 381, 395
Turpentine, oil of, 19
Turpeth mineral, 92
Two-volume formulæ, 276
Type theory, the newer (Gerhardt), 283, 288
 theory, the newer (Kekulé), 289 *et seq.*
 theory, the older (Dumas), 263-265

- Types, chemical, 264
 condensed, 284, 290
 duplicated, 284, 290
 (Gerhardt's), 283 *et seq.*
 mechanical, 264
 mixed, 289 *et seq.*
 real, as opposed to formal, 289, 302
- ULTRAMARINE, 511, 512
- Unitarism, beginnings of, 224 *et seq.*
 development of, 259 *et seq.*
- Universal medicine, 43
- Universities, establishment of, 63
- Uranium, 381, 395
- Urea, estimation of, 365
 synthesis of, 338
- Uric acid, 145, 252
- Urine, analysis of, 488
 chemistry of the, 487, 488
- VALENCY of the elements, 304 *et seq.*
 constant or varying, 319-324
 doctrine of, its influence on the
 development of chemistry, 313
et seq.
 maximum, 319, 320
 of alcohol radicals, 309-310
 of carbon, 309-312
 speculations as to its nature, 337
- Vanadium, 383, 396
- Vapour density determination, methods
 of, 258, 440-442
- Vegetable physiology, 477 *et seq.*
- Versuch über die Theorie der chemischen
 Proportionen* (Berzelius), 201
- Vitriols, 51
- Volumenisms, 293
- Volume theory (Berzelius), 203, 216
- Volumes, law of, extension by
 Avogadro, 202
 law of Gay-Lussac, 188, 201, 202
 law of its appreciation by Berzelius,
 203
- Volumetric analysis, 364
 analysis, beginnings of, 189
 analysis, development of, 364
- Wahlverwandtschaft*, 130, 316
- Water as a type, 283 *et seq.*
 as the primary element (van Hel-
 mont), 74
 composition of, 361, 368
 culture, 481
 glass, 90, 512
 its supposed transformation into
 earth and air respectively, 20, 152
 recognition of its composition, 157-158
- Wood spirit, 402
- ZAFFRE, 87
- Zeitschrift für Chemie*, 534
- Zinc, 16, 47, 138
 dust as a reducing agent, 401
 preparations, 91
- Zirconium, 382
- Zoo-chemistry, 477, 484 *et seq.*

THE END

QD11.M6

Meyer, Ernest S. C. von

A history of chemistry from
earliest times to the present
day

Borrower's Name

Date Due

Scholasticate Library
Inisfada, Manhasset
Long Island, New York

QD11.M6

Meyer, Ernest S. C. von

A history of chemistry from earliest
times to the present day

